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INORGANIC SULFUR REAGENTS. I: THE THIONYL HALIDES

By

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Introduction

Our studies of organic sulfur compounds have made us increasingly aware of the inorganic sulfur reagents. While it is clear that many compounds of this group have now become of greatest theoretical and practical interest, yet there has been a lack of correlative and review papers which would serve to focus attention on them. Since we had much of the required literature at hand, it seemed desirable to prepare a series of brief articles, each of which would give a thorough entry to the literature on a selected group of these substances.

The present paper concerns the thionyl halides, SOX_2 . The known examples of this group of compounds are shown in Table I, together with the other known compounds of sulfur and halogen, and of sulfur, oxygen and halogen. Part II of this series will present the sulfuryl halides, SO_2X_2 .

THIONYL CHLORIDEPhysical Properties

Thionyl chloride, SOCl_2 , melts at -104.5° , boils at 78.8° , has a density (at 14.5°) of 1.656 g./ml. and a refractive index, n_D^{10} , of 1.517.¹

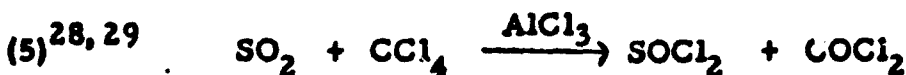
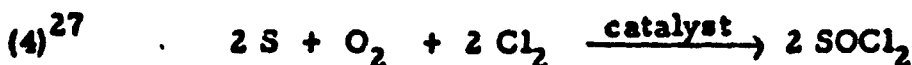
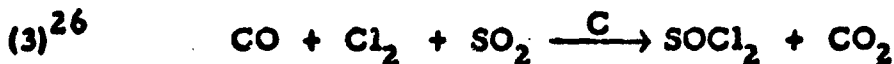
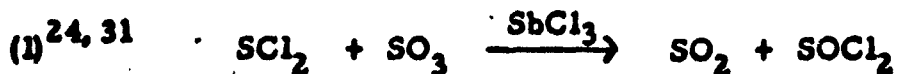
Pure thionyl chloride is a colorless, fuming, toxic, irritating liquid which is miscible with benzene, carbon tetrachloride, chloroform and most inert organic solvents. Studies involving force constants,^{2,3} infrared,^{4,5,6} Raman,⁷⁻¹⁰ and ultraviolet¹¹ spectra, magnetic rotatory power,¹² magnetic susceptibility¹³, and dielectric constants and dipole moments^{14,15} have been carried out. Its physical properties have suggested its frequent use as an ionizing solvent.¹⁶⁻²⁰ The heat of formation of thionyl chloride is -58.5 Kcal./mole.²¹

Preparation

The first preparations of thionyl chloride employed the reactions of sodium sulfite with phosphorus pentachloride²² and of calcium sulfite with phosphorus oxychloride.²³ Industrial methods, described in the patent literature,²⁵⁻³¹ have sought less expensive syntheses, examples of which are shown in the following equations:

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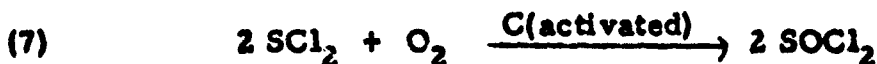
In the process of equation 1, sulfur dichloride may be prepared in situ from sulfur monochloride and chlorine ($\text{ClSSCl} + \text{Cl}_2 \longrightarrow 2 \text{ClSCl}$).^a

The sulfur trioxide is usually introduced by using fuming sulfuric acid. Excess sulfur trioxide must be avoided because of the easy formation of the so-called 'pyrosulfuryl chloride' (equation 6). If radioactive SCl_2 is used



in the method of equation 1, the tagged sulfur is found mainly in the thionyl chloride.³²

Considering the reaction of equation 2: sulfur dichloride may be substituted for sulfur monochloride and part of the chlorine, and sulfuryl chloride may be used as a source of SO_2 and Cl_2 . The reaction is best carried out at 150-200°. With recycling of the sulfur chlorides, SO_2 and SO_2Cl_2 which are formed, nearly 100% conversion is reported. At high temperatures (500°), the decomposition of thionyl chloride is nearly complete. At intermediate temperatures (198° and 238°), the system, $\text{SO}_2\text{Cl}_2 + \text{SCl}_2 \rightleftharpoons 2 \text{SOCl}_2$, represents a reversible reaction.³⁰ The oxidation of SCl_2 with oxygen, to give sulfuryl chloride is irreversible under these conditions. Equation 7 represents an overall procedure which has been used to prepare thionyl chloride.³⁰

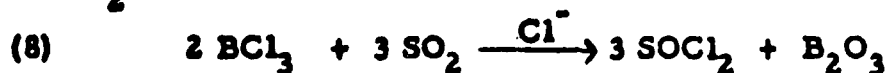


Phosgene may also be employed in reaction 3.

In reaction 5, the yield of thionyl chloride is 72%. Substitution of chloroform or pentachloroethane for carbon tetrachloride lowers the yield to 40% and 6%, respectively.

^a This chlorinolysis of the sulfur-sulfur bond is a very general reaction, and has found many applications in organic chemistry. The products obtained are sulfenyl chlorides: $2 \text{RSSR} + \text{Cl}_2 \longrightarrow 2 \text{RSCl}$. Cf. e.g., Organic Sulfur Compounds, Vol. 1, edited by N. Kharasch, Pergamon Press (1960); Chs. 31 and 32.

The reaction of boron trichloride with the sulfur dioxide, in which the latter acts as solvent, is extremely slow unless chloride ion is present.³³ It has been suggested that the catalytic effect of chloride ion may involve the ion, SO_2Cl^- .



Technical grade thionyl chloride contains about 7% impurities, mainly sulfur monochloride, sulfuryl chloride and sulfur dioxide.³⁴ Sulfur monochloride may be removed after reaction with sulfur, and the sulfuryl chloride may be taken out by means of its selective reaction with other substances.³⁴ Reagent grade thionyl chloride is also available.

