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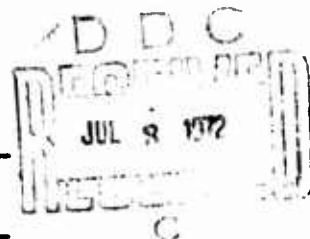
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Computer Modeling and Parametric Study for a Pulsed $H_2 + F_2$ Laser

Prepared by R. L. KERBER, G. EMANUEL, and J. S. WHITTIER
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FOREWORD

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This report, which documents research carried out from December 1970 to June 1971, was submitted 13 March 1972 to Captain Karl J. Hoch, SYAE, for review and approval.

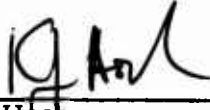
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Approved



W. R. Warren, Jr., Director
Aerodynamics and Propulsion Research
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Publication of this report does not constitute Air Force approval of the report's findings or conclusions. It is published only for the exchange and stimulation of ideas.



Karl J. Hoch
Captain, United States Air Force
Project Officer

ABSTRACT

A computer simulation of a pulsed HF laser pumped by the $\text{H}_2 + \text{F}_2$ chain reaction is presented. A chemical kinetic model encompassing 68 reactions is used to approximate the reacting mixture contained within an optical cavity. For each vibrational level, a Boltzmann distribution for the rotational levels is assumed with lasing on the vibrational band at line center of the transition having maximum gain.

An analysis of cavity and chemical mechanisms yields a simple relationship between the pumping and depletion rates of the vibrational levels. This relationship is used to make a comprehensive study of the effects of cavity and chemical parameters on the laser pulse. The effect of changes in uncertain chemical reaction rate coefficients is assessed. A competition exists between vibration-translation deactivation reactions and stimulated emission. Laser performance is most sensitive to the relative rates of the vibrational-translational and pumping reactions. In addition, the effect of the photon flux on the chemical mechanisms is significant.

CONTENTS

FOREWORD	ii
ABSTRACT	iii
NOMENCLATURE	ix
I. INTRODUCTION	1
II. THE CHEMICAL LASER COMPUTER SIMULATION	5
III. MECHANISMS OF THE PULSED HF LASER	9
A. General Energy Interpretation	9
B. A Detailed Interpretation of Mechanisms for the Standard Case	15
IV. PARAMETRIC VARIATION	23
A. Threshold	23
B. Diluents	26
C. Initial F ₂ Dissociation	29
D. Initial Temperature	32
E. Initial H ₂ Concentration	32
F. Variation of Important Rate Constants	32
V. CHEMICAL EFFICIENCY	37
VI. SUMMARY AND CONCLUSIONS	39
REFERENCES	41
APPENDIX. THE REACTION SYSTEM	A-1

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FIGURES

1.	Power History of the Standard Case	10
2.	Laser Power and Energy Emission Rate	11
3.	Time Histories of the Concentrations of H_2 , F_2 , H, and F for the Standard Case	12
4.	Relationship of Dominant Mechanisms for the 2-1 Band of HF in a Chemical Laser	14
5.	Production and Reduction Mechanisms for the Gain of the 2-1 Line	16
6.	Temperature History of the Standard Case, with and without Radiation Suppressed	18
7.	Gain History for Standard Case with Radiation Suppressed	19
8.	Mechanisms Producing and Reducing F-Atom Concentration for the Standard Case	21
9.	Time Histories of Important Rate Coefficients for the Standard Case	22
10.	Effect of Threshold Gain on Pulse Energy	24
11.	Effect of Threshold Gain on t_1 and $t_{1/2}$	25
12.	Effect of Diluent Concentration on Pulse Energy with Two Different Initial F-Atom Concentrations	27
13.	Effect of Diluent Concentration on $t_{1/2}$ and t_1	28
14.	Effect of Initial F_2 Dissociation on Pulse Length	30
15.	Effect of Initial F_2 Dissociation on Pulse Energy	31

FIGURES (cont.)

16.	Effect of Initial Temperature	33
17.	Effect of Initial H ₂ Concentration	34

TABLES

1.	Effect of Rate Changes on Pulse Characteristics	36
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NOMENCLATURE

$B(v, J)$	Einstein isotropic absorption coefficient
c	speed of light, 2.9979×10^{10} cm/sec
C_{pi}	specific heat at constant pressure of species i
$D_{v+1, v}$	chemical deactivation rate from level $v + 1$ to v
$E_L(t)$	energy emitted during laser pulse to time t
E_J^v	rotational energy associated with vibrational level v and rotational level J
$E_{v+1, v}$	total energy emitted during lasing on a $v + 1 \rightarrow v$ band
h	Planck's constant, 6.6256×10^{-34} joule-sec
H_i	molar enthalpy of species i
J	rotational quantum number
J_L	rotational quantum number associated with the lower level of a lasing transition
k	Boltzmann's constant, 1.3805×10^{-23} joule/ $^{\circ}$ K
k_n, k_{fr}	forward rate coefficients
k_{-n}, k_{br}	backward rate coefficients
L	length of active medium
L_{fr}, L_{br}	forward and backward rates for reaction r
M_j	catalytic species of type j

N_A	Avogadro's number, 6.0225×10^{23} molecules/mole
N_i	molar concentration of species i
$N(v)$	molar concentration of lasing species with vibrational level v
$N(v, J)$	molar concentration of lasing species with vibrational level v and rotational level J
p	pressure
p_i	initial pressure of the reacting mixture
P_L	output power
$P_{v+1, v}$	output power for lasing into level v
$Q_r^v(T)$	rotational partition function for vibrational level v
R_0, R_L	mirror reflectivities
t	time
$t_{1/2}$	time to emit first 1/2 of the total pulse energy
t_1	pulse duration
T	temperature
T_i	initial temperature of the reacting mixture
v	vibrational quantum number
α_{th}	threshold gain
α_{ri}, β_{ri}	stoichiometric coefficients
$\alpha(v, J)$	optical gain for transitions to state (v, J)

η	laser efficiency
κ_v	rate of pumping into vibrational level v
ν_r	reaction rate weighting factor
$\phi(v, J)$	Voigt profile for level (v, J)
ψ_{ch}	net chemical production rate ($\text{mole}/\text{cm}^3\text{-sec}$)
$\psi_{rad}(v, J)$	rate of stimulated emission from $(v + 1, J - 1)$ to (v, J) , ($\text{mole}/\text{cm}^3\text{-sec}$)
$\omega_c(v, J)$	wave number for P-branch transitions into level (v, J)

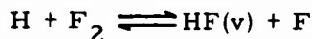
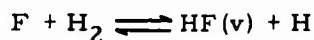
I. INTRODUCTION

Increased interest in conversion of chemical energy into laser power and the scientific value of quantum-chemistry data available from the use of the chemical laser have stimulated recent experimental¹⁻⁴ and theoretical⁴⁻⁷ studies. A current review of chemical laser technology has been made by Dzhidzhov, Platonenko, and Khokhlov.⁸ The multiplicity of phenomena operating in a chemical laser makes quantitative predictions of laser performance or detailed interpretation of experimental data difficult without the aid of comprehensive modeling of these phenomena. Early modeling studies estimated the performance of a pulsed laser by considering the time history of gain on a given band, while the effect of the cavity and induced emission on the chemistry were neglected.^{6,7} More recently, computer simulations have included these effects,^{4,9} and such a code is used in this study. The simulation is used to study the effects of various parameters on the behavior of a pulsed HF laser. The principal objectives of this study are to determine:

1. Relative importance of the chemical kinetic mechanisms involved in the pulsed HF laser
2. Effect of the optical cavity on the reacting $H_2 + F_2$ mixture
3. Effect of various parameters such as the initial concentrations and the initial temperature.

The method of simulation is similar to that used by Airey⁴ for an HCl laser. The major provisions of the laser model are summarized as follows:

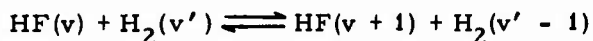
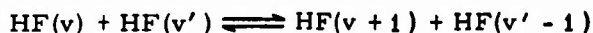
1. The dominant reactions used to represent the chemical processes in the reacting mixture are:
 - a. the $H_2 - F_2$ chain



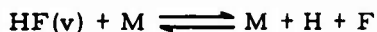
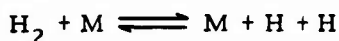
- b. vibrational-translational (VT) deactivation



- c. vibrational-vibrational (VV) quantum exchange



- d. dissociation-recombination



This reaction system is approximated by the 68 reactions listed in the Appendix.

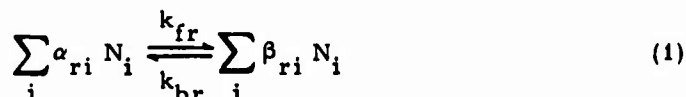
2. The reacting mixture is homogeneous and is contained in a Fabry-Perot cavity.
3. For each vibrational level, the rotational levels have a Boltzmann distribution at the translational temperature.
4. Only one transition within a band can lase at a given time. This is the transition with maximum gain in the band and is always in the P-branch. During lasing on this line, the gain at line center is held constant at its threshold value.
5. Initiation is modeled by the instantaneous introduction of a finite concentration of F atoms into an initial gas mixture of F_2 , H_2 , and Ar diluent, as might be caused by a very fast flashlamp or electrical discharge. A finite time history of F_2 dissociation such as that produced by most initiation methods can be modeled and, indeed, might be of interest for modifying laser performance. Because of the additional parameters involved, such modeling was deemed outside the scope of the present study.

With similar assumptions and a simpler form of the reaction system, an analytical solution for a two-level model of a continuous chemical laser was obtained.⁵ The present study extends that model to its multi-level equivalent and considers numerical solutions for a pulsed HF laser. The simulation is used to investigate the effect of the optical and chemical parameters on the laser's performance.

II. THE CHEMICAL LASER COMPUTER SIMULATION

This section describes briefly the formulation of the chemical laser computer simulation. A detailed description is given in Ref. 9.

The chemical reactions may be written as



where N_i is the molar concentration of species i , α_{ri} and β_{ri} are stoichiometric coefficients, and k_{fr} and k_{br} are forward and backward rate coefficients. (The Nomenclature gives detailed definitions for symbols used in the text.)

The rate of change of concentration $N(v)$ for vibrational level v is given by

$$\frac{dN(v)}{dt} = \psi_{ch}(v) + \psi_{rad}(v, J) - \psi_{rad}(v-1, J_L) \quad (2)$$

where the ψ_{rad} terms are rates of concentration change resulting from lasing into and out of level v . Lower level rotational quantum numbers J and J_L are selected as those giving maximum gain for transitions $v+1 \rightarrow v$ and $v \rightarrow v-1$, respectively. Chemical reactions cause a rate of change $\psi_{ch}(v)$ which is

$$\psi_{ch}(v) = \sum_r (\beta_{ri} - \alpha_{ri})(L_{fr} - L_{br}) \quad (3)$$

where

$$L_{fr} = k_{fr} \prod_j N_j \alpha_{rj}$$

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and

$$L_{br} = k_{br} \prod_j N_j \beta_{rj}$$

The laser cavity is assumed to have a uniform photon flux with active medium length L and mirror reflectivities R_0 and R_L . Lasing initiates for any $v + 1 \rightarrow v$ band when the gain α for the highest gain transition of that band is equal to the threshold gain α_{th} , where

$$\alpha_{th} = -\frac{1}{2L} \ln(R_0 R_L) \quad (4)$$

The gain is assumed to be constant over length L . Only P-branch transitions need be considered; thus, the gain is given by Ref. 9

$$\alpha(v, J) = \frac{hN_A}{4\pi} \omega_c(v, J) \phi(v, J) B(v, J) \left[\frac{2J+1}{2J-1} N(v+1, J-1) - N(v, J) \right] \quad (5)$$

where v and J refer to the lower level of the transition. The wave number of the transition is $\omega_c(v, J)$, and $B(v, J)$ is the Einstein isotropic absorption coefficient based on the intensity. Line-broadening constants and resonance constants used in the Voigt profile at line center $\phi(v, J)$ are those of Ref. 9.

