WOODS HOLE OCEANOGRAPHIC INSTITUTION

Reference No. 55-19

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By

A. H. Woodcock and D. C. Blanchard

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Observations supporting the salt-nuclei hypothesis of raindrop formation[1]

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Abstract

Atmospheric chlorides in sea-salt nuclei and the chlorides dissolved in shower rainwaters were recently measured in Hawaii. A comparison of these measurements reveals the remarkable fact that the weight of chloride present in a certain number of nuclei in a cubic meter of clear air tends to be equal to the weight of chloride dissolved in an equal number of raindrops in a cubic meter of rainy air. This result is explained as an indication that the raindrops grow on the salt nuclei in some manner which prevents a marked change in the distribution of these nuclei during the drop-growth process.

The data presented add new evidence in further support of the salt-nuclei raindrop hypothesis previously proposed by the first author (11).

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1. Introduction

The earlier study of atmospheric salt particles at cloud levels over Florida (Woodcock, 1952) showed that the number of the larger of these particles per cubic meter of air was similar to the average number of raindrops found during rain storms in other locations. This study also showed that the chloride in these nuclei, assuming that each raindrop forms around a single salt nucleus, was sufficient to account for the range of chloride concentrations usually observed in rains. This similarity of the number of chloride content of salt nuclei and the number of raindrops and the chloride content of rain waters led to the suggestion that raindrops form on the large salt nuclei. It was suggested that further studies should reveal a close relation between the salt in the clear air and salt in the rain falling from clouds subsequently formed in this air.

The purpose of this paper is to present the results of a first test of the salt-particle raindrop hypothesis. This test was carried out during 1952 in Hawaii, where the salt nuclei content of the trade winds is well known and where shower rains occur with great regularity from orographic clouds over the mountains. These clouds form in the wind stream as it flows over the islands, and are usually restricted in vertical development to the lower atmosphere up to the trade-wind inversion at about 2,000 meters.

Rain samples were taken within the clouds at postitons 1, 3, 4, and 5 (see fig. 1) on the island of Hawaii. Sampling of rain within the clouds made it possible to avoid the difficult problems of evaporation of

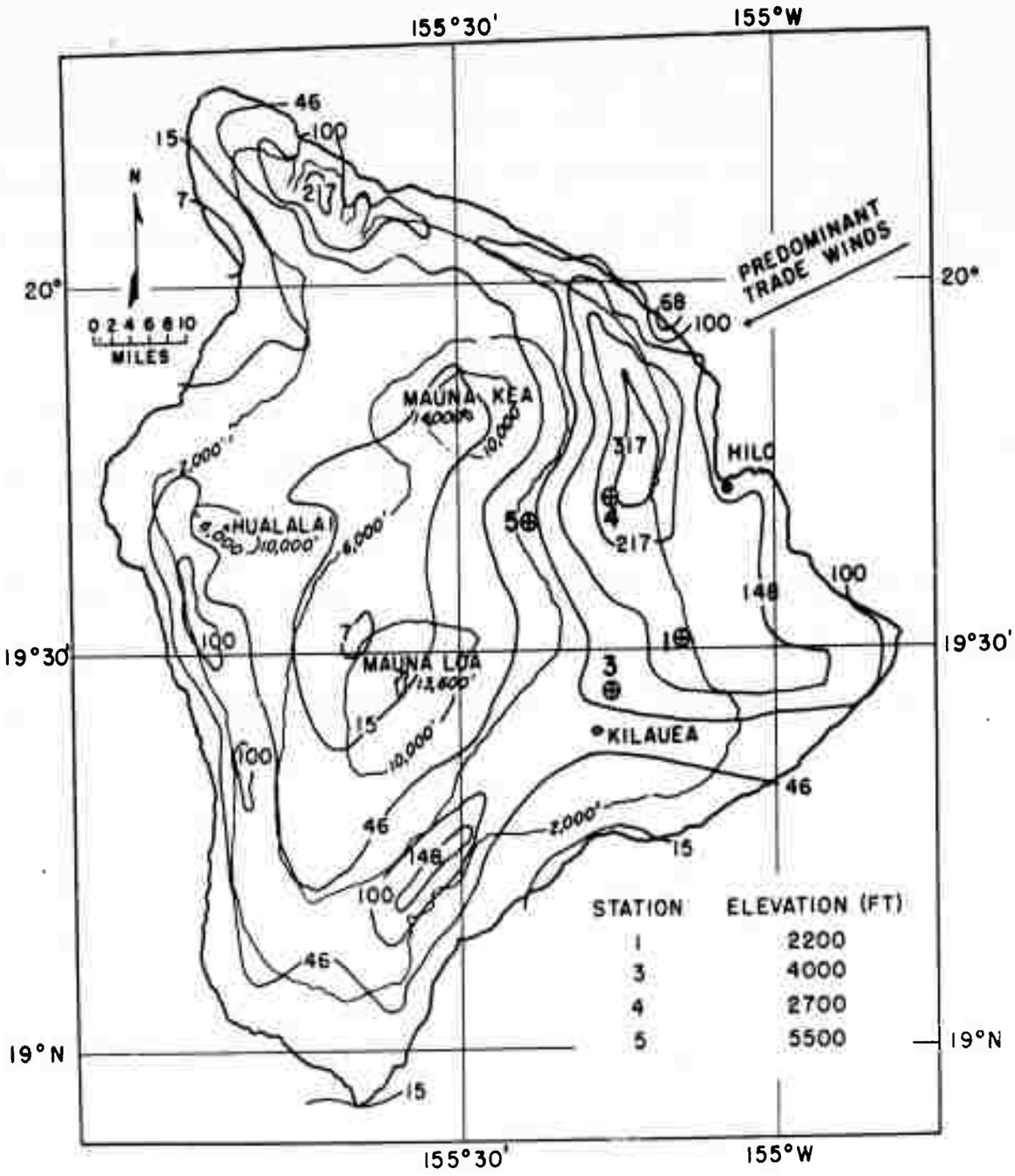

Figure 1

Isohyetal map of the island of Hawaii, with the locations of the rain-sampling stations.

raindrops and of the coalescence of drops with salt particles, which exists in rains which are sampled far below cloud base (6). The sampling positions were selected because positions 1 and 4 are within the bases of the clouds near the area where many of the first shower rains reach the ground, and because positions 3 and 5 are approximately downwind from 1 and 4 respectively and also within the clouds.