Chemical Properties

Above its boiling point, thionyl chloride decomposes to sulfur chloride, sulfuryl chloride, sulfur dioxide, and chlorine. The photochemistry of thionyl chloride does not seem to have been investigated in detail, but seems worthy of study, especially in comparison with sulfuryl chloride.

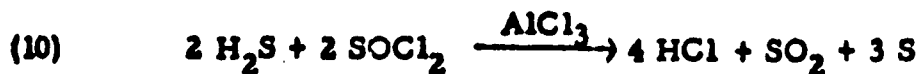
Rapid sulfur exchange between S^*OCl_2 - SOBr_2 and SOCl_2 - S^*OBr_2 in SO_2 solution at -50° or in mixtures of pure compounds at -20° has been observed.³⁵ Similarly, chlorine exchanges rapidly between Cl^* -labelled pyridinium chloride and SOCl_2 in chloroform at 20° .³⁶ These results may be accounted for by the ionization of the thionyl halides or the formation of addition complexes and are in marked contrast to results wherein the lack of exchange between S^*OCl_2 and SO_2Cl_2 has been noted.³⁵

Thionyl chloride reacts with metals, such as antimony, sodium, aluminium, magnesium, and iron. The reactions are generally slow in the absence of acids³⁷ (except in the case of antimony) and this explains the use of iron containers for storage of the thionyl chloride. Glass, lead, and nickel containers are also used for storage.

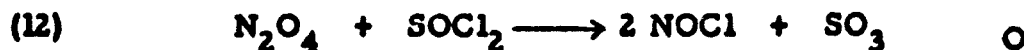
With phosphorus, selenium and tellurium, reaction with thionyl chloride yields the respective chlorides.

Hydrated chlorides, such as $\text{MgCl}_2 \cdot 6 \text{H}_2\text{O}$, are conveniently dehydrated with SOCl_2 , since the products, HCl and SO_2 , are gases.³⁹ Mustard gas may also be dried in this way.⁴⁰

In some of its inorganic reactions, thionyl chloride acts as an oxidizing agent, e.g.:

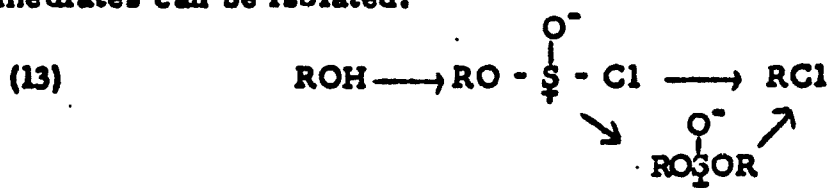


In other reactions, however, it functions as the reducing agent:



The reaction with silver nitrate yields silver chloride and $\text{Cl}\overset{\text{O}}{\underset{\text{O}}{\text{S}}}\text{ONO}_2$, while with ammonia, polymeric thionylimine, $(\text{HNSO})_x$, is reported as the product.⁴¹ The preparation of chromyl chloride using CrO_3 or $\text{K}_2\text{Cr}_2\text{O}_7$ has been mentioned.⁴²

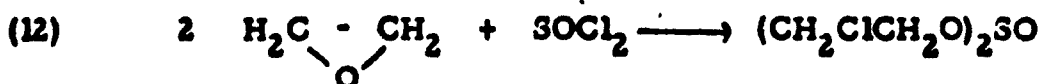
The best known reactions of thionyl chloride in organic chemistry involve the replacement of hydroxyl groups, in alcohols and carboxylic acids, by chlorine. The conversion of alcohols to alkyl halides often occurs with retention of configuration. This is not, however, always the case. The steps involved in the reaction converts the alcohol to chlorosulfonates and, in some instances, to sulfites. Under certain conditions these intermediates can be isolated.



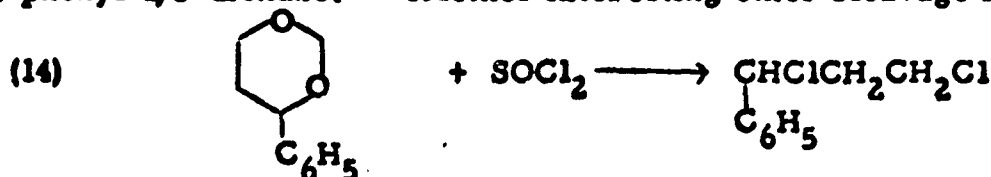
The elimination of sulfur dioxide from the chlorosulfonates may occur through different transition states.

The detailed mechanism and stereochemistry has been studied in a number of investigations with alcohols,^{32, 46, 47, 48} with small ring alcohols,⁴⁹ with chlorocyclohexanols,⁵⁰ with allylic alcohols,⁵¹ with acylamino alcohols,⁵² and with sterols.^{53, 54} The widely used reaction for conversion of acids to acid chlorides has been shown to occur through the carboxylic anhydride when pyridine is employed.⁵⁵ Sulfonic acids and carboxylic acids, which do not normally react with thionyl chloride, do so in the presence of dimethylformamide.⁵⁶

The ionized salt $(\text{CH}_3)_2\text{NCHCl}^+\text{Cl}^-$, is formed. Another interesting application involves the estimation of the number of carboxyl groups in gelatin with SOCl_2 and CH_3OH .⁵⁷ 2-Naphthoic acid, heated with thionyl chloride at 180-200°, yields 1,6-dichloronaphthalene. Mercapto groups are replaced by chlorine and some ketones are converted to vinyl chlorides by SOCl_2 .⁵⁹ Ethylene oxides yield 2-chloro sulfites⁶⁰ while aliphatic β -lactones give 3-chloroalkanoyl chlorides.⁶¹