A Boltzmann distribution at the translational temperature is taken for the rotational populations, hence

$$N(v, J) = N(v) \frac{(2J+1)}{Q_r^v(T)} e^{-hcE_J^v/kT} \quad (6)$$

where values of rotational partition function $Q_r^v(T)$ and rotational energy E_J^v are from data of Mann, et al.¹⁰

The energy equation for a constant density gas is written in the form

$$\sum N_i C_{pi} \frac{dT}{dt} - \frac{dp}{dt} = - P_L - \sum \frac{dN_i}{dt} H_i \quad (7)$$

where C_{pi} is the specific heat at constant pressure and H_i the molar enthalpy of species i , p is the pressure, and P_L is the output lasing power per unit volume. This power is given by

$$P_L(t) = \sum_{v=0}^5 h c N_A \omega_c(v, J) \psi_{rad}(v, J) \quad (8)$$

and the pulse energy released in time t is

$$E_L(t) = \int_0^t P_L(t') dt' \quad (9)$$

Chemistry and thermodynamics determine the production of excited species until, at some time t_0 and for some $J = J_0$, the gain on a given vibration-rotation transition reaches α_{th} . At this time, the laser pulse begins. Then Eqs. (2), (4), and (7), along with the state equation, are solved for the temperature, pressure, concentrations, and the J for each band that is lasing as a function of time. During lasing, the value of J may change; this process will be referred to as a "J-shift." Lasing terminates when all gains drop below α_{th} .

III. MECHANISMS OF THE PULSED HF LASER

A. GENERAL ENERGY INTERPRETATION

The initial gas mixture in the laser cavity has a given concentration ratio $F_2:H_2:F:Ar$, where the F atoms represent varying degrees of F_2 dissociation by some initiation scheme. The diluent Ar is used mainly for temperature control. A typical time history of the total pulse power is shown in Fig. 1, and curves depicting rate of laser energy release for the bands lasing are shown in Fig. 2(a). For this case, the initial $F_2:H_2:F:Ar$ is 1:1:0.1:50, where one unit of concentration is equivalent to 9.4072×10^{-7} mole/cc. This unit is used throughout this work. Initial temperature and pressure are 300° K and 1.207 atm, respectively. Mirror reflectivities are 0.8 and 1.0, and the active medium length is 100 cm. This case serves as a standard, and excursions from it comprise the parametric study.

In the course of the reaction, the temperature and the concentrations of HF(v) increase, causing the J value for maximum gain on a given vibrational band to increase. In the calculation when lasing shifts to the next higher J, there is a discontinuous, but small, increase in the rate of photon emission. In Fig. 2(b), the laser energy emission rate for the 2-1 band is shown as a function of time with all the J-shift details included. The J's used to label J-shifts on the 2-1 band are associated with the $v=1$ level. The discontinuities without labels are caused by J-shifts on adjacent bands. For simplicity, discontinuities have been smoothed in Figs. 1 and 2(a).

The following paragraphs give an interpretation of the major mechanisms of the standard case. The initial emission "spike" results from the pumping of HF(v) levels by Reaction 8-10.¹¹ During this spike, approximately 80% of the initial F-atom concentration is consumed (Fig. 3). The increase in the rate of photon emission after the initial spike is due to an increase in temperature and to the increasing importance of Reactions 14-20 as a result of the production of H by Reactions 7-10 and, hence, the operation of the chain.

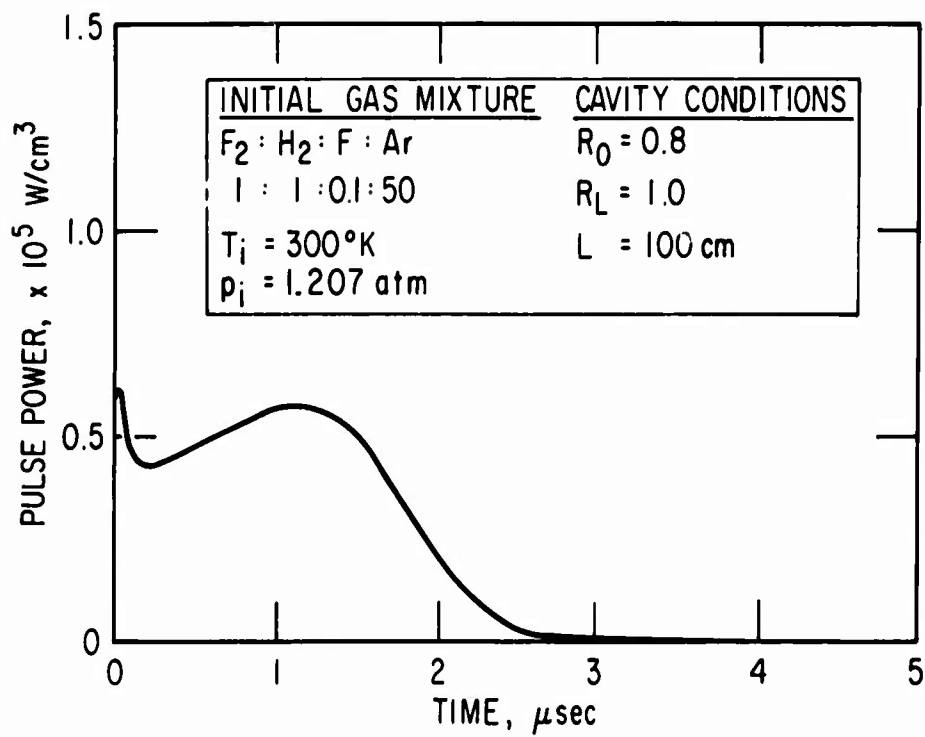


Figure 1. Power History of the Standard Case

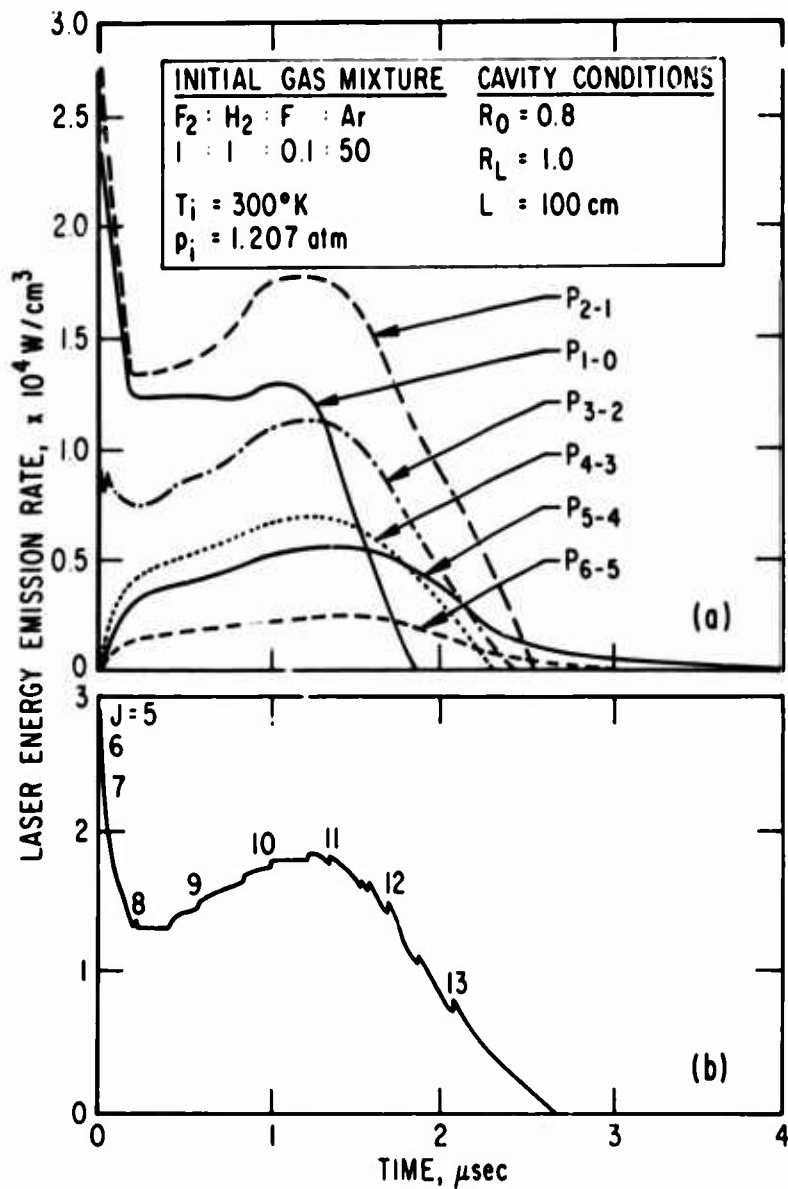


Figure 2. Laser Power and Energy Emission Rate

(a) Laser power from the first six levels of HF. (b) Laser energy emission rate from the 2-1 band for the standard case. The new value of J for the lower level is shown for each J -shift location.

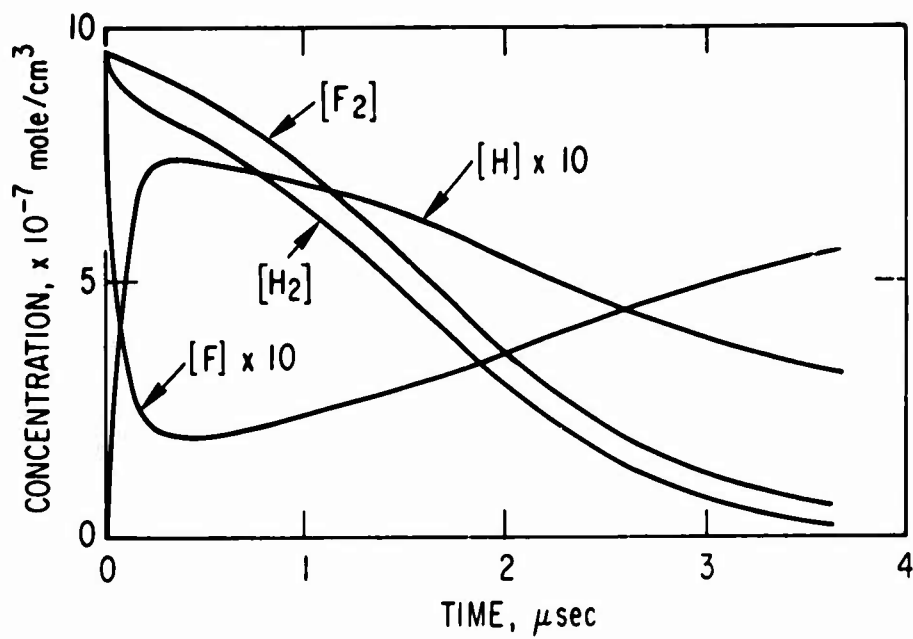


Figure 3. Time Histories of the Concentrations of H_2 , F_2 , H , and F for the Standard Case

Termination of the lasing pulse is caused in part by the depletion of the H_2 and F_2 (Fig. 3). In addition, as the fuel supply is consumed, the pumping rate decreases until chemical deactivation of $HF(v)$ forces the gain below threshold.

