

AD623839

TECHNICAL REPORT 31

FORMATION OF H^- BY ELECTRON IMPACT ON H_2 AT LOW ENERGY

G. J. Schulz
R. K. Asundi

CLEARINGHOUSE FOR FEDERAL SCIENTIFIC AND TECHNICAL INFORMATION			
Hardcopy	Microfiche		
\$ 1.00	\$ 0.50	13	02
ARCHIVE COPY			

Page 1

ARPA Order Number: 125-62 (Amd. 11)

Contract Number: NONR-2584(00)

Project Number: 4720

Atomic and Molecular Sciences
Research and Development
Westinghouse Research Laboratories
Pittsburgh, Pennsylvania 15235

October 29, 1965

This research is a part of Project DEFENDER sponsored by the Advanced
Research Projects Agency, Department of Defense.

TECHNICAL REPORT 31

FORMATION OF H^- BY ELECTRON IMPACT ON H_2 AT LOW ENERGY

G. J. Schulz
R. K. Adundi

ARPA Order Number: 125-63 (Amd. 11)

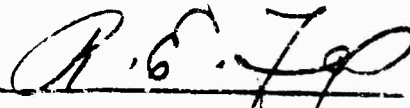
Contract Number: NONR-2584(00)

Project Number: 4720

Atomic and Molecular Sciences
Research and Development
Westinghouse Research Laboratories
Pittsburgh, Pennsylvania 15235

October 29, 1965

This research is a part of Project DEFENDER, sponsored by the Advanced
Research Projects Agency, Department of Defense.

Approved 
L. E. Fox, Director
Atomic and Molecular Sciences
Research and Development

October 29, 1965

Formation of H^- by Electron Impact on H_2 at Low Energy^{*}

G. J. Schulz and R. K. Asundi
Westinghouse Research Laboratories
Pittsburgh, Pa. 15235

This letter reports the formation of the atomic hydrogen negative ion by electron impact on H_2 very close to the theoretical threshold for the production of H^- and H with zero kinetic energy via the process $e + H_2 \rightarrow H^- + H$. We find H^- formation with a very sharp onset at an electron energy of 3.73 ± 0.07 eV and a very steeply rising, albeit small, cross section which peaks very close to its onset. This observation leads to a better understanding of the potential-energy curves of the H_2^- system and has a bearing on the interpretation of vibrational excitation in H_2 .

Dissociative attachment by electron impact on H_2 has been the subject of repeated studies, both without^{1,2} and with³ mass spectrometry.

1. D. Rapp, T. E. Sharp, D. D. Briglia, Phys. Rev. Letters 14, 533 (1965).
2. G. J. Schulz, Phys. Rev. 113, 816 (1959).
3. V. I. Khvostenko and V. M. Dukel'nikii, JETP 6, 657 (1958).

* This work has been supported in part by the Advanced Research Projects Agency through the Office of Naval Research.

The identity of the negative ions as well as the energy dependence of the cross section are rather well established^{1,2} above electron energies of about 6 eV. One observes a broad peak in the energy range about 11 eV with a cross section in the range 1.3 to 1.5×10^{-20} cm². The first value is from Ref. (1) and the second Ref. (2); the agreement in this energy range is therefore extremely good. A sharply rising peak in H⁻ formation at an energy 14.0 - 14.2 eV has a cross section^{1,2} in the range 2.1 to 3.5×10^{-20} cm². In the energy range 6 to 13.6 eV, the formation of H⁻ proceeds via a repulsive energy curve, leading to H⁻ ions and H atoms with kinetic energy. In the region of the peak near 14 eV, hydrogen atoms and H⁻ with low kinetic energy are produced, the hydrogen atoms being in the $n = 2$ excited states.

In previous experiments on H⁻ formation in this¹ and in other laboratories,⁴ a small negative signal was observed at an energy about 3.7 eV, but these results were not published because doubts persisted whether scattered electrons falsify the collected currents at these low energies. It should be noted that the early theory⁵ predicted that no state of H₂⁻ could lie at these low energies. The predictions of this theory have recently been found to be in error.^{6,7} Curran,

4. D. Rapp and D. D. Briglia, Lockheed Report LMSC6-74-64-40, state that the "peak at 3.9 eV in H₂ may possibly be caused by something other than negative ions."
5. H. Eyring, J. O. Hirschfelder and H. S. Taylor, J. Chem. Phys. 4, 479 (1936) found that the lowest lying potential energy curve of H₂ traverses the Franck-Condon region at an energy about 5 eV.
6. H. S. Taylor and F. E. Harris, J. Chem. Phys. 39, 1012 (1963).
7. Yu. N. Demkov, JETP 19, 762 (1964).

working in this laboratory,⁸ used a mass spectrometer to identify the 3.7 eV process as the formation of H^- from H_2 but was unable to establish that the H^- current is proportional to gas pressure, as required for a two body collision process.