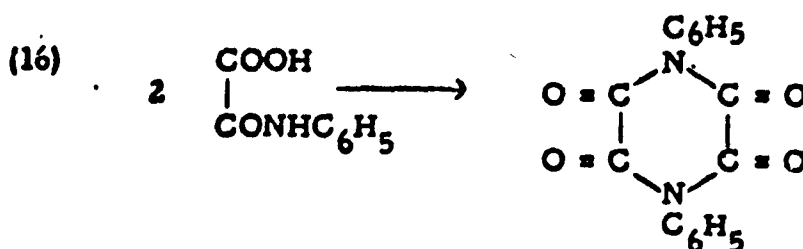
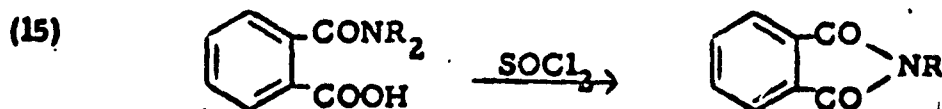
Thionyl chloride cleaves cyclic ethers and related compounds to afford dichloro derivatives. Examples are the reaction with tetrahydrofuran⁶² and 4-phenyl-1,3-dioxane.⁶³ Another interesting ether cleavage involves the



conversion of γ -methoxyvaleric acid to methyl δ -chlorovalerate with inversion at the γ -carbon atom.⁶⁴

The products of the reaction of SOCl_2 with phenols are dependent on the conditions. In some instances sulfites can be obtained and under other conditions ring substitution occurs. The reactions of this type have been reviewed.⁶⁵ In general, the Friedel-Crafts reaction with SOCl_2 is not very successful, but 9-fluorenyl⁶⁶ and *p*-acetamidophenyl⁶⁷ sulfoxide may be prepared in this way. In the presence of copper, SOCl_2 and substituted phenols (4-hydroxycoumarin,⁶⁸ 4-propionylresorcinol,⁶⁹ 2-hydroxy-5-methylacetophenone,⁷⁰ and the naphthols⁷¹) yield symmetrical sulfides by ring substitution and attendant reduction.

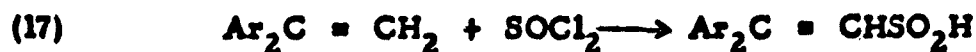
Thionyl chloride is an effective agent for intramolecular removal of the elements of water. Examples are the dehydration of 1,1,1-trichloro-2-methyl-2-propanol, which occurs *via* the chlorosulfinate and which is catalyzed by amines,⁷² and ethyl 2-carbomethoxy-1-hydroxy-2-methyl cyclohexylacetate.⁷³ The units of alcohol are similarly removed from a substituted phthalamic acid⁷⁴ and water is removed intermolecularly from oxanilic acid.⁷⁵ Urea is converted to cyanuric acid and other products.^{76,77}



The intermolecular condensation of glucose is also effected through the use of thionyl chloride.⁷⁸

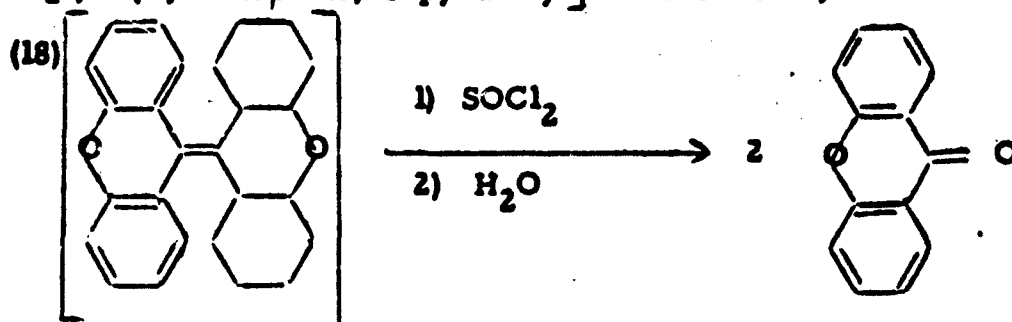
Chlorination of compounds with thionyl chloride is illustrated by the reaction with dimethyl sulfide.⁷⁹ For these purposes, thionyl chloride is a less vigorous agent than sulfuryl chloride and promotes formation of mono and dichloro derivatives, rather than higher chlorinated products. The

addition of chlorine to benzene is activated by the presence of thionyl chloride and the amount of the gamma isomer of hexachlorocyclohexane may be affected by its presence^{80, 81} In the special case of 1, 1-diarylethylenes, sulfinic acid is obtained⁸² after hydrolysis of the product. Chloropheno



thiazines are chlorinated by thionyl chloride.⁸³ Alternatively, the phenothiazine ring may be closed with the chlorination by the action of thionyl chloride on diphenylamine.

The oxidative character of thionyl chloride is sometimes exhibited in organic reactions. The cleavage of substituted $\Delta^{9, 9'}$ -bixanthenes to xanthenes^{84, 85, 86} and the dehydrogenation of a highly hindered acrylic acid, cis-o- [4, 5-(1', 2'-naphtho)-3-pyrazolyl] cinnamic acid, to the corresponding



propionic acid⁸⁷ are examples.

Thionyl chloride is not usually the best reagent for effecting the Beckmann rearrangement because it initiates cleavage reactions.⁸⁸ The conversion of cyclohexanone oxime to epsilon-caprolactam with thionyl chloride occurs with fair results, however.⁸⁹

Primary aromatic amines are converted to thionylimines, $\text{ArN} = \text{SO}$, with SOCl_2 .⁹⁰ 8-Thiapurines are obtained by the reaction of thionyl chloride with 5, 6-diaminouracils.⁹¹ Other thiadiazoles are obtained through the use of orthodiamines.⁹² The cleavage of sulfinyl esters by SOCl_2 ⁹³ and its addition to ketene to give after methanolysis $(\text{MeOOCCH}_2)_2\text{SO}$ ⁹⁴ are further interesting applications. Its reaction with phenylarsine to give $\text{C}_6\text{H}_5\text{AsS}$ ⁹⁵ contrasts with the reaction of phenylphosphine with SOCl_2 , which yields $(\text{C}_6\text{H}_5\text{PSO})_3$.⁹⁶ $\text{C}_6\text{H}_5\text{As}(\text{MgBr})_2$ and SOCl_2 gives $\text{C}_6\text{H}_5\text{AsSO}$.