The set of equations (2) may be solved for emission on a particular band $\psi_{rad}(v, J)$ and then put in the form

$$\sum_{v'=v+1}^6 \kappa_{v'} = \psi_{rad}(v, J) + \sum_{v'=v+1}^6 \frac{dHF(v')}{dt} + D_{v+1, v} \quad (10)$$

where $\kappa_{v'}$ is the rate of pumping into level v' by the Reactions 8-10 and 15-20, and $D_{v+1, v}$ is the chemical deactivation rate of the $v+1, v$ inversion. In general, the VV HF-HF deactivation reactions are not very important when compared to VT deactivation; hence, the major contribution to $D_{v+1, v}$ is deactivation of $HF(v+1)$ by VT reactions and VV reactions with $H_2(v')$. In fact, these are the only contributions to $D_{v+1, v}$ when anharmonic effects are negligible. Of course, Eq. (10) is a result of conservation of number density of $HF(v)$, and terms in the equation may be interpreted as

$$\left[\begin{array}{c} \text{Pumping} \\ \text{Rate} \end{array} \right] = \left[\begin{array}{c} \text{Photon} \\ \text{Emission} \\ \text{Rate} \end{array} \right] + \left[\begin{array}{c} \text{Level} \\ \text{Filling} \\ \text{Rate} \end{array} \right] + \left[\begin{array}{c} \text{Chemical} \\ \text{Deactivation} \\ \text{Rate} \end{array} \right]$$

where level filling is the filling of vibrational levels as needed to maintain the required gain. The relative sizes of the terms in Eq. (10) for the 2-1 band are shown in Fig. 4, where the rate of photon emission on that band is also shown as a solid line. Note that Eq. (10) is essentially a representation of the distribution of pumping energy when anharmonic effects are small.

Distribution of the pumping rates among the vibrational levels has a significant effect on the energy available for lasing. For example, the production of one $HF(6)$ molecule makes up to six times more energy available for lasing than the production of one $HF(1)$ molecule. Pumping reactions

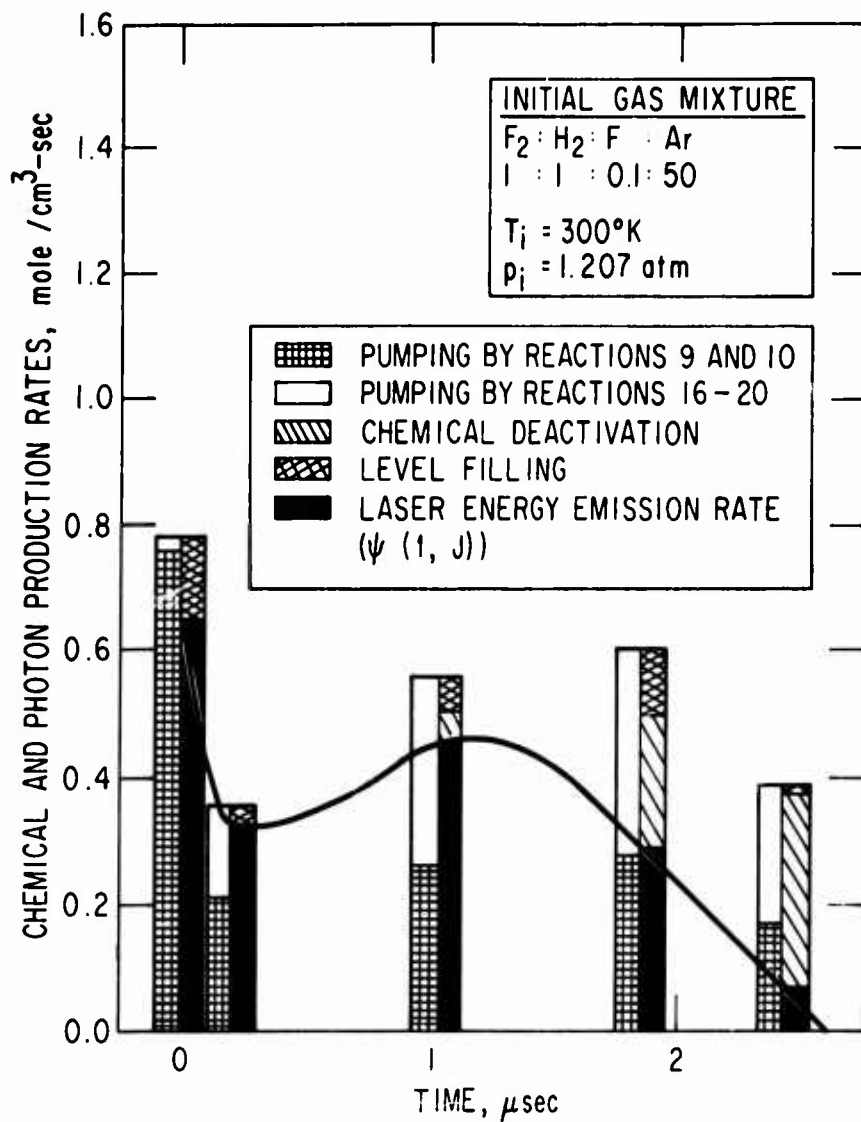


Figure 4. Relationship of Dominant Mechanisms for the 2-1 Band of HF in a Chemical Laser

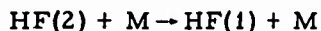
The solid curve represents the instantaneous rate of release of lasing photons for the 2-1 band

such as 8 and 9 that promote a total inversion are more favorable for meeting threshold conditions than the more even pumping distribution of Reactions 15-20. In regimes where threshold is easily attained and in the absence of deactivation, changes that increase the pumping of Reactions 15-20 relative to Reactions 8-10, such as H or F₂ rich mixtures, can be expected to increase pulse energy, since they supply more vibrational energy per mole of fuel.

B. A DETAILED INTERPRETATION OF MECHANISMS FOR THE STANDARD CASE

This section describes in detail the mechanisms that promote the partial inversion and contribute to laser power from the 2-1 band. In addition, we shall assess the effect of the photon flux on the chemical mechanisms associated with this band.

The concentration difference HF(2)-HF(1) is nearly proportional to the gain on the 2-1 line, Eq. (5). The relative importance of the chemical reactions and photon emissions that alter the $v = 2$ to $v = 1$ inversion is depicted in Fig. 5(a). The rate of mass transport by a given reaction is $L_{fr} - L_{br}$. A weighting factor ν_r is used to give the appropriate contribution of each reaction to the increase of HF(2)-HF(1). For example, the VT reaction



decreases HF(2) and increases HF(1). The effective mass transported is twice ($L_{fr} - L_{br}$) for the reaction; hence, $\nu_r = 2$. The major contribution to mass transport during the initial stages of the reaction results from lasing on the 2-1 and adjacent bands. At $t = 0$, only the pumping Reactions 8 and 9 are present. In the figure, the first bar shown is for $t = 0+$. The initial spike in the photon emission rate is a result of the high initial F concentration, and the remainder of the curve (Fig. 1) may be interpreted by noting the effect of temperature rise and action of the chain, as will be shown later.

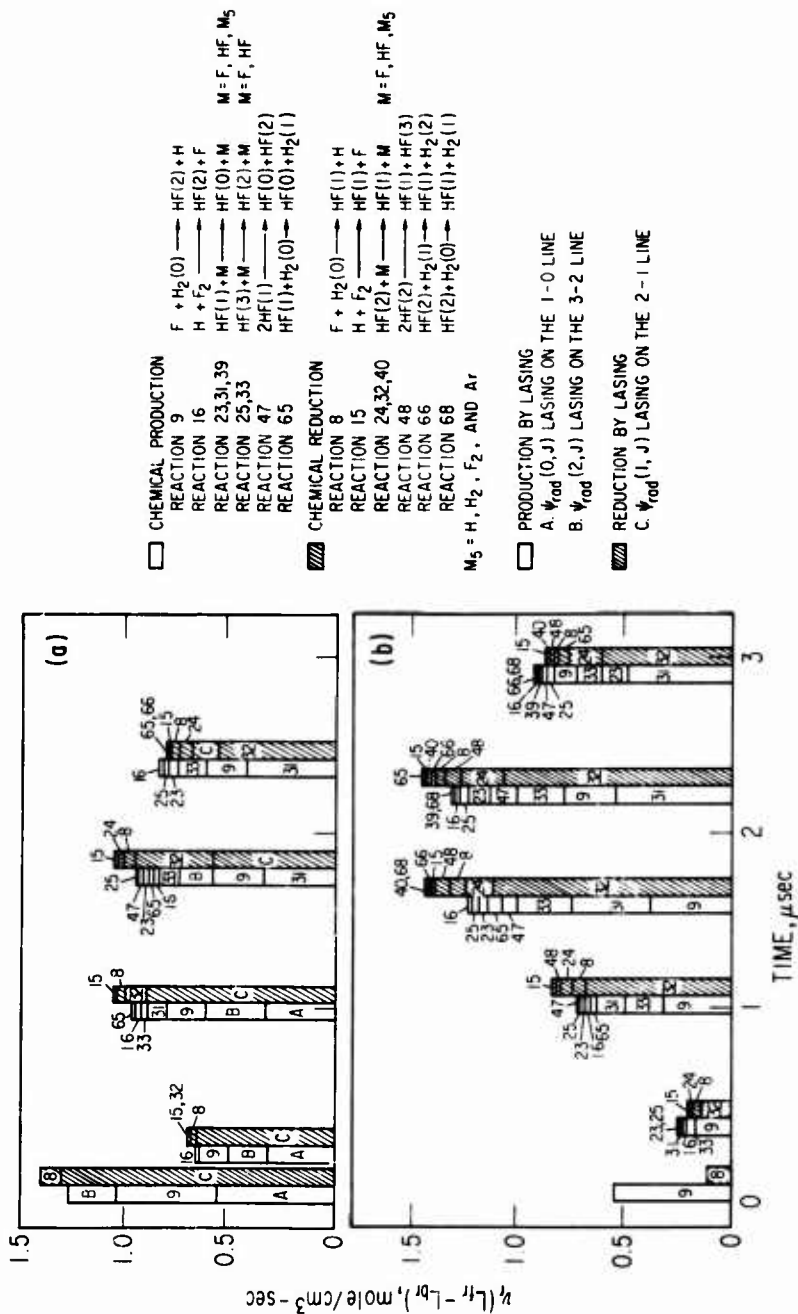
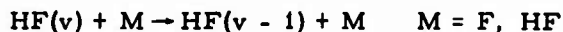


Figure 5. Production and Reduction Mechanisms for the Gain of the 2-1 Line
 (a) Standard case. (b) Standard case with radiation suppressed.