If the salt nuclei were growing to raindrop sizes in these shower clouds, we supposed that the larger less numerous nuclei, which are initially the larger of the cloud droplets (5), might become larger raindrops and fall out of the clouds at the windward stations 1 and 4. The remaining small and numerous drops, formed upon the smaller and more numerous nuclei, would be expected to fall more slowly and to reach the ground further downwind near postions 3 and 5. Thus it was anticipated that the raindrop number and the rain water chlorinity at stations 1 and 4 might be found to be related to the number of the larger salt nuclei in the air and that the rain at stations 3 and 5 might be similarly related to the smaller nuclei in the air.

The observations presented below give reasonably consistent support to this anticipated pattern of interrelationship of salt nuclei and rain.

2. Sampling rains and atmospheric salt in Hawaii

In this phase of our work in Hawaii the field problems were first, to measure from aircraft the number and the weight of the salt particles in the air stream before it arrived over the islands and second, to measure on the mountainside the drop size distribution, the

salt content and intensities of the rains within the orographic clouds.

The filter-paper technique was used for measuring the sizes of the raindrops. Exposure of these filter papers was made at about the same time that bulk rain-water samples were taken for chloride analysis. One-half square meter stainless steel funnels were used to obtain these bulk-water samples. Adequate volumes of water for chloride analysis could be taken with these funnels in from 10 to 300 seconds, the time depending upon the rain intensity. The filter-paper and funnel samples were taken only at times when the rain rate appeared to be constant. During the sampling periods rain showers were not observed to be drifting in over the land from the windward sea. Thus the samples were taken in orographic showers, with little rain falling to windward of stations 1 and 4.

The results of the rain sampling at stations 1 and 4 are shown on Table 1, and those at stations 3 and 5 on Table 2. In these tables averages are made from data which is separately tabulated into "parts" having similar rain intensities. The average raindrop distribution from the different rain intensities shown on Table 1 were plotted cumulatively as shown on the lower left part of figure 2.

In the present study it is also necessary to know the liquid water content and the rate of fall of the rain sampled with the filter papers. These quantities were derived as illustrated on Table 4, where rate of fall and liquid water content (columns 6 and 7) are determined from the raindrop distributions found in the highest intensity rains at stations 1 and 4 (see Table 1, part 4).

The chlorinities of the rains, which are given on Tables 1 and 2, were determined by the standard Mohr method, using a microburette and

Table I

Intensities, chlorinities, and drop size distributions in rains at stations 1 and 4.
Sampling by filter-paper and bulk-water techniques.

Position	Date (1952)	Time (LCT)	Bulk water samples: Rate mm hr^{-1}	Bulk water samples: Chloride mg l^{-1}	Rain drops sampled by filter paper technique: Rate mm hr^{-1}	Number of rain drops m^{-3} within various diameter ranges (mm): .2-.4	.4-.6	.6-.8	.8-1.	1.-1.2	1.2-1.4	1.4-1.6	1.6-1.8	1.8-2.	2.-2.2
Part 1 - Rain intensity range 0.4 to 2.0 mm/hour															
1	May 4	1810	1.9	6.3	0.71	4320	855	21							
1	" 4	1815	1.9	6.2	0.90	3200	1300	43							
1	" 4	1827	0.5	5.7	0.21	2620	19								
1	" 4	1832	1.8	2.8	1.00	6900	971	21	8						
4	" 6	1655	1.9	13.7	1.30	3210	890	287	20						
4	" 6	1706	0.4	24.8	0.71	4050	790	41							
4	" 6	1716	0.4	22.9	0.38	1600	580								
			1.26	11.8	0.74	3700	770	59	4						
Part 2 - Intensity range 2.0 to 5.0 mm/hour															
1	May 1	1339	2.9	6.8	3.4	1500	385	390	265	52	26				
1	" 4	1835	2.3	2.9	3.2	1700	850	620	225	40					
1	" 4	1842	2.4	3.2	2.0	1800	1770	390	41	6					
4	" 6	1722	3.2	8.9	2.3	2170	524	590	168						
4	" 6	1734	3.9	8.0	3.8	1080	137	355	550	31	5				
			2.6	6.0	2.9	1650	733	468	250	26					
Part 3 - Intensity range 5 to 10 mm/hour															
1	May 4	1848	7.0	1.3	8.0	111	520	690	670	194	54				
4	" 6	1735	7.9	7.7	8.4	790	280	82	87	121	65	125	14		
			7.45	4.5	8.2	450	400	386	379	158	60	63	7		
Part 4 - Intensity range > 10 mm/hour															
1	May 1	1344			20.6	1282	369	76	206	164	198	202	23	55	7
1	" 4	1846	12.9	1.5	25.0	290	196	251	710	850	410	85	11		
			12.9	1.5	22.8	786	282	163	458	507	304	144	17	28	4

Table II

Intensities, chlorinities and drop size distributions in rains at stations 3 and 5. Sampling by filter-paper and bulk-water techniques.

Position	Date (1952)	Time (LCT)	Bulk water samples: Rate mm hr^{-1}	Bulk water samples: Chloride mg l^{-1}	Rain drops sampled by filter paper technique: Rate mm hr^{-1}	Number drops m^{-3} in various diameter ranges (mm): <0.2	0.2-0.4	0.4-0.6	0.6-0.8	0.8-1.0	1.0-1.2
Part 1 - Intensity range 0.1 to 1.0 mm/hr^{-1}											
5	Apr. 28	1702	0.13	2.10	0.09	48,700	85				
5	" 28	1740	0.22	2.3	0.19	41,500	2,040				
5	" 28	1812	0.23	6.8	0.06	25,400	243				
5	" 28	1819	0.23	6.8	0.19	78,500	930				
5	" 28	1834	0.33	2.9	0.18	110,000					
3	May 1	1513	1.00	0.8	0.58	800	5,800	485			
3	" 4	1710	0.74	0.8	0.87	20,000	8,600	56			
3	" 4	1737	0.29	0.4	0.37	28,500	5,400				
5	" 5	2043	0.80	0.6	0.65	101,000	7,030	144			
5	" 5	2116	1.00	0.3	0.53	80,800	6,400	33			
5	" 5	2138	0.22	0.4	0.19	66,200	1,360				
5	" 5	2149	0.11	1.4	0.25	149,000					
5	" 5	2222	1.00	0.2	0.26	29,500	3,600				
5	" 5	2235	0.92	0.2	0.45	51,000	6,100				
			0.52	1.9	0.35	59,400	3,400	51			
Part 2 - Intensity range 1.0 to 3.0 mm/hour											
5	Apr. 28	1727	1.09	2.00	1.43	116,000	11,500	1150			
5	" 29	1759	2.28	1.2	1.95	2,450	5,400	2060	270	24	
3	May 1	1455	2.70	0.7	2.40		8,300	3860	31		
5	" 5	2053	1.10	0.4	1.34	65,000	11,200	1120	27		
			1.79	1.1	1.78	45,800	9,100	2050	82	6	

a one-hundredth normal silver nitrate solution.