A possible application of thionyl chloride, which has not been exploited fully as yet, is its chain terminating properties. It has been shown that SOCl_2 acts in this way in the sulfoxidation of cyclohexane (reaction with SO_2 and O_2 , under irradiation).⁹⁷

Reviews on the early chemistry of SOCl_2 ,²³ the reaction with phenol,⁶⁵ with p-aminobenzoic acid,⁹⁸ synthetic applications,⁹⁹ and technology¹⁰⁰ have appeared.

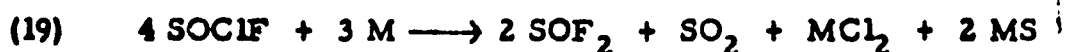
THIONYL FLUORIDE

Thionyl fluoride, SOF_2 , melts at -129.5° and boils at -43.8° ¹⁰¹. Raman,^{102, 103} infrared,¹⁰⁴ and microwave¹⁰⁵ spectra, dipole moment¹⁰⁵ and force constant studies have been carried out. Thionyl fluoride is obtained by the reaction of thionyl chloride with AsF_3 , SbF_3 ,^{106, 107} HF ,¹⁰⁸ or KSO_2F ,^{109, 110} by reaction of S_4N_4 with aqueous HF and CuO ,¹¹¹ and by reaction of SO_2 with VF_5 .¹¹² A 77% yield was obtained recently by reaction of sodium fluoride with thionyl chloride in acetonitrile.¹¹³

Thionyl fluoride hydrolyzes slowly, but completely at room temperature, forming sulfurous and hydrofluoric acid. Thermally, however, it is quite stable.¹⁰⁶ With SiO_2 at 400° , thionyl fluoride gives SiF_4 and SO_2 . Mixtures of SOF_2 and O_2 , submitted to an electric discharge, give peroxy compounds.¹¹⁴

THIONYL CHLOROFUORIDE

Thionyl chlorofluoride, SOClF , boils at 12.2° , melts at -139.5° , and has a liquid density of 1.580 g./ml.¹⁰¹ It is prepared by the action of SbF_3 and SbCl_5 ¹⁰¹ or IF_5 .¹¹⁵ on SOCl_2 . The chlorofluoride reacts with mercury and copper, at room temperature, and with iron at 70° . The overall reaction is represented by equation 19. (M = metal).



Disproportionation may occur first; if so, the products are those of the reaction of SOCl_2 with the metal.¹⁰¹

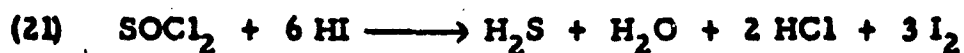
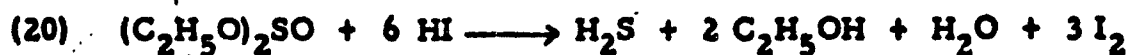
THIONYL BROMIDE

Thionyl bromide, SOBr_2 , melts at -52° , boils at 138° (decompn.), at atmospheric pressure or at 42.5° at 16 mm. pressure, has a density at 15° of 2.698 and is an orange-yellow liquid.¹¹⁶ Raman spectrum.¹¹⁷ Thionyl bromide is prepared by the reaction of HBr or KBr with SOCl_2 .^{116, 118, 119} Reports of the existence of SOClBr are probably incorrect.¹¹⁶ Thionyl bromide decomposes slowly at room temperature ($4 \text{SOBr}_2 \longrightarrow \text{S}_2\text{Br}_2 + 2 \text{SO}_2 + 3 \text{Br}_2$) and is rapidly hydrolyzed by water. Thionyl bromide converts carboxylic acids to acid bromides and substituted alcohols to the corresponding bromides.^{119, 120} A general method involving conversion of alcohols to sulfitates with SOCl_2 and cleavage of the $(\text{RO})_2\text{SO}$ with thionyl bromide is recommended for the preparation of alkyl bromides.¹²¹ Thionyl bromide reacts with pyridinium halides to give $\text{C}_5\text{H}_5\text{NH}^+ \text{XBr}_2^-$ along with S and SO_2 .^{122, 123} Sulfur exchange between sulfur dioxide and thionyl bromide is slow¹²⁴ unless ionic halides or catalysts are present.^{125, 126}

A review of the synthetic uses of sulfuryl bromide has been made.⁹⁹

THIONYL IODIDE

Thionyl iodide, SOI_2 , is unstable. Its formation by reaction of thionyl chloride in carbon tetrachloride solution with potassium iodide,^{127, 128} has been reported and its decomposition was followed by spectroscopic means. It is decomposed by light. Dilute solutions in the dark are claimed to be stable for eight hours.¹²⁷ It has been considered to be an intermediate in the following reactions:¹²⁹



The results probably bear reinvestigation, since quaternary ammonium iodides, when treated with thionyl chloride, yield salts as follows:¹³⁰



Iodine is liberated when potassium iodide is added to these salts.

