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AEDC-TDR-63-190



MOLLIER DIAGRAM FOR AIR

By

Wanda J. Little
von Kármán Gas Dynamics Facility
ARO, Inc.

TECHNICAL DOCUMENTARY REPORT NO. AEDC-TDR-63-190

September 1963

Program Element 61405014/8951, Task 89603

(Prepared under Contract No. AF 40(600)-1000 by ARO, Inc.,
contract operator of AEDC, Arnold Air Force Station, Tenn.)

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AIR FORCE SYSTEMS COMMAND
UNITED STATES AIR FORCE**

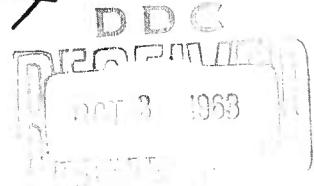
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MOLLIER DIAGRAM FOR EQUILIBRIUM AIR
ARO, INC. MARCH 1964

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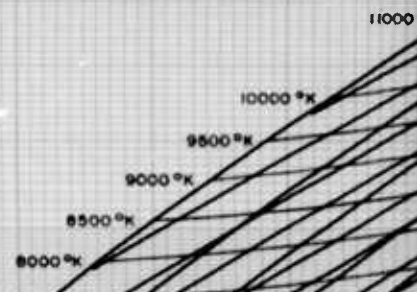
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MOLLIER DIAGRAM FOR EQ

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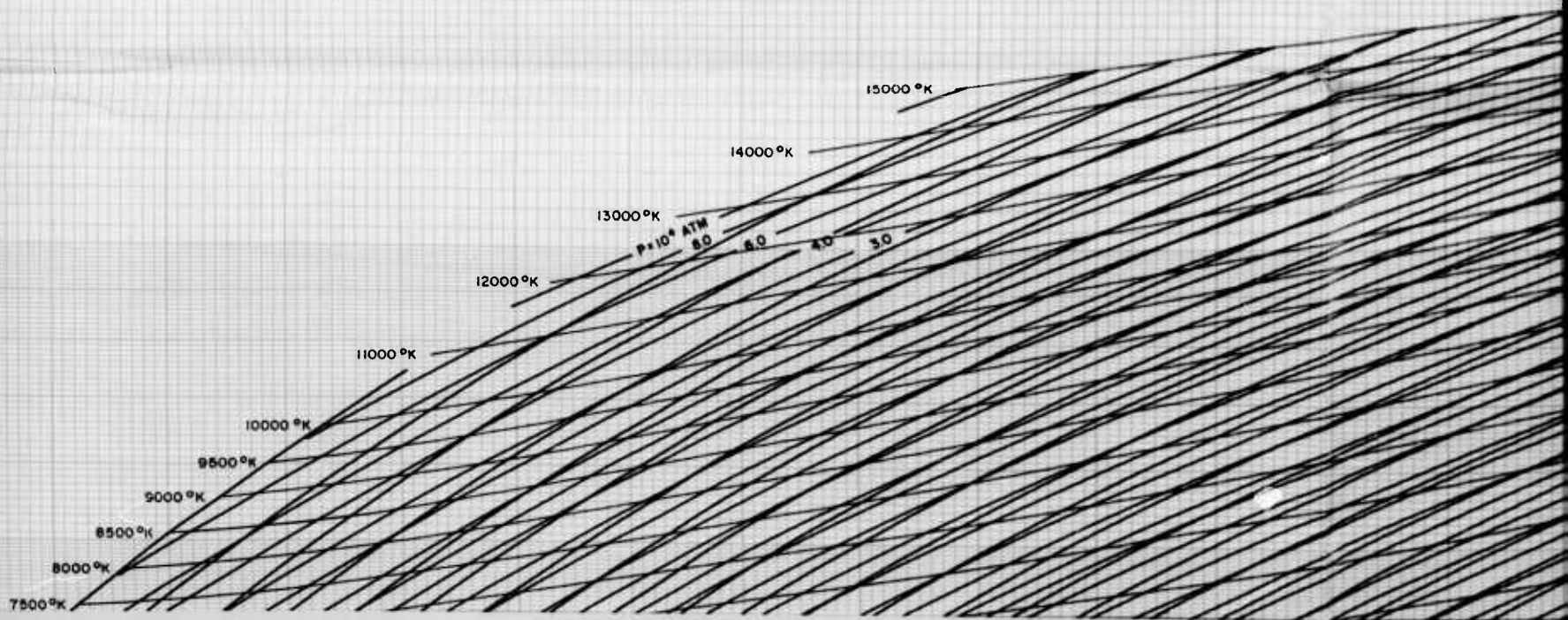
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M FOR EQUILIBRIUM AIR

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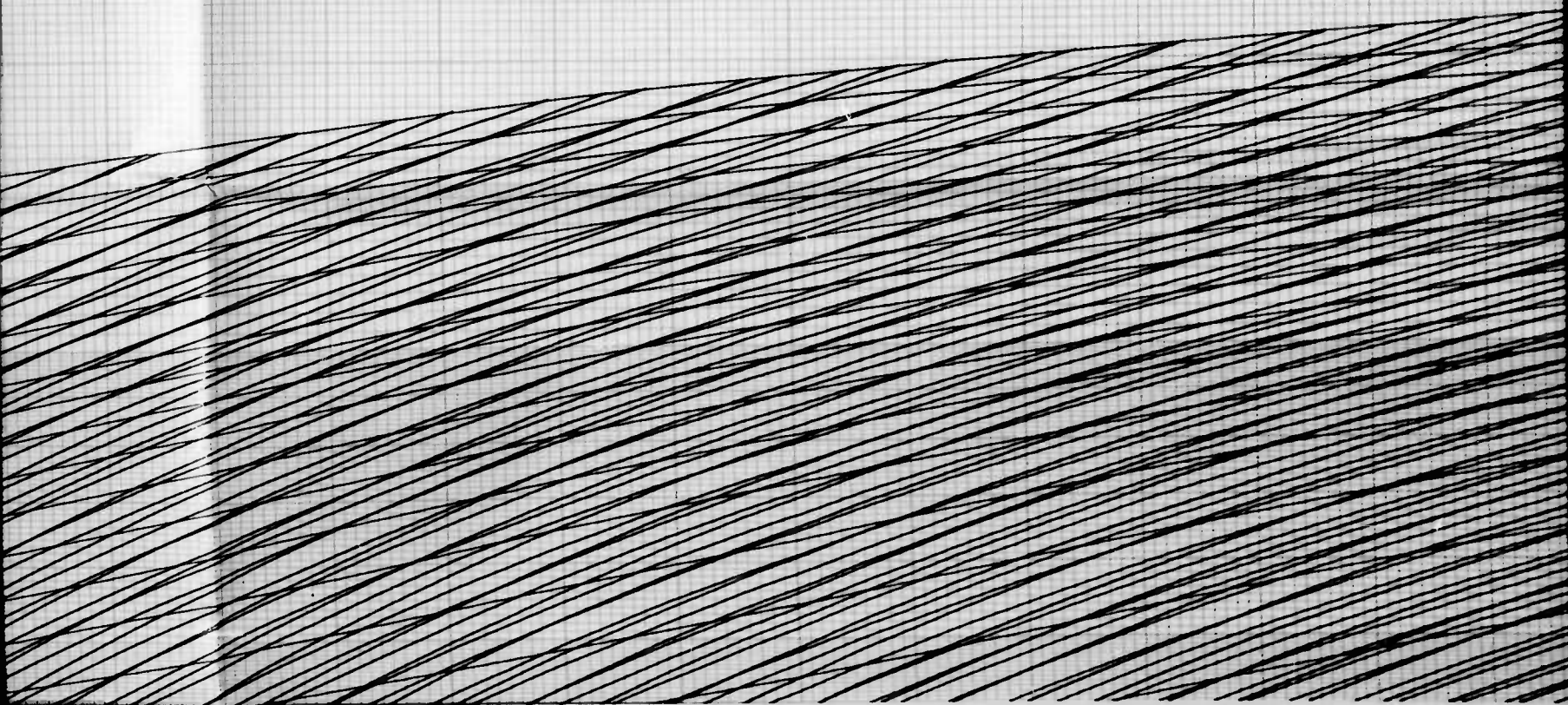
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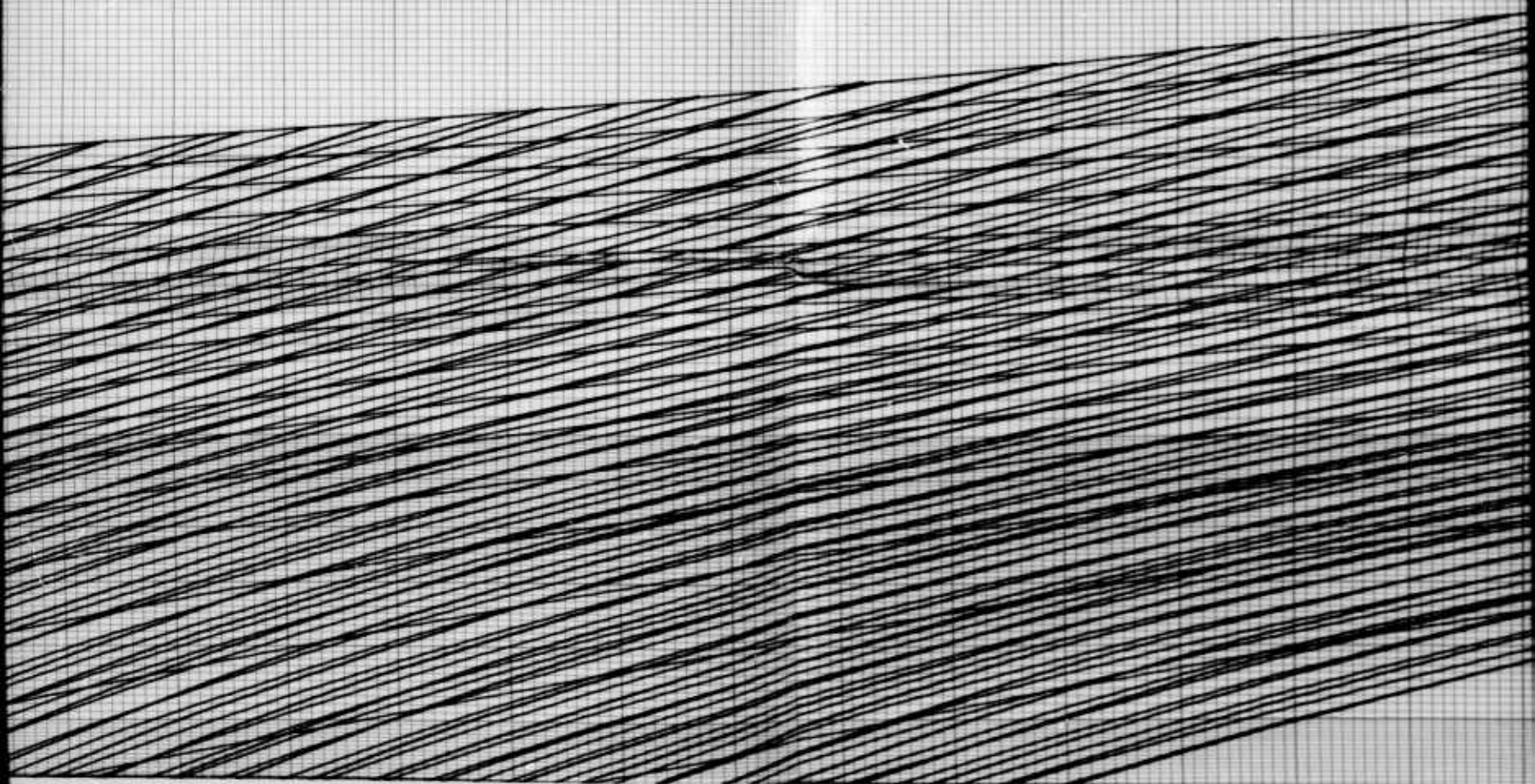
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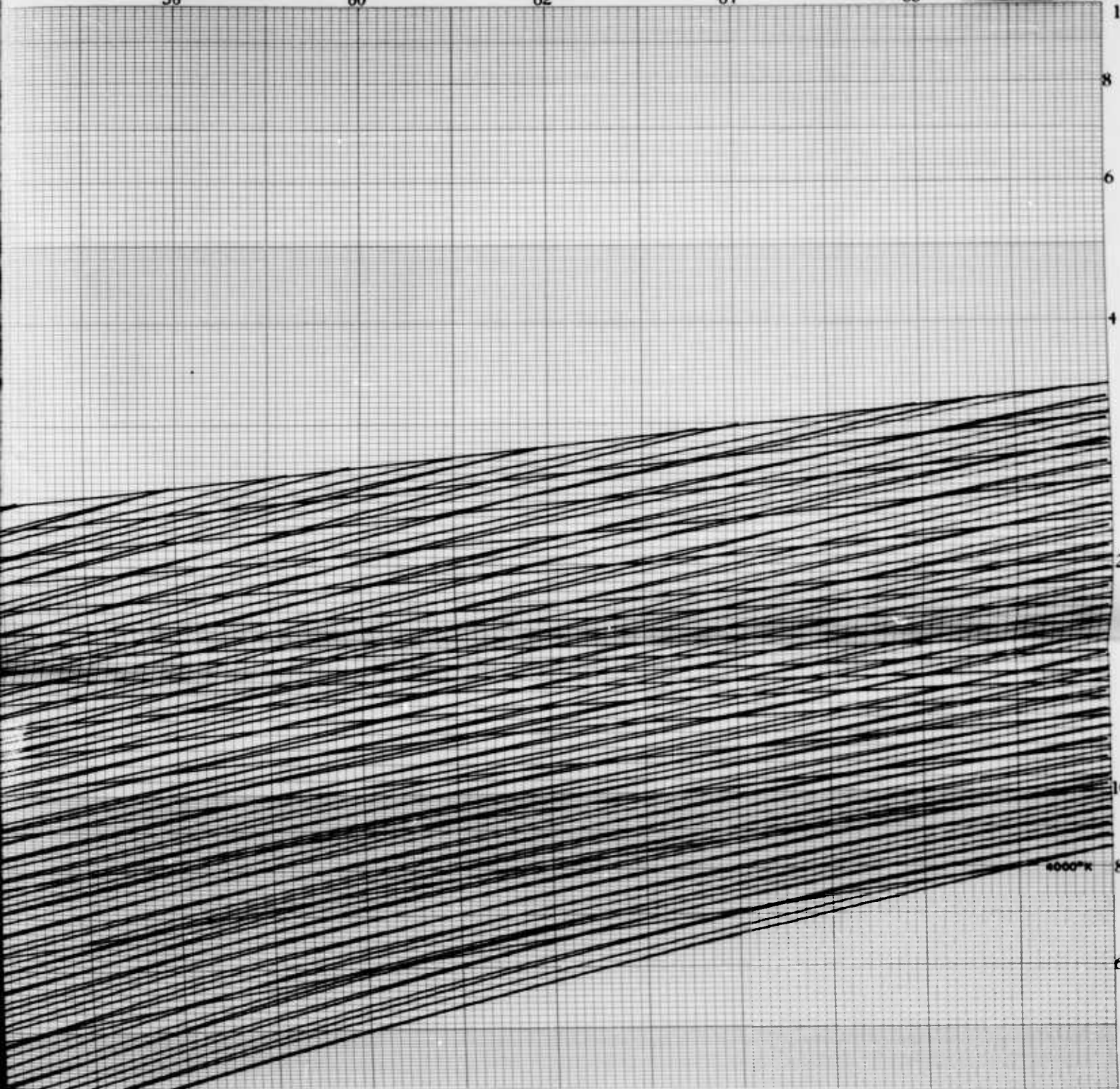
H/R [K]

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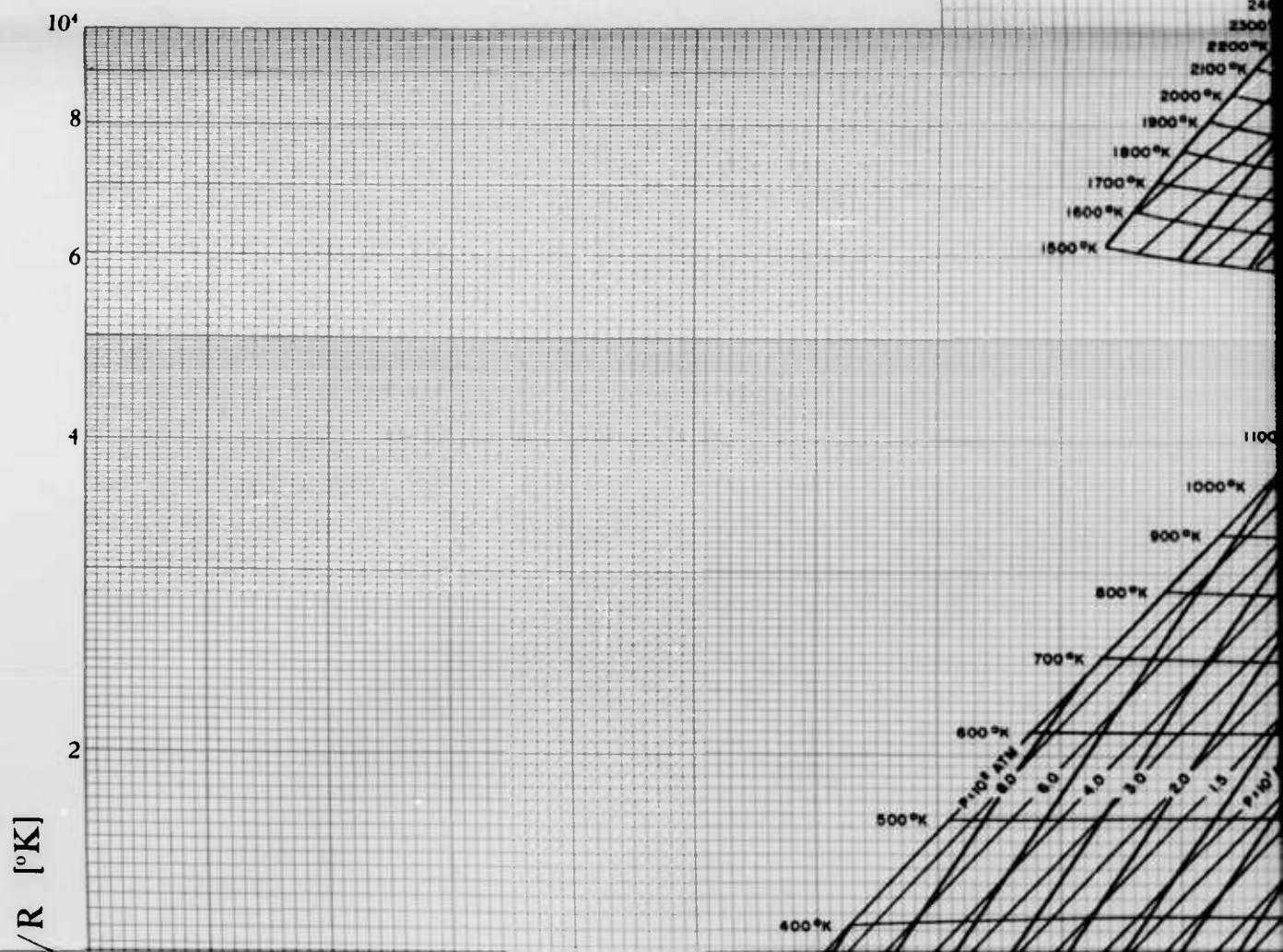
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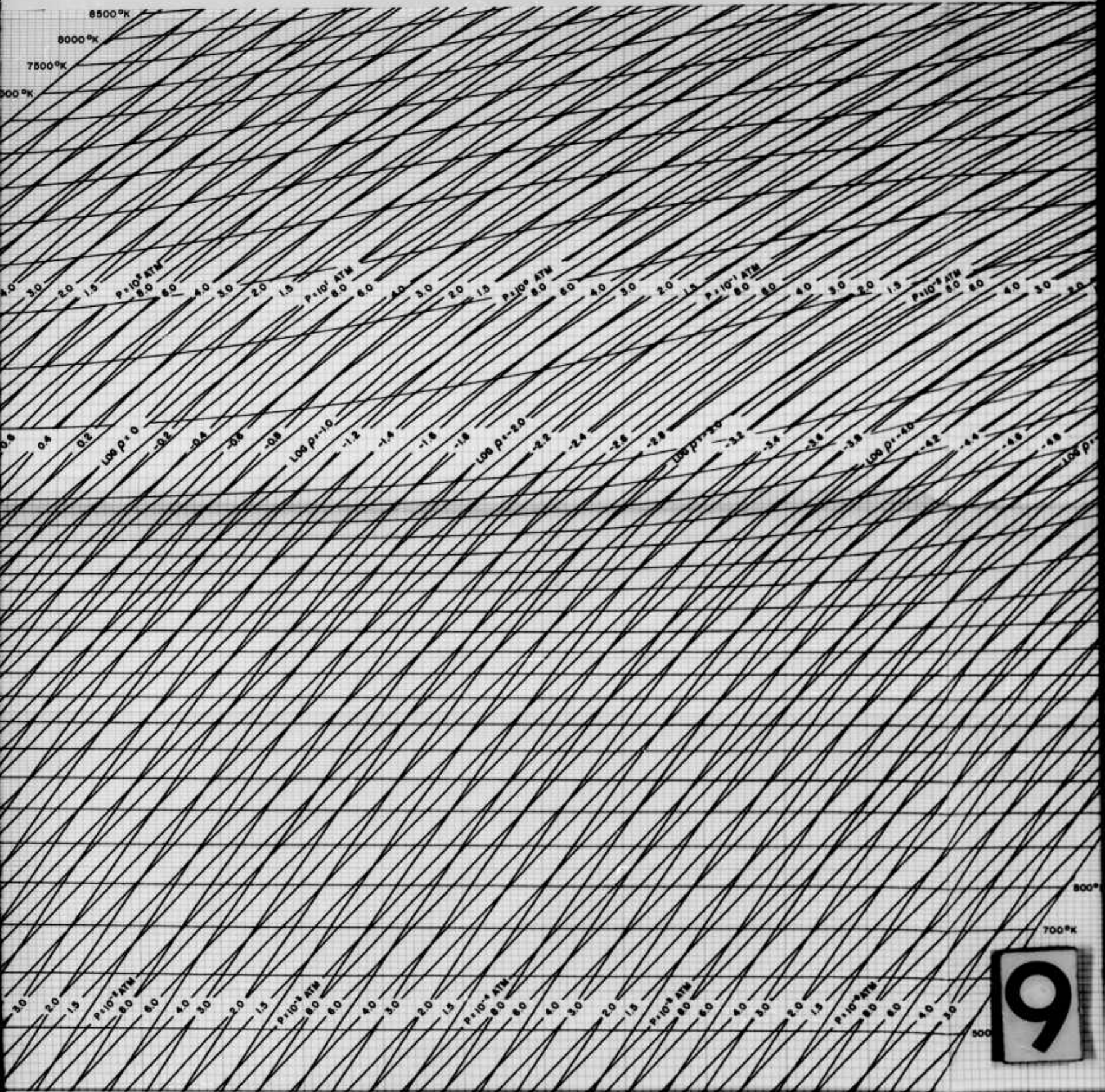