Atmospheric salt particles were sampled in clear air at cloud levels over the sea by exposing small glass slides from aircraft. The methods used to sample atmospheric salt and those used to sample the rains have been described previously (1) (10) (11), and will not be discussed further in this paper.

It proved to impracticable, during the 1952 field trip, to obtain simultaneous observations of airborne salt and of rain. This difficulty was eliminated because it had been found, during many previous salt-sampling airplane flights, that a clear relationship existed between the wind force and the quantity of salt nuclei in the air at cloud levels. Figure 3 shows the averaged results from the salt nuclei samples taken in Hawaii on twenty-three different days between June 1951 and July 1952. Differences in the distribution of mass of salt among the nuclei on these days are shown cumulatively on figure 4. This figure also shows that the smoothed curves on figure 3 are based upon from nine to twenty-seven measured size-range categories.

Thus by using this salt-nuclei data and the average wind force observations from U. S. Weather Bureau surface maps (see table 3), we were able to make a quantitative estimate of the average weight of salt as nuclei in the air during the interval of days when we sampled rain. This average salt amount was compared to the average weight of salt in the rain on these days.

3. Relating the salt in the nuclei to the salt in the rain at the windward stations 1 and 4

As previously stated, the purpose of the present study is to test

Table 3

Surface winds over the sea, during the days when rain was sampled, as reported by ships and planes within 400 nautical miles of Hilo, Hawaii. Taken from the four daily U. S. Weather Bureau surface maps of the North Pacific.

Date 1952	Number of observations reported	Average wind force
April 28	14	4.9
April 29	5	3.6
May 1	11	3.3
May 4	8	3.8
May 5	14	3.8
May 6	9	3.8
		av. = 3.9

the idea that a close relationship exists between the weight of the larger salt nuclei per cubic meter of clear air and the weight of salt dissolved in the rain in a cubic meter of rainy air. Expressed differently the above idea would mean that the largest salt nuclei form the raindrops, and that the individual drops which make up the first rains coming from a shower cloud should contain the salt found in an equal number of the largest salt nuclei.

The average numbers and sizes of raindrops for various rain intensities at stations 1 and 4 are plotted cumulatively on figure 2 (see curves, lower left), starting with the largest drops observed. On the same figure the average distribution of number and weight of salt nuclei is similarly presented at the lower right. This salt nuclei curve, taken from the data shown on figure 3 is used because the observed average wind force during the days when the rains samples were taken was very nearly force 4 (see table 3).

Figure 2 and table 4 contain the information required to compute the chloride concentration which should be observed in rains if each of the larger salt particles has become a raindrop and if, in the process, little change has occurred in their distribution in the air. For example: rains averaging 6328 mm^3 m^{-2} sec^{-1} intensity were made up of a definite number of drops of each size range falling per square meter per second (see table 4, columns 4 and 5). The approximate weight of sea salt in the nucleus of each of these raindrops is taken from the force 4 cumulative salt nuclei curve on figure 2, using the appropriate cumulative raindrop curve to establish, in each case, this nucleus weight. These weights are shown on table 4, column 8. As an indication of the derivation of these weights,