KNOWN SULFUR HALIDES AND SULFUR-OXYGEN-HALOGEN COMPOUNDS¹

BINARY		TERTIARY			
Fluorides	Chlorides	Bromides and Iodides	Fluorides	Chlorides	Bromides and Iodides
S_2F_2 2	S_2Cl_2 7,8	S_2Br_2 11	SOF_2 12	$SOCl_2$ 12	$SOBr_2$ 12
		(S_2I_2, SI_2) 12	$SOClF$ 13	SO_2ClF 20	(SOI_2) 13
SF_2 3	SCl_2 9		SO_2F_2 14	SO_2Cl_2 21	SO_2Br_2 24
SF_4 4	SCl_4 9		$S_2O_2F_2$ 15	$S_2O_2Cl_2$ 22	(SO_2I_2) 25
SF_6 5			SO_2FCl 16	$S_2O_2Cl_2$ 23	"Bromine and iodine sulfates" 26
S_2F_{10} 6	S_2Cl_{10} 10	S_2Br_{10} 10	$S_2O_3F_2$ 17		
			$S_2F_{10}O_2$ 18		
			F_4SO 19		
			F_2SOF 19		
			$SO_2(O_2F)$ 19		

REFERENCES TO TABLE I

- (1). General references: (a) T. Moeller, 'Inorganic Chemistry', John Wiley, New York, 1952 pp. 518 ff.; (b) A.B. Burg, 'Fluorine Chemistry', J.H. Simons, ed., Academic Press, New York, 1950, Vol. I, pp. 77 ff.; (c) D.M. Yost and H. Russell, Jr., 'Systematic Inorganic Chemistry', Prentice-Hall, New York 1944, pp. 295 ff.
- (2). M. Centnerszwer and C. Strenk, Chem. Ber. **58**, 914 (1925); M. Troutz and K. Ehrmann, J. prakt. Chem. **142**, 79 (1935); L.M. Dubnikov and N.I. Zorin, J. Gen. Chem., U.S.S.R., **17**, 185, (1947); Chem. Abstr. **42**, 57 (1936).
- (3). L.M. Dubnikov and N.I. Zorin, reference 2; D. Edelson, C.A. Bieling, and G.T. Kohman, Ind. Eng. Chem. **45**, 2094 (1953).
- (4). F. Brown and P.L. Robinson, J. Chem. Soc. **1955**, 3147; M. Bartlett and P.L. Robinson, Proc. Chem. Soc. **1957**, 230; F. Seel and O. Detmer, Z. anorg. u. allgem. Chem. **301**, 113 (1959).
- (5). W.C. Schumb, Inorg. Synth. **III**, 119 (1950).
- (6). K.G. Denbigh and R. Whytlaw-Gray, J. Chem. Soc. **1934**, 1346; W.R. Trost and R.L. McIntosh, Can. J. Chem. **29**, 508 (1951).
- (7). K.J. Palmer, J. Am. Chem. Soc. **60**, 2360 (1938); M. Schmidt and B. Wirwoll, Z. anorg. u. allgem. Chem. **303**, 177 (1960).
- (8). G.H. Whiting, J. Applied Chem. **2**, 390 (1952).
- (9). T.M. Lowry and G. Jessop, J. Chem. Soc. **1930**, 782.
- (10). F. Fehér and G. Rempe, Z. Naturforsch. **8b**, 688 (1953); F. Fehér, K. Naused, and H. Weber, Z. anorg. u. allgem. Chem. **290**, 303 (1957) and preceding articles.
- (11). F. Fehér, J. Kraemer and G. Rempe, Z. anorg. u. allgem. Chem. **279**, 18, (1955).
- (12). A.R. Vasudeva Murthy, Proc. Indian Acad. Sci. **37A**, 17 (1953); Chem. Abstr. **47**, 12084. As intermediates.
- (13). See text.
- (14). F. Seel and L. Riel, Z. anorg. u. allgem. Chem. **282**, 293 (1955).
- (15). E. Hayek and A. Czaloun, Monatsh. **87**, 790 (1956); E. Hayek, A. Aignesberger, and A. Engelbrecht, Monatsh. **86**, 735 (1955).
- (16a) S.S. Sandhu, B.S. Chakkul, and G.S. Sandhu, J. Indian Chem. Soc. **37**, 329 (1960).

- (16). H.S. Booth and C.V. Hermann, J. Am. Chem. Soc. **58**, 63 (1936); F. Seel and L. Riehl, Z. anorg. u. allgem. Chem. **282**, 293 (1955).
- (17). F.B.D. Dudley and G.H. Cady, J. Am. Chem. Soc. **79**, 513 (1957).
- (18). R.B. Harvey and S.H. Bauer, J. Am. Chem. Soc. **76**, 859 (1954).
- (19). (a) F.B. Dudley, G.H. Cady, and D.F. Eggers, J. Am. Chem. Soc. **78**, 290, 1553 (1956); (b) F.B. Dudley, J.N. Shoolery, and G.H. Cady, J. Am. Chem. Soc. **78**, 568 (1956); (c) F. Seel and O. Detmer, Z. anorg. u. allgem. Chem. **301**, 113 (1959).
- (20). H.S. Booth and C.V. Hermann, J. Am. Chem. Soc. **58**, 63 (1936).
- (21). H.R. Allen and R.H. Maxson, Inorg. Synth. **I**, 114 (1939).
- (22). M. Sveda, Inorg. Synth. **III**, 124 (1950).
- (23). H.A. Lehmann and G. Ladwig, Chem. Tech. **5**, 455 (1953).
- (24). H. Jonas, Z. anorg. u. allgem. Chem. **265**, 273 (1951).
- (25). M.R.A. Rao, Proc. Indian Acad. Sci. **12A**, 354 (1940); Chem. Abstr. **35**, 2807 (1941).
- (26). I. Wasson, J. Chem. Soc. **1938**, 1708; I. Wasson and C. Arguement, J. Chem. Soc. **1938**, 1702; M.T. Kikindai, Compt. rend. **240**, 873 (1955); A.A. Woolf, Chemistry and Industry, **1954**, 1320.

