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INTRAMOLECULAR REACTIONS OF UNSATURATED PEROXIDES AND PEROXY RA--ETC(U)

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DAHC04-75-G-0117

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(11) (18) ARO (19) 12782.5-C

(6) Intramolecular Reactions of Unsaturated Peroxides
and Peroxy Radicals

(9) Final Report
1 May 75 - 3 Apr 78

(10) Ned A. Porter

(11) Jul 78

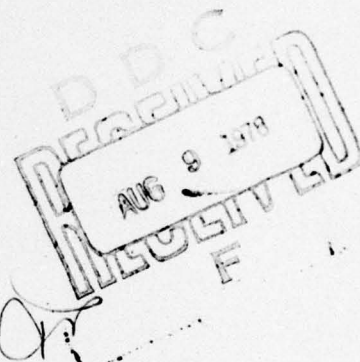
Grant # DAHC04-75-G-0117

(15) Duke University

LEVEL II

(12) 1 p.

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SECURITY CLASSIFICATION OF THIS PAGE (When Data Entered)

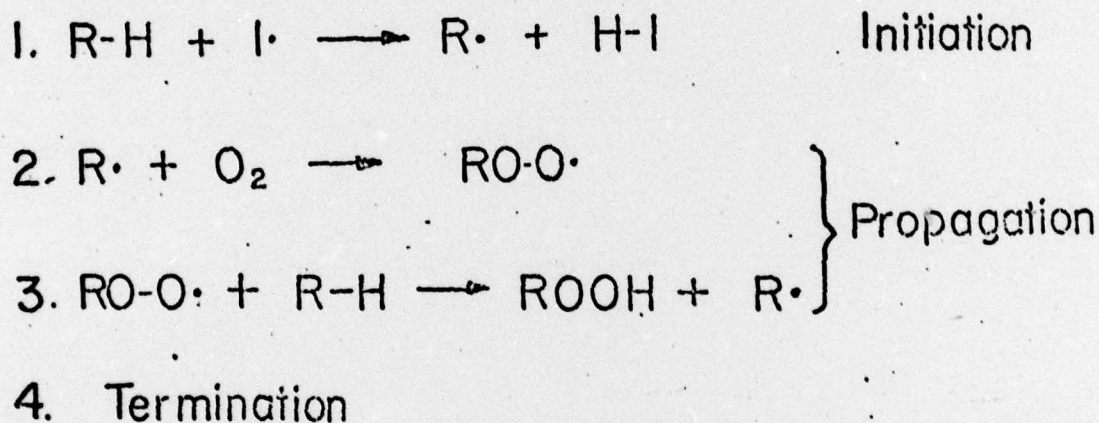
REPORT DOCUMENTATION PAGE		READ INSTRUCTIONS BEFORE COMPLETING FORM
1. REPORT NUMBER DAHCO 75 G 0117	2. JOINT ACCESSION NO.	3. RECIPIENT'S CATALOG NUMBER
4. TITLE (and Subtitle) Intramolecular Reactions of Unsaturated Peroxides and Peroxy Radicals		5. TYPE OF REPORT & PERIOD COVERED Final 1 May, 1975-30 April, 1978
		6. PERFORMING ORG. REPORT NUMBER
7. AUTHOR(s) Ned A. Porter		8. CONTRACT OR GRANT NUMBER(s) DAHCO 4-75-G-0117
9. PERFORMING ORGANIZATION NAME AND ADDRESS Duke University Durham, NC 27706		10. PROGRAM ELEMENT, PROJECT, TASK AREA & WORK UNIT NUMBERS
11. CONTROLLING OFFICE NAME AND ADDRESS U. S. Army Research Office P. O. Box 12211 Research Triangle Park, NC 27709		12. REPORT DATE July 8, 1978
		13. NUMBER OF PAGES ten
14. MONITORING AGENCY NAME & ADDRESS (if different from Controlling Office)		15. SECURITY CLASS. (of this report) unclassified
		15a. DECLASSIFICATION/DOWNGRADING SCHEDULE
16. DISTRIBUTION STATEMENT (of this Report) Approved for public release; distribution unlimited.		
17. DISTRIBUTION STATEMENT (of the abstract entered in Block 20, if different from Report)		
18. SUPPLEMENTARY NOTES The findings in this report are not to be construed as an official Department of the Army position, unless so designated by other authorized documents.		
19. KEY WORDS (Continue on reverse side if necessary and identify by block number) Material degradation, lipid oxidation, free radicals, peroxides, peroxy radicals, molecular oxygen		
20. ABSTRACT (Continue on reverse side if necessary and identify by block number) The mechanisms of free radical oxidation of organic compounds have been studied. It has been determined that unsaturated peroxy radicals will undergo cyclization reactions leading to a variety of cyclic peroxide compounds. It appears that peroxy radicals derived from naturally occurring organic compounds such as rubbers, fats, and oils, are converted into a variety of compounds that contain the cyclic peroxide nucleus. Simple model compounds that permit the study of the peroxy radical cyclization have been examined.		