In the preceding section, the significance of pumping by Reactions 16-20 was noted. The format of Fig. 5(a) tends to mask this significance, for at first glance Reactions 8 and 9 appear to be the dominant chemical pumping mechanisms. However, Reactions 16-20 pump the upper vibrational levels and contribute to the pumping of $v=2$ by cascading. We note that the major pumping and deactivation of the 2-1 inversion arises from lasing on the bands 1-0, 2-1, and 3-2. Lasing terminates before the fuel supply (F_2 and H_2) is consumed because the VT chemical deactivation



becomes large enough to force the gain below threshold. The rate of mass transport by all other reactions is small.

We examine the standard case with radiation suppressed by operating the code with reflectivities set equal to zero. Production and reduction mechanisms of the 2-1 inversion for this zero-power case are shown in Fig. 5(b). Comparing Figs. 5(a) and 5(b), we note that in the absence of a photon flux both pumping of the $HF(v)$ levels and their deactivation rates are larger. Pumping is larger because the gas is hotter (hence, the reactions are faster), as is shown in Fig. 6, where the temperature history of the standard case and the zero-power calculations are compared. The increase in deactivation is caused by the presence of higher concentrations of $HF(v)$. The higher populations in the $HF(v)$ levels for the zero-power case also result in an increased significance of VV reactions (47, 48, 65, 66, and 68) of $HF(v)$ with $HF(v')$ and $H_2(v'')$. In addition, we see in Fig. 5(b) that VT deactivation of $HF(v)$ by the diluent (Reactions 39 and 40) begins to appear.

For the lasing case, the gain reaches threshold and remains at that value, but in a zero-power calculation the maximum gains for the lower transitions reach values nearly 3 orders of magnitude larger than the threshold value (Fig. 7). In the figure, locations where the gain for a given band becomes larger for a new value of J are denoted by the new J . A comparison of the gain on the 4-3 band in Fig. 7 with the 4-3 curve in Fig. 2(a) shows the

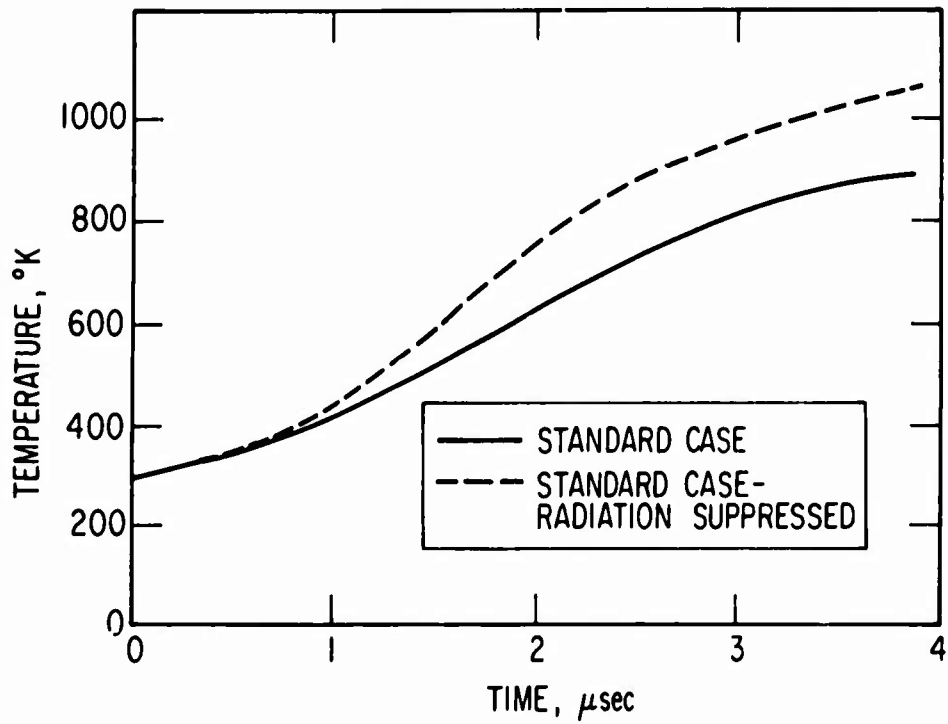


Figure 6. Temperature History of the Standard Case, with and without Radiation Suppressed

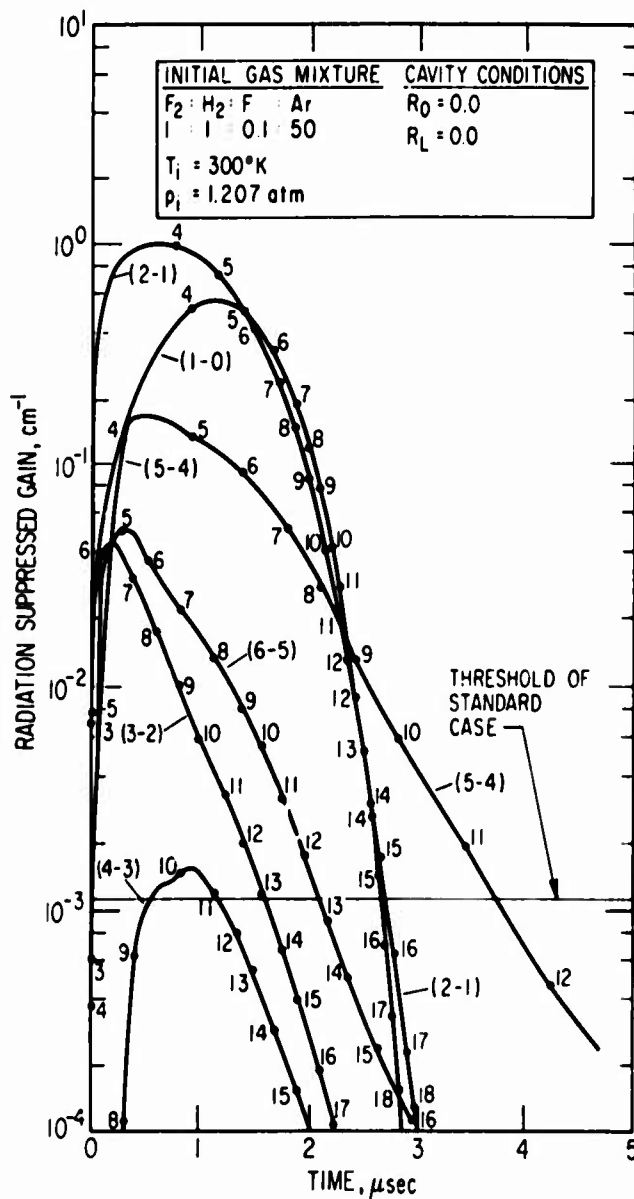


Figure 7. Gain History for Standard Case with Radiation Suppressed

importance of radiative cascading and the limited usefulness of zero-power gain calculations for a multi-level system such as this.

In Fig. 8, production and reduction mechanisms for the F-atom population are shown. After a transient period of consumption of F atoms by Reactions 7-10, the production and reduction mechanisms are nearly the same, which illustrates the presence of a quasi-steady chain.

As indicated earlier, the second hump in the laser photon emission curve may be explained by noting the effect of the temperature increase. Rate coefficients characteristic of the important chemical mechanisms are depicted as a function of time for the standard case in Fig. 9. The pumping rates increase, while the principal deactivation rate decreases with time. Hence, temperature variations of rate coefficients during the pulse tend to increase the instantaneous rate of photon emission and reduce chemical deactivation losses. At still higher temperatures, the VT deactivation rate of HF begins to increase with temperature, as is illustrated for Reaction 32 in the insert in Fig. 9. The increase in the photon emission curve (Fig. 2) between 0.15 and 1.0 μ sec may be attributed to an increase in pumping with the rise in temperature. This may be seen by comparing Fig. 9 with Fig. 3. During this time interval, the reactant concentrations decrease slightly, whereas, due to the temperature rise, the pumping reaction rates increase by a factor of 2.

From the previous discussion, it is clear that the important rate coefficients give a more favorable competition between stimulated emission and deactivation with a temperature somewhat above 300° K. With pumping increased relative to deactivation, the pulse energy is larger because more of the fuel is consumed while the reacting mixture is above threshold. The effect of a temperature increase on the pulse duration is less certain. While the overall reaction is faster, the increased pumping relative to deactivation tends to hold the mixture above threshold for a longer time.

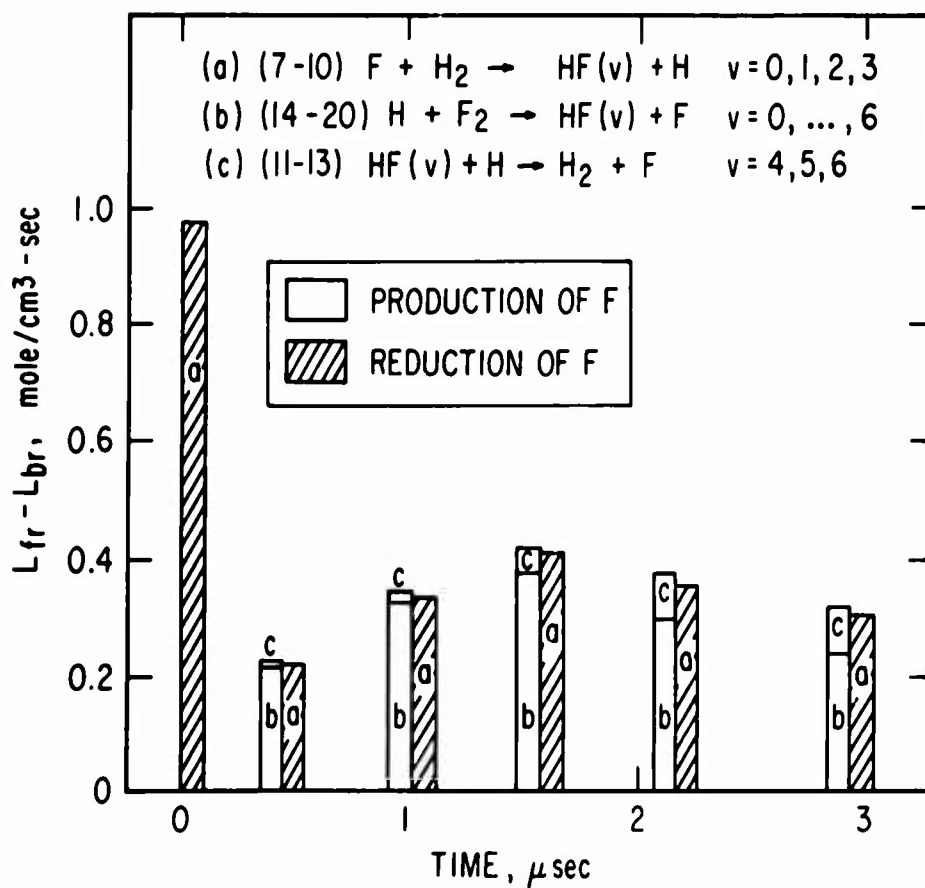


Figure 8. Mechanisms Producing and Reducing F-atom Concentration for the Standard Case