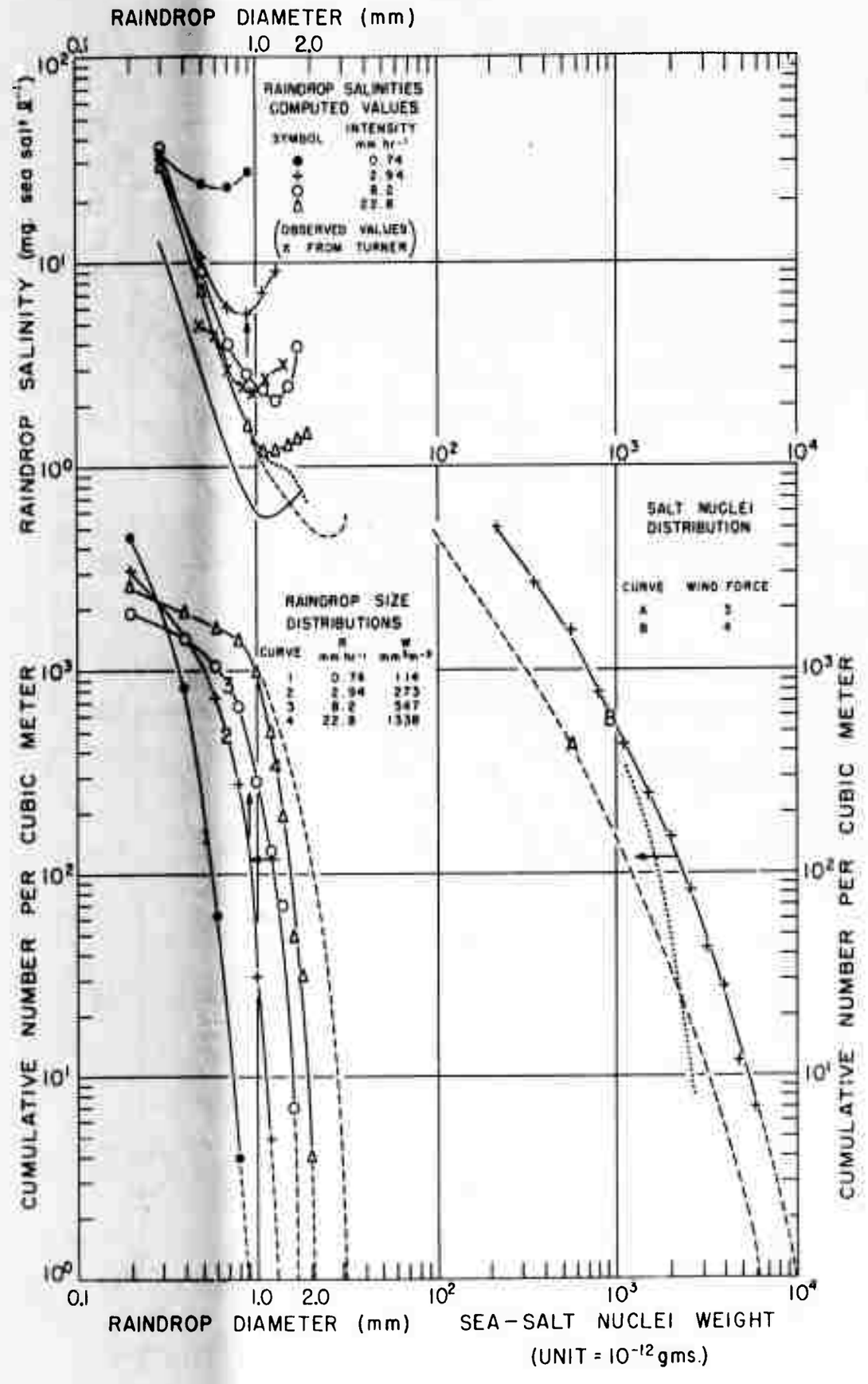

Figure 2

The lower part of this figure shows cumulative average number distribution curves for salt nuclei and for drops in rains of various intensities. These curves are used, as described in the text, to compute the average chlorinity of the rain waters (fig. 5) and to compute the raindrop salinity values which are given above

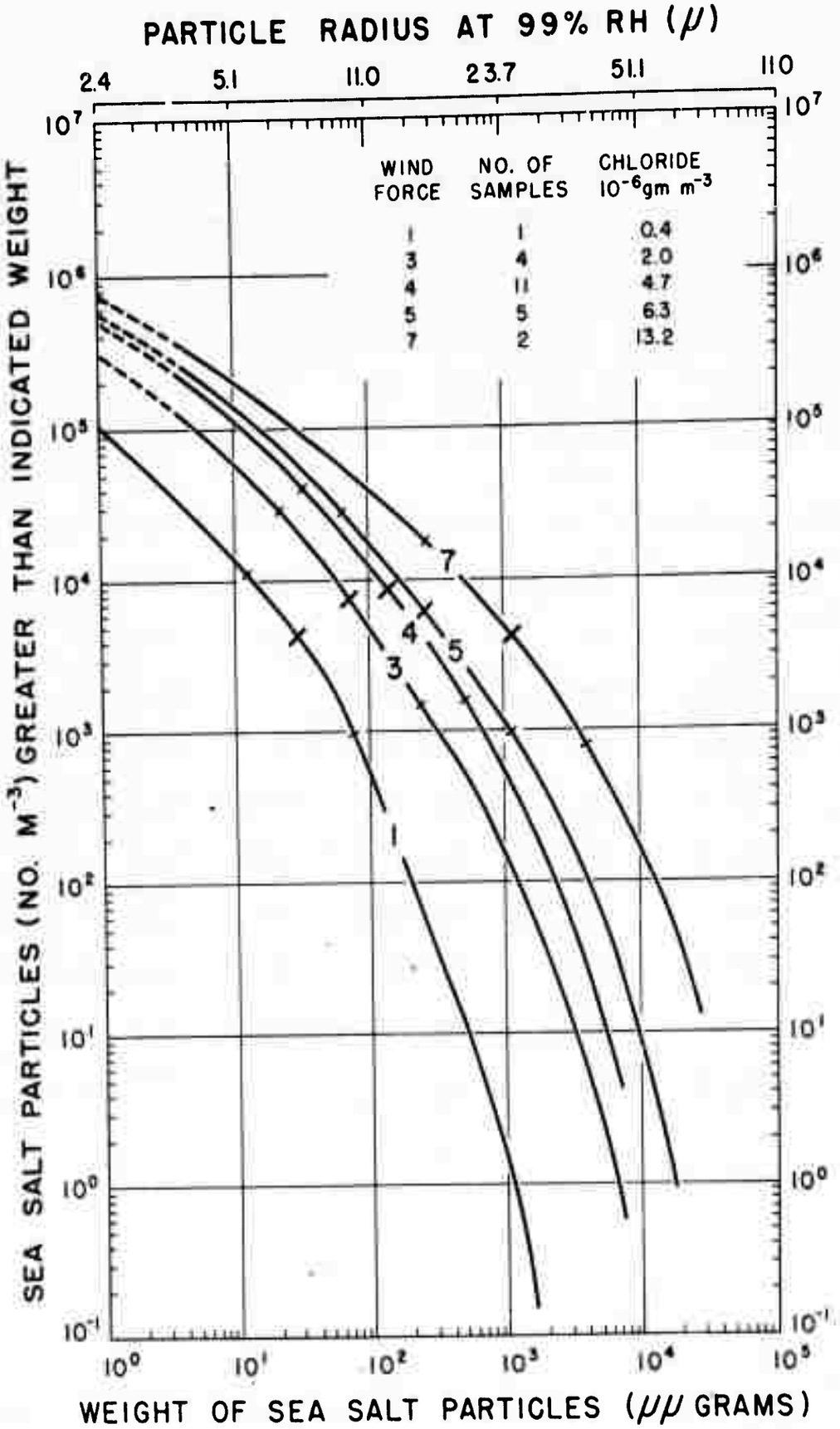

Figure 3

Cumulative distribution curves for averaged salt nuclei number during winds of various forces in Hawaii. (taken from Woodcock, ref. 12)