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These model studies show that, in addition to cyclic peroxides, a variety of compounds that include epoxides and alcohols are formed by free radical mechanisms. Thus, primary products to be expected from autoxidation of natural materials such as fats and rubbers are mono and bicyclic peroxides and epoxy alcohols.

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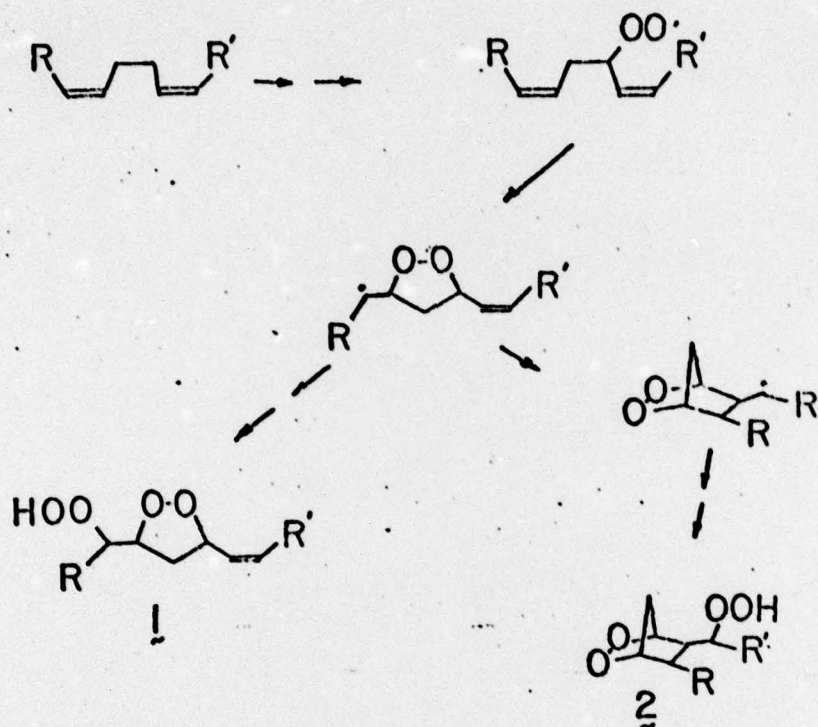
The destruction of organic substrates by molecular oxygen is perhaps one of the most important organic reactions. Natural and synthetic polymers, biological materials such as lipids, and for that matter, any organic molecule exposed to the atmosphere may react with ground state molecular oxygen. It is no surprise then, that extensive investigations have focused on the kinetics of organic substrate oxidation by O_2 .¹ The "hydroperoxide mechanism" for autoxidation which is now generally accepted is outlined below:



Despite the numerous studies on autoxidation and inhibition, surprisingly little research has been directed towards product study in autoxidation. It is true that for simple molecules like isopropylbenzene² the primary autoxidation product is the tertiary hydroperoxide. However, if the organic substrate has several centers of unsaturation, the simple hydroperoxide product may not be found and in its place, cyclic peroxides are formed. For example, squalene³ absorbs two moles of oxygen upon autoxidation and