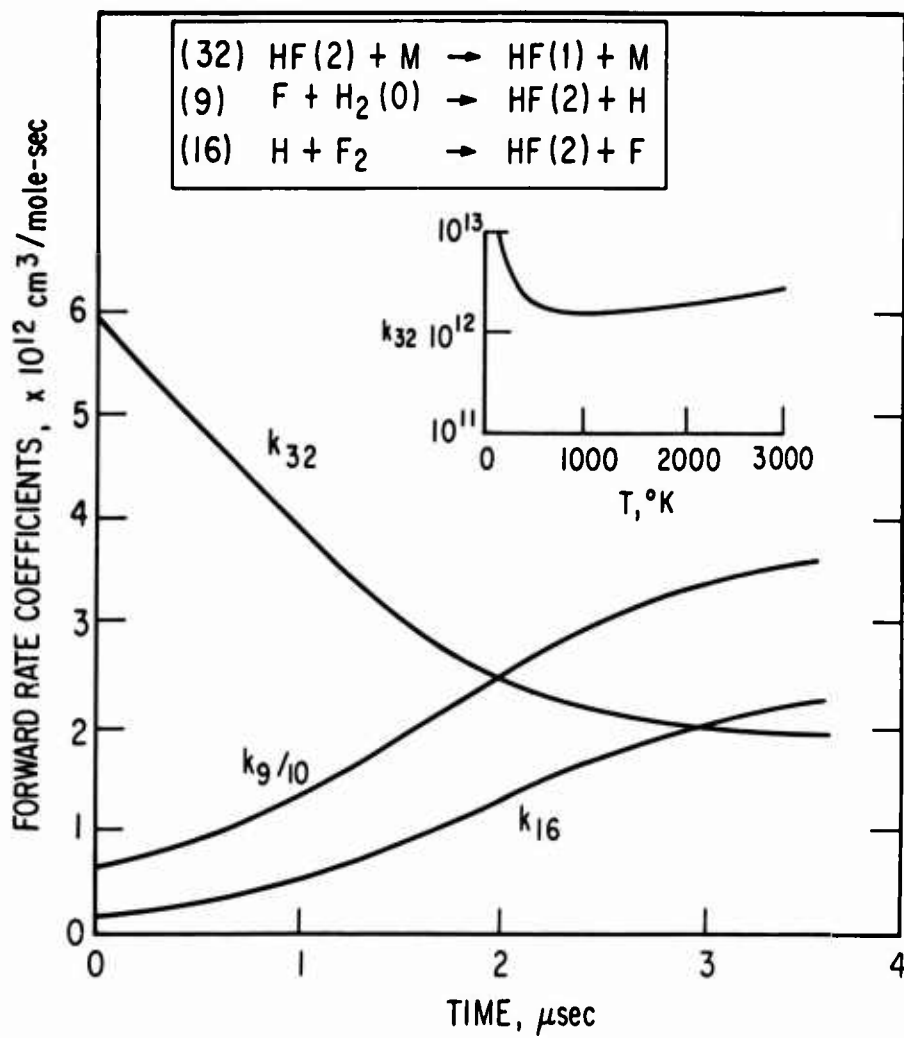


Figure 9. Time Histories of Important Rate Coefficients for the Standard Case

IV. PARAMETRIC VARIATION

The laser pulse may be characterized in part by the pulse duration t_1 and total pulse energy $E_L(t_1)$. In the parametric study, we encounter many pulses with long low-power tails; hence, the time $t_{1/2}$ required to release one-half the pulse energy $E_L(t_1)/2$ is also a useful characteristic. The effect of varying the initial chemical concentrations, cavity conditions, and rate coefficients is discussed by interpreting the variations in these pulse characteristics.

A. THRESHOLD

Concentrations and temperature are given the same initial values as in the standard case, while the characteristics of the optical cavity (mirror reflectivities and cavity length) are varied. This is equivalent to varying α_{th} , as can be seen from Eq. (4). In Fig. 10, the variation of pulse energy with α_{th} is shown. Additional abscissa scales for equivalent values of R_0 and L are also shown. In Fig. 10, the pulse energy is zero for $\alpha_{th} \geq 1.0$. This is consistent with expectations based on the maximum gain shown in Fig. 7. As α_{th} approaches zero, the pulse energy approaches a maximum; of course, practical limitations on cavity losses and cavity length can make realization of this maximum difficult. One aspect of decreasing α_{th} is that, in addition to increasing $E_L(t_1)$, the pulse length characterized by $t_{1/2}$ or t_1 is also extended, as shown in Fig. 11. For decreasing values of α_{th} , the pronounced increase in t_1 relative to $t_{1/2}$ is the result of a long low-power tail. The time required for release of the majority of the pulse energy is relatively insensitive to α_{th} , as seen by the small variation of $t_{1/2}$.

Since more energy is extracted for low α_{th} , the temperature of the reacting mixture is also lower. As α_{th} is decreased, the vibrational populations are smaller and the rate of VT chemical deactivation is reduced, even though the rate coefficients for these processes are higher because of the lower temperature. Examination of the reaction rates for the 2-1 band shows that for smaller values of α_{th} the rate of pumping is larger, the instantaneous

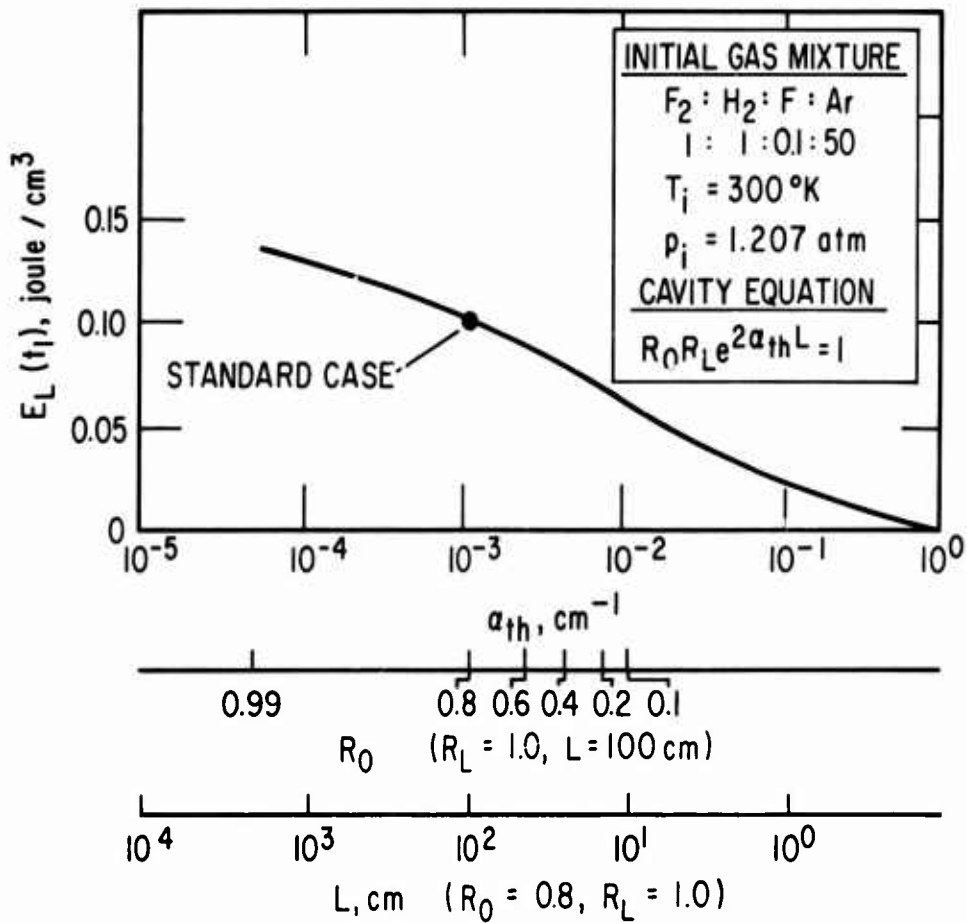


Figure 10. Effect of Threshold Gain on Pulse Energy

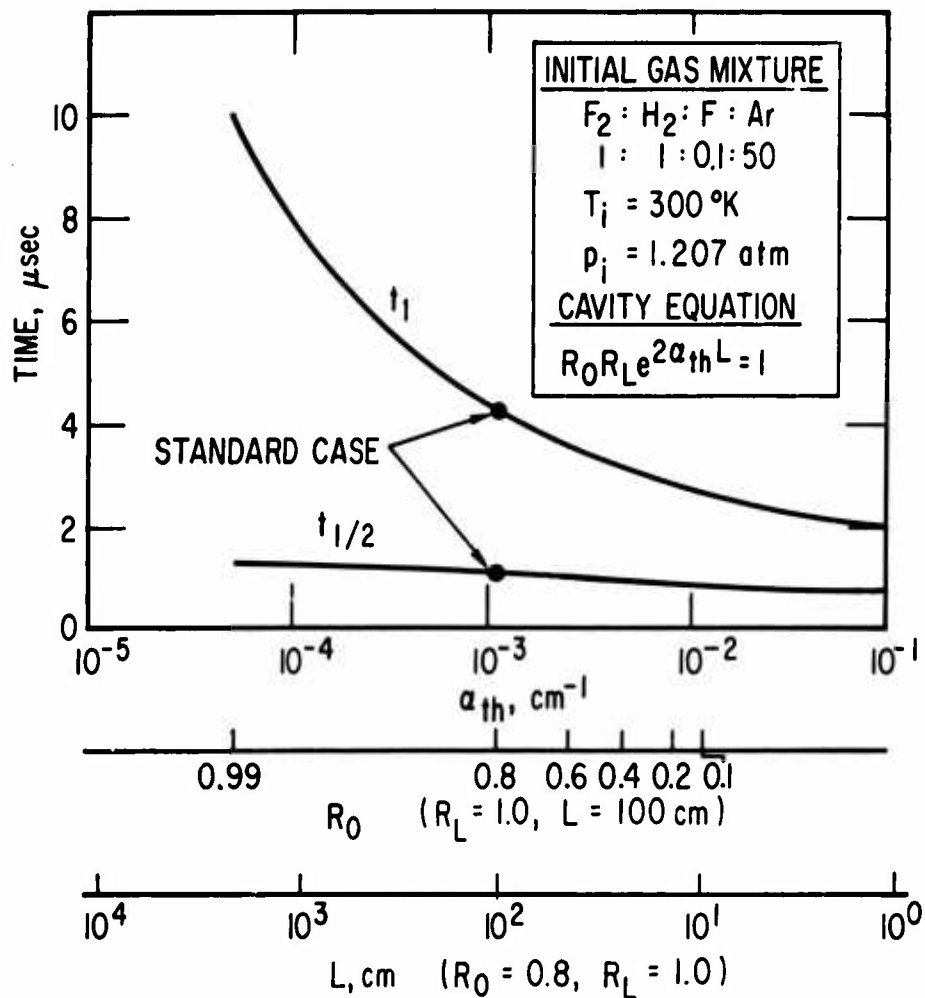


Figure 11. Effect of Threshold Gain on t_1 and $t_{1/2}$

power is larger, and the rate of level filling (energy storage) becomes smaller. The mixture reaches threshold at an earlier time and, thus, the F-atom concentration is larger, and pumping at pulse initiation is higher and continues to be so throughout the run.

B. DILUENTS

Diluents contribute to the heat capacity of the total system and act as catalysts for the VT deactivation reactions and recombination reactions. Since the VT deactivation rate of HF(v) with Ar is about 10^5 times smaller at 300°K than the VT rate of HF(v) with itself, Ar contributes little to chemical deactivation; its major role is temperature control.

The effect of variations in dilution is studied for two initial concentrations of F. For initial F equal to 10% of initial F_2 , the following trends are noted as the concentration of Ar is increased.