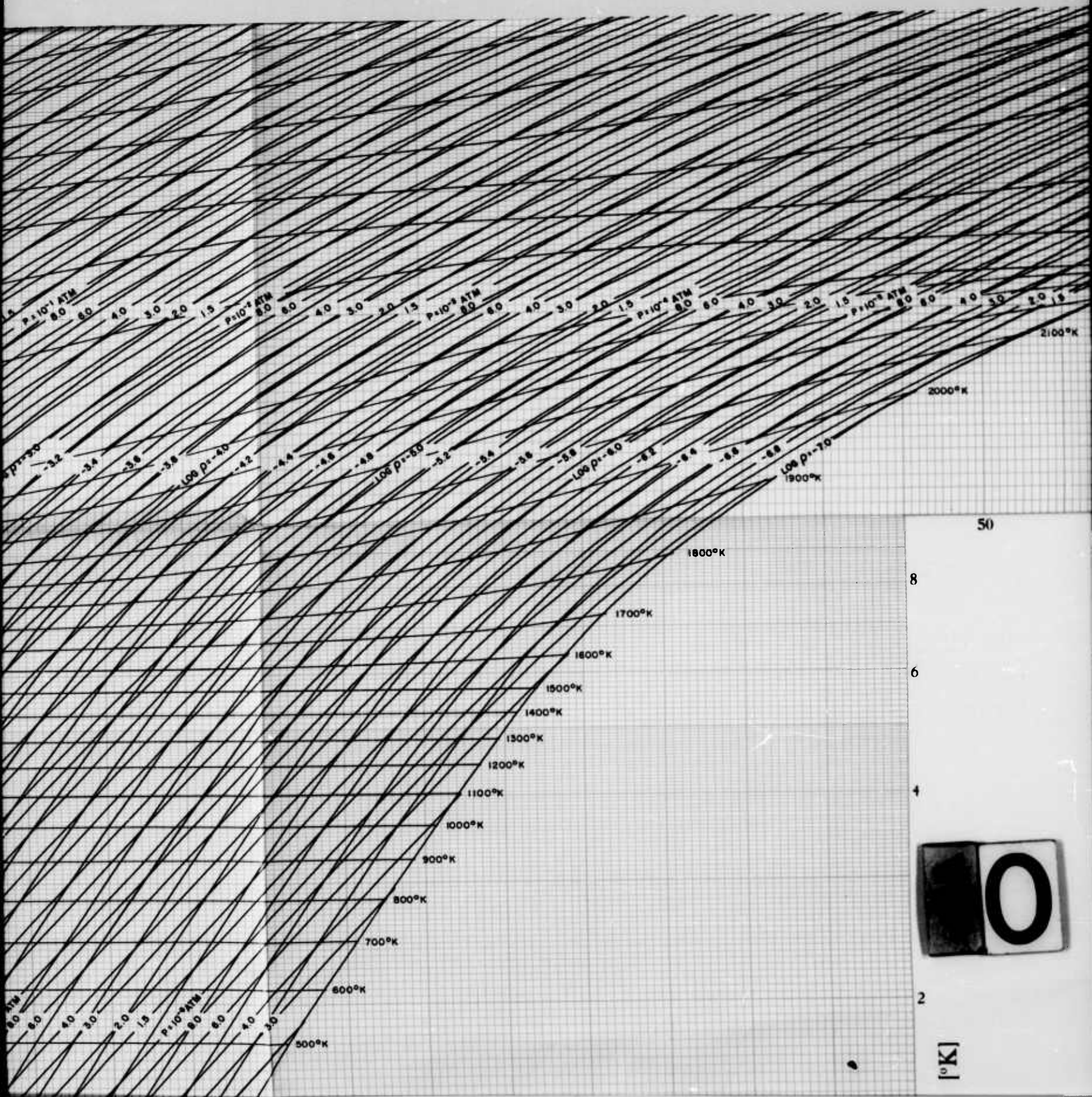
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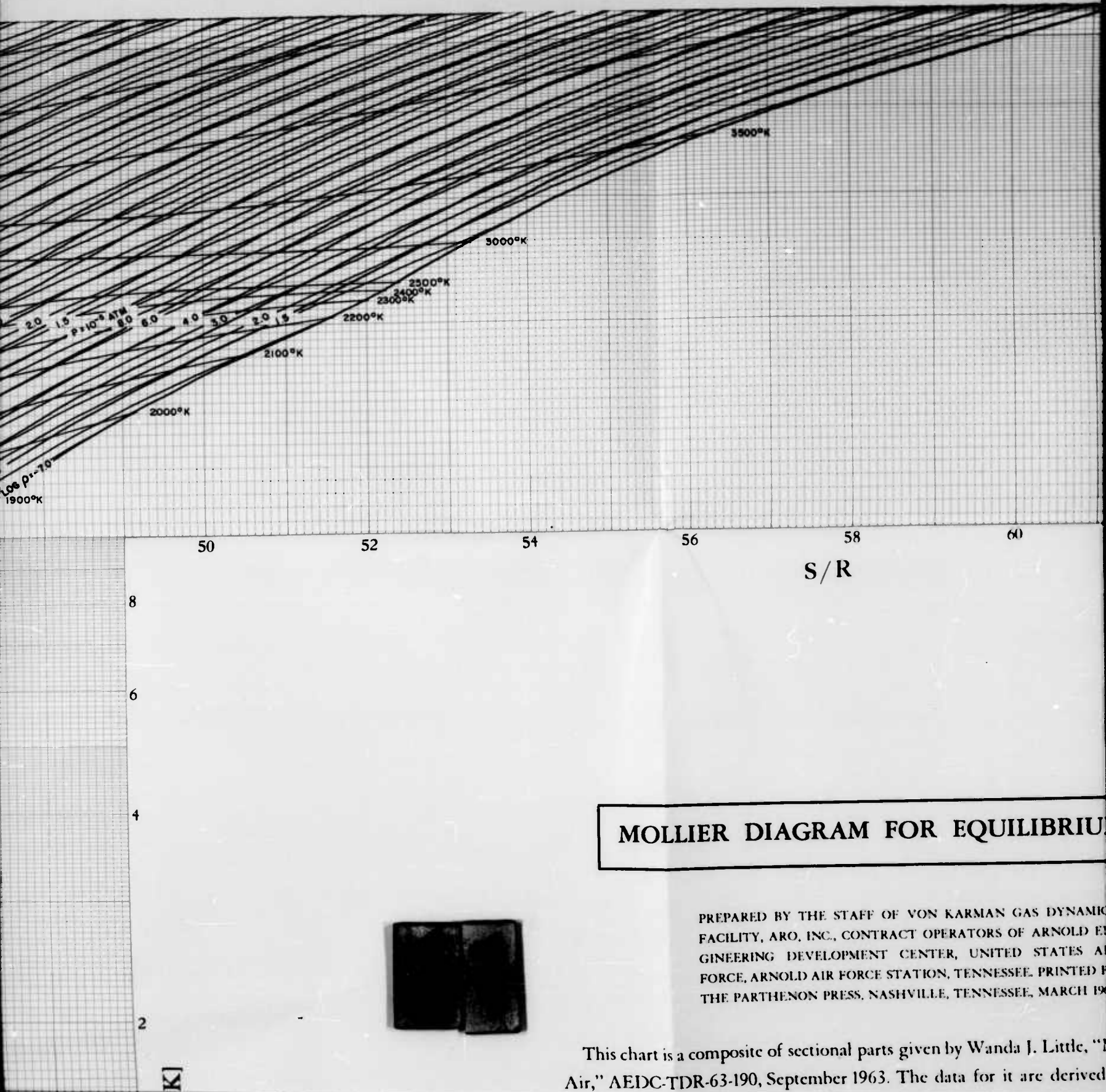
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[°K]



MOLLIER DIAGRAM FOR EQUILIBRIUM

PREPARED BY THE STAFF OF VON KARMAN GAS DYNAMIC
FACILITY, ARO, INC., CONTRACT OPERATORS OF ARNOLD EN-
GINEERING DEVELOPMENT CENTER, UNITED STATES AIR
FORCE, ARNOLD AIR FORCE STATION, TENNESSEE. PRINTED BY
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This chart is a composite of sectional parts given by Wanda J. Little, "Non-
Equilibrium Air," AEDC-TDR-63-190, September 1963. The data for it are derived

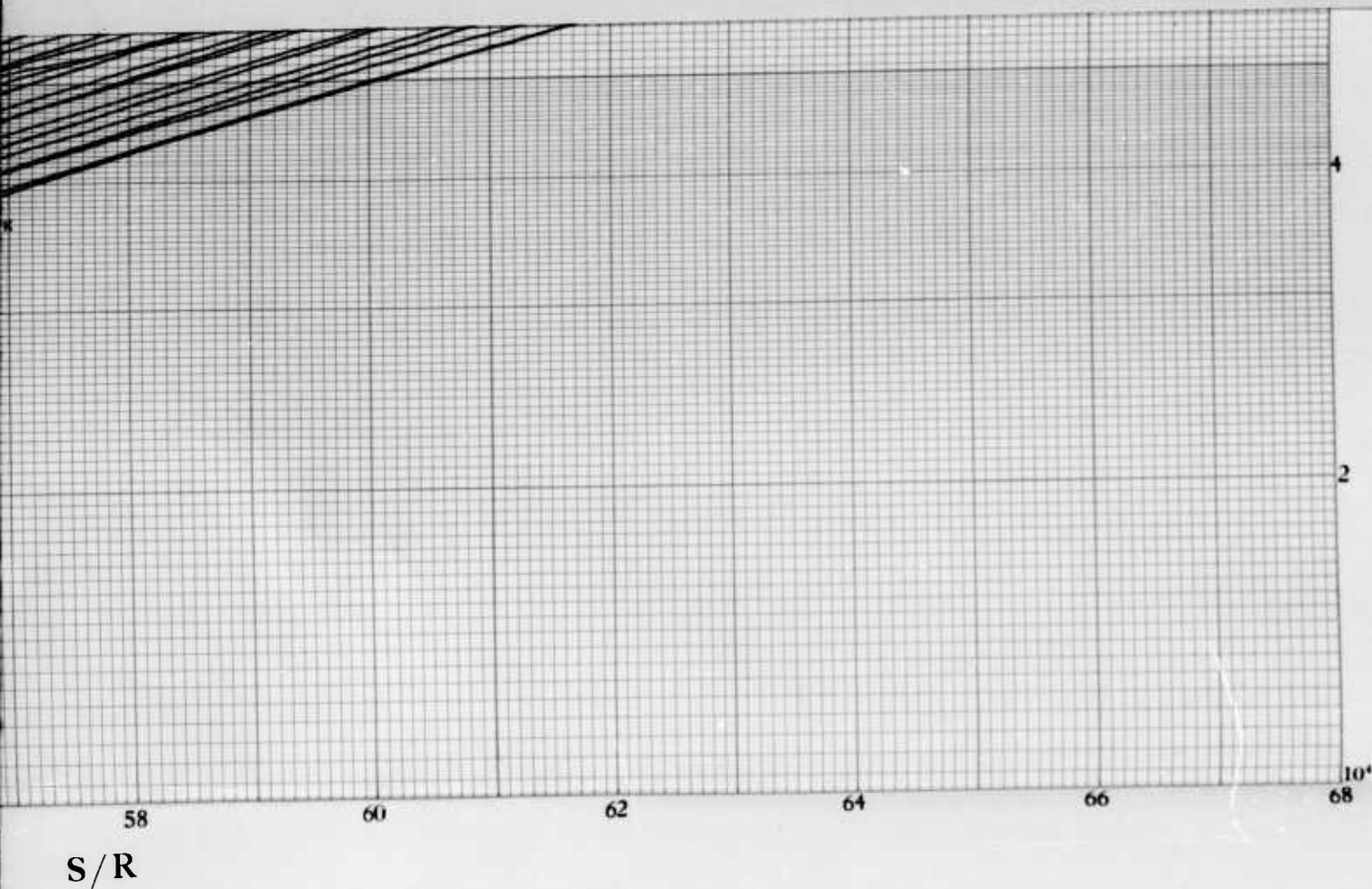
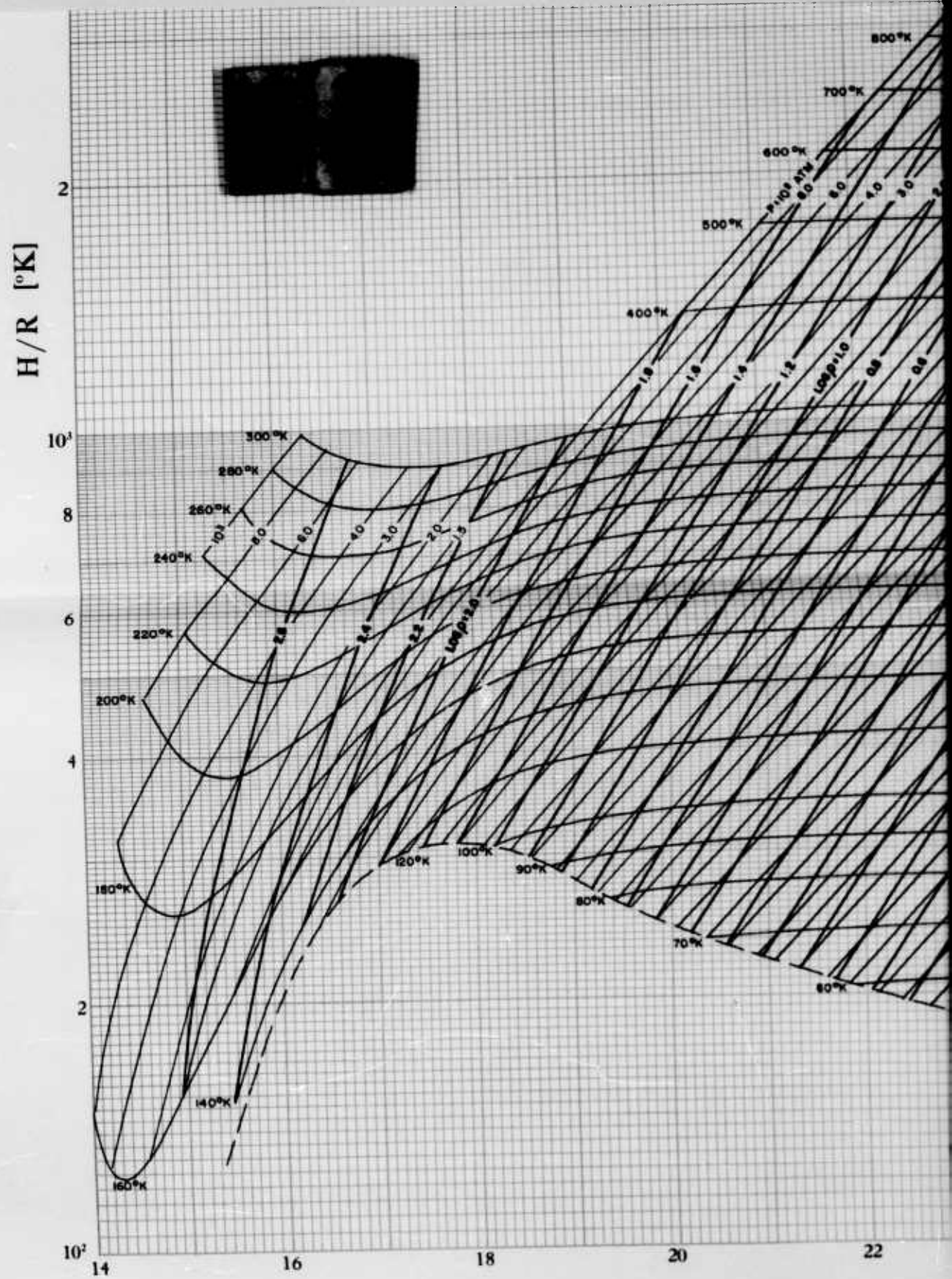
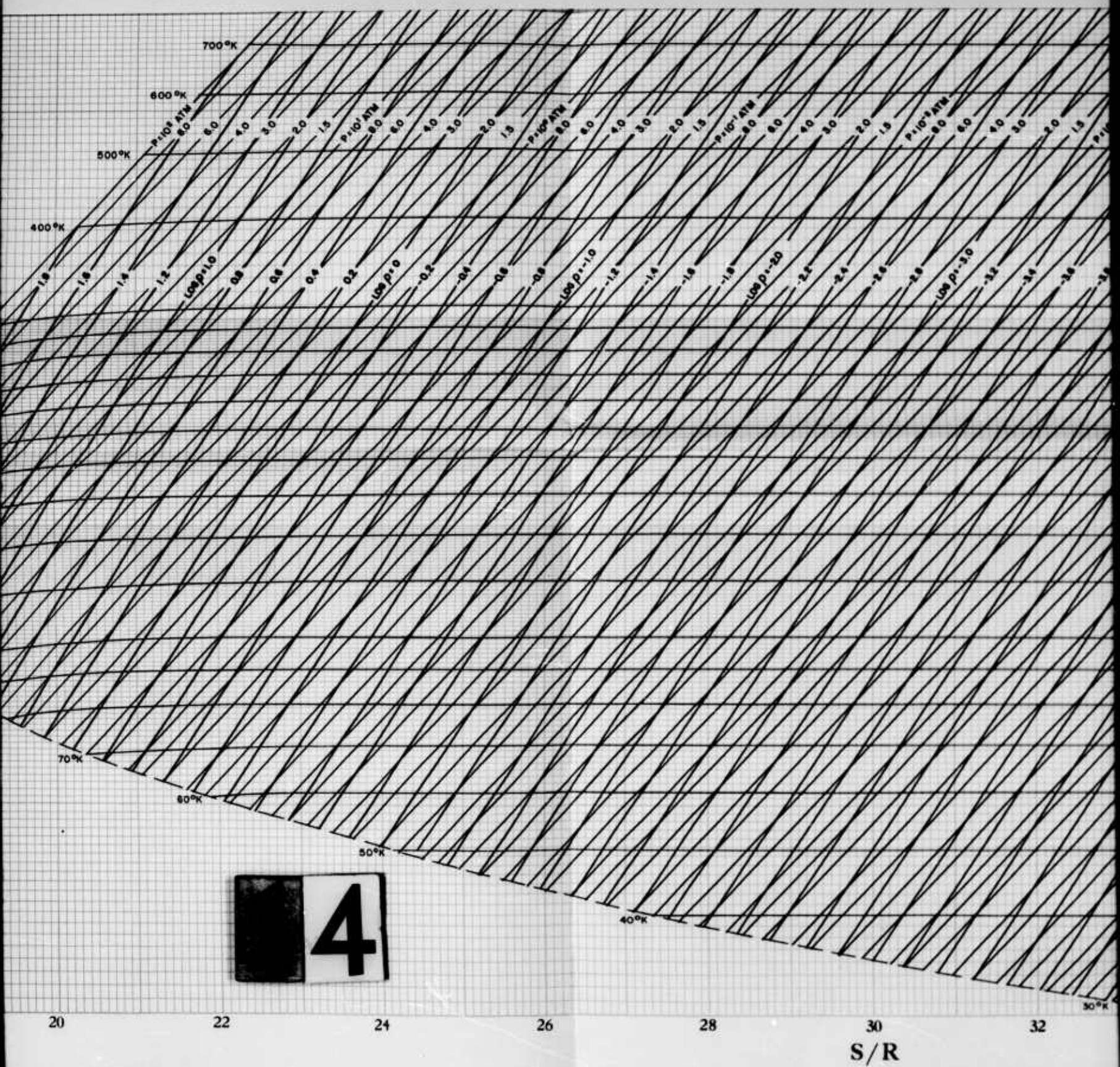