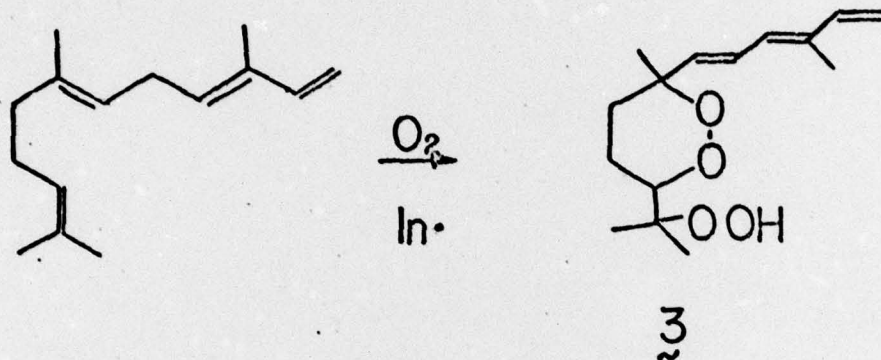
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one of these oxygen molecules is incorporated as a peroxide, rather than a hydroperoxide group. Although no rigorous structure proof was offered for the squalene oxidation product, it seems likely on the basis of subsequent work⁴ that the principal product of squalene autoxidation is one of the cyclic peroxides (1 or 2) formed by the mechanism described below:



In this mechanism, the first formed peroxy radical can undergo cyclization to give either a monocyclic peroxide 1 or the bicyclic species 2. Peroxy radical cyclization, then, probably plays an important role in determining the products of squalene autoxidation.

Another example of peroxy radical cyclization is provided by the work of Anet *et. al.*⁵ who found the monocyclic peroxide **3** as the primary product of autoxidation of α -farnesene. Peroxy radical cyclization is again implicated in this example of autoxidation and thus cyclic peroxides must be considered as potential primary autoxidation products of polyunsaturated materials such as natural and synthetic rubbers and other commercially important polymers.



It should also be noted that bicyclic endoperoxides like **2** have been shown to be intermediates in prostaglandin biosynthesis and two classes of prostaglandin, PGG and PGH, have the bicyclic endoperoxide functionality incorporated in their structure.^{6,7} The chemistry of monocyclic and bicyclic endoperoxides is thus of interest with reference to uncontrolled autoxidation of naturally occurring compounds as well as in the enzymatically controlled oxidation of polyunsaturated fatty acids. Very little is known however, about the chemistry of mono and bicyclic endoperoxides like **1** and **2** since compounds of this type have not been synthetically available. Recently, however, we^{8,9} and others¹⁰ have

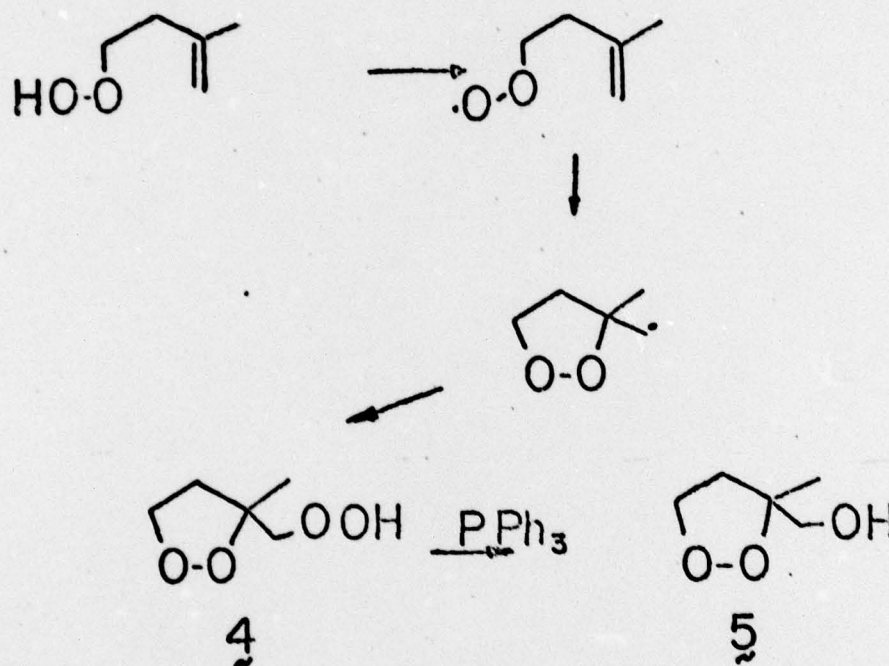
developed synthetic methods that readily provide compounds like **1** and **2** along with a variety of analogs.

B. The Synthesis of Cyclic Peroxides

1. *Radical Cyclization*

We have developed a general method for the synthesis of cyclic peroxides based on the free radical cyclization of unsaturated hydroperoxides^{4,8}. The method is described in Figure 1 and proceeds by an intermediate peroxy radical, generated from the corresponding hydroperoxide.