1. The temperature rise throughout the reaction is less, as expected.
2. The laser power decreases and the second peak in the power curve disappears. Also, pumping and level filling rates decrease while deactivation rates increase as a result of the lower temperature.
3. The pulse energy decreases gradually (Fig. 12) because of the less favorable competition described in Item 2.
4. The pulse duration t_1 and time parameter $t_{1/2}$ increase, as expected, for lower temperatures (Fig. 13).

Initial F concentrations in this 10% range and above emphasize Reactions 8-10 relative to Reactions 15-20.

For initial F equal to 1% of initial F_2 , pumping rates throughout the reaction are much lower. Deactivation is more significant and the pulse energy is lower. As the concentration of Ar is increased, the temperature throughout the pulse is lower and the pulse length is increased. Further

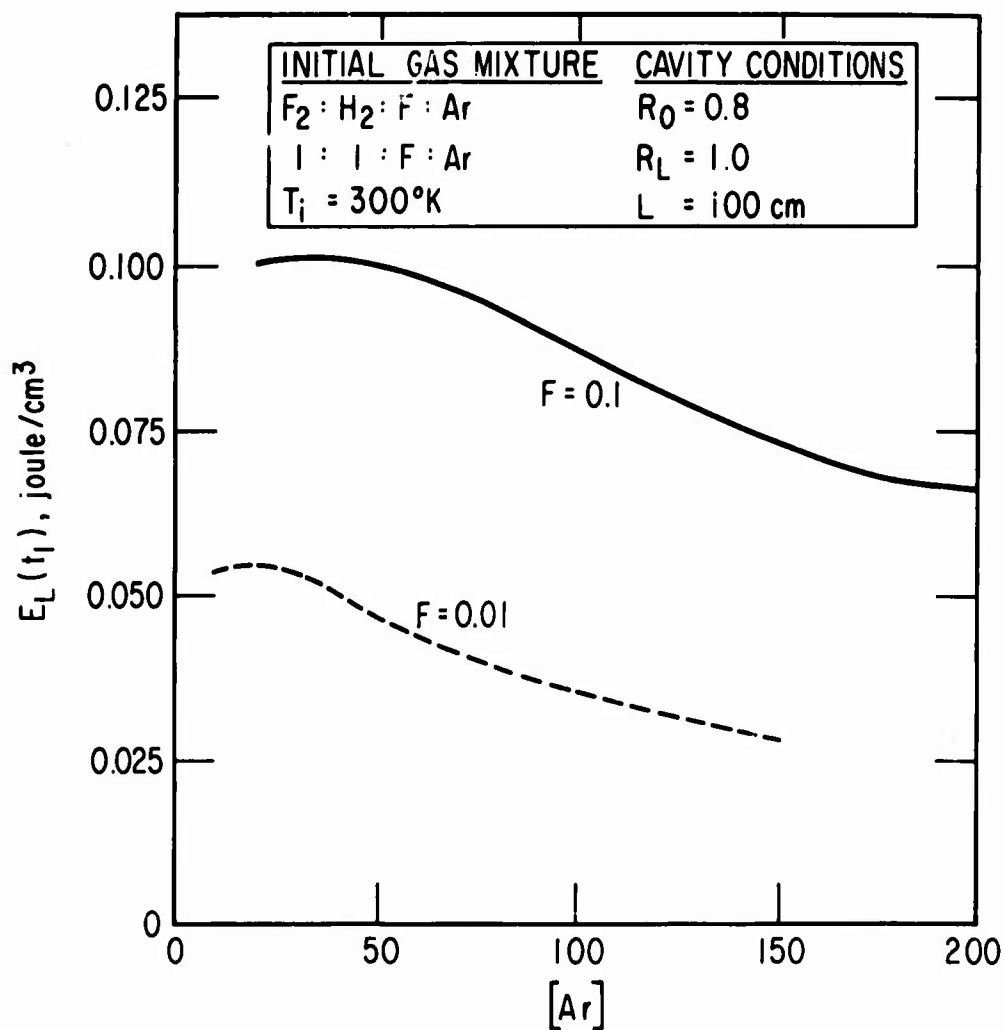


Figure 12. Effect of Diluent Concentration on Pulse Energy with Two Different Initial F-atom Concentrations

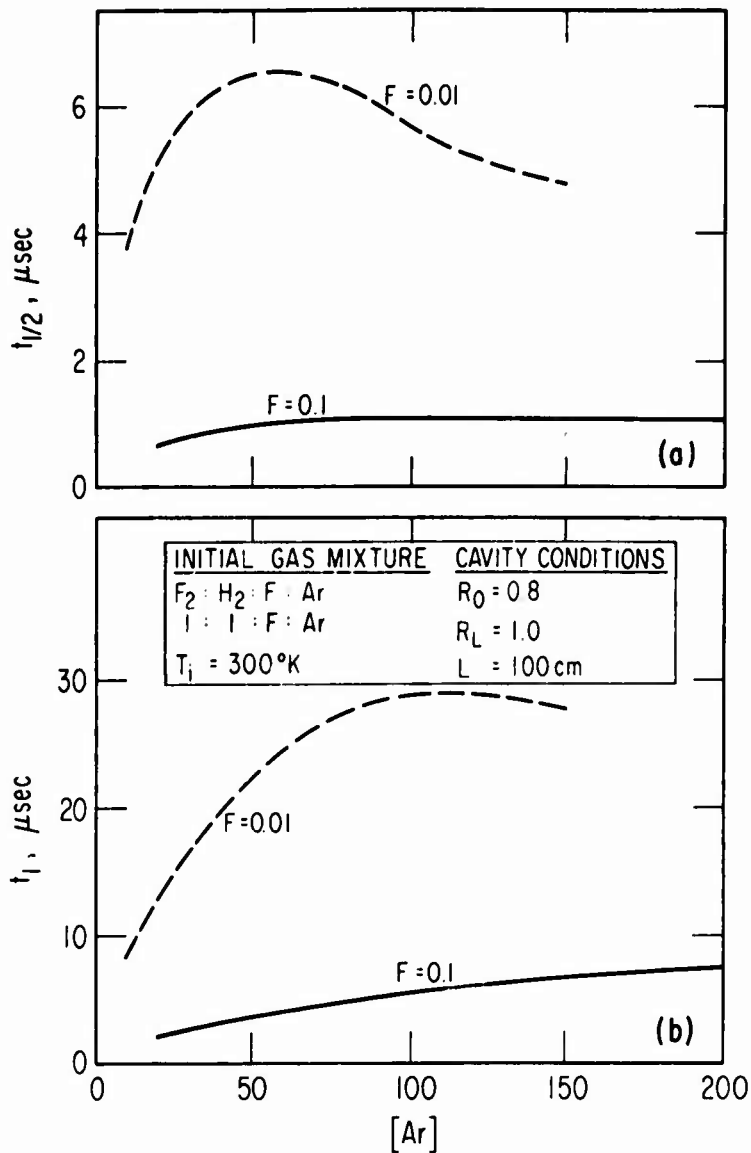


Figure 13. Effect of Diluent Concentration on $t_{1/2}$ and t_1
 (a) Effect on $t_{1/2}$ with two initial F-atom concentrations. (b) Effect on pulse length t_1 with two initial F-atom concentrations.

increases in Ar lower the temperature sufficiently such that the stimulated emission-deactivation competition forces the pulse length to decrease and a maximum pulse duration is found.

C. INITIAL F₂ DISSOCIATION

The initial F-atom concentration is increased while the stoichiometric balance is maintained between hydrogen and fluorine. Other initial conditions are held constant. As the concentration of F is increased from very low values, the following trends are noted.

1. The rate of pumping increases, laser power increases, and the rate of level filling increases.
2. The faster pumping rates result in smaller pulse durations t_1 and smaller values of $t_{1/2}$ (Fig. 14).
3. The pulse energy increases until the mass fraction of F reaches 0.2 (Fig. 15). Further increases in F result in a dominance of reactions 7-10 throughout the pulse. The second peak in the power curve disappears, and $E_L(t_1)$ and $E_{v+1,v}$ for each band decrease. Finally, for total dissociation of F₂, Reactions 14-20 do not contribute to the pulse energy and E_{43} , E_{54} , and E_{65} are zero.
4. Since the pumping Reactions 7-10 release only about one-third as much energy as Reactions 14-20, the reaction temperature is lower for large initial F. Such behavior leads to the stimulated emission-deactivation competition suggested in the discussion of the standard case.
5. For small concentrations of F, there is a strong competition between stimulated emission and deactivation. As initial F₂ dissociation increases, the relative importance of VT deactivation decreases until Reactions 23-30 make a significant contribution.