Experiment

An electron beam, aligned by a magnetic field of about 150 gauss, traverses a differentially pumped collision chamber where the ions are produced by dissociative attachment. The effective energy distribution of the electron beam is reduced by use of the retarding potential difference method.⁹ The ions produced in the collision chamber are expelled by a repeller which is usually operated about 2.7 volt negative with respect to the collision chamber and are analyzed by a 90 degree magnetic mass spectrometer equipped with an electron multiplier. Since the collision chamber, as well as the mass spectrometer, are differentially pumped, it is possible to reach pressures up to about 0.1 Torr in the collision chamber without appreciably scattering ions in the mass spectrometer focussing and analyzer regions.

The pressure in the collision chamber is determined by measuring the total saturated ion current at an electron energy of 70 eV on the repeller electrode; the pressure is then calculated using the known ionization cross section.

8. R. K. Curran, private communication.

9. R. E. Fox, W. M. Hickam, D. J. Grove, and J. Kjellaas, Rev. Sci. Instr. 26, 1101 (1955).

Results in Hydrogen

Figure 1 shows a plot of the H^- ion current as a function of electron energy. A retarding curve on the electrons, shown by the dashed curve in Fig. 1, determines the energy scale.¹⁰ The onset of the negative ion current is determined from the high energy tail of the retarding curve, and we obtain a value of 3.73 eV for the onset. The peak of the negative ion curve is determined from the half-power point of the retarding curve and the mean of 10 determination gives a value 3.75 ± 0.07 eV. These results indicate that the H^- curve rises more steeply than can be determined from the presently available energy distribution, and the data, therefore, are not inconsistent with a much more steeply rising excitation function at threshold. Using accepted values for the dissociation energy, D , of H_2 and the electron affinity, A , for H , the onset of H^- with zero kinetic energy occurs at an energy $(D-A) = (4.48 - 0.75) = 3.73$ eV, in agreement with the onset observed in the present experiment.

The peak current at 3.75 eV is found to be 8% of the peak current at 14 eV. Since both these processes lead to H^- with essentially zero kinetic energy, kinetic energy discrimination should be absent, and one can obtain a cross section for the 3.75 eV process by normalizing to the previously measured cross section at the 14.0 eV peak. Depending on the value one chooses to accept for the cross section^{1,2} at 14 eV, one obtains

10. When a negative potential with respect to the collision chamber is applied to the repeller--as is customary in negative ion studies--then the retarding curve on the electrons yields the correct energy scale calibration. [G. J. Schulz, J. Appl. Phys. 31, 1134 (1960)].

a value of 1.6×10^{-21} or $2.6 \times 10^{-21} \text{ cm}^2$ for the cross section at the 3.75 eV peak. It should be noted that, since the actual energy dependence near the peak could not be traced out in the present experiment because of insufficient energy resolution, the actual cross section at 3.75 eV is probably somewhat larger.

Figure 2 shows the pressure dependence of the ion currents at an electron energy, W , of 3.75 and 14.0 eV showing that both processes are linear with pressure indicating that we are dealing with two-body processes.

Isotope Effect

We have attempted to observe the isotope effect by measuring the cross section of D^- formation from D_2 in the energy range of Figure 1; in our experiments to date, we were unable to observe a D^- signal around 3.8 eV although we could observe a D^- signal at 14 eV (where the cross section for D^- formation is $1 \times 10^{-20} \text{ cm}^2$) with a signal to noise ratio of more than 100. This observation places an upper limit of $1 \times 10^{-22} \text{ cm}^2$ on the D^- cross section in 3.8 eV range. An expression for the negative ion formation via a compound state, where "backward decay" is present, is given^{7,11} by $Q_- = Q_0 \exp(-2 \bar{\Gamma} \tau / \hbar)$ here Q_- is the cross section for negative ion formation, Q_0 is the cross section for formation of the compound state (H_2^-), $\bar{\Gamma}$ is the mean width of the potential energy curve and τ is the time for the atoms to move to a distance where they are stabilized. This stabilization time for D^- is $\sqrt{2}$ times the stabilization time for H^- ; from these considerations we arrive at a limit $2 \bar{\Gamma} \tau \geq 3 \times 10^{-14} \text{ eV sec.}$ Assuming τ to be of the order of a vibration time, $\tau \sim 10^{-14} \text{ sec.}$, we obtain