Table 4

Averaged rain data derived from
filter paper samples, table 1, part 4

1	2	3	4	5	6	7	8	9
Drop dia.	Drop no.	Drop vol.	Drop fall rate	Drop no. falling	Rain Rate mm^3	Rain water in air	Wt. sea salt in each drop unit =	Total sea salt falling in drops
mm	m^{-3}	mm^3	$m\ sec^{-1}$	$m^{-2}sec^{-1}$	$m^{-2}sec^{-1}$	mm^3m^{-3}	$10^{-12}g$	$10^{-6}gm\ m^{-2}sec^{-1}$
0.3	786	.011	1.17	920	13	11.1	420	0.386
0.5	282	.065	2.06	581	38	18.3	485	0.282
0.7	163	.180	2.90	473	85	29.4	530	0.251
0.9	458	.375	3.65	1670	628	172	600	1.000
1.1	507	.700	4.30	2180	1523	354	830	1.810
1.3	304	1.14	4.90	1490	1700	347	1,350	2.010
1.5	143	1.73	5.45	780	1348	248	2.270	1.770
1.7	17	2.56	5.95	100	256	43	3,500	0.350
1.9	28	3.59	6.35	175	628	99	5,100	0.892
2.1	3	4.85	6.75	22	108	16	10,000	0.220
					6328	1338		8.971

note on salt curve B (fig. 2) that salt particles weighing 420$\mu\mu$gms., occur at the cumulative number 2150, which corresponds to the same number of 0.3 mm drops on raindrop curve number 4. Thus 420$\mu\mu$gms. is entered on table 4, column 8, as the weight of sea salt to be found in a 0.3 mm raindrop and in rains of the indicated average intensity. Column 9 gives the product of these nucleus weights and the numbers of raindrops (column 5). The computed salinity of the rain (S_c) is now simply the ratio of the total weight of sea salt falling, to the total amount of water falling.

$$S_c = \frac{Ms}{W} \qquad (1.)$$

where S_c = computed sea-salt concentration of rain (Mg l^{-1})

Ms = sea salt falling m^{-2} sec^{-1} (Mg)

W = liquid water falling m^{-2} sec^{-1} (liters)

For the high-intensity rains averaged on table 4, the amounts of sea salt and water falling are 8.97 μ g m^{-2} sec^{-1} and 6328 mm^3 m^{-2} sec^{-1} respectively. The ratio of these quantities, with the units adjusted, yields a computed sea-salt concentration of 1.42 mg l^{-1}, or a chloride concentration 0.8 mg l^{-1}.

The above computed value for the average chloride concentration, plus other concentrations similarly computed from the averaged data concerning rains of lower intensity (see table 1) are plotted on figure 5.* This

*These computed chloride concentrations for the rains of other intensities were derived in the manner shown on table 4, but are not tabulated here in order to save space and to avoid unnecessary repetetion.

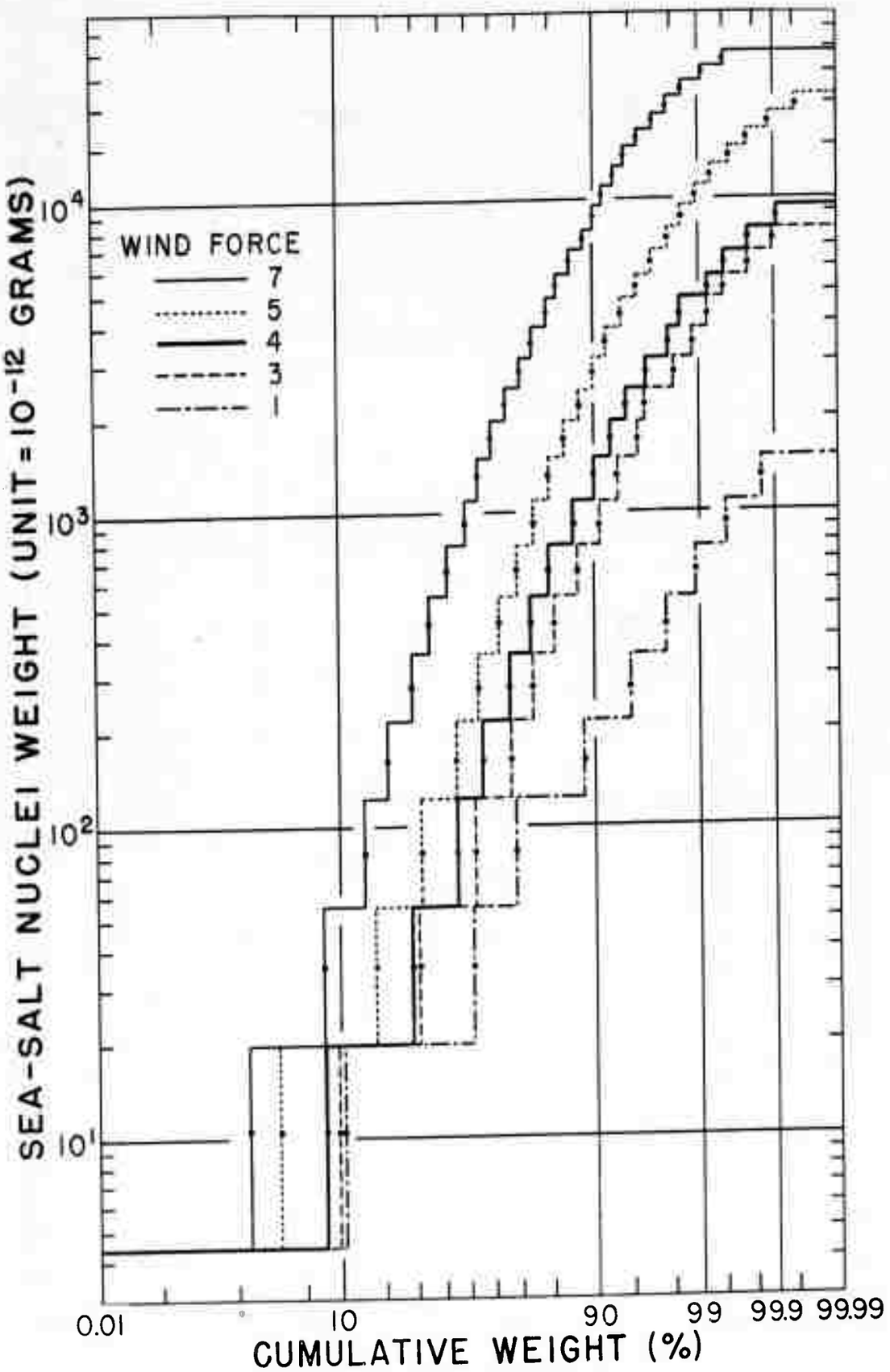

Figure 4

Averaged cumulative percent weight distribution curves for salt nuclei present in the atmosphere in Hawaii under varying wind conditions (forces 1 through 7).