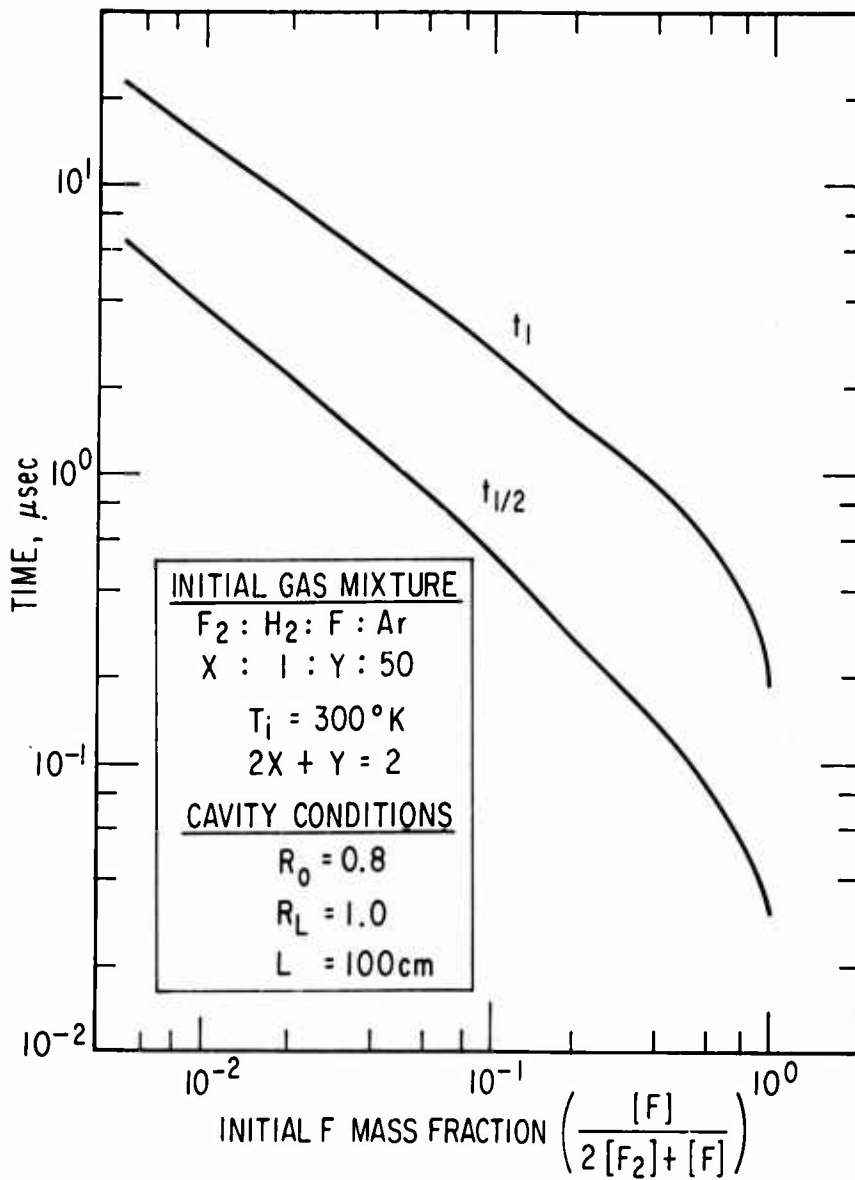


Figure 14. Effect of Initial F_2 Dissociation on Pulse Length

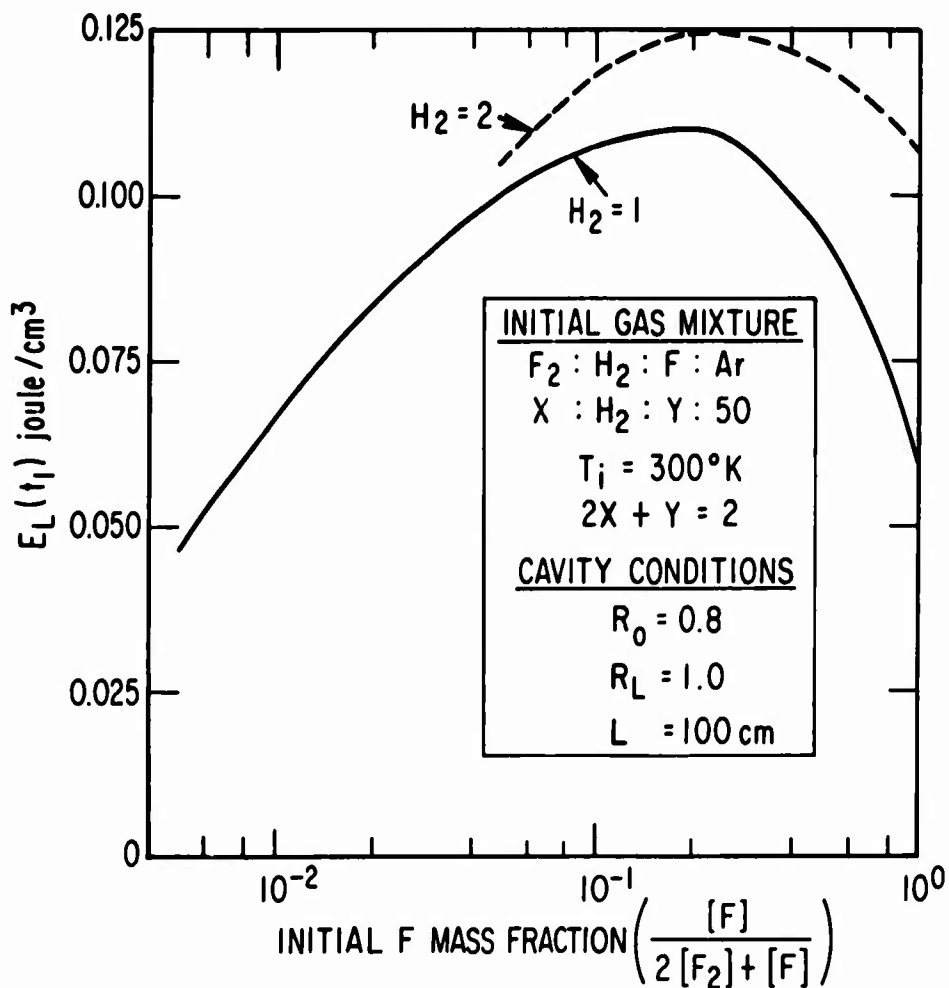


Figure 15. Effect of Initial F_2 Dissociation on Pulse Energy

For the larger initial F concentrations, the mixture is below stoichiometric for H_2 . To assess this effect, the same variations of F were made with the initial H_2 concentration doubled. The results of these computations are shown in Fig. 15 as a dashed curve. There is a marked increase in pulse energy for high initial F. There is also a slight decrease in t_1 and $t_{1/2}$ as a result of the increased reaction rates.

D. INITIAL TEMPERATURE

The initial temperature T_i was varied while all other conditions were maintained at the values of the standard case. As T_i increases, the temperature throughout the pulse is higher, reaction rates are faster, and characteristic times become shorter (Fig. 16). For T_i above 400°K, the hump in the power curve disappears; i.e., throughout the pulse the rate of pumping decreases monotonically, since the effect of temperature increase is less significant. In addition, the pulse energy begins to decrease slightly; however, the overall effect of T_i on pulse energy is small.

E. INITIAL H_2 CONCENTRATION

The initial concentration of H_2 is varied while other initial conditions are those of the standard case. The pulse energy and pulse duration are both quite sensitive to the initial H_2 concentration for H_2 -lean mixtures (Fig. 17). In this regime, increases in H_2 are accompanied by increased pumping, increased pulse energy, and increased pulse duration.

The effect of variations in H_2 is very small for H_2 -rich mixtures. As the H_2 concentration is increased, the pumping rate increases, the pulse energy increases, and the pulse duration decreases because of the faster reaction rate. Note that the time $t_{1/2}$ required to deliver one-half the pulse energy is virtually independent of initial H_2 . Of course, very high H_2 concentrations give rise to VV and VT deactivation of $HF(v)$ with $H_2(v')$.

F. VARIATION OF IMPORTANT RATE CONSTANTS

The effect of variations in dominant rate constants was studied with variations from the standard (see Appendix) selected to include the major

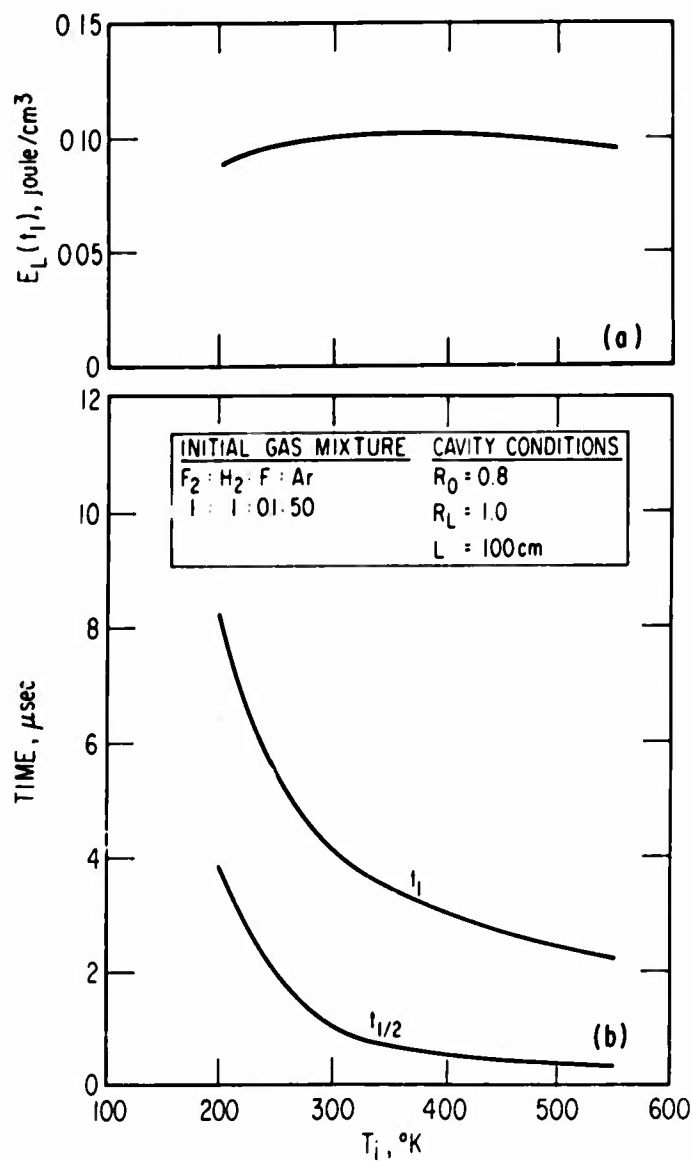


Figure 16. Effect of Initial Temperature

(a) Effect on pulse energy. (b) Effect on t_1 and $t_{1/2}$

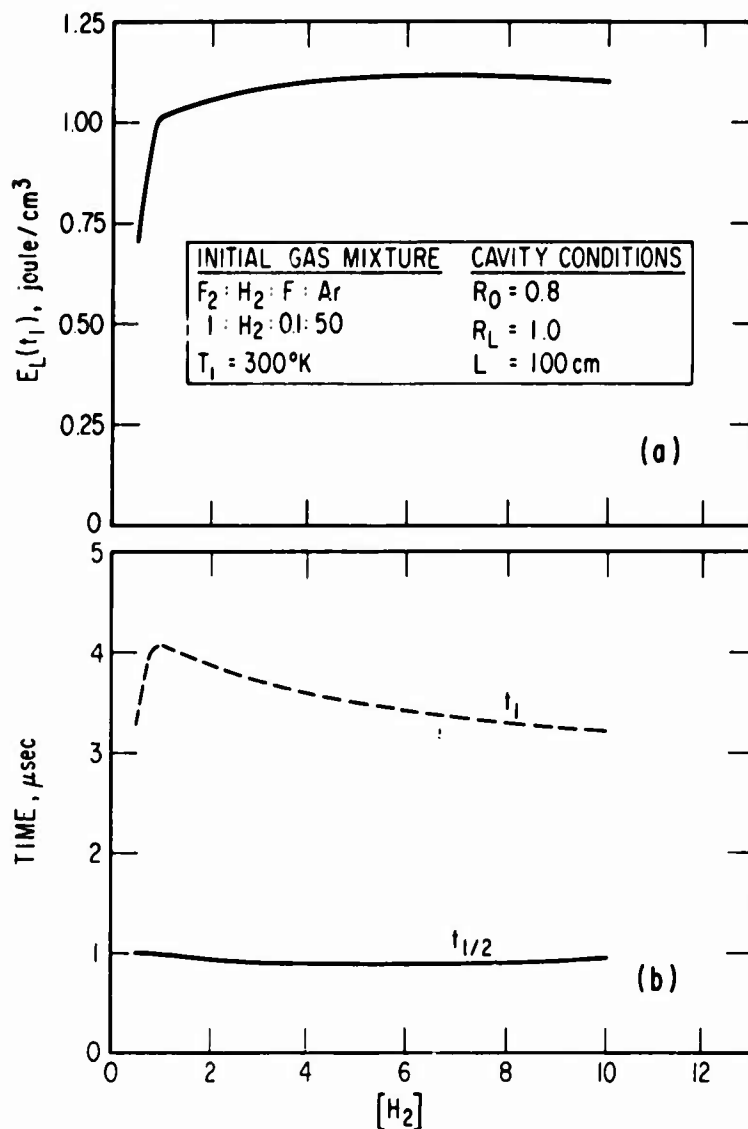


Figure 17. Effect of Initial H_2 Concentration

(a) Effect on $E_L(t_1)$. (b) Effect on the characteristic times t_1 and $t_{1/2}$.

uncertainties in the rates at the time of the study. The effect of these changes on the pulse characteristics and the energy released from the band $v + 1 \rightarrow v$, $E_{v+1,v}$, may be seen in Table I. The variations in the rates studied are described below.