11. T. D. Holstein, Conference on Gaseous Electronics, Schenectady, N.Y. (1951).

$\bar{\Gamma} \gg 1.5$ eV. Although this width seems large, it is not inconsistent with the hypothesis that the same compound state of H_2^- leads to vibrational excitation.¹² The large width of the H_2^- state causes the vibrational cross section to be broad and without appreciable structure, in agreement with previous experimental observations.^{13,14} The large isotope effect observed in the present work at low energy (~ 3.8 eV) is in striking contrast with the small isotope effect previously reported¹ in the energy range 7-18 eV. It indicates that the level width $\bar{\Gamma}$ is considerably smaller for the high energy process of negative ion formation.

Conclusions

The fact that we observe H^- negative ion formation at 3.7 eV indicates that a potential energy curve traverses the Franck-Condon region at this energy. The large isotope effect observed in the present experiment leads us to the conclusion that the state involved here is short lived and therefore the potential energy "curve" is broad. This state may be responsible¹² not only for the H^- production observed in this experiment but also for vibrational excitation of the H_2 molecule observed previously^{13,14,15}

12. J. N. Bardsley, A. Herzenberg, F. Mandl, IVth International Conference on the Physics of Ionized Gases (Science Bookcrafters, New York) 1965, p. 359, and private communication.

13. G. J. Schulz, Phys. Rev. 135, A988 (1964).

14. A. G. Engelhardt and A. V. Phelps, Phys. Rev. 131, 2115 (1963).

15. It should be noted, however, that K. Takayanagi, J. Phys. Soc. Japan 20, 562 (1965) calculates the vibrational excitation in H_2 without invoking a compound state and obtains agreement with experiment. The two viewpoints, namely "direct" excitation of vibration with the polarization force dominant, and the "compound state" viewpoint described above may not be irreconcilable.

as well as the peak in the elastic cross section¹⁶ around 2 eV. These observations are substantially in agreement with the viewpoint of Taylor and Harris⁶ and of Bardsley, Herzenberg and Mandl¹² who point out that a short-lived state of H_2^- traverses the Franck-Condon region at energies below the $H^- + H$ dissociation limit and that a small portion of this curve extends above this limit.

Acknowledgment

The authors wish to express their thanks to A. V. Phelps and P. J. Chantry for many helpful comments.

16. R. Kollath, *Physikalische Zeitschrift* 31, 925 (1930) and H.S.W. Massey and E.H.S. Burhop, *Electronic and Ionic Impact Phenomena*, Clarendon Press, Oxford, 1952, p. 215.