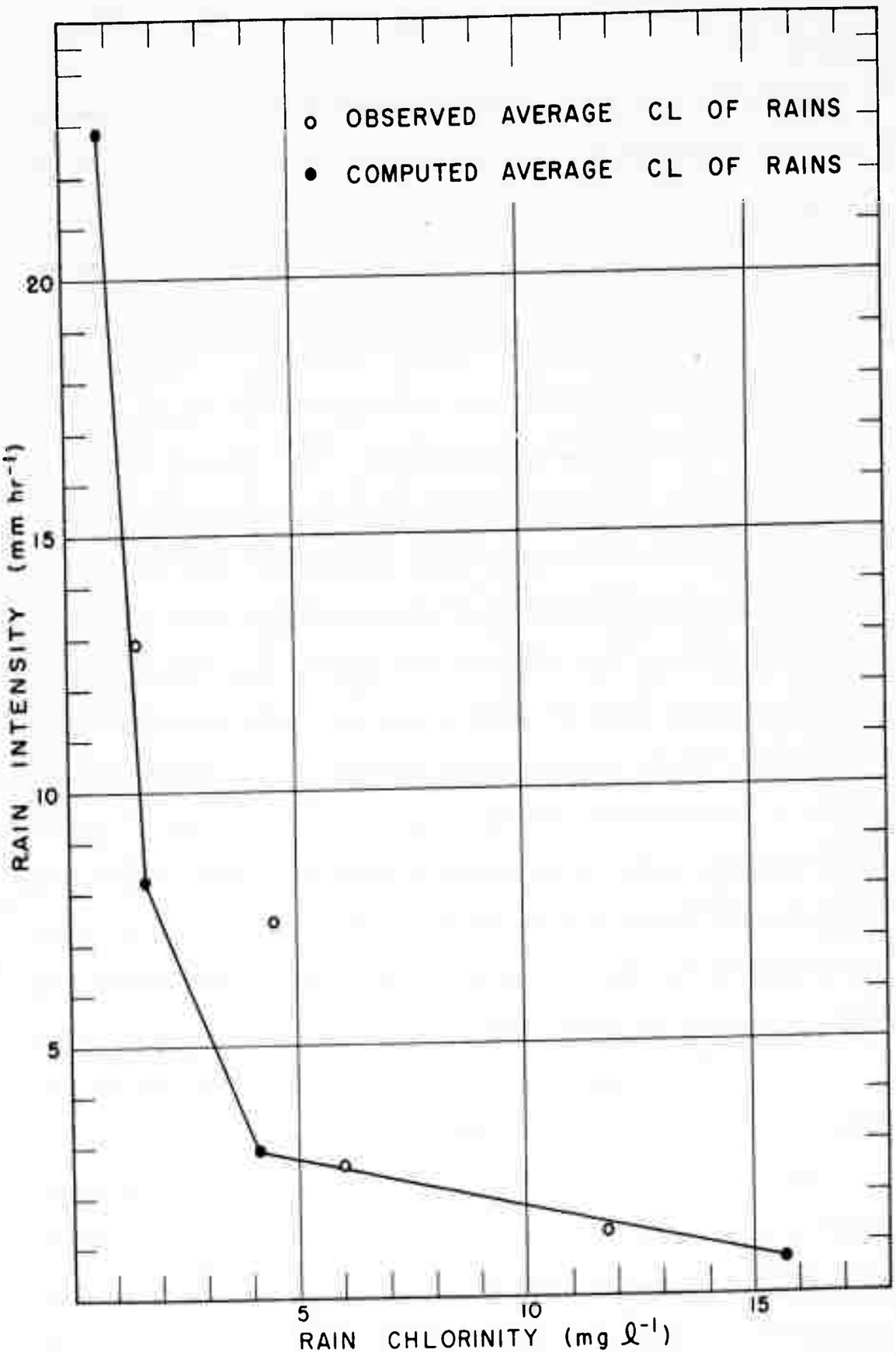

Figure 5

Observed and computed chlorinities of rains related to rain intensity.

also shows the observed average chlorinities of the rains at stations 1 and 4. These observed chloride concentrations are remarkably similar to the computed concentrations, and this similarity is regarded as suggesting that each large salt particle has formed a raindrop.

Relating the salt nuclei to the salinity of the rains at the leeward stations 3 and 5

As previously suggested on pages 2 and 3, the rains falling at the leeward stations 3 and 5 were found to be made up of great numbers of small drops containing relatively little chloride (see table 2). This character of these leeward rains tends to support the idea that they are composed of drops which have formed on the very numerous salt nuclei found among those particles weighing from about 5 to 20$\mu\mu$ gms. The average chloride in these rains (see table 6, column F) equals 6.2 percent of the total chlorides found during force 4 winds (total = 4.7×10^{-6} gms m^{-3}). On figure 4 it can be seen that about 7.5 percent of the chlorides in the air are between 4.5 and 20$\mu\mu$ grams in weight, and in figure 3 the salt nuclei curve for force 4 winds shows that there are about 1.4×10^{5} particles present in this weight range. Thus the chloride content and the numbers of the small salt nuclei are quite adequate to account for the chloride content and the numbers of the raindrops in the leeward rains.

Due to the nature of the techniques used to measure these small raindrops and small nuclei, it is not possible, with the present data, to relate the raindrops to the nuclei more precisely. Such precision could be obtained in future measurements by dividing the smaller raindrop and salt nuclei number distributions during the measuring process into greater numbers of size-range categories. Note on table 2 that the greatest numbers of the raindrops at stations 3 and 5 are in the first size category,

Table 5

Summary of averaged rain data from windward stations 1 and 4, Hawaii.
May 1, 4, and 6, 1952.