Figure 1.

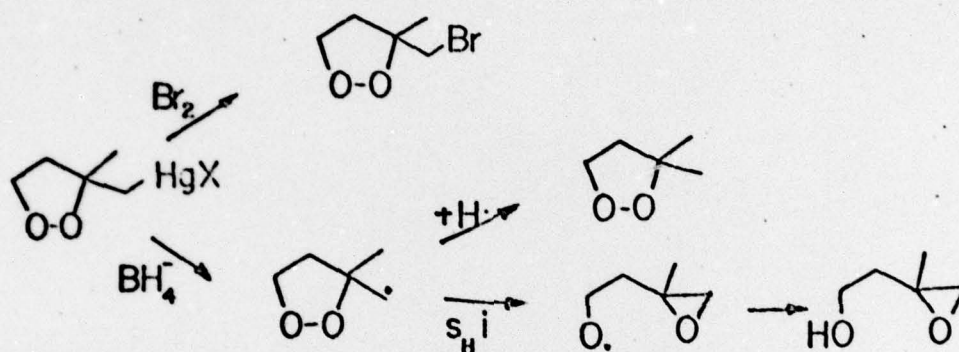


The hydroperoxide **4** and the alcohol **5** are available from this reaction scheme as are several other five and six membered ring analogs. Systematic investigation has shown that the ease of cyclization is (in the terminology proposed by Baldwin¹¹) 5 ring exocyclic (as in Fig. 1) > 6 ring exo > 7 ring exo > endocyclic cyclization.⁴

membered endocyclic products that are separated by HPLC.

All of the peroxymercuric compounds can be readily converted to the β -bromoperoxides by reaction with bromine (Fig. 2).¹³

Borohydride reduction leads to saturated cyclic peroxides and ring cleavage products, oxirane alcohols. The reduction of alkylmercuric compounds by BH_4^- proceeds via a radical mechanism¹⁴ and the mechanism of epoxide formation is shown below in Figure 2.



β -Alkyl radical attack on the peroxide linkage leads to the formation of the epoxide and the alkoxy radical. This radical attack on the peroxide linkage has ample precedent in the studies of the "unzipping" of styrene-oxygen copolymer^{13,15} and the reaction can be categorized as an intramolecular example of an $\text{S}_{\text{H}}2$ reaction (commonly referred to as $\text{S}_{\text{H}}\text{i}$).

The research supported by DAH CO 4-75-G-0117 has been a blend of mechanistic and synthetic peroxide chemistry. We have studied peroxy radical cyclizations mechanistically, and we have prepared the compounds that are favored from these cyclization reactions by independent synthetic routes.

Bibliography

1. See, for example, J. A. Howard, "Advances in Free Radical Chemistry", Vol. 4, G. H. Williams, ed., Academic Press, New York/London, (1972).
2. R. R. Hiatt, "Organic Peroxides", Vol. 2, pg. 1, D. Swern, ed., Wiley-Interscience, New York (1971).
3. R. Bateman, Quart. Rev., 8, 164 (1954).
4. N. A. Porter and M. O. Funk, J. Amer. Chem. Soc., 97, 1281(1975).
5. E. F. Anet, Aust. J. Chem., 22, 2403(1969).
6. M. Hamberg, E. Granstrom, and B. Samuelsson, P. Nat. Acad. Sci., 72, 2994(1975).
7. N. Hamberg, J. Svensson, T. Wakabayashi, and B. Samuelsson, ibid., 71, 345(1974).
8. N. A. Porter, et al., J. Amer. Chem. Soc., 98, 6000(1976).
9. N. A. Porter and D. W. Gilmore, ibid., 99, 3505(1977).
10. R. G. Salomon and M. F. Salomon, ibid., 99, 3501(1977) and references cited therein.
11. J. E. Baldwin, Chem. Commun., 734(1976).
12. N. A. Porter and J. R. Nixon, to be published.
13. A. J. Bloodworth and M. E. Loveitt, Chem. Commun., 94(1976).
14. a) B. Giese, Angew. Chem. Int. Ed., 15, 173(1976).
b) C. L. Hill and G. M. Whitesides, J. Amer. Chem. Soc., 96, 870(1974).
15. See, for example, W. G. Lloyd, "Methods in Free Radical Chemistry", 4, pg. 49, E. S. Hayser, ed., Marcel Dekker, New York (1973).