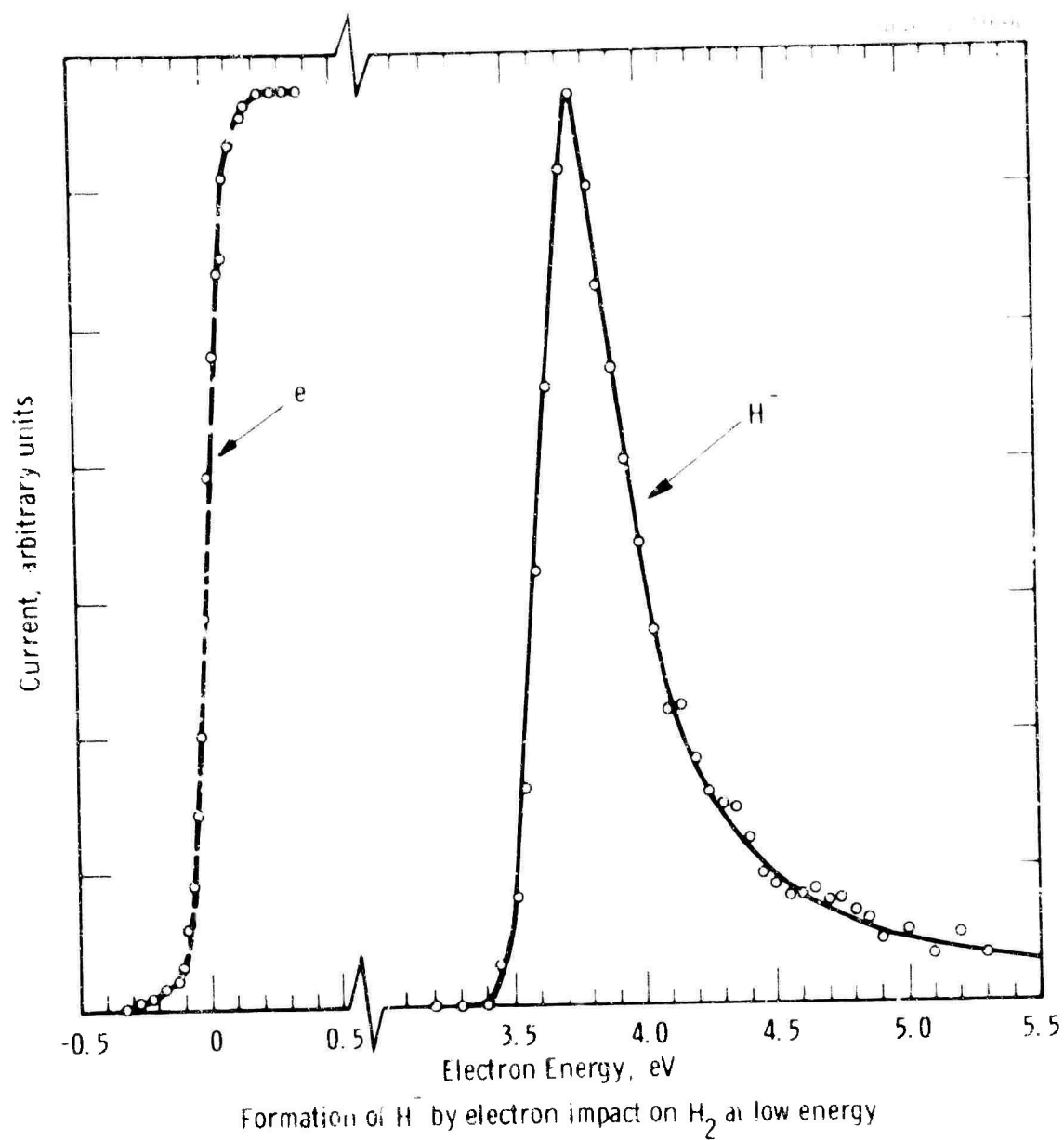


Fig. 1 Energy dependence of H^- formation from H_2 at low energies. The dashed curve is the electron retarding curve used to obtain the energy scale. The onset of H^- formation occurs at 3.73 ± 0.07 eV and the peak H^- current occurs at 3.75 ± 0.07 eV. The value of the cross section at 3.75 eV is about $2 \times 10^{-21} \text{cm}^2$. See text.