A	B	C	D	E	F	G	H	J	K	L
Rain from funnel sampler		Rain from filter paper samples			Rain Cl computed from nuclei	Rain Cl m^{-3} of air (ExF)	Wt. of Cl falling in rain 10^{-6}gm	Ratio of Cl in rain to Cl in air	Rain-drop size range	Total number of rain-drops
Rate	Cl conc.	Rate		Water content						
mm hr^{-1}	mg l^{-1}	mm hr^{-1}	mm $m^{-2}sec^{-1}$	mm^3m^{-3}	mg l^{-1}	10^{-6}gm	$m^{-2}sec^{-1}$	G/4.7	mm dia.	m^{-3}
1.26	11.8	0.74	206	114	15.7	1.79	5.68	0.38	0.3-0.9	4533
2.60	6.0	2.94	821	273	4.1	1.12	6.09	0.24	0.3-1.3	3132
7.45	4.5	8.20	2279	547	1.7	0.93	6.80	0.19	0.3-1.7	1903
12.90	1.5	22.80	6329	1338	0.8	1.07	9.0	0.23	0.3-2.1	2693

Table 6

Summary of averaged rain data from leeward stations 3 and 5, Hawaii.
April 28 and 29, and May 1, 4, and 5, 1952

A	B	C	D	E	F	G	H	J
Rain-water samples								
funnel collector		filter-paper samples						
Rate	Chloride	Rate		Rain in air	Rain Cl m^{-3} of air (BxE)	Ratio of Cl in rain to Cl in air	Raindrop size range	Raindrop number m^{-3} of air
mm hr^{-1}	mg l^{-1}	mm hr^{-1}	$mm^3 m^{-2} sec^{-1}$	$mm^3 m^{-3}$	10^{-6}gm	F/4.7	mm dia.	
0.52	1.9	0.35	97	119	0.22	.047	0.1-0.3	62,850
1.79	1.1	1.78	494	331	0.36 / 0.29	.077	0.1-0.9	57,038

while figure 4 shows that all of the appropriate salt nuclei fall within a single measured size-range category.

Most of the raindrops sampled at stations 3 and 5 are so small that due to evaporation it is very unlikely that they would survive a fall through many hundreds of meters of clear air. This probably explains why these small drops have not been reported in such great numbers by other observers.

Computed raindrop salinity compared to observed salinity

During the spring of 1954, Turner (9) measured the salinity of individual raindrop size ranges in Hawaii. For obtaining rain waters from different drop size-range categories, he used a simple modification of the raindrop spectrograph which has been described by Bowen and Davidson (3). Turner determined the sodium chloride content of drops in many rain samples at cloud base near station 4, and it is useful here to compare his interesting observations to the raindrop salinities computed from our raindrop size distributions obtained at the same location. For this purpose his sodium chloride concentration values were converted to total sea-salt concentration by multiplying by 1.2, the approximate ratio of total salts in the sea to the sodium chloride present.

On figure 2 (top) his average raindrop salinity values for 39 samples are compared to similar values computed from the average rain and salt nuclei data (lower curves on figure 2) used in the present paper. The method used to derive these latter raindrop salinities is given in equation (1) and is applied, for example, as follows. A nucleus, having a mass of approximately 2,100 $\mu\mu$gms., corresponds in cumulative number

per cubic meter of air, to the number of 0.9 mm diameter raindrops shown on one of the raindrop distribution curves (see arrow marks near lower curves on fig. 2). The ratio of this nucleus mass (2.1×10^{-6} mg) to the volume of a 0.9 mm raindrop ($.375 \times 10^{-6}$ liters), equals a concentration of 5.6 mg l^{-1}. This drop salinity is indicated by the arrow mark beside the second raindrop salinity curve at the top of figure 2. The other computed values for raindrop salinity were similarly derived.

It is interesting that the form of Turner's average observed raindrop salinity curve and those computed from our raindrop and salt nuclei data is similar, despite the differences in the sampling times and in the methods used to derive the chloride values. The similarity of the actual average values of the drops to some of our computed values is perhaps coincidental, but the occurrence of the pronounced salinity minimum in both results is apparently a reflection of the average slopes and relative values of the cumulative raindrop and salt nuclei curves. For instance, if the observed cumulative average raindrop distribution curve represented by curve 4 (see figure 2, lower left) had actually followed the adjacent dashed line, then the computed raindrop salinity minimum would not have occurred at a drop diameter of 1.1 mm (see triangular symbols, upper diagram, fig. 2). Instead a different salinity curve would have occurred (see extended dashed line), with the salinity minimum falling at a drop diameter of about 2.5 mm. As an indication of the effects of changing nuclei distribution, the assumption of a force 3 salt-nuclei curve instead of the average force 4 curve (see curve A, fig. 2), simply shifts the computed drop salinity curves down. The unbroken drop salinity curve on figure 2 shows the dilution effects of the reduced nucleus

sizes from the force 3 salt curve upon drop salt concentrations for the highest intensity rains. Thus the form, and, of course, the slope of the raindrop salinity curves is clearly dependent upon the relative slopes and magnitudes of the cumulative raindrop and salt-nuclei curves. Note that the use of a salt nuclei distribution curve represented by the dotted-line modification of curve B, in computing raindrop salinity, removes the salinity minimum from the drop salinity curve for the highest intensity rains (see dotted line at top of fig. 2).

Turner's raindrop salinity curve is based upon many hours of sampling, during which time the rain intensity probably varied greatly. Had he been able to collect his rain water during short periods of time (i.e., in about 100 seconds) and, in each case during rains of nearly constant intensity, it is supposed that he would have obtained salinity curves similar to ours. Thus Turner's values for the salinity of raindrops of each size range, may represent an integrated quantity. Note for instance on figure 2 that the salinity of the 0.9 mm diameter raindrops computed from our curves may vary from 1.6 to 28 mg l^{-1}, depending upon the rain intensity in which they are sampled. The raindrop salinity values which will be obtained during a prolonged sampling period in orographic rains at station 4 would seem to depend, in part at least, upon the relative frequency of occurrence of rains of different intensities. They will also presumably depend upon the extent to which his samples were taken in the first rains from the showers or in rains which were the winnowed remnants of showers occurring to windward of the raindrop spectrograph sampling position.

In future physical-chemical studies of rain, it would be very

useful to combine raindrop spectrograph measurements of drop salinity, with simultaneous determinations of raindrop size distribution and of salt nucleus size distributions.

Discussion and conclusions

Clearly many more observations, such as those given on table 1 and 2, are needed in order to thoroughly test the consistancy of the pattern of interrelationship of nuclei to raindrops which has evolved from the present study. Data deficiencies are especially great among the higher intensity rains (see table 1), and the lack of simultaniety of the rain and salt-nuclei observations introduces a further element of uncertainty. However, the authors feel that the data and ideas presented here, when considered as a whole, are sufficiently useful to justify publication at this time.