References

- 1). D. M. Yost and H. Russell, Jr., "Systematic Inorganic Chemistry," Prentice-Hall, Inc., New York, p. 307 (1944).
- 2). H. Siebert, Z. anorg. u. allgem. Chem. **274**, 24 (1953).
- 3). K. Venkateswarlu and S. Sundaram, J. chim. phys. **54**, 202 (1957); CA **51**, 12633.
- 4). K. C. Schreiber, Anal. Chem. **21**, 1168 (1949).
- 5). D. E. Martz and R. T. Lagemann, J. Chem. Phys. **22**, 1193 (1954).
- 5). R. N. Haszeldine and J. M. Kidd, J. Chem. Soc. 1955, 2901.
- 7). M. Goehring, Chem. Ber. **80**, 2 19 (1947).
- 3). R. Vogel-Högler, Acta Phys. Auriaca **1**, 323 (1948); Chem. Abst. **42**, 6663 (1948).
- 9). C. A. McDowell, Trans. Faraday Soc. **49**, 371 (1953).
- 9). G. Allen and C. A. McDowell, J. Chem. Phys. **23**, 209 (1955).
- 1). H. P. Koch, J. Chem. Soc. 1949, 394.
- 2). D. Voigt, Ann. chim. (12) **4**, 393 (1949); Chem. Abst. **45**, 4980 (1939).
- 3). S. S. Dharmatti, Proc. Indian Acad. Sci. **13A**, 359 (1941); Chem. Abst. **36**, 361 (1942).
- 4). M. Beyaert and F. Govaert, Natuurw. Tijdschr. **20**, 119 (1938); Chem. Abst. **32**, 6517 (1938).
- 5). I. E. Coop and L. E. Sutton, Trans. Faraday Soc. **35**, 505 (1939).
- 5). H. Balczó, Mitt. chem. Forsch.-Insts. Wirtsch. Österr. **7**, 104 (1953); Chem. Abst. **48**, 5013 (1954).
- 7). H. Spaniau and E. Brunneck, Z. anorg. u. allgem. Chem. **270**, 201 (1952).
- 3). B. J. Masters, N. D. Potter, D. R. Asher, and T. H. Norris, J. Am. Chem. Soc. **78**, 4252 (1956).
- 9). L. E. D. Pease, Jr., and W. E. Luder, J. Am. Chem. Soc. **75**, 5195 (1953).
- 9). V. Gutmann, Monatsh. **85**, 286 (1954).
- 9). E. Neale and L. T. D. Williams, J. Chem. Soc. 1954, 2156.
- 9). M. Persoz, Compt. rend. **28**, 86 (1849); L. Carius, Ann. **70**, 297 (1849).
- 9). The early chemistry of SOCl_2 is excellently summarized by C. A. Silberrad, Chemistry and Industry, **45**, 36, 55 (1926); Cf. also R. Cherton, Bull. soc. roy. sci. Liege **11**, 54 (1942); Chem. Abst. **38**, 3209 (1944). Cf. also, "Interscience Encycl. of Chem. Tech." Vol. 13, pp. 406-410 (1954).
- 9). J. P. Edwards (to Hooker Electrochem. Co.), U.S. 2,362,057, Nov. 7, 1944; Chem. Abst. **39**, 2386 (1945); Farbenfabriken Bayer A.-G. (H. Jonas and P. Lueg, inventors) Ger. 939,571, Feb. 23, 1956; Chem. Abst. **52**, 14113 (1958).
- 9). A. Pechukas (to Pittsburgh Plate Glass Co.), U.S. 2,431,823, Dec. 2, 1947; Chem. Abst. **42**, 2409 (1948).
- 9). E. F. Fricke (to Allied Chem. & Dye Corp.), U.S. 2,471,946, May 31, 1949; Chem. Abst. **42**, 7201 (1948).
- 9). Badische Anilin- & Soda- Fabrik (by K. Wintersberger and G. Koudela), Ger. 803,411, Apr. 2, 1951; Chem. Abst. **45**, 6357 (1951).
- 9). A. T. Hallowell and G. T. Vaala (to E. I. du Pont de Nemours & Co.), U.S. 2,393,247, Jan. 22, 1946; Chem. Abst. **40**, 1980 (1946).
- 9). C. E. Frank, A. T. Hallowell, C. W. Theobald, and G. T. Vaala, Ind. Eng. Chem. **41**, 2061 (1949).
- 9). A. G. Evans and G. W. Meadows, Trans. Faraday Soc. **43**, 667 (1947).
- 9). W. H. Salzenberg and M. Sveda (to E. I. du Pont de Nemours & Co.), U.S. 2,420,623, May 13, 1947; Chem. Abst. **41**, 5696 (1947).
- 9). R. Daudel, Bull. soc. chim. France **19**, D23 (1952); R. Muxart, P. Daudel and B. Boscardin, J. chim. phys. **46**, 466 (1949); Chem. Abst. **44**, 2831 (1950).
- 9). A. B. Burg and E. R. Birnbaum, J. Inorg. & Nuclear Chem. **7**, 146 (1958); Cf. H. Hecht, R. Greese, and G. Jander, Z. anorg. u. allgem. Chem. **269**, 265 (1952).
- 9). T. Harrington and T. H. Boyd, J. Soc. Chem. Ind. **64**, 209 (1945); Chem. Abst. **39**, 5206 (1945).
- 9). L. F. Jonson, Jr., and T. H. Norris, J. Am. Chem. Soc. **79**, 1584 (1957); D. E. Burge and T. H. Norris, J. Am. Chem. Soc. **81**, 2324, 2329 (1959).