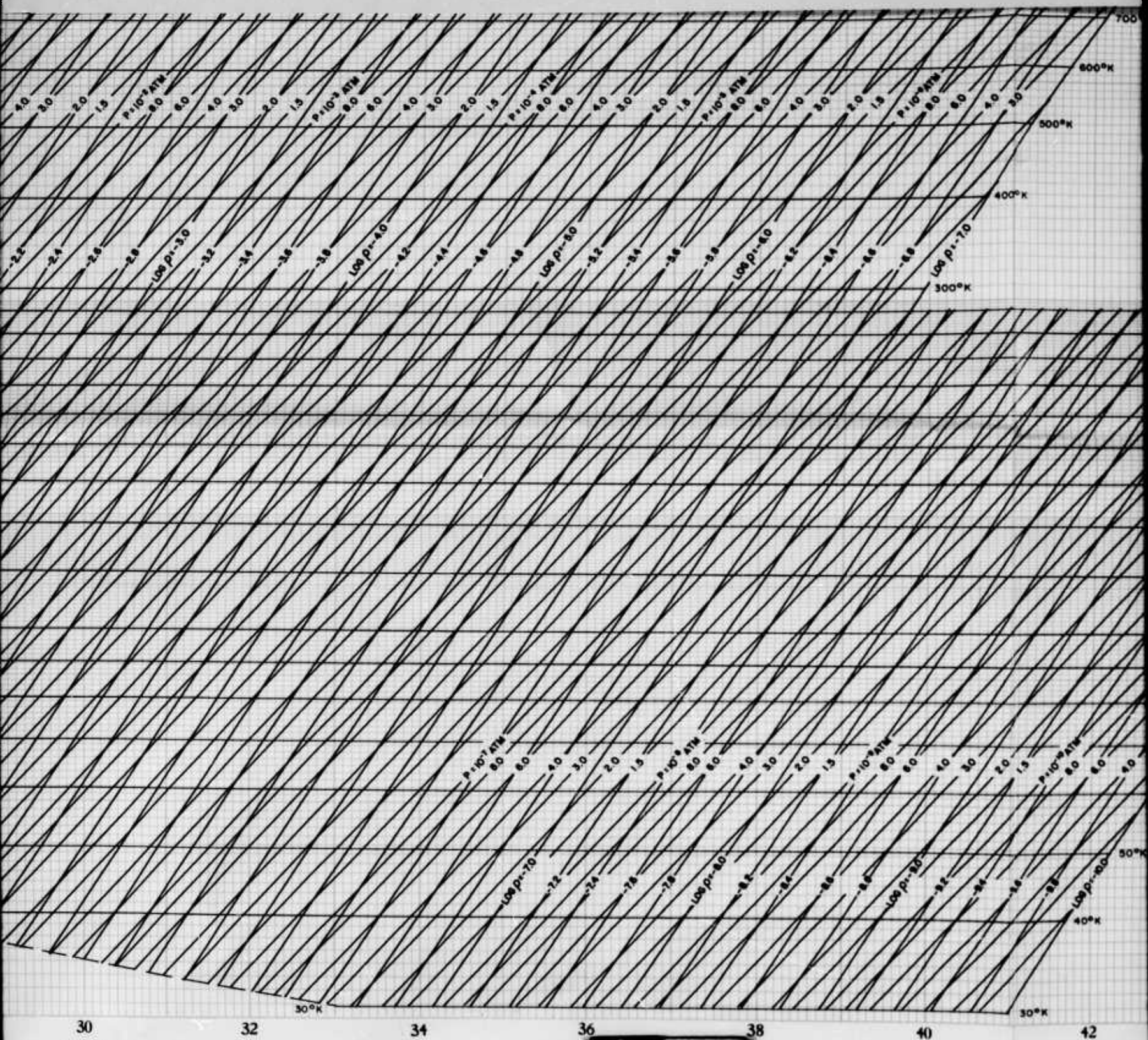
DIAGRAM FOR EQUILIBRIUM AIR

ED BY THE STAFF OF VON KARMAN GAS DYNAMICS
Y, ARO, INC., CONTRACT OPERATORS OF ARNOLD EN-
NG DEVELOPMENT CENTER, UNITED STATES AIR
ARNOLD AIR FORCE STATION, TENNESSEE. PRINTED BY
ATHENON PRESS, NASHVILLE, TENNESSEE, MARCH 1964.

e of sectional parts given by Wanda J. Little, "Mollier Diagram for
September 1963. The data for it are derived from J. Hilsenrath







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S/R

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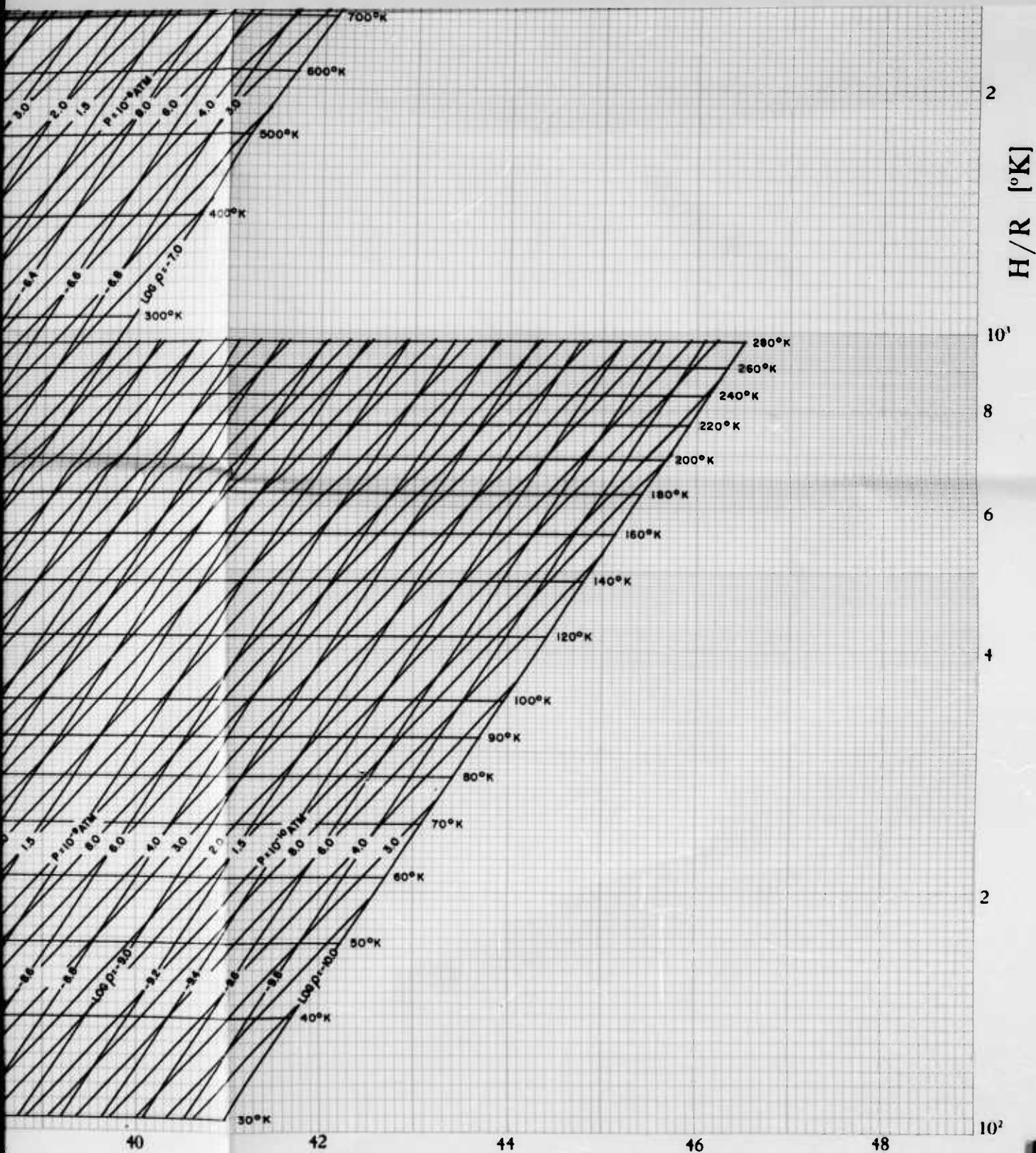
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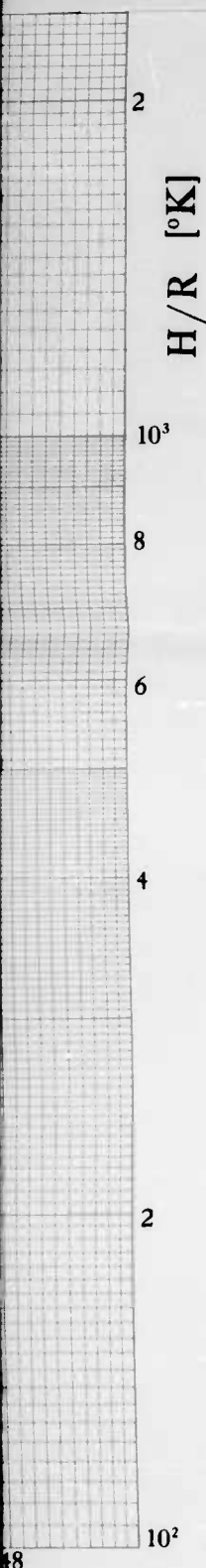
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The vertical coordinate is the enthalpy function H/R expressed in units of degrees Kelvin. Its values are obtained by multiplying the dimensionless function H/R by the temperature T . The horizontal coordinate is the dimensionless entropy function S/R . The temperature T is measured in degrees Kelvin. The pressure p is given in atmospheres. The density ρ in amagat units based on the density of air at 0° C and a pressure of one atmosphere. The logarithms are common logarithms to base ten.

For air, one amagat is the equivalent of 0.00129 grams per cubic centimeter or 0.075 pounds per cubic foot. Enthalpy in calories per gram is equal to $H/R(^{\circ}\text{K})$ multiplied by 0.0686 and in British thermal units per pound it is equal to $H/R(^{\circ}\text{K})$ multiplied by 0.00018. The specific enthalpy in feet-squared per second-squared is equal to $H/R(^{\circ}\text{K})$ multiplied by 3092.0 and in meters-squared per second-squared to $H/R(^{\circ}\text{K})$ multiplied by 0.0003092.

This chart is a composite of sectional parts given by Wanda J. Little, "Mollier Diagram for Air," AEDC-TDR-63-190, September 1963. The data for it are derived from J. Hilsenrath and M. Klein, "Tables of Thermodynamic Properties of Air in Chemical Equilibrium Including Second Virial Corrections from 1500 to 15,000°K," AEDC-TDR-63-161, August 1963, and from R. L. Humphrey and C. A. Neel, "Tables of Thermodynamic Properties of Air from 90 to 1500°K," AEDC-TR-61-103, August 1961. The latter tables are based mainly on J. Hilsenrath, C. W. Beckett, et al., *Tables of Thermal Properties of Gases*, NBS Circular 564, November 1955, and on F. Din, *Thermodynamic Functions of Gases*, Vol. 2, Butterworths, London, 1956. The tables of Din also provide data for the saturation line at pressures above one atmosphere. At pressures below one atmosphere, the approximate saturation line is based on G. T. Furukawa and R. E. McCoskey, "The Condensation Line of Air and the Heats of Vaporization of Oxygen and Nitrogen," NBS Technical Report 1775, August 1952, and on W. D. Erickson and H. E. Creekmore, "Study of Equilibrium Real-Gas Effects in Hypersonic Air Nozzles, Including Charts of Thermodynamic Properties of Equilibrium Air," NASA-TN-D-231, April 1960.

The vertical coordinate is the enthalpy function H/R expressed in units of degrees Kelvin. Its values are obtained by multiplying the dimensionless function H/RT by the temperature. The horizontal coordinate is the dimensionless entropy function S/R . The temperature T is measured in degrees Kelvin. The pressure p is given in atmospheres, and the density ρ in amagat units based on the density of air at 0° C and a pressure of one atmosphere. The logarithms are common logarithms to base ten.

For air, one amagat is the equivalent of 0.00129 grams per cubic centimeter or 0.0807 pounds per cubic foot. Enthalpy in calories per gram is equal to $H/R(^{\circ}\text{K})$ multiplied by 0.0686 and in British thermal units per pound it is equal to $H/R(^{\circ}\text{K})$ multiplied by 0.1234. The specific enthalpy in feet-squared per second-squared is equal to $H/R(^{\circ}\text{K})$ multiplied by 3092.0 and in meters-squared per second-squared to $H/R(^{\circ}\text{K})$ multiplied by 287.2.

AEDC-TDR-63-190

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Fig. 1



MOLLIER DIAGRAM FOR AIR

By

**Wanda J. Little
von Kármán Gas Dynamics Facility
ARO, Inc.**

TECHNICAL DOCUMENTARY REPORT NO. AEDC-TDR-63-190

September 1963

Program Element 61405014/8951, Task 89603

**(Prepared under Contract No. AF 40(600)-1000 by ARO, Inc.,
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ABSTRACT

A Mollier diagram for air, including the effects of dissociation and ionization, as well as of intermolecular potentials to the second virial correction, is presented. The range of temperatures extends from the saturation line to 15,000°K and the range of densities, from 10^{-7} to approximately 200 amagats.

PUBLICATION REVIEW

This report has been reviewed and publication is approved.


Larry D. Fitzgerald
Capt, USAF
Aerospace Sciences Division
DCS/Research


Donald R. Eastman, Jr.
DCS/Research

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NOMENCLATURE

H/R	Enthalpy in °K, obtained by multiplying the dimensionless function H/RT by the temperature in °K
$\log$	Common logarithm to base 10
p	Pressure, atm
R	Gas constant
S/R	Entropy, dimensionless
T	Temperature, °K
ρ/ρ_0	Density in amagats, based on ρ_0 at 0°C and one atm of pressure

1.0 INTRODUCTION

This air Mollier diagram was prepared for the purposes of data reduction in the operation of wind tunnels at the Arnold Engineering Development Center (AEDC), Air Force Systems Command (AFSC), USAF. This diagram represents the most recently available information on the thermodynamic properties of air, including the effects of dissociation, ionization, and of intermolecular potentials to the second virial correction. The diagram covers the range of temperatures from the saturation line to 15,000°K and the range of densities from 10^{-7} to from 200 to 250 amagats. Above 1500°K, the plotted data are taken from the tables of Hilsenrath and Klein (Ref. 1), whereas the data below 1500°K are derived from the tables of Humphrey and Neel (Ref. 2) which combine information from various sources, notably the tables of Din (Ref. 3). Data below 90°K, as well as all data at densities below 10^{-7} amagat, were generated mainly by extrapolation, using the method described in Ref. 4. The tables of Din (Ref. 3) provide explicit data for the saturation line at pressures from one atmosphere upwards. At pressures below one atmosphere, the saturation line is based on the work of Furukawa and McCoskey (Ref. 5) and of Erickson and Creekmore (Ref. 6), the latter being based on vapor-pressure data for pure nitrogen and oxygen.

Enthalpy, expressed as H/R having dimensions of degrees Kelvin, is plotted vertically on a logarithmic scale through four decimal orders of magnitude. The dimensionless entropy, S/R , is the other basic variable. Lines at constant density are identified by their values of $\log \rho/\rho_0$, in which the ratio ρ/ρ_0 is the density in amagats. The constant pressure lines are marked in atmospheres and the constant temperature in degrees Kelvin.

The appendix contains a short list of useful conversion factors.

REFERENCES

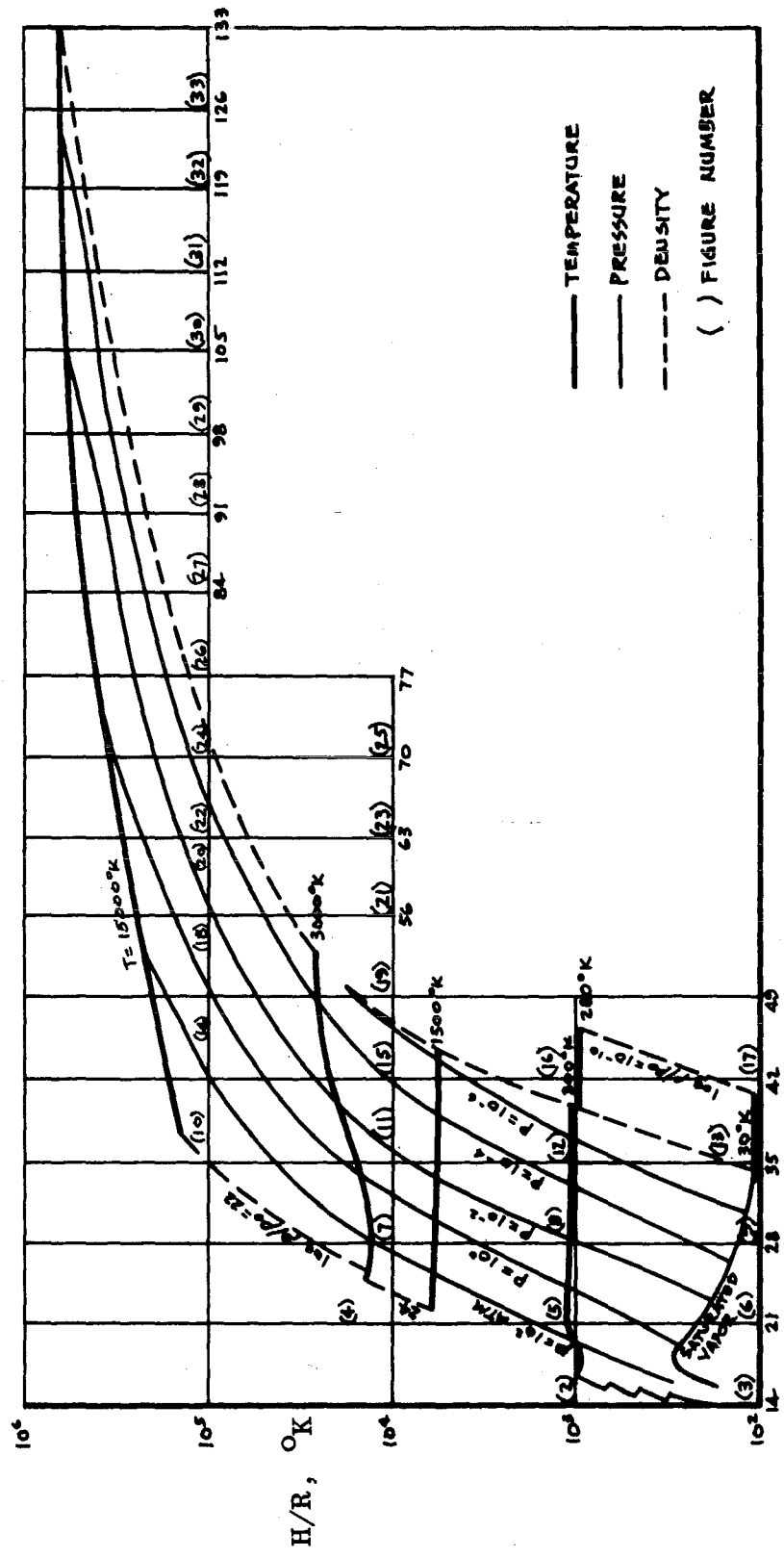
1. Hilsenrath, J., and Klein, M. "Tables of Equilibrium Properties of Air Including Second Virial Corrections." AEDC-TDR-63-161. September 1963.
2. Humphrey, R. L., and Neel, C. A. "Tables of the Thermodynamic Properties of Air from 90 to 1500°K." AEDC-TN-61-103, August 1961.

Manuscript received August 1963.