Rate 1: Pumping into vibrational levels 1-6 is reduced by a factor of 2, while the total reaction rate is held constant by increasing the rate of production of HF(0). As expected, the pulse energy and the pulse length decreased with the lower pumping. Also, the energy in the 1-0 band is much less because the rate of populating HF(0) is much higher.

Rate 2: The rates for HF-HF VV reactions are increased by a factor of 100. This change has a very small effect on the pulse energy and the characteristic times; however, it does redistribute the energy emitted by the various bands.

Rate 3: The rates for HF-H₂ VV reactions are increased by a factor of 100. In this case, the pulse energy is essentially constant, the energy in the 2-1 band is higher while the energy in the 1-0 band is lower. Reaction 65 becomes a significant deactivation mechanism for the 1-0 inversion, and for this mixture the combination of Reactions 65 and 66 aids the 2-1 inversion. In addition, since the 1-0 band contributes energy early in the pulse, $t_{1/2}$ increases slightly.

Rate 4: The changes in Rates 1, 2, and 3 are combined. A much lower pulse energy and shorter pulse duration are observed, as expected.

Rate 5: The backward rates for $F + H_2 \rightleftharpoons HF(v) + H$, $v = 4, 5, 6$ are reduced by a factor of 100. The pulse energy increases 10%. A marked increase in the energy released in the bands 6-5, 5-4, and 4-3 is noted as a result of the rate change, and the energy released in the lower bands is higher as a result of cascading.

Table I. Effect of Rate Changes on Pulse Characteristics

Rate	Energy, joule/cc							Time, μ sec	
	$\underline{E}_{10}$	$\underline{E}_{21}$	$\underline{E}_{32}$	$\underline{E}_{43}$	$\underline{E}_{54}$	$\underline{E}_{65}$	$\underline{E}_L(t_1)$	$\underline{t}_1$	$\underline{t}_{1/2}$
Standard									
1-Low pumping	.0208	.0337	.0195	.0115	.0109	.0048	.101	4.09	1.0
2-High VV-HF	.0028	.0115	.0069	.0038	.0044	.0019	.0313	2.86	0.7
3-High VV-H ₂ *	.0151	.0400	.0235	.0123	.0112	.0049	.107	3.96	1.0
4-Combination of 1, 2, and 3	.0177	.0347	.0193	.0114	>.0109	.0048	.0988	>3.7	1.1
5-Reduced backward cold reaction	.0003	.0109	.0082	.0041	.0043	.0019	.0297	2.83	0.8
	.0213	.0347	.0221	.0153	.0124	.0053	.111	3.57	1.1

* This computation was extremely long and the 5-4 band was still lasing at a very low level at termination of the run.

V. CHEMICAL EFFICIENCY

Of the 31.56 kcal/mole released by Reactions 7-10, 22.6 kcal/mole is supplied as vibrational energy of HF(v) with the pumping rates listed in the Appendix. Reactions 14-20 release 98.04 kcal/mole of which 43.1 kcal of vibrational energy is available. By combining these reactions, we find 50.7% of the chemical energy converted to vibrational energy. After the threshold condition and chemical deactivation are considered, only a fraction of this vibrational energy is available for laser emission.

A thermodynamic efficiency of the laser system would account for the energy required to dissociate F_2 upon pulse initiation. However, for simplicity, efficiency η will be taken as

$$\eta = \frac{E_L(t_1) \times 100/\text{mole } F_2}{129.6 \text{ kcal/mole } F_2} \quad (11)$$

Thus, pumping rates alone require η to be less than 50.7% for the chain. For the standard case, η is 20.6%. Optimization of η has not been attempted. However, we note that the effects of the parameters studied in this report indicate the probable direction to be taken for optimization, and that the variation of η is approximately like that of $E_L(t_1)$.

VI. SUMMARY AND CONCLUSIONS

A computer simulation is used to study the interaction of the chemistry and the optical cavity of a pulsed HF laser. A detailed study of the effect of the optical cavity, the initial chemical composition, the initial temperature, and variations in uncertain chemical rate coefficients on the pulse length and pulse energy is made. Variations about a standard case are used to study the effect of these parameters.

Pumping, vibrational-translational deactivation reactions, temperature control, and threshold gain are found to be important parameters governing the behavior of the laser. A competition between stimulated emission and deactivation is observed, and conditions where the effect of deactivation is minimized are indicated. A decrease in the threshold gain is accompanied by an increase in pulse energy and pulse duration. The photon flux is found to have a large effect on the populating mechanisms of the vibrational levels.

Argon is used as a diluent for temperature control. An optimum concentration of Ar is found that maximizes pulse energy, and at low concentrations of initial F , an optimum concentration of Ar is found for maximizing pulse duration.

Pulse duration decreases with increasing initial F_2 dissociation, while an optimum dissociation of F_2 is found that maximizes the pulse energy.

The effect of varying the initial temperature of the reactants on the pulse energy is small. An increase in the initial temperature decreases the pulse duration, as expected.

An H_2 -rich mixture is desirable to provide an increase in pulse energy and pulse duration. The effect of varying the initial H_2 concentration for H_2 -rich mixtures is small.

The variations of the more uncertain rate coefficients at the time of this study reveal that the pulse characteristics are sensitive to changes in the total pumping reaction rates and to the distribution of pumping in the vibrational levels. These characteristics are not very sensitive to changes in the VV rates of $HF(v)$ with $HF(v')$ or $H_2(v'')$. The major effect of these changes is redistribution of energy emitted by the lasing bands.

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APPENDIX

THE REACTION SYSTEM

The chemical kinetic model and corresponding rate coefficients used in this study are listed in the following table. Chemical rate coefficients for the HF laser system are from a compilation prepared by N. Cohen of our laboratory in December 1970. Subsequently, Cohen has reported a more final review of these rates.¹² Implications of the difference in rates between the table and Ref. 12 are not serious and are covered in the discussion. Some of the catalytic species listed at the bottom of the table are multiplied by a constant. This is equivalent to using a rate coefficient whose value is larger by this factor. The thermodynamic data are taken from the JANAF (Joint-Army-Navy-Air Force) publication distributed by Dow Chemical Corporation. However, the thermodynamic data of excited species are generated at this laboratory.

Table of Reactions and Rate Coefficients

Reaction No.	Reaction ¹	Rate Coefficient ²
1	$H_2(0) + M_2 \rightleftharpoons H + H + M_2$	$k_{-1} = 10^{18} T^{-1.0}$
2 ³	$F_2 + M_4 \rightleftharpoons F + F + M_4$	$k_2 = 5.0 \times 10^{13} e^{35.30}$
3, ..., 6	$HF(v) + M_4 \rightleftharpoons H + F + M_4$	$k_{3+v} = 1.2 \times 10^{19} T^{-1.0} e^{135.80} \quad v = 0, \dots, 3$
7	$F + H_2(0) \rightleftharpoons HF(0) + H$	$k_7 = 9.0 \times 10^{12} e^{1.60}$
8	$F + H_2(0) \rightleftharpoons HF(1) + H$	$k_8 = 2k_7$
9	$F + H_2(0) \rightleftharpoons HF(2) + H$	$k_9 = 10k_7$
10	$F + H_2(0) \rightleftharpoons HF(3) + H$	$k_{10} = 5k_7$
11, 12, 13	$F + H_2(0) \rightleftharpoons HF(v) + H$	$k_{-7-v} = 10^{12} T^{0.670} \quad v = 4, 5, 6$
14	$H + F_2 \rightleftharpoons HF(0) + F$	$k_{14} = 6.0 \times 10^{14} e^{2.40}$
15	$H + F_2 \rightleftharpoons HF(1) + F$	$k_{15} = k_{14}$
16	$H + F_2 \rightleftharpoons HF(2) + F$	$k_{16} = 1.5k_{14}$
17	$H + F_2 \rightleftharpoons HF(3) + F$	$k_{17} = 2.66667k_{14}$
18	$H + F_2 \rightleftharpoons HF(4) + F$	$k_{18} = 3.33333k_{14}$
19	$H + F_2 \rightleftharpoons HF(5) + F$	$k_{19} = 5.5k_{14}$
20	$H + F_2 \rightleftharpoons HF(6) + F$	$k_{20} = 5k_{14}$
21, 22	$H_2(v) + M_1 \rightleftharpoons H_2(v-1) + M_1$	$k_{20+v} = 2.5 \times 10^{-4} v T^{4.3} \quad v = 1, 2$
23, ..., 30	$HF(v) + M_3 \rightleftharpoons HF(v-1) + M_3$	$k_{22+v} = 9.0 \times 10^8 v T^{1.30} \quad v = 1, \dots, 8$
31, ..., 38	$HF(v) + M_6 \rightleftharpoons HF(v-1) + M_6$	$k_{30+v} = 5.0 \times 10^7 v T^{1.30} + 10^{16} v T^{-1.43} \quad v = 1, \dots, 8$
39, ..., 46	$HF(v) + M_5 \rightleftharpoons HF(v-1) + M_5$	$k_{38+v} = 1.3 \times 10^{-2} v T^{3.60} \quad v = 1, \dots, 8$
47, ..., 53	$2HF(v) \rightleftharpoons HF(v-1) + HF(v+1)$	$k_{46+v} = 4.0 \times 10^5 T^{2.2} \quad v = 1, \dots, 7$
54, ..., 59	$HF(v) + HF(v+1) \rightleftharpoons HF(v-1) + HF(v+2)$	$k_{53+v} = 1.2 \times 10^3 T^{2.80} \quad v = 1, \dots, 6$
60, ..., 64	$HF(v) + HF(v+2) \rightleftharpoons HF(v-1) + HF(v+3)$	$k_{59+v} = 6.0 \times 10^{-2} T^{3.90} \quad v = 1, \dots, 5$
65, 66	$HF(v) + H_2(v-1) \rightleftharpoons HF(v-1) + H_2(v)$	$k_{-64-v} = 8.0 \times 10^5 T^{2.20} \quad v = 1, 2$
67	$HF(3) + H_2(0) \rightleftharpoons HF(2) + H_2(1)$	$k_{-67} = 1.2 \times 10^{-1} T^{3.90}$
68	$HF(2) + H_2(0) \rightleftharpoons HF(1) + H_2(1)$	$k_{-68} = 2.4 \times 10^3 T^{2.8}$

¹Catalytic species:

$M_1 = H, F, Ar, HF(0), \dots, HF(8), 4^*H_2(0), 4^*H_2(1), 4^*H_2(2), F_2$
 $M_2 = 20^*H, F, Ar, HF(0), \dots, HF(8), 2.5^*H_2(0), 2.5^*H_2(1), 2.5^*H_2(2), F_2$
 $M_3 = F$
 $M_4 = H, F, Ar, HF(0), \dots, HF(8), H_2(0), H_2(1), H_2(2), F_2$
 $M_5 = H, Ar, H_2(0), H_2(1), H_2(2), F_2$
 $M_6 = HF(0), \dots, HF(8)$

²The rate coefficients k_f and k_b designate forward and backward rates, respectively, with units in terms of moles, cm³, and sec. For each reaction, the missing rate coefficient is determined from the equilibrium constant.

³The quantity $\theta = -(10^3/RT)$, where the temperature T is in °K and R is 1.987 cal/mole-°K.

44