Some of the above observational deficiencies were eliminated in further measurements recently made (1954) in Hawaii in connection with the co-operative cloud physics studies of "Project Shower". These numerous measurements, which will require much time to completely analyse, will be presented in a later more complete study of this problem.

A rather obvious result of the present study is the emphasis which is given to the importance of detailed knowledge of the salt nuclei distribution in the clear air and the distribution of raindrop size and raindrop salinity within the clouds. It is also important to know where, in time, rain samples are taken in relation to the stage of development of the showers. In the present study an effort was made to sample the first and the last rains falling from showers moving up the mountain slope. The authors feel that this detailed knowledge is essential to a further understanding of the evolution of raindrops in the shower rains of marine air masses. The techniques used here should also prove useful in exploring

the role of sea-salt nuclei in the formation of raindrops in continental rains.

It should be pointed out that in the analysis presented here the quantities of salt falling in the rains at stations 1 and 4 amount to only about 20 to 40 percent of the total salt in the air (see summary table 5, column J). Since, of the remaining 60 to 80 percent of the atmospheric salt, only about 6 percent falls at stations 3 and 5 (see summary table 6, column G), it is supposed that the remaining salt rained-out in the drops falling between these stations, as the rain showers were carried by the winds up the mountain from the lower to the upper stations. In a later publication (of "Project Shower" observations) the authors plan to include an analysis of the effect of wind and wind shear upon the distribution of raindrops at various locations on the mountainside.

In the Hawaiian shower rains, rapid changes in intensity with time are very common, as shown on tables 1 and 2, and by Blanchard (see fig. 2, ref. 1). Most of this change is thought to be due to winds transporting the showers over the sampling stations. For instance the average clouds in the study were probably moving at about 7 meters per second up the mountain slope. During a sixty second interval between rain samples a shower-producing cloud could thus move almost one-half kilometer. The showers are seldom large in diameter, and in aircraft a flight of one-half kilometer is often sufficient to carry the observer from the misty peripheral regions to the relatively intense rains of the center.

The interpretation of the measurements presented here, in terms of the details of the raindrop-forming processes operating in these Hawaiian rains, is uncertain. They seem, however, to be clearly useful in narrowing

the areas of search for these "details". One can say from the present measurements, that a drop-forming process seems to be required which will, (a), involve most of the larger salt nuclei in a simultaneous growth, (b) cause this growth to occur, presumably very rapidly, in a manner which prevents marked changes in the distribution of the nuclei (or drops), and (c) add water to the nuclei without at the same time greatly adding to their salt content.

In the present study the quantities of salt in the individual large salt nuclei are adequate to account for all of the salt found in the rains. This suggests that the drops have grown entirely by condensation processes, or through coalescence with relatively chloride-free cloud droplets.* If one assumes, for instance, that the 0.9 mm diameter raindrops (which, on rain curve 2, fig. 2, are supposed to have formed on 2100 $\mu\mu$ gram salt nuclei) have grown through coalescence with 10 μ radius cloud droplets, about 91,000 of these droplets would be required. Table 7 shows amounts of salt which would be added to the initial giant (2100 $\mu\mu$ gram) nucleus due to coalescence with cloud droplets which have formed on sea-salt nuclei of various sub-micron sizes. On this table it is evident that the quantities of salt added through accretional growth might in some instances equal or greatly exceed the weight of the giant nucleus upon which the raindrop initially formed. The very small nuclei would, of course, require rather large water vapor supersaturations in the clouds to become activated.

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* Hawaiian clouds, which are not producing rain, have been found to contain from 30 to 140 cloud droplets per cubic centimeter.

Table 7

Weights of sea salt added to a 0.9 mm diameter raindrop due to coalescence with 10 and 20 μ radius cloud droplets containing sea-salt nuclei of various sub-micron sizes.

Cloud droplet size μ	Cloud droplet no. required to form 0.9 mm diameter raindrop	Weight sea salt (10^{-12} gms.) in 0.9 dia. raindrops (cloud droplet no. x nucleus wt.)				
		(0.5 μ) 10^{-12} gm	(0.2 μ) 10^{-13} gm	(0.1 μ) 10^{-14} gm	(0.05 μ) 10^{-15} gm	(0.02 μ) 10^{-16} gm
10	91,000	91,000	9,100	910	91	9.1
20	11,400	11,400	1,140	114	11.4	1.1

Junge's work (4) indicates that most of the numerous nuclei in marine air which are less than about 0.8 μ in radius (or $< 5 \times 10^{-12}$ gms., as shown on fig. 2) are composed largely of ammonium sulfate. If Junge is right and if these ammonium sulfate particles are the nuclei upon which most of the cloud droplets form, little or no chloride would be added to a raindrop, formed upon a giant sea-salt nucleus, which had grown by coalsecence with these cloud droplets.

Thus one cannot say, from the results of this paper, that the raindrops in Hawaii do or do not grow by accretion. If the nuclei smaller than about 10^{-12} grams are sea salt, then accretion would seem to be ruled out unless the nuclei are very small (see table 7). If they are ammonium sulfate, as Junge indicates, then one should find appropriate quantities of this compound in solution in the Hawaiian rain waters. Ammonium sulfate is actually found in these waters (Junge), but we cannot know to what extent the quantities present are appropriate to the assumption of raindrop growth by accretion because the numbers and the masses of the original nuclei are not yet known.

Thus we feel that one of the important steps in future physical-chemical studies of raindrop formation in shower clouds is to determine the chemical nature and number of the majority of the individual cloud-droplet nuclei which are smaller than about 0.8 microns and to relate these particles to the solutes found in the cloud and rain waters.

If the Hawaiian raindrops do in fact grow largely by accretion, the authors feel that it becomes necessary to suppose, from the results of the present study, that this growth occurs in turbulent volumes of air which ascend within clouds much as do the "bubble" parcels pictured by

Scorer and Ludlam (7). The conservation of the number of the larger nuclei per unit volume of air, which is required in the results given here, seems to rule out the type of raindrop motion through the clouds which is pictured by Bowen (2) as being associated with accretional growth.

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