3. Din, F. Thermodynamic Functions of Gases. Vol. 1, Butterworths Scientific Publications, London, 1961.
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5. Furukawa, G. T., and McCoskey, R. E. "The Condensation Line of Air and the Heats of Vaporization of Oxygen and Nitrogen." National Bureau of Standards Technical Report 1775, August 1, 1952.
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APPENDIX

Conversion Factors To Change	To	With Units of	Multiply by
$\frac{H}{R}$	$\frac{H}{RT}$	None	$\frac{1}{T(^{\circ}K)}$
	$\frac{H}{R}$	$^{\circ}R$	1.8
	H	$\frac{ft^2}{sec^2}$	$3.09235(10)^3$
		$\frac{Btu}{lb}$	$1.23406(10)^{-1}$
		$\frac{cal}{gm}$	$6.86042(10)^{-2}$
$\frac{S}{R}$	S	$\frac{ft^2}{sec^2^{\circ}K}$	$3.09235(10)^3$
		$\frac{ft^2}{sec^2^{\circ}R}$	$1.71797(10)^3$
		$\frac{Btu}{lb^{\circ}R}$	$6.85590(10)^{-2}$
		$\frac{Btu}{lb^{\circ}K}$	$1.23406(10)^{-1}$
		$\frac{cal}{gm^{\circ}K}$	$6.86402(10)^{-2}$
$\frac{\rho}{\rho_0}$	ρ	$\frac{gm}{cm^3}$	$1.29304(10)^{-3}$
		$\frac{lb}{in.^3}$	$4.67143(10)^{-5}$
		$\frac{lb}{ft^3}$	$8.07223(10)^{-2}$



S/R

Fig. 1 Index Sheet

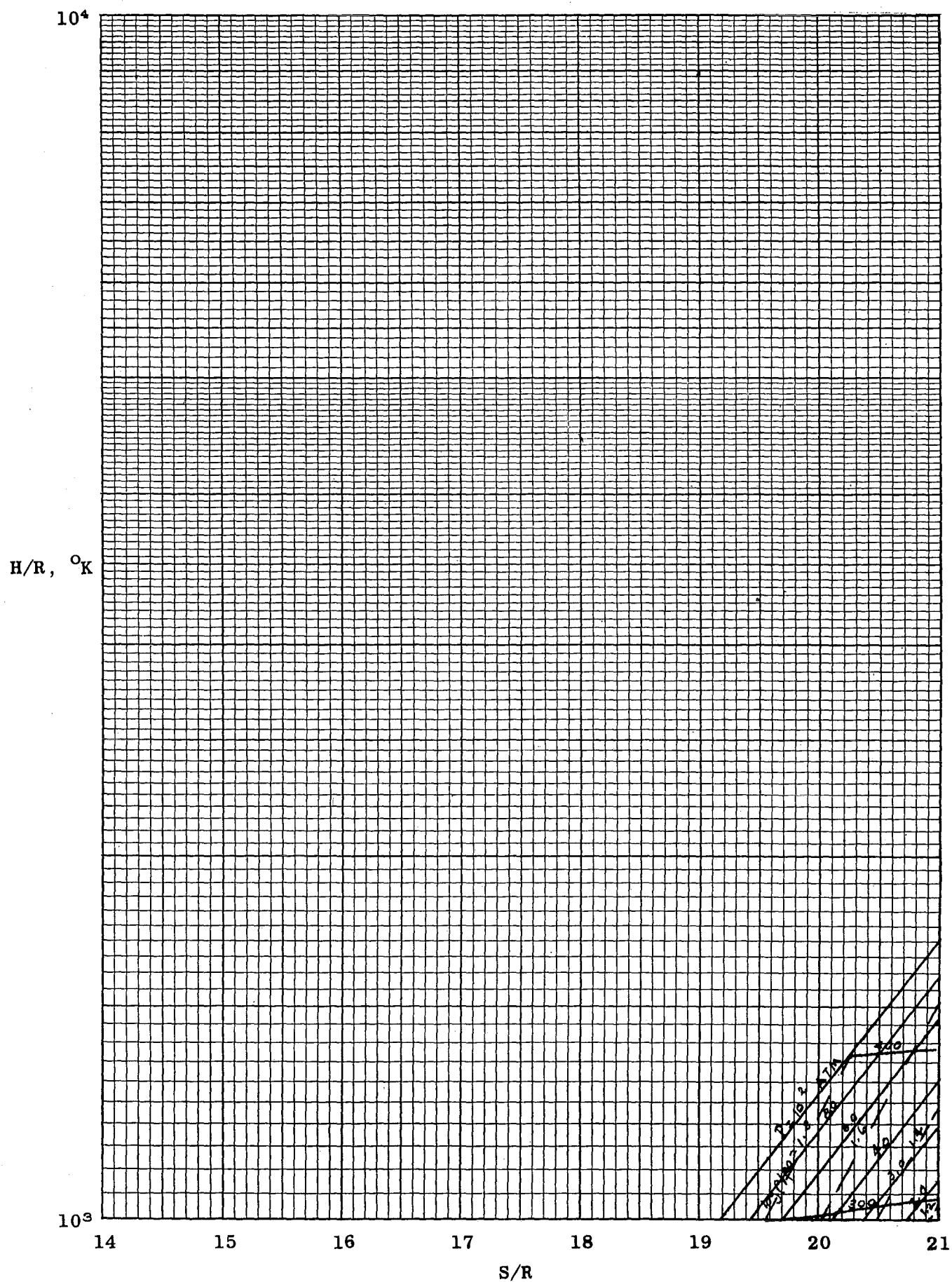


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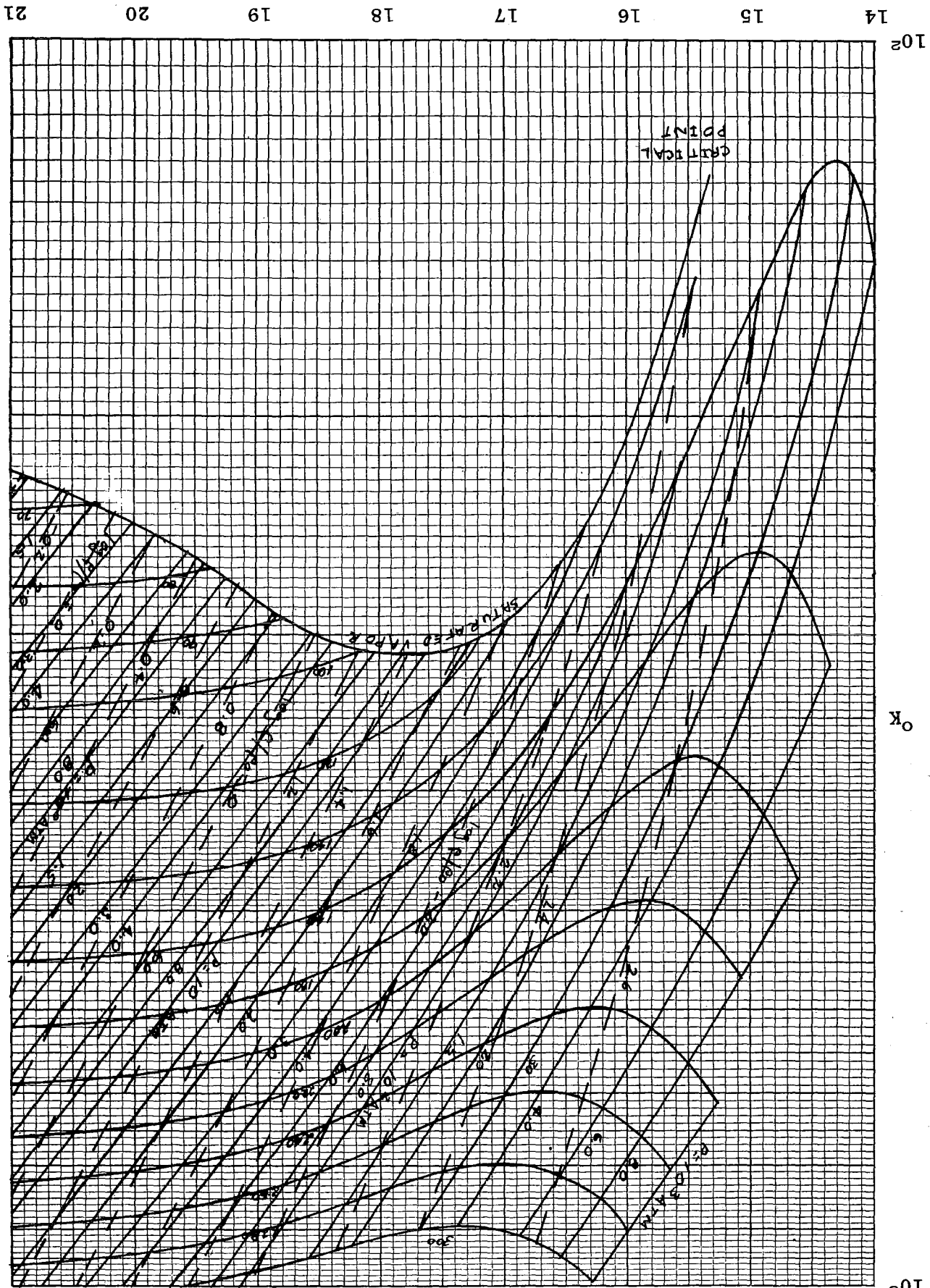


Fig. 3 Air Mollier Diagram

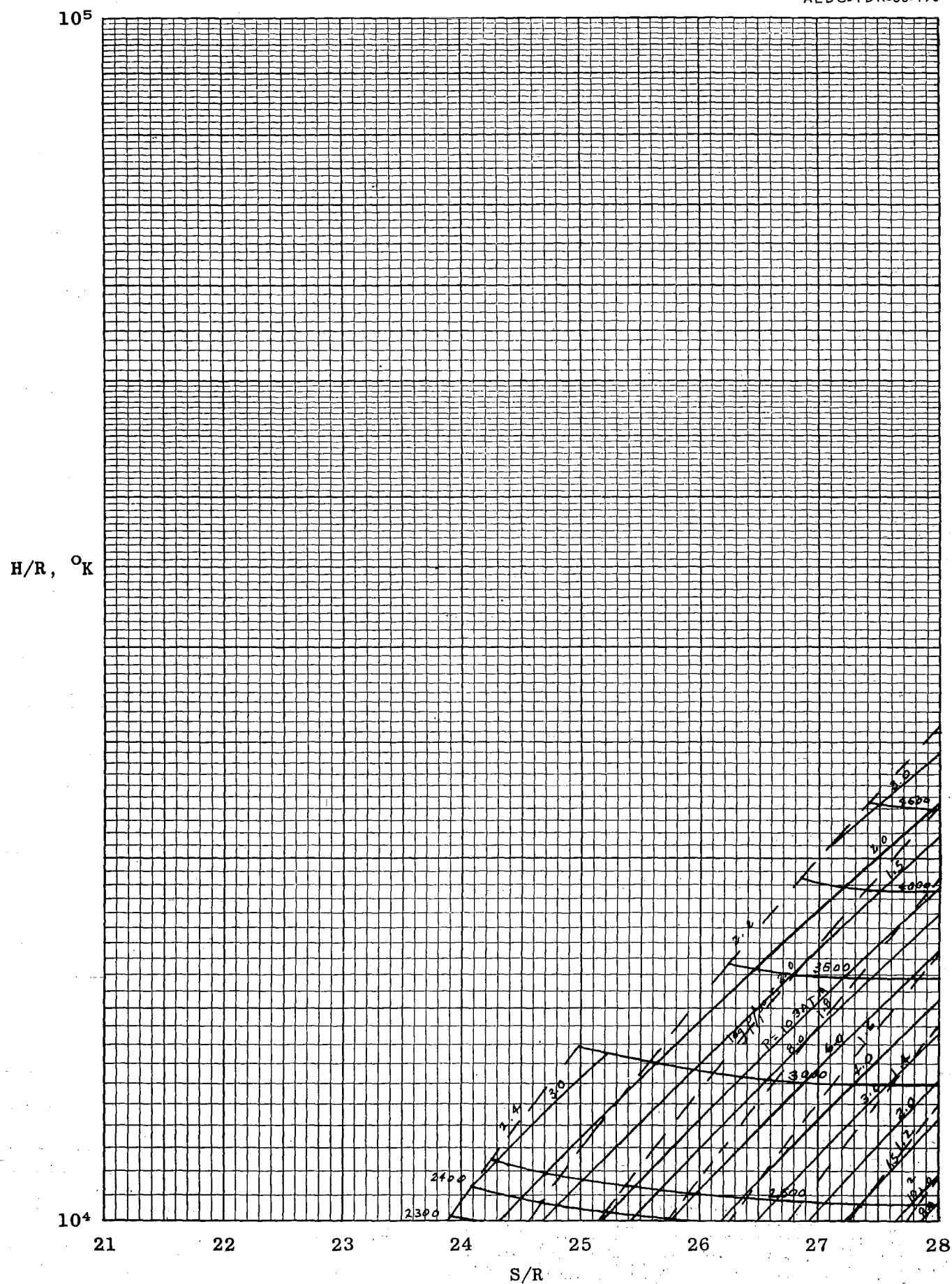


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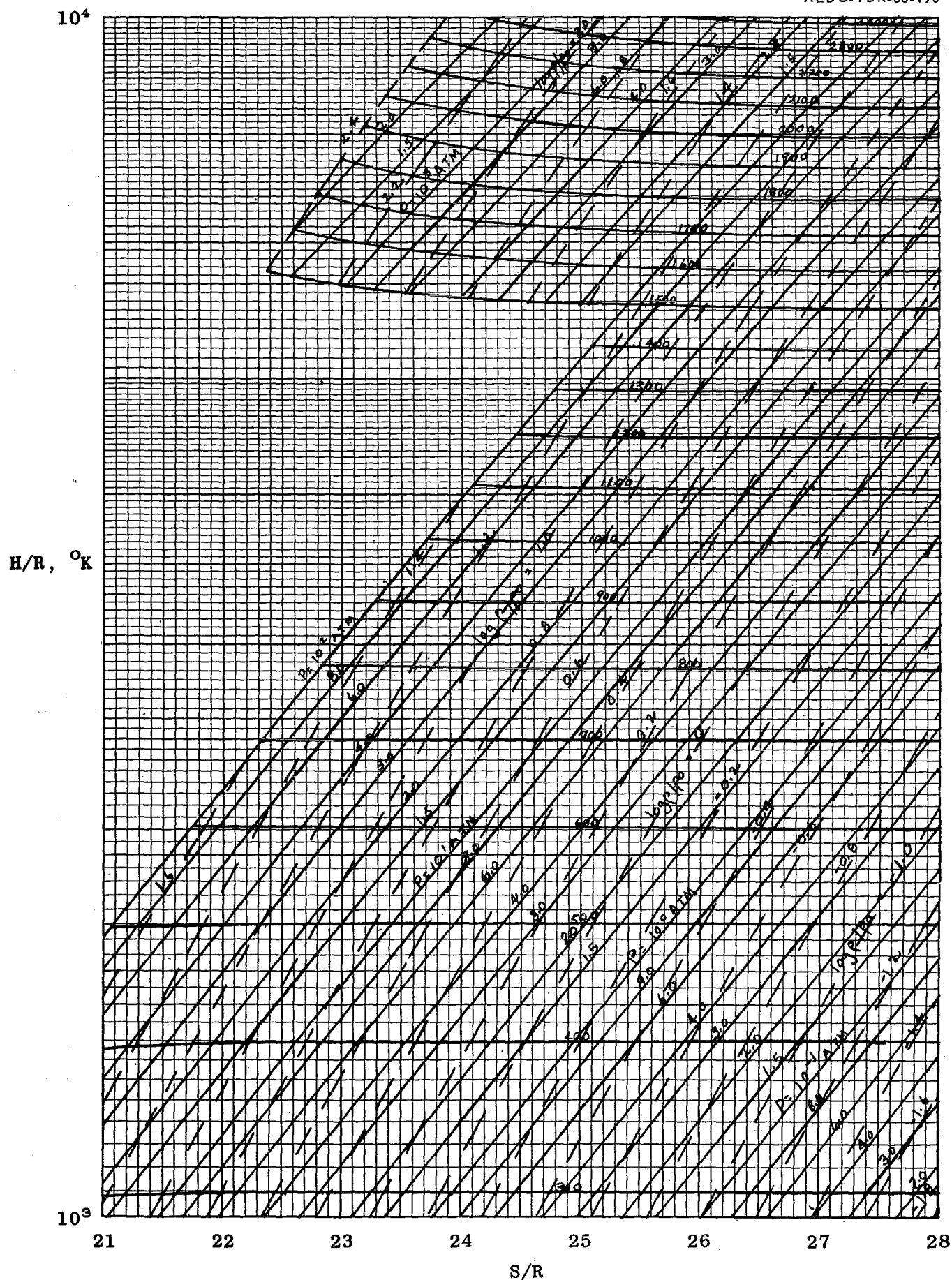


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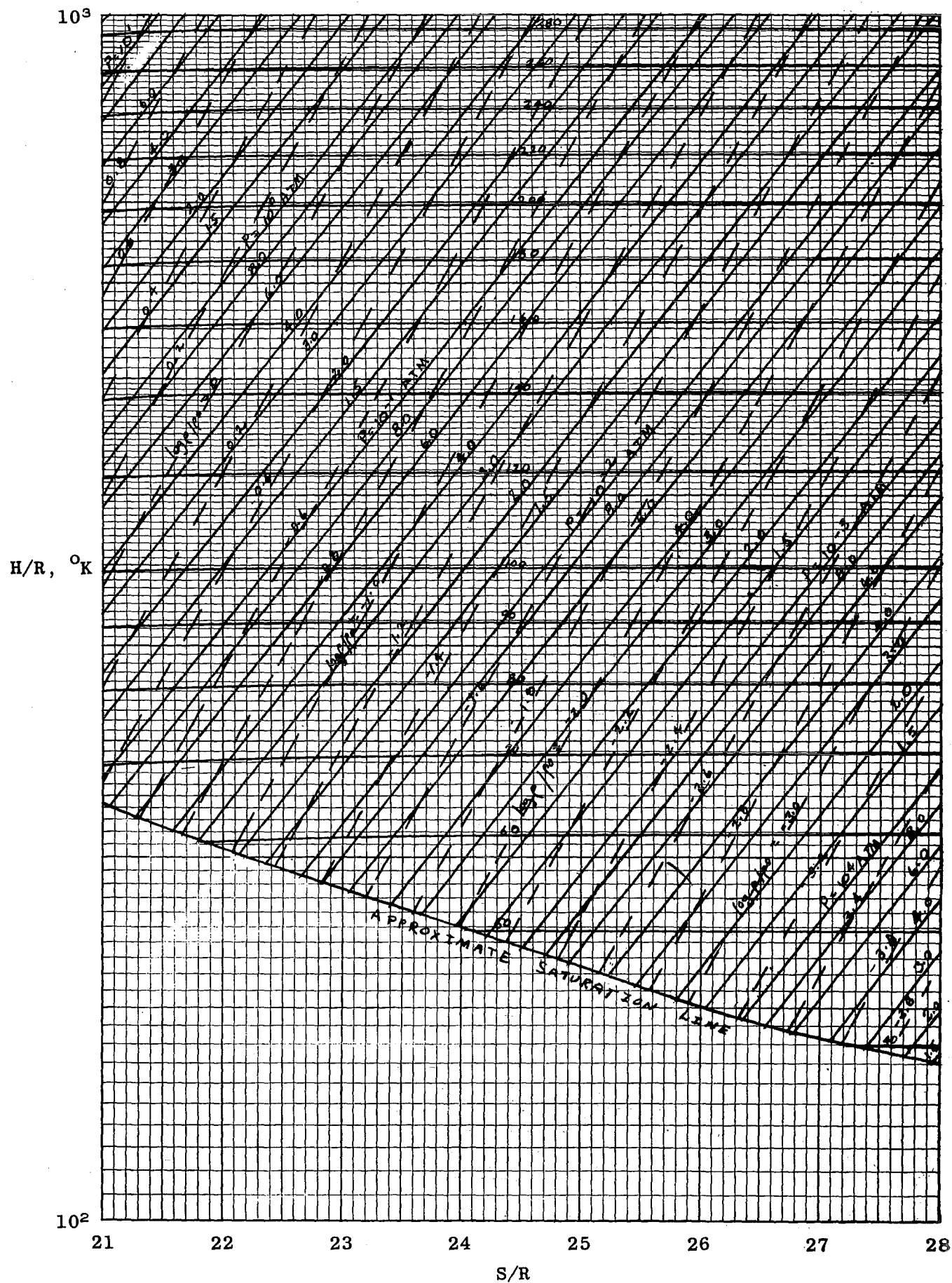