Curve 575395-A

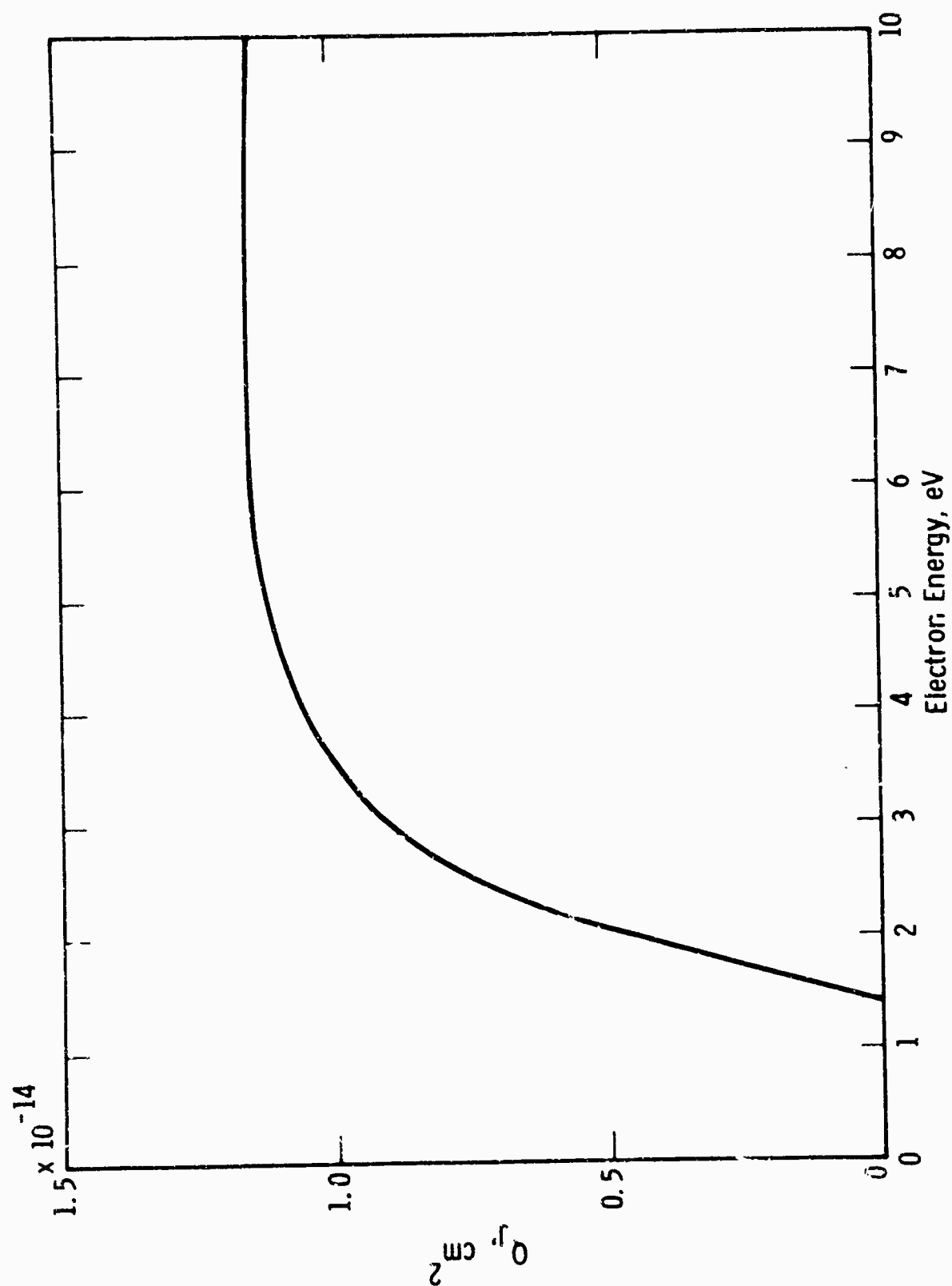


Fig. 2. Pressure dependence of two peaks involving H^- formation from H_2 with zero kinetic energy. The ratio of the ion peak at an electron energy, W , of 14 eV to that at $W = 3.75$ eV is 12.5.

DOCUMENT CONTROL DATA - R&D

(Security classification of title, body of abstract and indexing annotation must be entered when the overall report is classified)

1 ORIGINATING ACTIVITY (Corporate author) Westinghouse Research Laboratories Pittsburgh, Pennsylvania 15235		2a REPORT SECURITY CLASSIFICATION UNCLASSIFIED	
		2b GROUP	
3 REPORT TITLE FORMATION OF H^- BY ELECTRON IMPACT ON H_2 AT LOW ENERGY			
4 DESCRIPTIVE NOTES (Type of report and inclusive dates) Technical Report 31			
5 AUTHOR(S) (Last name, first name, initial) Schulz, G. J., Dr. PI Asundi, R. K., Dr.			
6 REPORT DATE October 29, 1965	7a TOTAL NO OF PAGES 9	7b NO OF REFS 16	
8a CONTRACT OR GRANT NO. NONR-2584(00)	9a ORIGINATOR'S REPORT NUMBER(S) 65-9E3-113-P3		
b. PROJECT NO 4720			
c	9b OTHER REPORT NO(S) (Any other numbers that may be assigned this report)		
d			
10 AVAILABILITY/LIMITATION NOTICES			
11 SUPPLEMENTARY NOTES 2 figures		12 SPONSORING MILITARY ACTIVITY ARPA Department of Defense Washington 25, D. C.	
13 ABSTRACT This letter reports the formation of the atomic hydrogen negative ion by electron impact on H_2 very close to the theoretical threshold for the production of H^- and H with zero kinetic energy via the process $e + H_2 \rightarrow H^- + H$. We find H^- formation with a very sharp onset at an electron energy of 3.73 ± 0.07 eV and a very steeply rising, albeit small, cross section which peaks very close to its onset. This observation leads to a better understanding of the potential-energy curves of the H_2^- system and has a bearing on the interpretation of vibrational excitation in H_2 . Dissociative attachment by electron impact on H_2 has been the subject of repeated studies, both without 1,2 and with mass spectrometry.			