- (36). M. J. Frazer, J. Chem. Soc. 1957, 3319.
- (37). R. A. Hubbard, 2nd. and W. F. Luder, J. Am. Chem. Soc. 73, 1327 (1951).
- (38). H. Hecht, G. Jander, and H. Schlapmann, Z. anorg. Chem. 254, 255 (1947).
- (39). H. Hecht, Z. anorg. Chem. 254, 37 (1947).
- (40). L. E. Simerl (to U.S.A.), U.S. 2,681,307, June 15, 1954; Chem. Abst. 48, 10959
- (41). P. W. Schenk, Ber. 75, 94 (1942).
- (42). J. H. Freeman and C. E. C. Richards, J. Inorg. and Nuclear Chem. 7, 287 (1958).
- (43). P. D. Bartlett and H. F. Herbrandson, J. Am. Chem. Soc. 74, 5971 (1952).
- (44). W. Gerrard and B. D. Shepherd, J. Chem. Soc. 1953, 2069.
- (45). G. Fodor and J. Kiss, J. Chem. Soc. 1952, 1589.
- (46). D. J. Cram, J. Am. Chem. Soc. 75, 332 (1953).
- (47). D. Y. Curtin and D. B. Kellom, J. Am. Chem. Soc. 75, 6011 (1953).
- (48). C. C. Lee and J. W. T. Spinks, Can. J. Chem. 32, 1005 (1954).
- (49). J. D. Roberts and R. H. Mazur, J. Am. Chem. Soc. 73, 2509 (1951).
- (50). H. C. Stevens and O. Grummitt, J. Am. Chem. Soc. 74, 4876 (1952).
- (51). F. F. Caserio, G. E. Dennis, R. H. DeWolfe and W. G. Young, J. Am. Chem. Soc. 77, 4182 (1955).
- (52). J. Sicher and M. Pankova, Chem. Listy 49, 1336 (1955).
- (53). H. B. Henbest and R. A. L. Wilson, J. Chem. Soc. 1956, 3289.
- (54). R. E. Ireland, T. I. Wrigley and W. G. Young, J. Am. Chem. Soc. 80, 4604 (1958).
- (55). W. Gerrard and A. M. Thrush, J. Chem. Soc. 1952, 741; Cf. M. Y. Kraft and V. V. Katyashkina, Doklady Akad. Nauk S.S.S.R. 109, 312 (1956), Proc. Acad. Sci. U.S.S.R., Sect. Chem. 109, 381 (1956); Chem. Abst. 51, 1832 (1957).
- (56). H. H. Bosshard, R. Mory, M. Schmid, and H. Zollinger, Helv. Chim. Acta 42, 1653 (1959).
- (57). J. Bello, Biochim. Biophys. Acta 20, 426 (1956).
- (58). V. G. Kulkarni and G. V. Jadhav, J. Karnatak Univ. 1, 19 (1956); Chem. Abst. 52, 8107 (1958).
- (59). G. Lloyd, W. B. Whalley and R. B. Travers, J. Chem. Soc. 1956, 3209.
- (60). A. Pechukas (to Pittsburgh Plate Glass Co.), U.S. 2,576,138, Nov. 27, 1951; Chem. Abst. 46, 5615 (1952).
- (61). T. L. Gresham, J. E. Jansen and F. W. Shaver, J. Am. Chem. Soc. 72, 72 (1950).
- (62). V. I. Lutkova, N. I. Kutsenko, and M. I. Itkina, Zhur. Obshchei Khim. 25, 2101 (1955); Chem. Abst. 50, 8584 (1956).
- (63). N. V. Shorygina, Zhur. Obshchei Khim. 26, 1460 (1956); Chem. Abst. 50, 14673 (1956).
- (64). K. B. Wiberg, J. Am. Chem. Soc. 74, 3957 (1952).
- (65). D. Libermann, Bull. soc. chim. France 1951, C141.
- (66). C. Courtot and N. Kozertchouk, Compt. rend. 218, 973 (1944).
- (67). S. Sugawara and K. Sakurai, J. Pharm. Soc. Japan 60, 22 (1940); Chem. Abst. 34, 3704.
- (68). J. Kloss, Arch. Pharm. 286, 348 (1953).
- (69). V. J. Dalvi and G. V. Jadhav, J. Indian Chem. Soc. 33, 440 (1956); Chem. Abst. 51, 2625 (1957).
- (70). V. G. Kulkarni and G. V. Jadhav, J. Indian Chem. Soc. 33, 266, 519, 818 (1956); 34, 324 (1957); Chem. Abst. 51, 1068, 10411 (1957); 52, 2796, 9984 (1958).
- (71). Ibid., p. 738; Chem. Abst. 51, 8706 (1957).
- (72). D. G. Kundiger, H. Pledger and L. E. Ott, J. Am. Chem. Soc. 77, 6659 (1955).
- (73). C. Kimura and K. Takiura, Ann. Rept. Takeda Research Lab. 8, 47 (1949); Chem. Abst. 47, 6359 (1953).
- (74). V. Horak, Chem. Listy, 43, 209 (1949); Chem. Abst. 45, 542 (1951).
- (75). D. Buckley and H. B. Henbest, J. Chem. Soc. 1956, 1888.
- (76). A. E. A. Werner and J. Gray, Sci. Proc. Roy. Dublin Soc. 24, 111 (1946); Chem. Abst. 41, 710 (1947).
- (77). H. S. Gable and G. W. Kerr, U.S. 2,729,637, Jan. 3, 1956; Chem. Abst. 50, 10766 (1956).