Fig. 6 Air Mollier Diagram

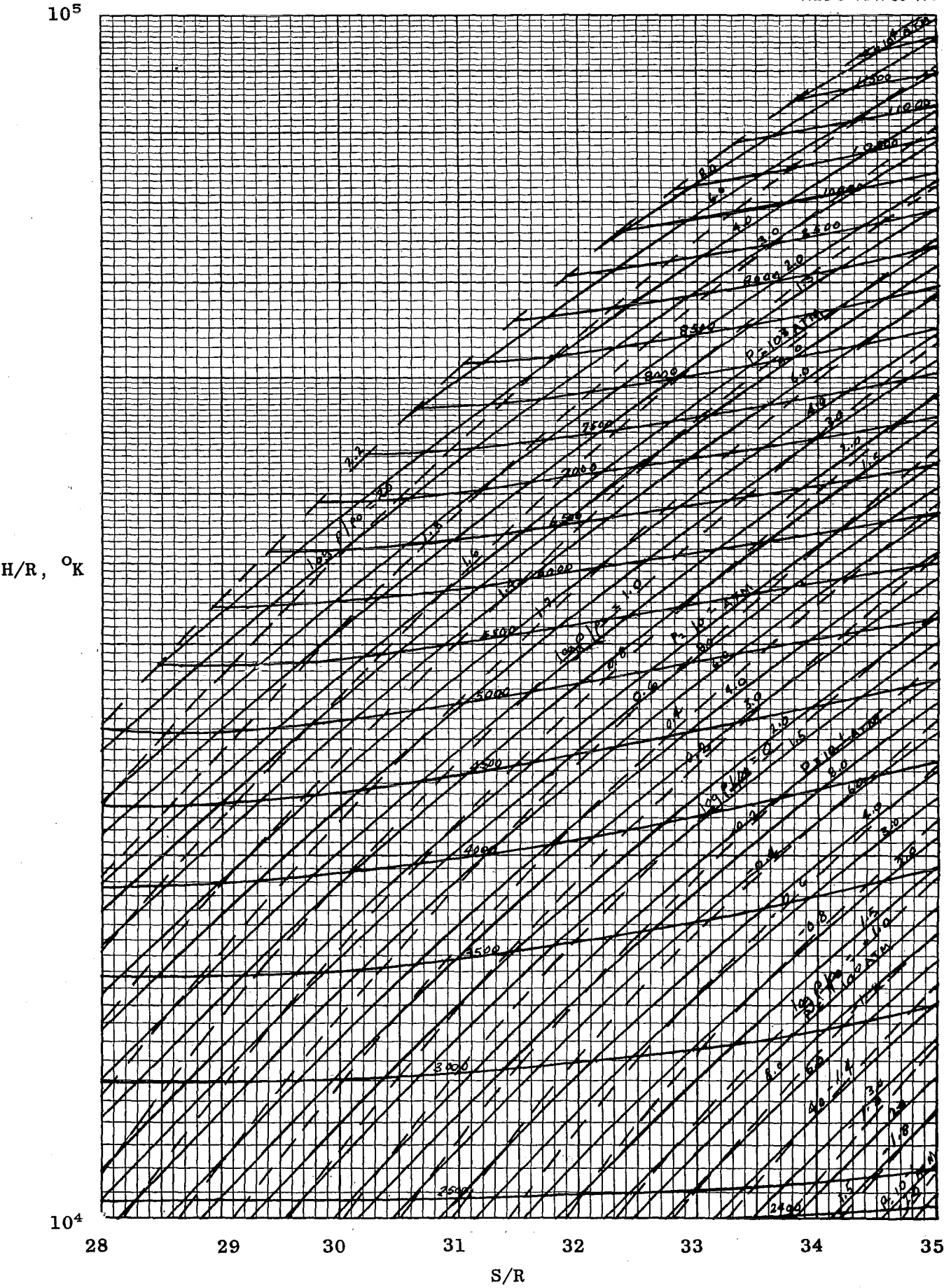


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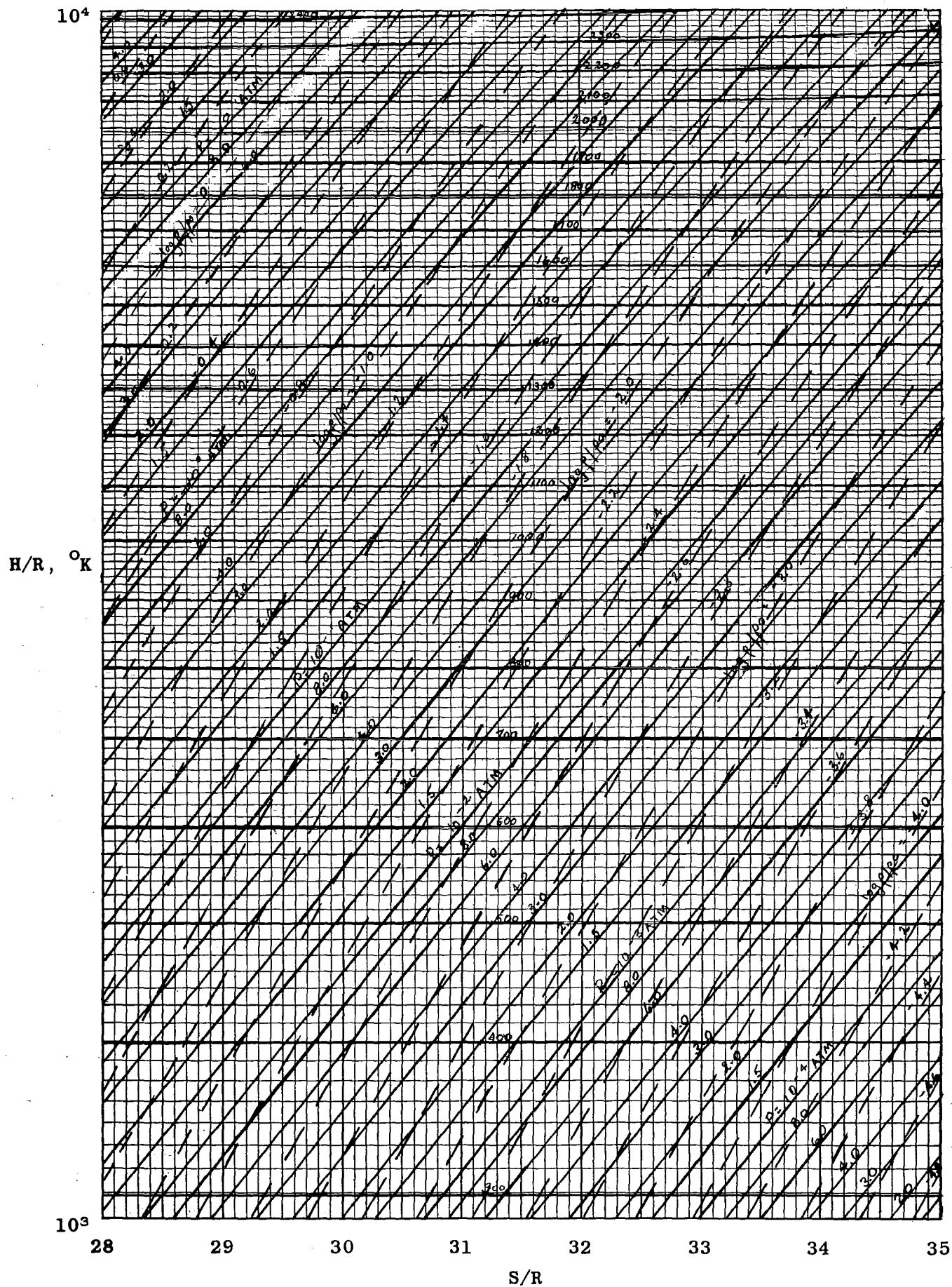


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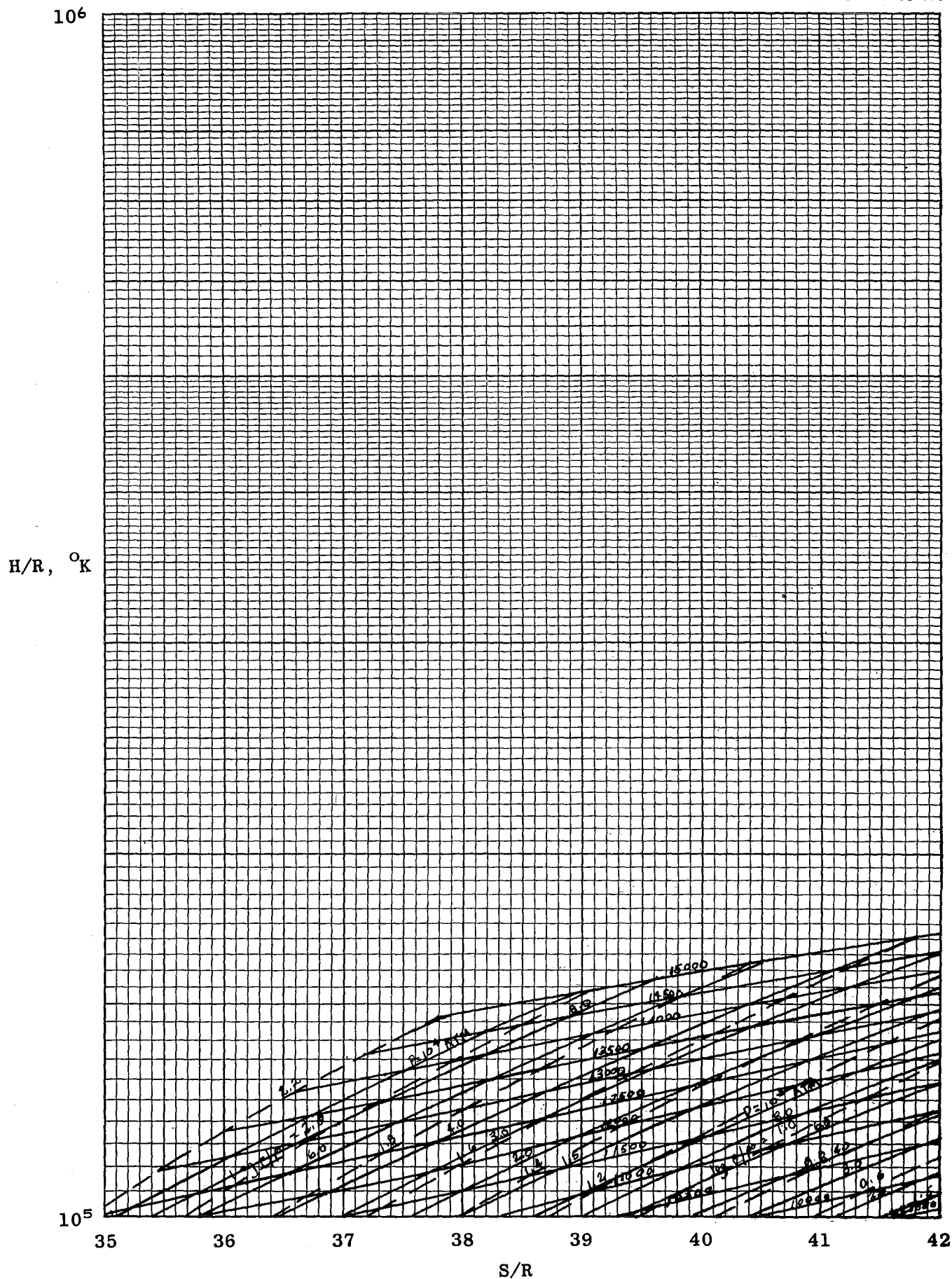


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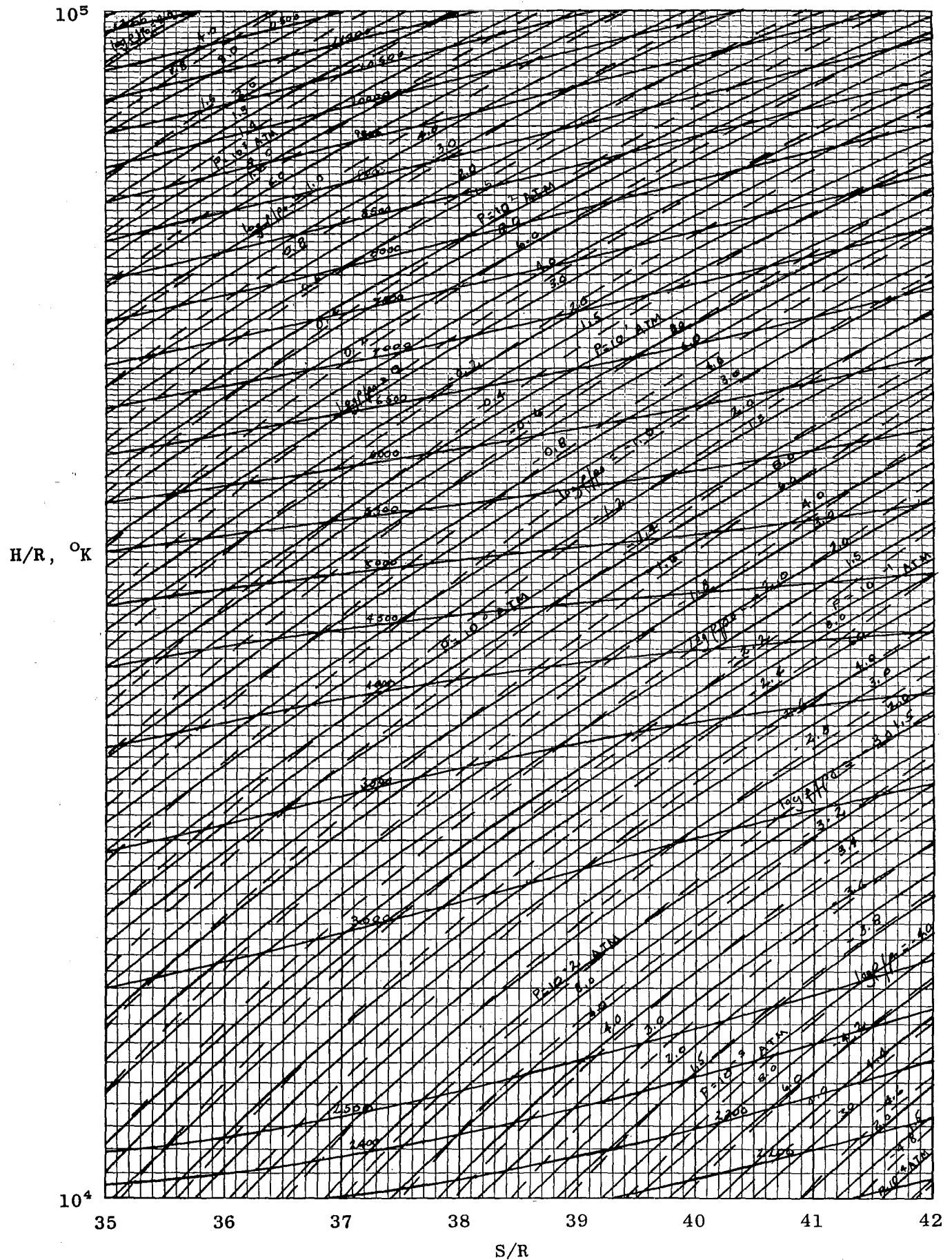


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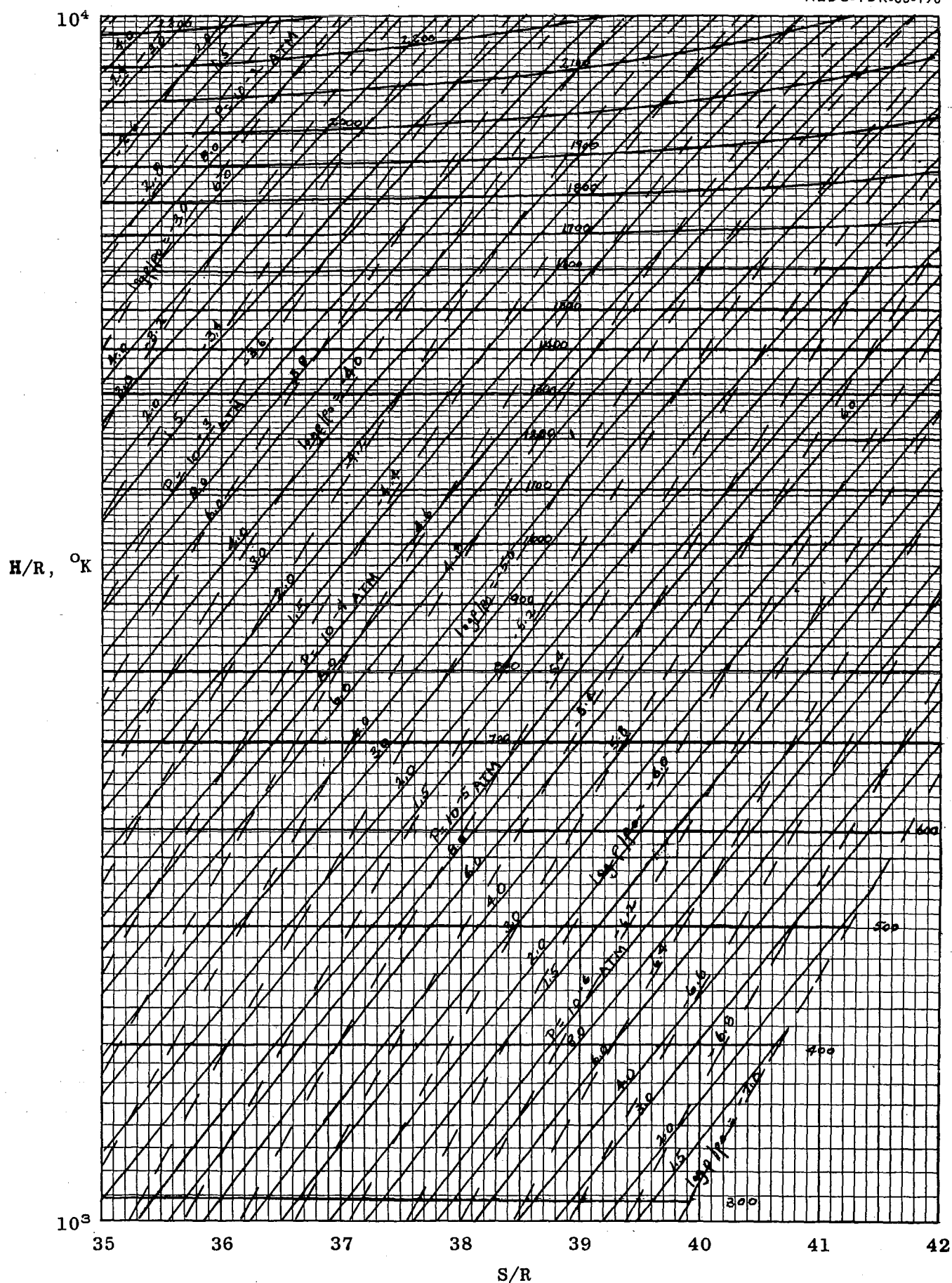