KEY WORDS

LINK A

LINK B

LINK C

ROLE

WT

ROLE

WT

ROLE

WT

hydrogen
deuterium
isotopes
cross sections
negative ions
electrons
vibration
dissociation
electron beams

INSTRUCTIONS

1. **ORIGINATING ACTIVITY:** Enter the name and address of the contractor, subcontractor, grantee, Department of Defense activity or other organization (*corporate author*) issuing the report.

2a. **REPORT SECURITY CLASSIFICATION:** Enter the overall security classification of the report. Indicate whether "Restricted Data" is included. Marking is to be in accordance with appropriate security regulations.

2b. **GROUP:** Automatic downgrading is specified in DoD Directive 5200.10 and Armed Forces Industrial Manual. Enter the group number. Also, when applicable, show that optional markings have been used for Group 3 and Group 4 as authorized.

3. **REPORT TITLE:** Enter the complete report title in all capital letters. Titles in all cases should be unclassified. If a meaningful title cannot be selected without classification, show title classification in all capitals in parentheses immediately following the title.

4. **DESCRIPTIVE NOTES:** If appropriate, enter the type of report, e.g., interim, progress, summary, annual, or final. Give the inclusive dates when a specific reporting period is covered.

5. **AUTHOR(S):** Enter the name(s) of author(s) as shown on or in the report. Enter last name, first name, middle initial. If military, show rank and branch of service. The name of the principal author is an absolute minimum requirement.

6. **REPORT DATE:** Enter the date of the report as day, month, year, or month, year. If more than one date appears on the report, use date of publication.

7a. **TOTAL NUMBER OF PAGES:** The total page count should follow normal pagination procedures, i.e., enter the number of pages containing information.

7b. **NUMBER OF REFERENCES:** Enter the total number of references cited in the report.

8a. **CONTRACT OR GRANT NUMBER:** If appropriate, enter the applicable number of the contract or grant under which the report was written.

8b, 8c, & 8d. **PROJECT NUMBER:** Enter the appropriate military department identification, such as project number, subproject number, system numbers, task number, etc.

9a. **ORIGINATOR'S REPORT NUMBER(S):** Enter the official report number by which the document will be identified and controlled by the originating activity. This number must be unique to this report.

9b. **OTHER REPORT NUMBER(S):** If the report has been assigned any other report numbers (*either by the originator or by the sponsor*), also enter this number(s).

10. **AVAILABILITY LIMITATION NOTICES:** Enter any limitations on further dissemination of the report, other than those

imposed by security classification, using standard statements such as:

- (1) "Qualified requesters may obtain copies of this report from DDC."
- (2) "Foreign announcement and dissemination of this report by DDC is not authorized."
- (3) "U. S. Government agencies may obtain copies of this report directly from DDC. Other qualified DDC users shall request through _____."
- (4) "U. S. military agencies may obtain copies of this report directly from DDC. Other qualified users shall request through _____."
- (5) "All distribution of this report is controlled. Qualified DDC users shall request through _____."

If the report has been furnished to the Office of Technical Services, Department of Commerce, for sale to the public, indicate this fact and enter the price, if known.

11. **SUPPLEMENTARY NOTES:** Use for additional explanatory notes.

12. **SPONSORING MILITARY ACTIVITY:** Enter the name of the departmental project office or laboratory sponsoring (paying for) the research and development. Include address.

13. **ABSTRACT:** Enter an abstract giving a brief and factual summary of the document indicative of the report, even though it may also appear elsewhere in the body of the technical report. If additional space is required, a continuation sheet shall be attached.

It is highly desirable that the abstract of classified reports be unclassified. Each paragraph of the abstract shall end with an indication of the military security classification of the information in the paragraph, represented as (TS), (S), (C), or (U).

There is no limitation on the length of the abstract. However, the suggested length is from 150 to 225 words.

14. **KEY WORDS:** Key words are technically meaningful terms or short phrases that characterize a report and may be used as index entries for cataloging the report. Key words must be selected so that no security classification is required. Identifiers, such as equipment model designation, trade name, military project code name, geographic location, may be used as key words but will be followed by an indication of technical content. The assignment of links, rules, and weights is optional.