- (78). P. W. Kent, Biochem. J. 55, 361 (1953).
- (79). W. E. Truce, G. H. Birum and E. T. McBee, J. Am. Chem. Soc. 74, 3594 (1952).
- (80). H. Fürst, J. Lang, and P. P. Rammelt, Chem. Tech. 5, 359 (1953); Chem. Abst. 47, 12726 (1953).
- (81). B. H. Nicolaisen (to Olin Mathieson Chem. Corp), U.S. 2,731,409, Jan. 17, 1956.
- (82). S. Patai and A. Patchornik, J. Am. Chem. Soc. 74, 4494 (1952).
- (83). H. Kano and M. Fujimoto, Pharm. Bull. (Tokyo) 5, 389 (1951); Chem. Abst. 52, 5416 (1958).
- (84). A. Mustafa and M. E. D. Sobhy, J. Am. Chem. Soc. 77, 5124 (1955).
- (85). A. Schonberg and W. Asker, J. Chem. Soc. 1942, 725.
- (86). A. Mustafa, W. Asker, and M. E. D. Sobhy, J. Org. Chem. 25, 1519 (1960).
- (87). A. Corbellini, C. Ghioldi, and F. Chevillard, Gazz. chim. ital., 69, 291 (1939); Chem. Abst. 33, 8603 (1939).
- (88). R. E. Lyle and G. G. Lyle, J. Org. Chem. 18, 1058 (1953).
- (89). N. Tokura, R. Asamai, and R. Tada, Sci. Repts. Research Insts., Tohoku Univ., Ser. A 8, 149 (1956); Chem. Abst. 51, 4944. Chem. Abst. 51, 4944 (1957).
- (90). H. E. Fierz-David, L. Blangey and E. Merian, Helv. Chim. Acta, 34, 846 (1951) (Leading reference).
- (91). F. F. Blicke and H. C. Godt, J. Am. Chem. Soc. 76, 2798 (1954).
- (92). V. G. Pesin, A. M. Khaletskii, and C. Chao, Zhur. Obshchei Khim. 27, 1570 (1957); Chem. Abst. 52, 3790 (1958).
- (93). H. F. Herbrandson, R. T. Dickerson, Jr., and J. Weinstein, J. Am. Chem. Soc. 78, 2576 (1956).
- (94). F. Sörm, J. Smrt, and J. Beranek, Chem. Listy 42, 573 (1955); Chem. Abst. 50, 2429 (1956).
- (95). L. Anschütz and H. Wirth, Naturwissenschaften 43, 59 (1956).
- (96). Ibid., p. 16.
- (97). R. Graf, Ann. 578, 50 (1952).
- (98). E. Riesz, Rev. facultad humanidad y cienc. (Montevideo) 2, No. 3, 65 (1948); Chem. Abst. 43, 1744 (1949).
- (99). G. Machell, Chem. Products 19, 307, 356 (1956); Chem. Abst. 50, 14501, 15409 (1956).
- (100). "Encyclopedia of Chemical Technology," Vol. 13, Interscience Publishers, New York p. 406 (1954).
- (101). H. S. Booth and F. C. Mericola, J. Am. Chem. Soc. 62, 640 (1940).
- (102). D. M. Yost, Proc. Indian Acad. Sci. 8A, 333 (1938); Chem. Abst. 33, 4873 (1939).
- (103). P. Bender and J. M. Wood, Jr., J. Chem. Phys. 23, 1316 (1955).
- (104). J. K. Chloane and M. K. Wilson, J. Chem. Phys. 23, 1313 (1955).
- (105). R. C. Ferguson, J. Am. Chem. Soc. 76, 850 (1954).
- (106). H. Moissan and P. Lebeau, Compt. rend., 130, 1436 (1900).
- (107). W. Steinkopf and J. Herold, J. prakt. Chem. 101, 79 (1920).
- (108). K. Wiechert, Z. anorg. Chem. 261, 310 (1950).
- (109). F. Seel and L. Riehl, Z. anorg. u. allgem. Chem. 282, 293 (1955).
- (110). F. Seel, H. Jonas, L. Riehl, and J. Langer, Angew. Chem. 67, 32 (1955).
- (111). O. Ruff and K. Thiel, Ber., 38, 549 (1905).
- (112). H. C. Clark and H. J. Emeleus, J. Chem. Soc. 1958, 190.
- (113). C. W. Tullock and D. D. Coffmann, J. Org. Chem. 25, 2016 (1960).
- (114). U. Wannaga and F. Mennicken, Z. anorg. u. allgem. Chem. 278, 310 (1955).
- (115). H. Jonas, Z. anorg. u. allgem. Chem. 265, 273 (1951).
- (116). H. A. Hayes and J. R. Partington, J. Chem. Soc. 1926, 2594.
- (117). H. Stammreich, J. Chem. Phys. 25, 1277 (1956).
- (118). H. Hibbert and J. C. Pullman, "Inorganic Syntheses." Vol. I, H.S. Booth, ed., McGraw-Hill Book Co., N.Y. p. 113.
- (119). M. J. Frazer and W. Gerrard, Chemistry and Industry, 1954, 280.
- (120). R. C. Elderfield et al., J. Am. Chem. Soc. 68, 1579 (1946).

- (121). D. E. Ames and R. E. Bowman, J. Chem. Soc. 1950, 406.
- (122). M. J. Frazer, W. Gerrard, G. Machell, and B. D. Shepherd, Chemistry and Industry, 1954, 931.
- (123). M. J. Frazer and W. Gerrard, Research Correspondence, Suppl. to Research (London) 7, No. 4, 827-8 (1954); Chem. Abst. 49, 8962 (1955); J. Chem. Soc. 1955, 3624.
- (124). R. E. Johnson, T. H. Norris, and J. L. Huston, J. Am. Chem. Soc. 73, 3052 (1951).
- (125). R. H. Herber, T. H. Norris, and J. L. Huston, J. Am. Chem. Soc. 76, 2015 (1954).
- (126). B. J. Masters and T. H. Norris, J. Am. Chem. Soc. 77, 1346 (1955).
- (127). M. R. A. Rao, Proc. Indian Acad. Sci. 11A, 185, 201 (1940); Chem. Abst. 34, 7200 (1940).
- (128). H. Spandau and E. Brunneck, Z. anorg. u. allgem. Chem. 278, 197 (1955).
- (129). A. R. V. Murthy, Proc. Indian Acad. Sci. 37A, 11, 17 (1953); Chem. Abst. 47, 12084 (1953).
- (130). W. B. Brownell and L. C. King, J. Am. Chem. Soc. 71, 2926 (1949).