Fig. 12 Air Mollier Diagram

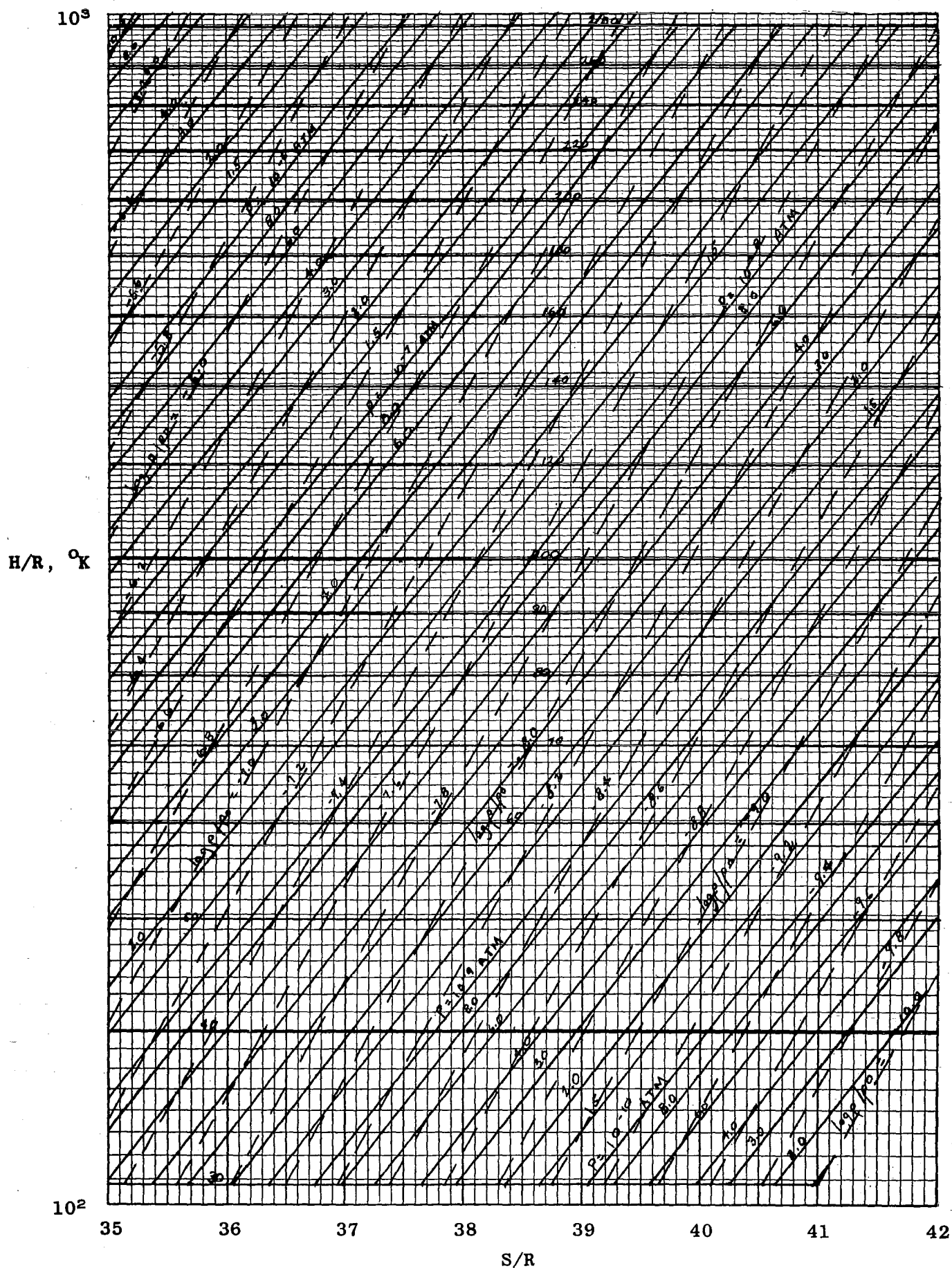


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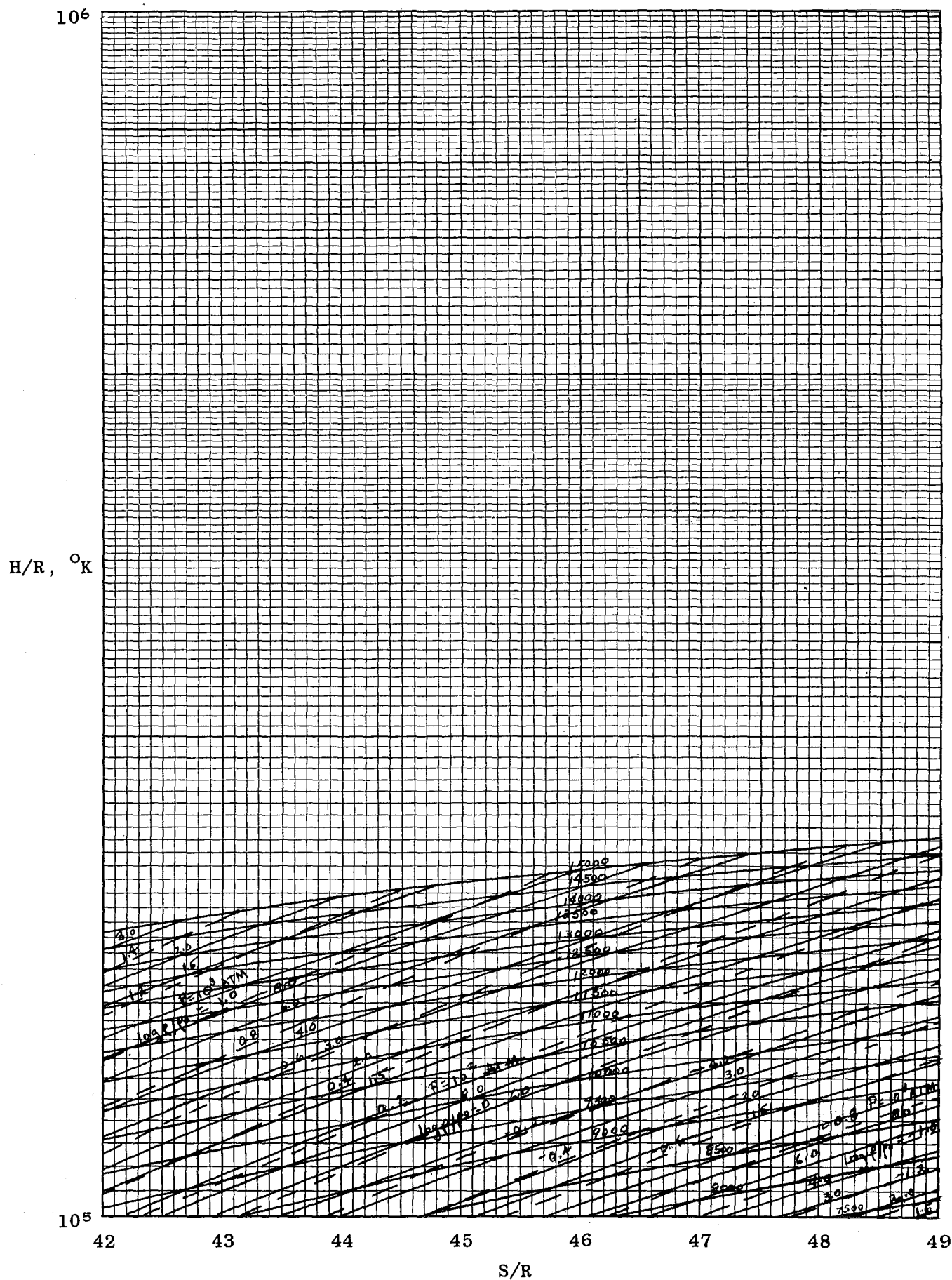
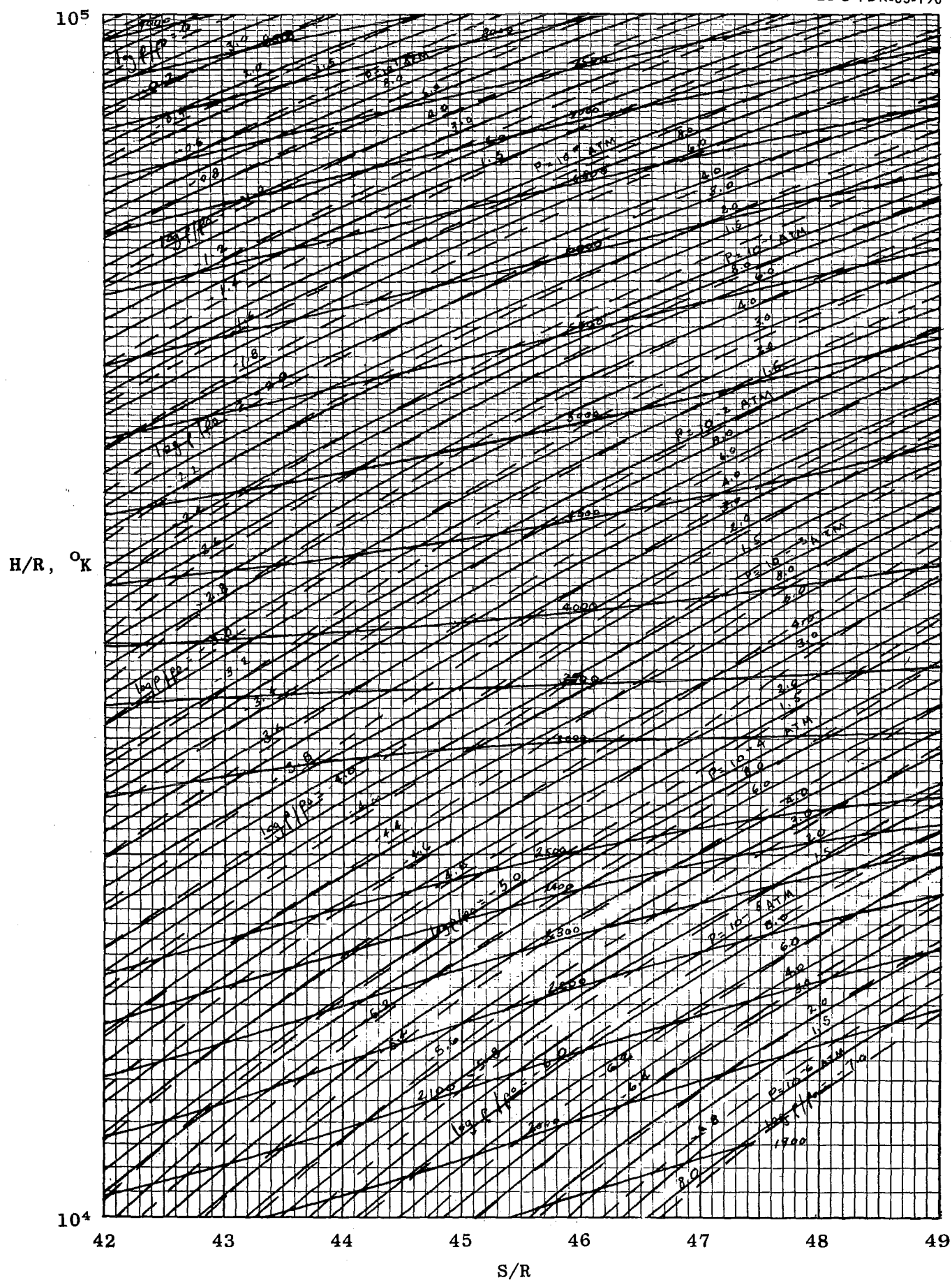


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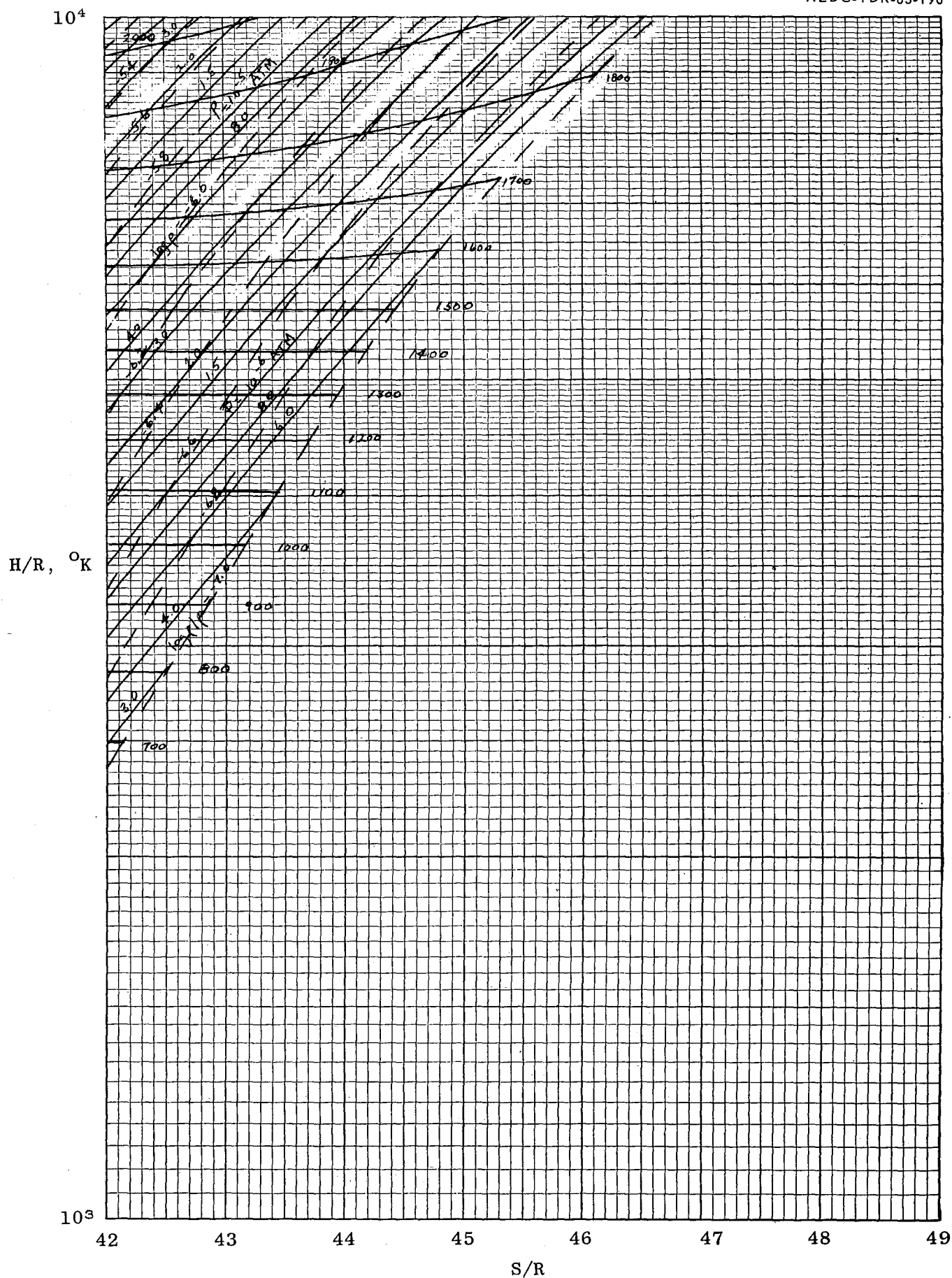


Fig. 16 Air Mollier Diagram

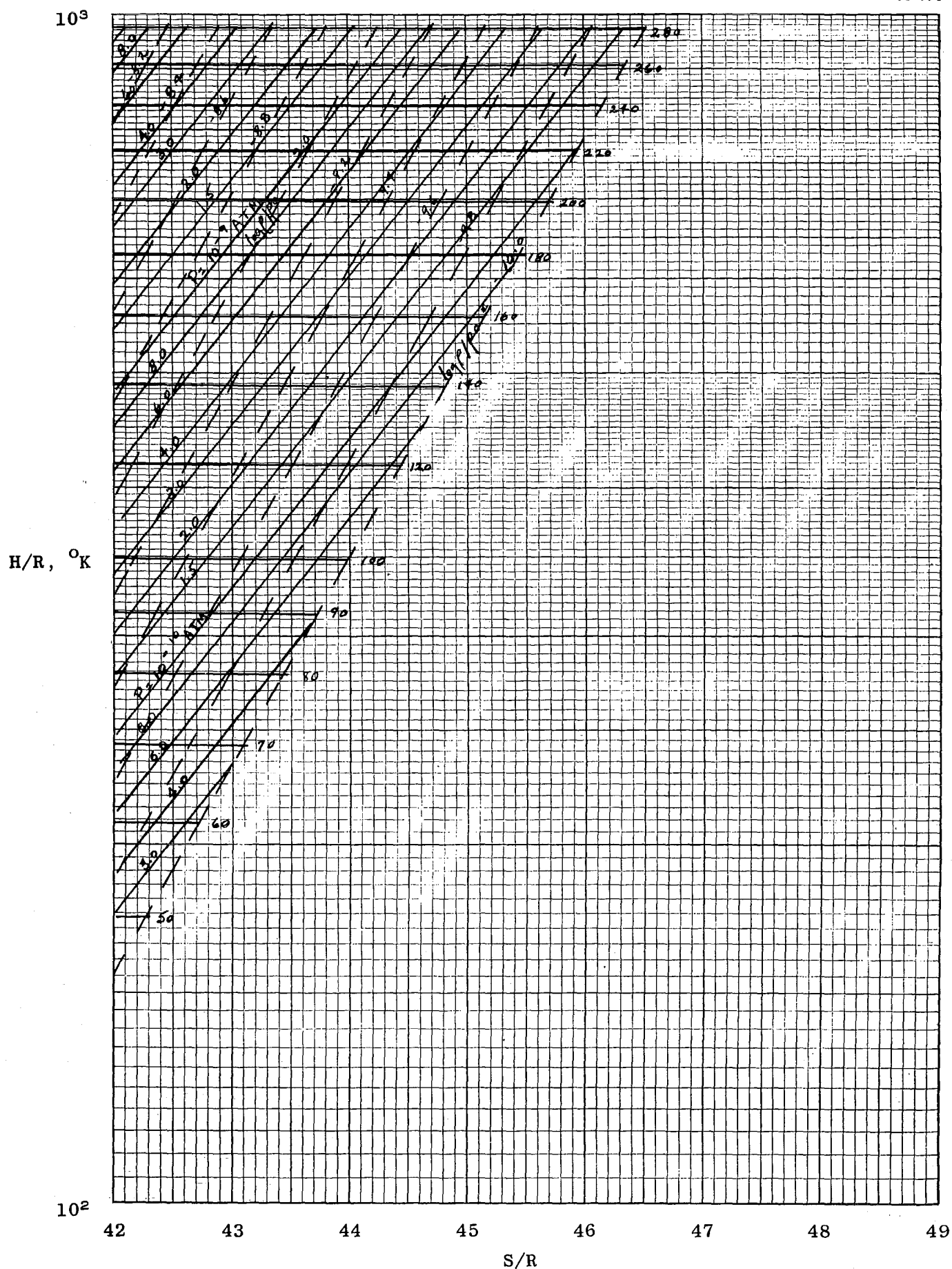
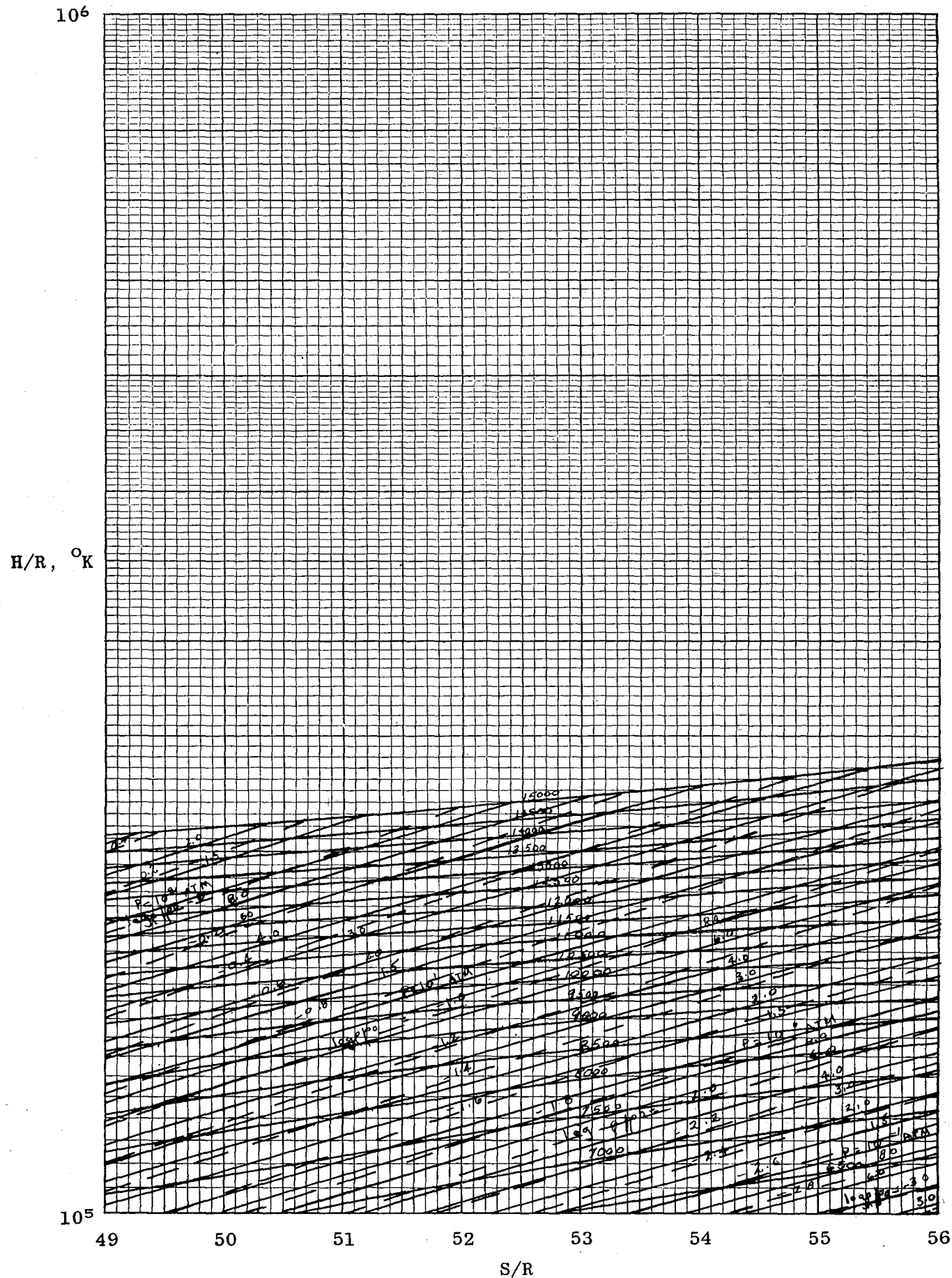
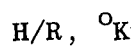


Fig. 17 Air Mollier Diagram



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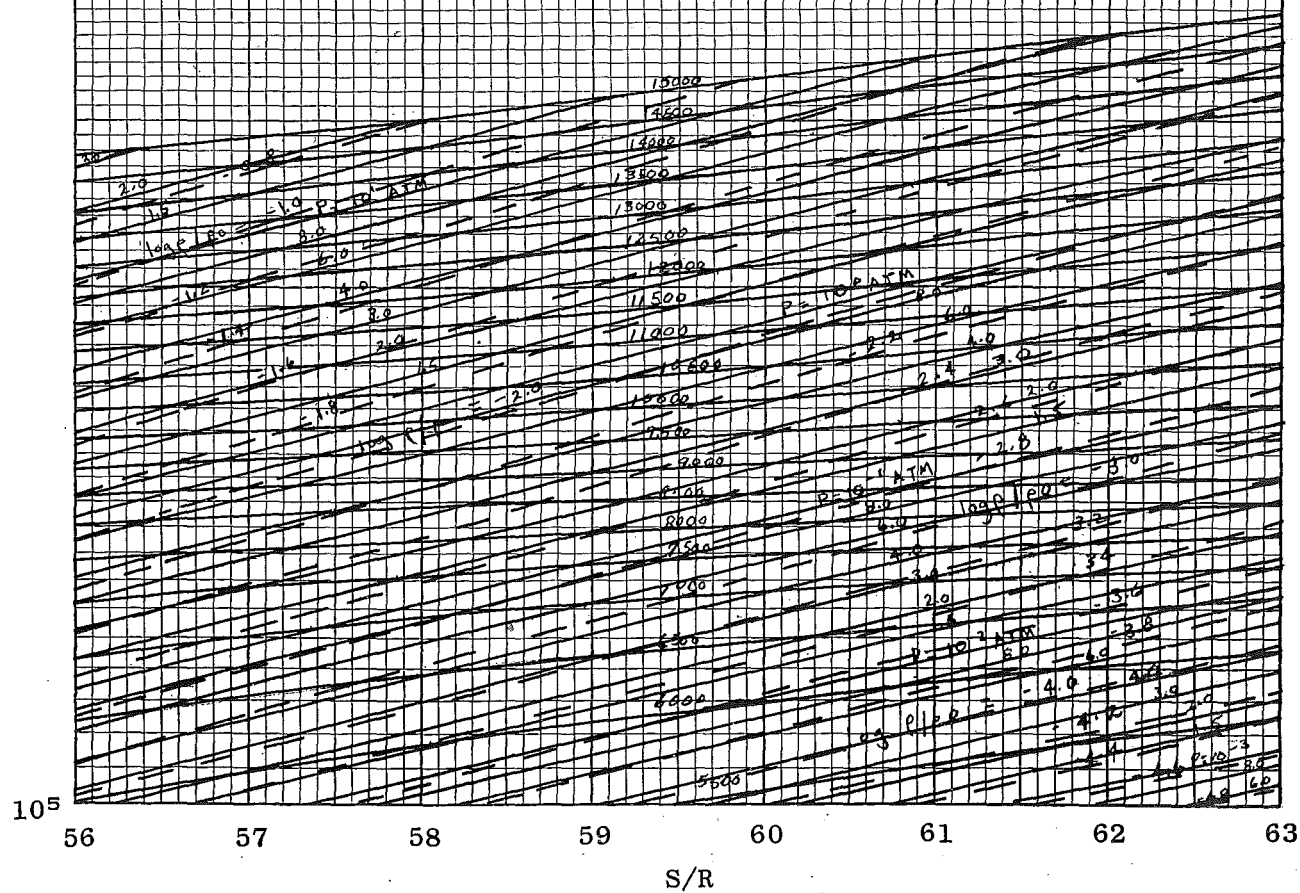
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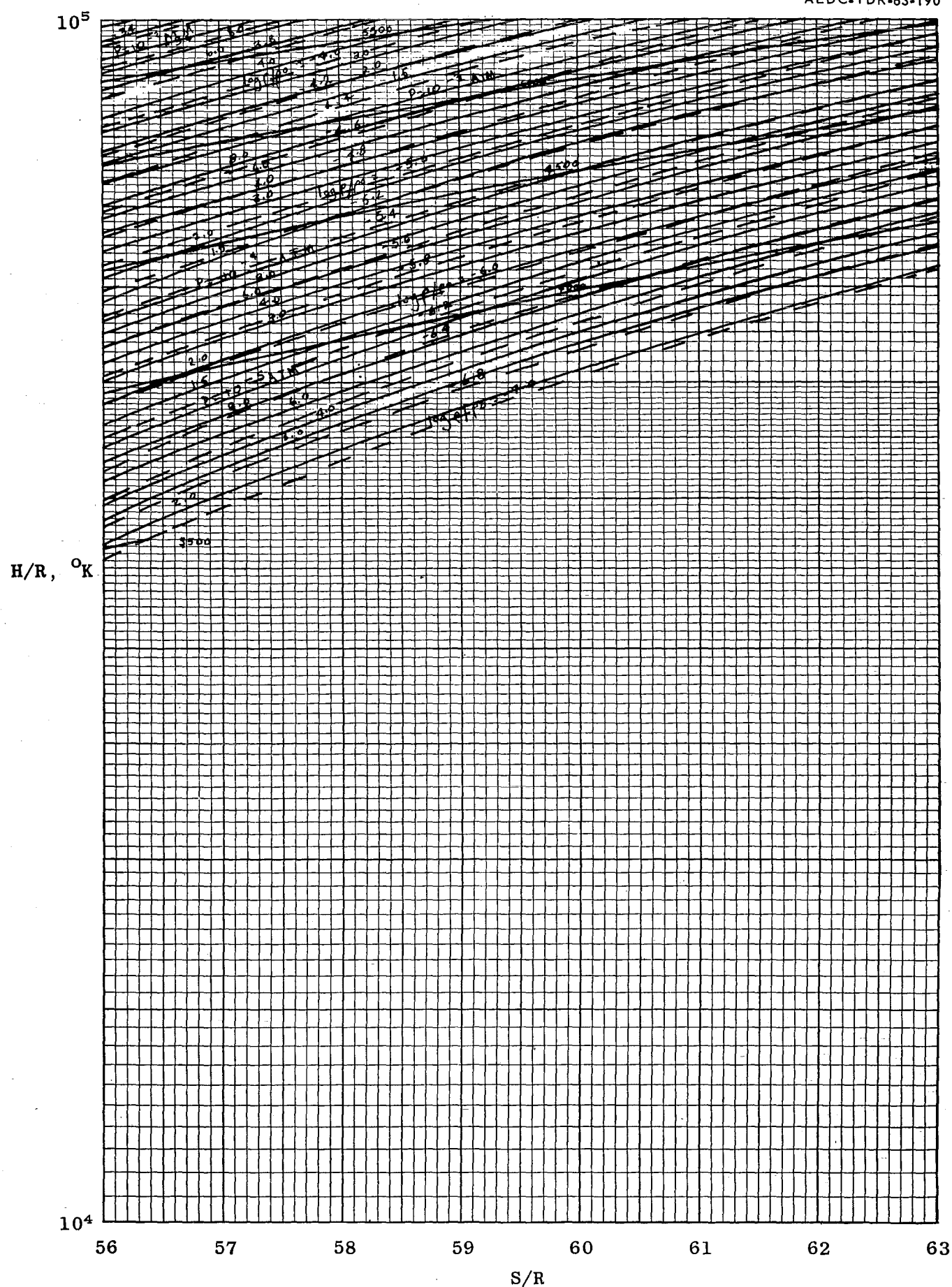
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S/R

Fig. 19 Air Mollier Diagram



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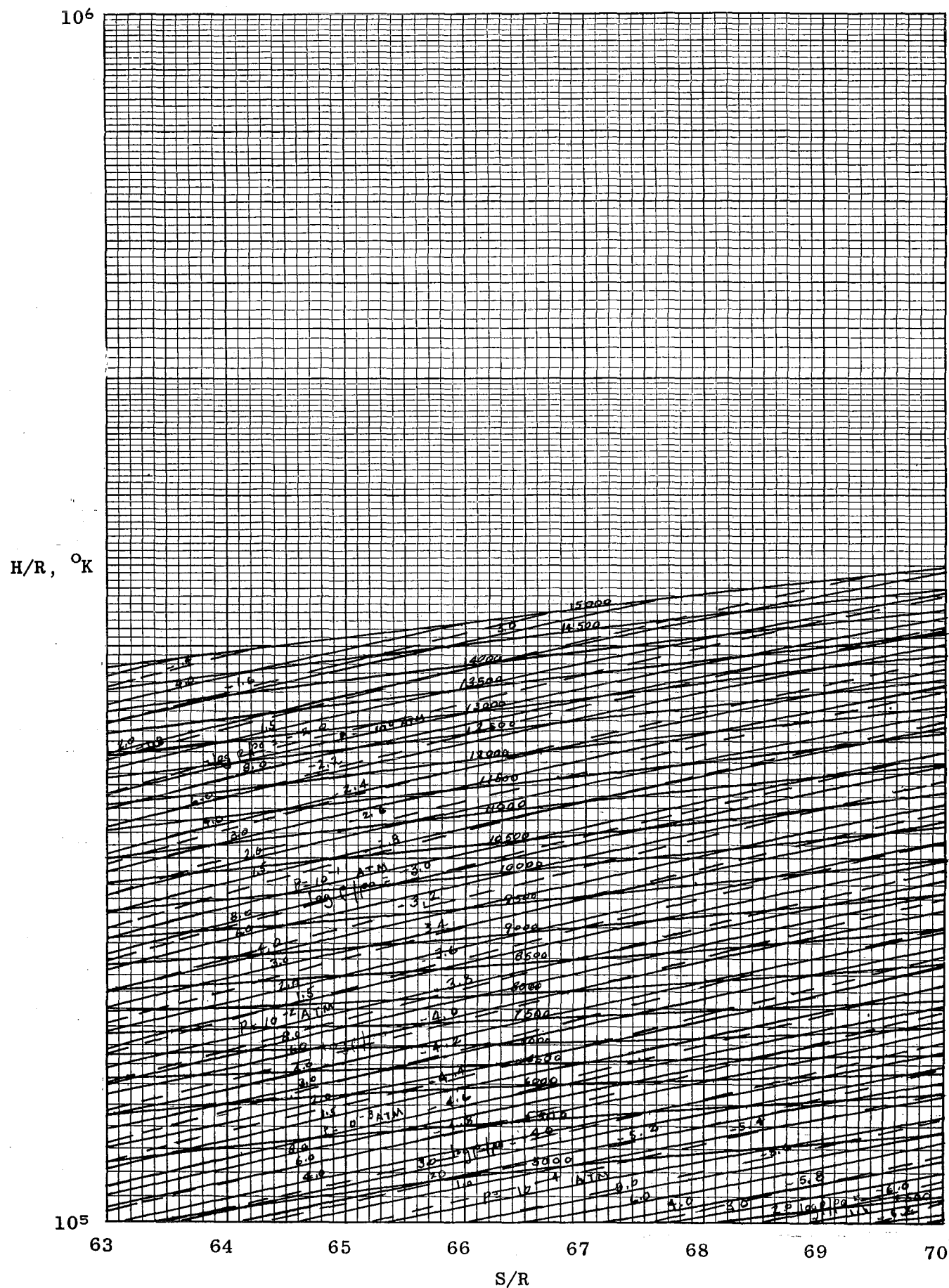


Fig. 22 Air Mollier Diagram

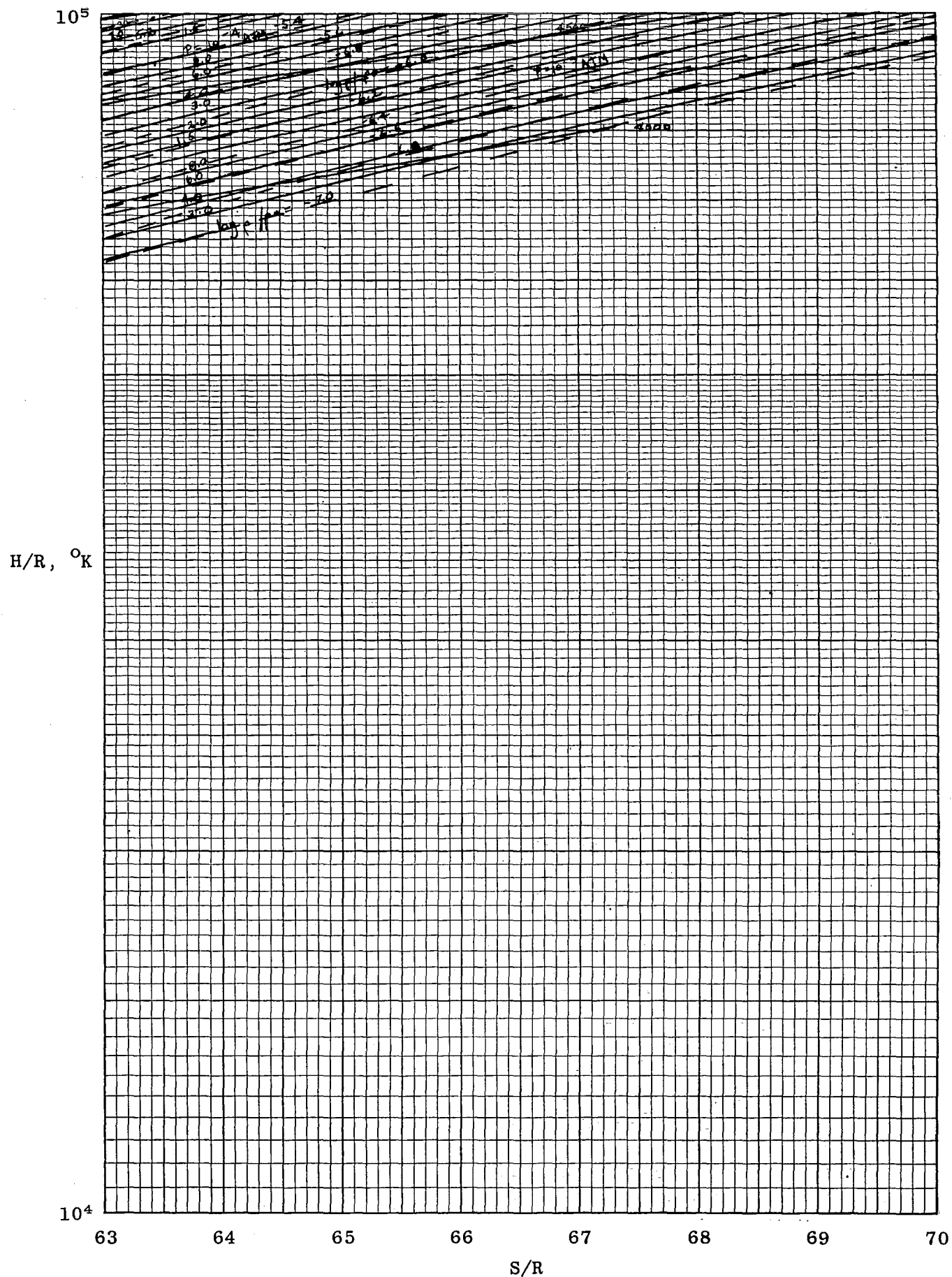


Fig. 23 Air Mollier Diagram

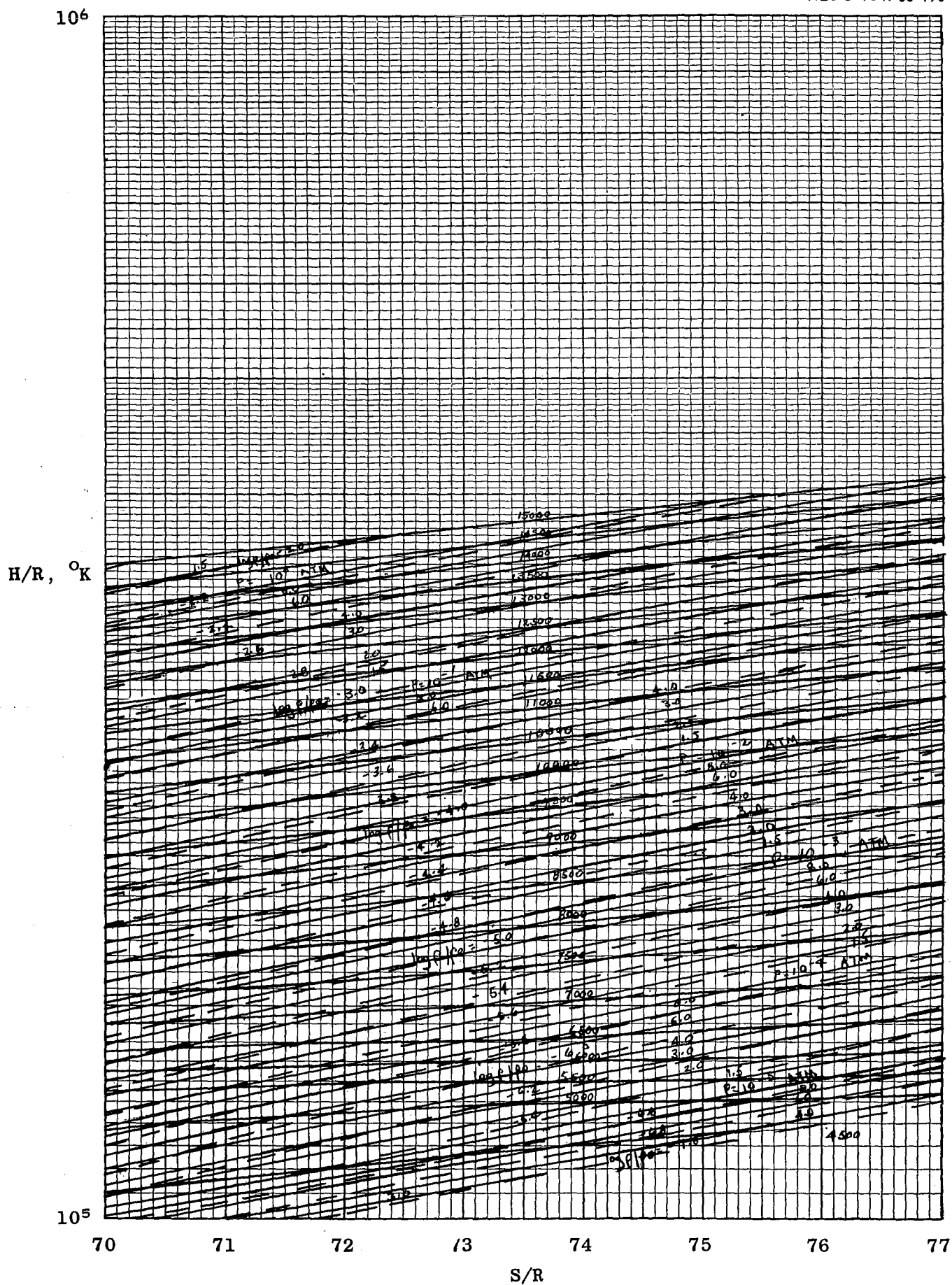


Fig. 24 Air Mollier Diagram

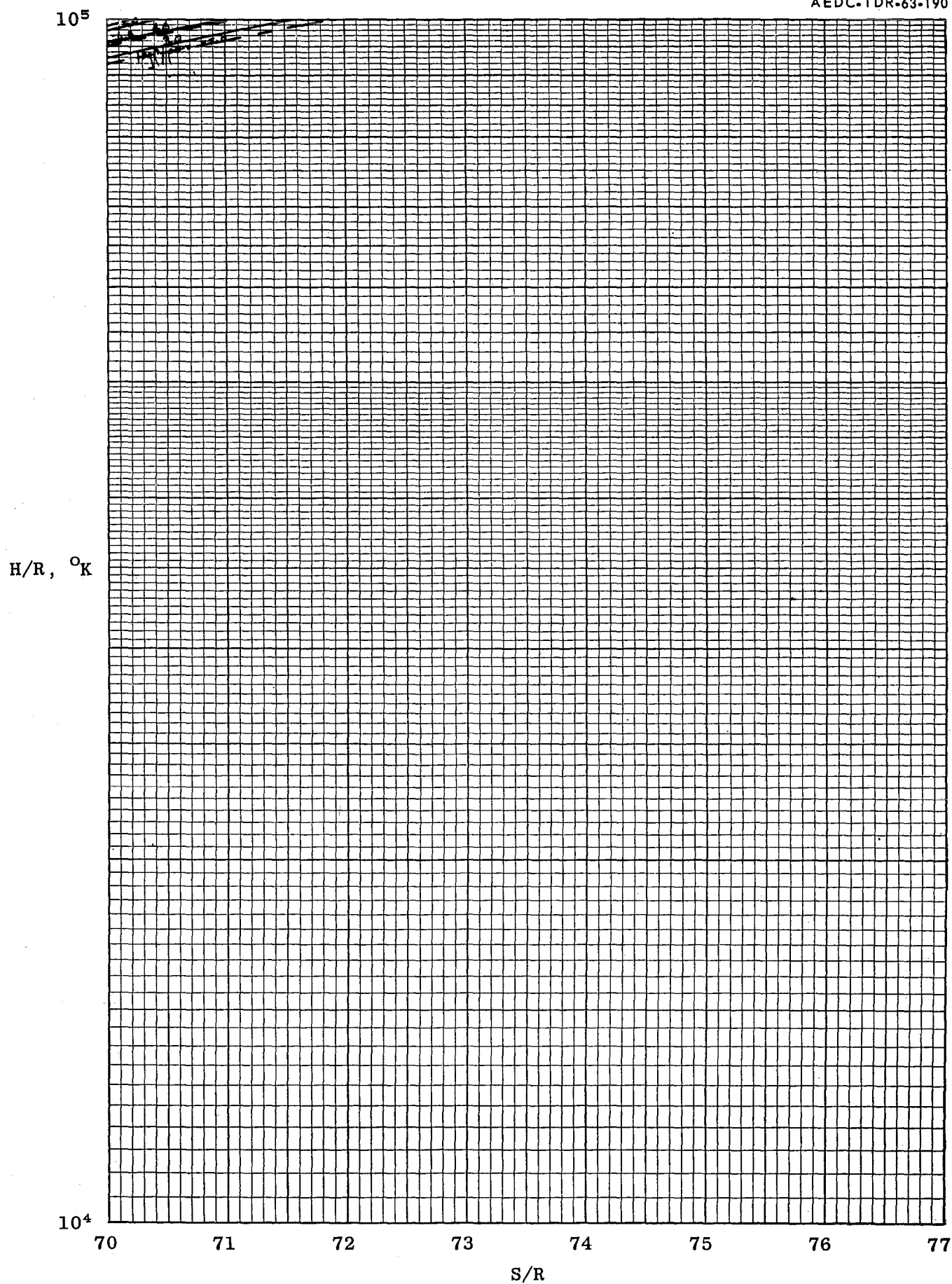


Fig. 25 Air Mollier Diagram

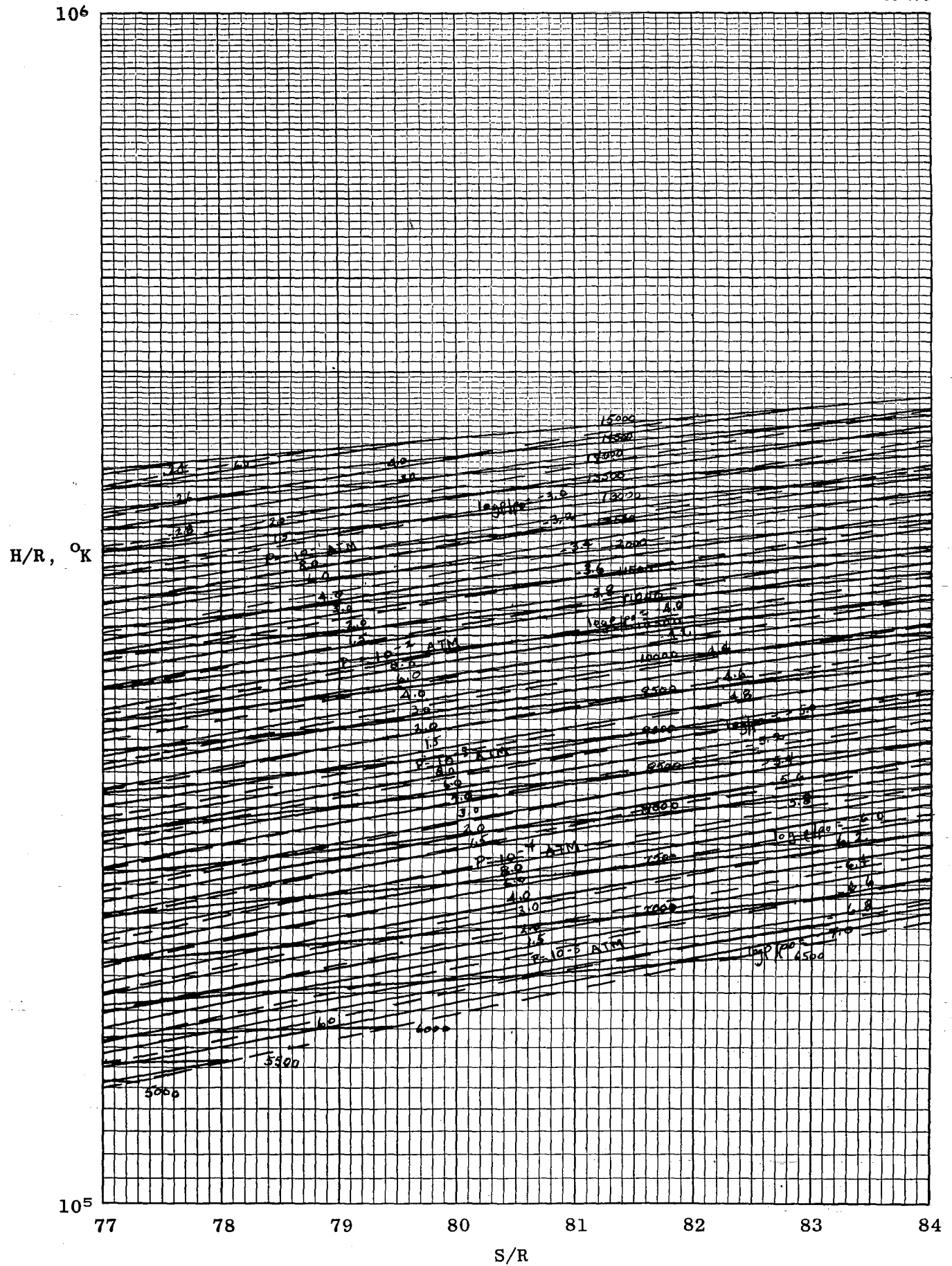


Fig. 26 Air Mollier Diagram

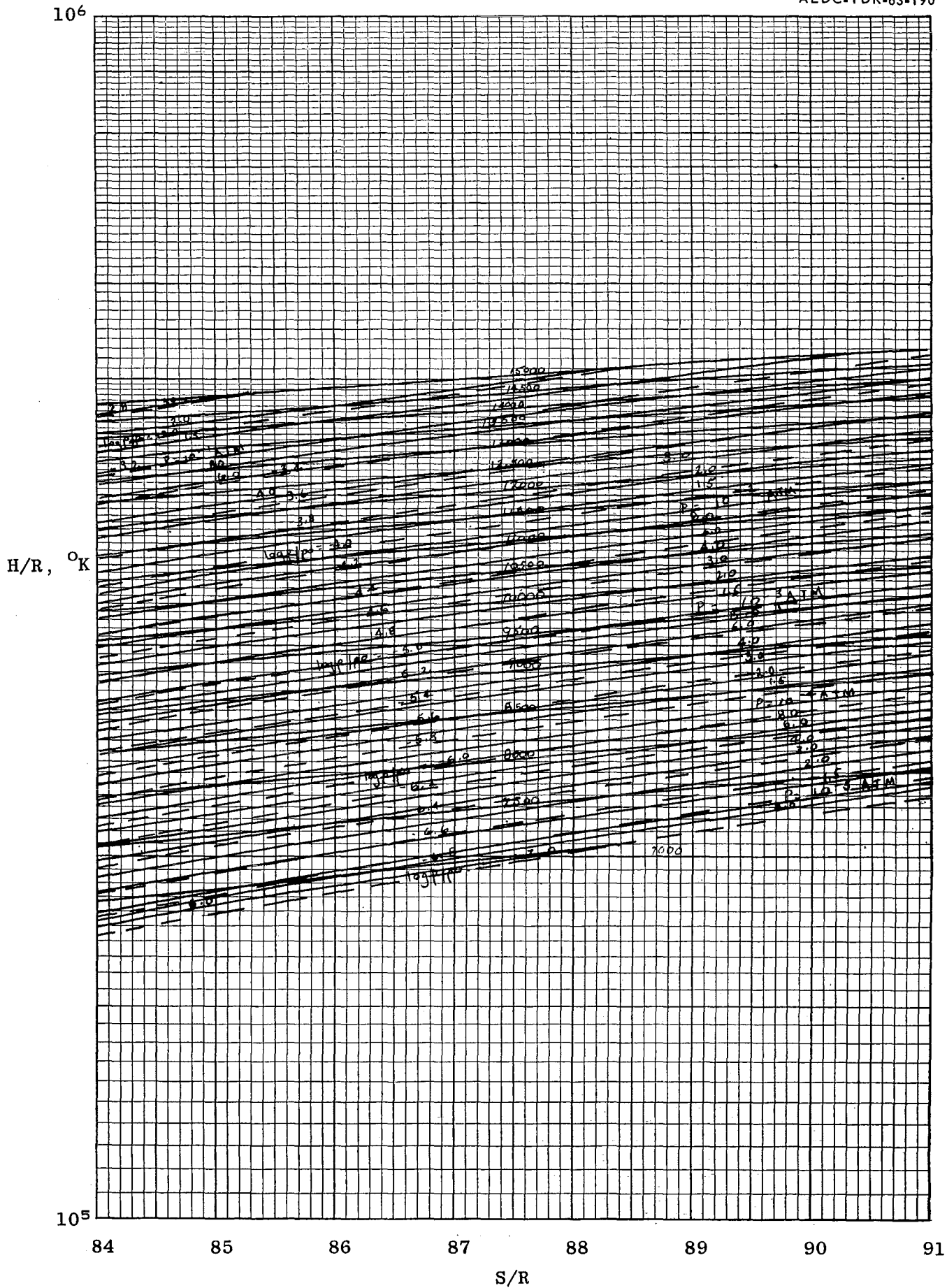


Fig. 27 Air Mollier Diagram

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$H/R, ^\circ K$

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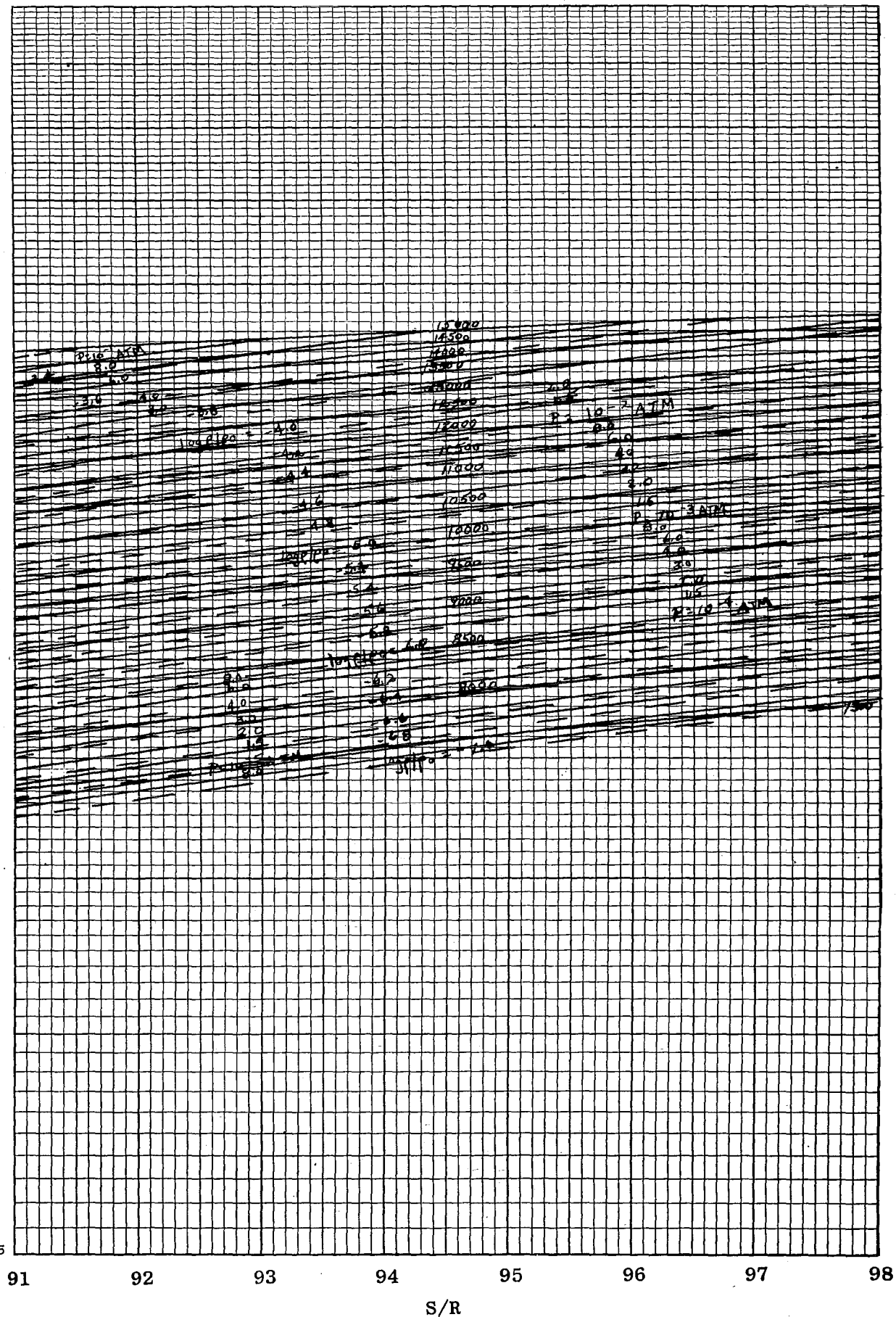


Fig. 28 Air Mollier Diagram

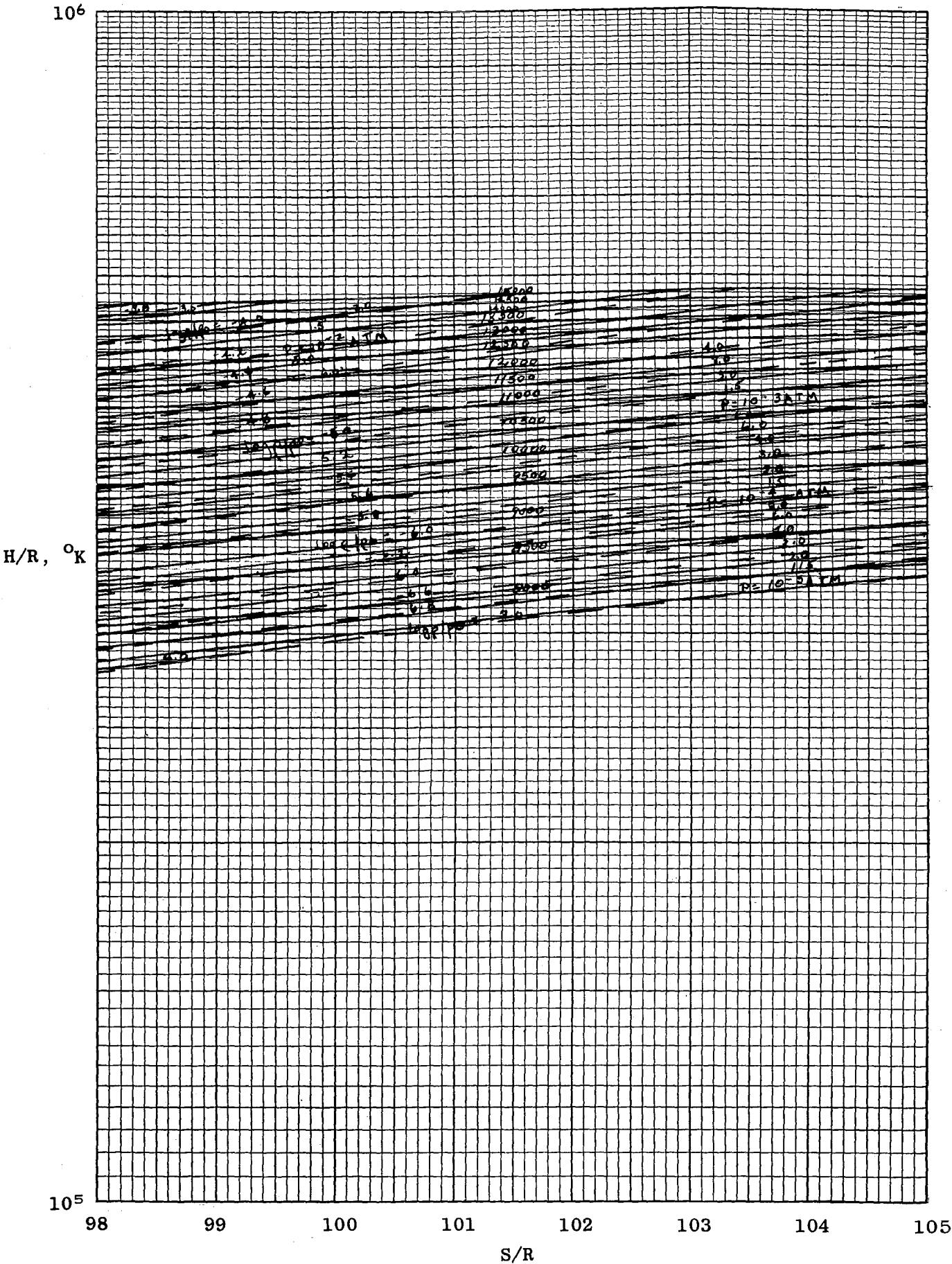


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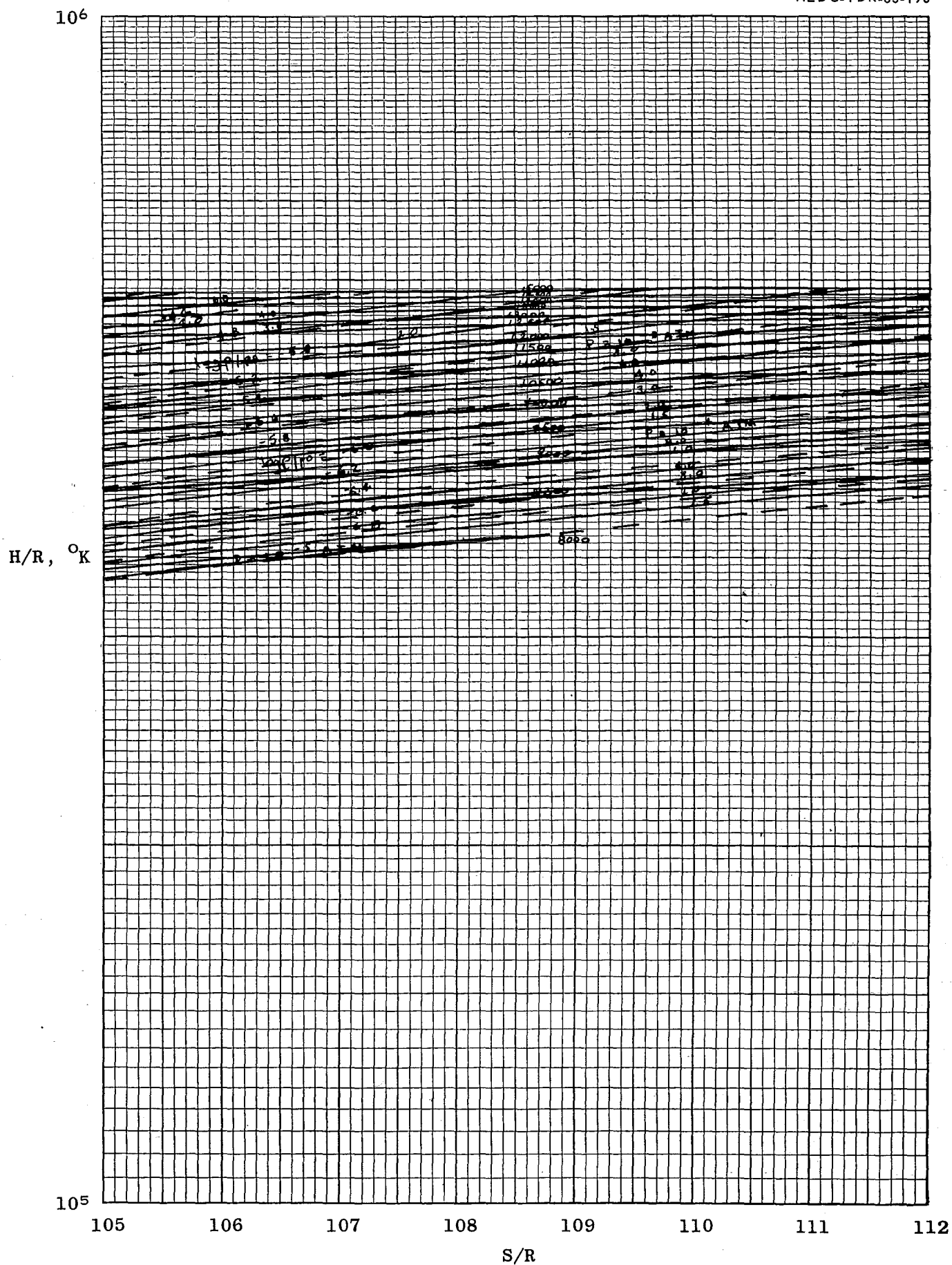
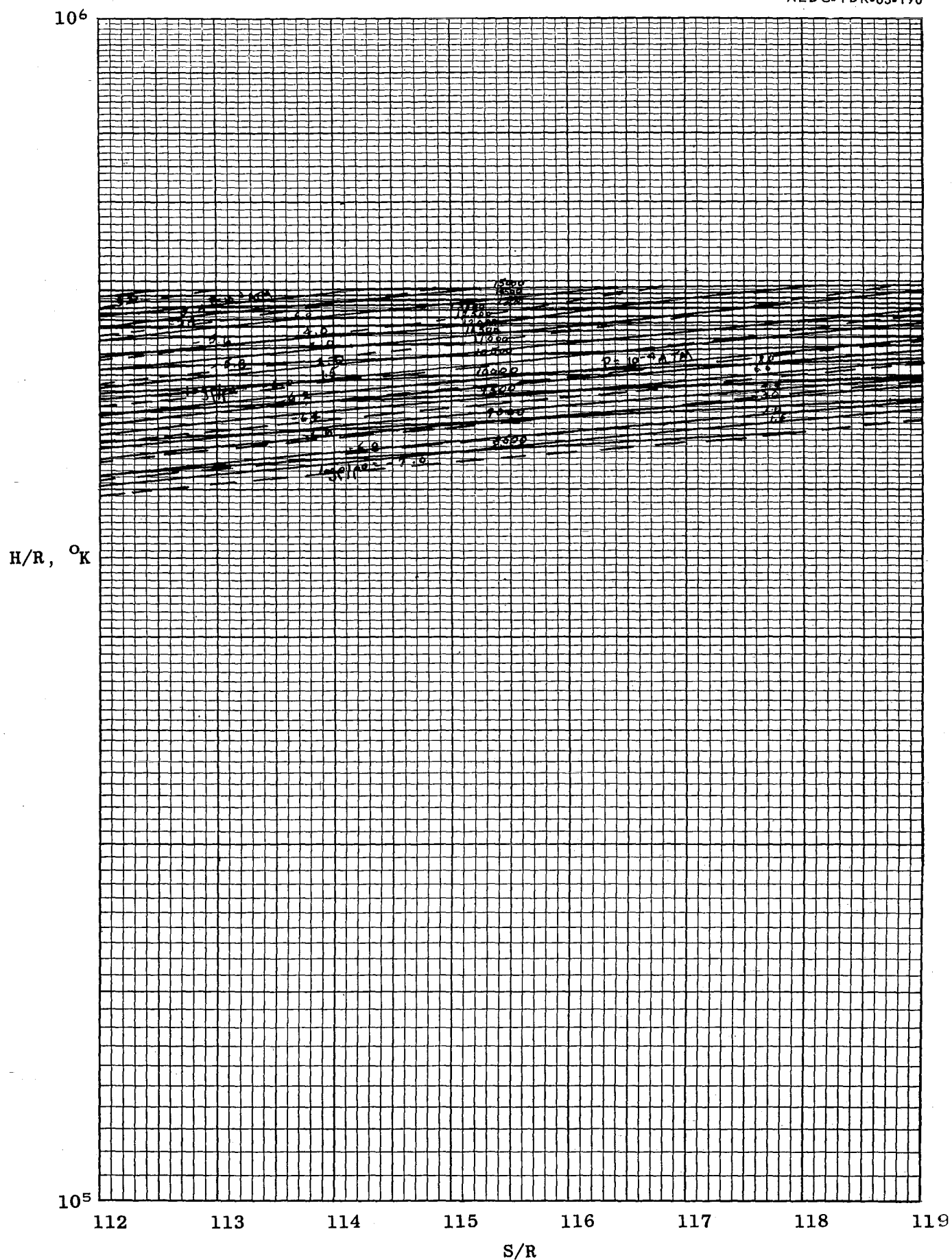


Fig. 30 Air Mollier Diagram



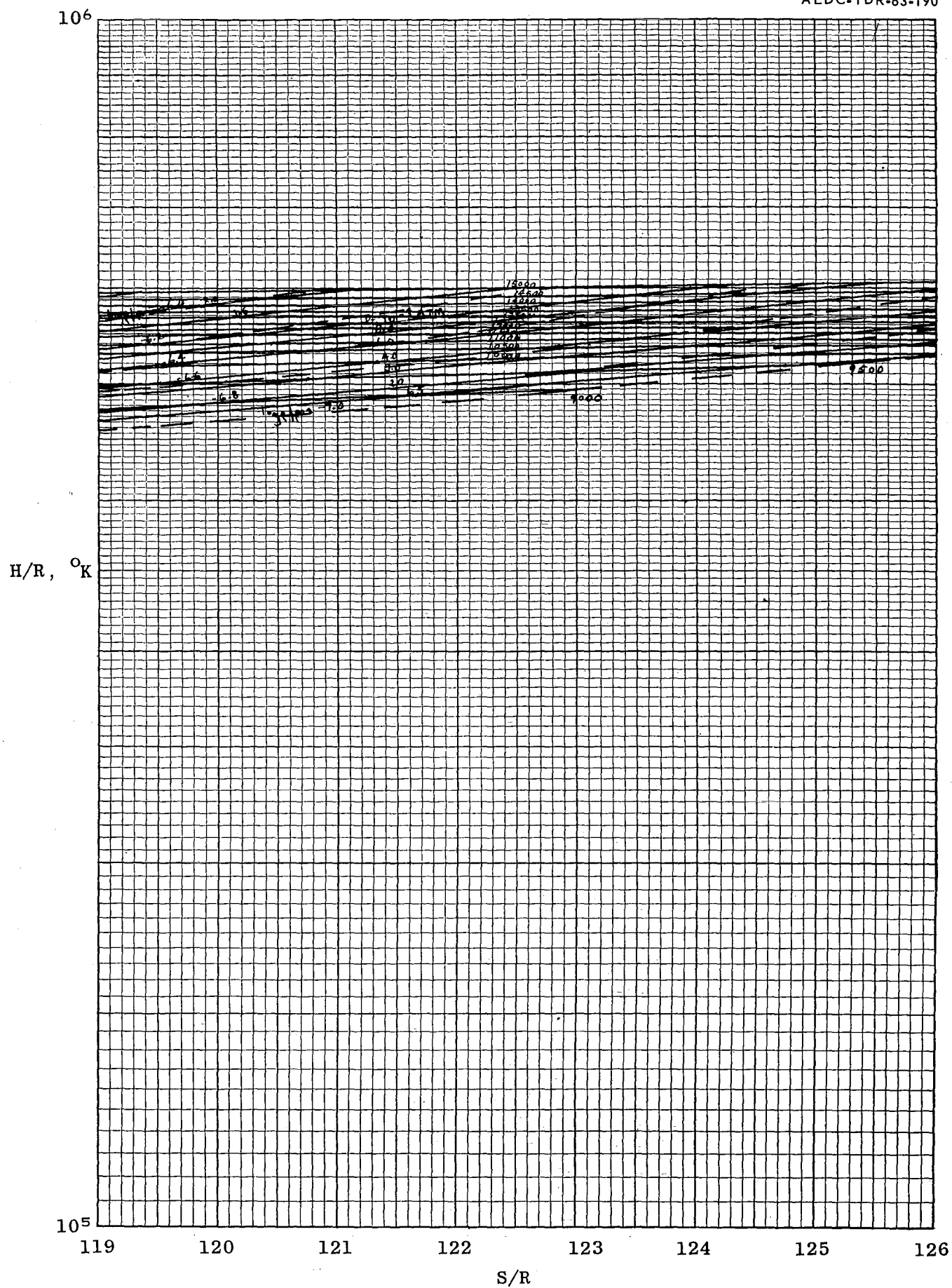


Fig. 32 Air Mollier Diagram

