

AD-A100 503

NORTH CAROLINA UNIV AT CHAPEL HILL DEPT OF CHEMISTRY F/6 7/3
THE SYNTHESIS OF 1,1'-BICOBALTOCENE SALTS OF TETRACYANO-P-QUINO--ETC(U)
JUN 81 C LAU, P SINGH, S J CLINE, R SEIDERS N00014-76-C-0816

UNCLASSIFIED TR-16

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OFFICE OF NAVAL RESEARCH

CONTRACT ¹⁵ N00014-76-C-0816

TASK NO. NR 053-617

(9) TECHNICAL REPORT NO. 16

THE SYNTHESIS OF 1,1'-BICOBALTOCENE SALTS OF TETRACYANO-P-QUINODIMETHANIDE AND
THE STRUCTURE OF 1,1'-BICOBALTOCENE[Co(III)Co(III)][TCNQ]₃

AD A100503

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PREPARED FOR PUBLICATION

IN

Inorganic Chemistry

UNIVERSITY OF NORTH CAROLINA
DEPARTMENT OF CHEMISTRY 045A
CHAPEL HILL, NORTH CAROLINA 27514

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REPORT DOCUMENTATION PAGE		READ INSTRUCTIONS BEFORE COMPLETING FORM
1. REPORT NUMBER 16	2. GOVT ACCESSION NO. AD-A100 503	3. RECIPIENT'S CATALOG NUMBER
4. TITLE (and Subtitle) The Synthesis of 1,1'-Bicobaltocene Salts of Tetracyano-p-quinodimethanide and the Structure of 1,1'-Bicobaltocene[Co(III)Co(III)][TCNQ] ₃		5. TYPE OF REPORT & PERIOD COVERED Technical Report, Interim
		6. PERFORMING ORG. REPORT NUMBER
7. AUTHOR(s) Chak-po Lau, Phirtu Singh, Susan J. Cline, Reginald Seiders, Maurice Brookhart, Wayne E. Marsh, Derek J. Hodgson, & William E. Hatfield		8. CONTRACT OR GRANT NUMBER(s) N00014-76-C-0816
9. PERFORMING ORGANIZATION NAME AND ADDRESS Department of Chemistry 045A University of North Carolina Chapel Hill, North Carolina 27514		10. PROGRAM ELEMENT, PROJECT, TASK AREA & WORK UNIT NUMBERS
11. CONTROLLING OFFICE NAME AND ADDRESS Office of Naval Research Department of the Navy Arlington, Virginia 22217		12. REPORT DATE June 3, 1981
		13. NUMBER OF PAGES 36
14. MONITORING AGENCY NAME & ADDRESS (if different from Controlling Office)		15. SECURITY CLASS. (of this report) Unclassified
		15a. DECLASSIFICATION/DOWNGRADING SCHEDULE
15. DISTRIBUTION STATEMENT (of this Report) Approved for Public Release, Distribution Unlimited		
17. DISTRIBUTION STATEMENT (of the abstract entered in Block 20, if different from Report) This document has been approved for public release and sale; its distribution is unlimited.		
18. SUPPLEMENTARY NOTES To be published in Inorganic Chemistry		
19. KEY WORDS (Continue on reverse side if necessary and identify by block number) bicobaltocene alternating chain mixed valence compounds electrical conductivity tetracyano-p-quinodimethanide Electron paramagnetic resonance (EPR) spectra magnetic susceptibilities X-ray Crystal Structure Studies		
20. ABSTRACT (Continue on reverse side if necessary and identify by block number) The reaction of 1,1'-bicobaltocene[Co(III)Co(III)](PF ₆) ₂ or the analogous mixed valence compound [Co(III)Co(II)] with excess [C ₆ H ₅] ₃ NH][TCNQ] ₂ and TCNQ yields the complex salt 1,1'-bicobaltocene[Co(III)Co(III)][TCNQ] ₃ . The complex salt crystallizes in the triclinic system, space group P $\bar{1}$ with unit cell dimensions $a = 7.475(4)$, $b = 10.113$ (5), and $c = 14.306(3)$ Å, and $\alpha = 92.51(3)^\circ$, $\beta = 98.84(3)^\circ$, and $\gamma = 92.64(4)^\circ$. The conventional and weighted R factors for 1979 reflections for which I was greater than $3\sigma(I)$ are $R = 0.047$ and $R_w = 0.044$. The TCNQ units stack to form an alternating chain with two short, adjacent interplanar spacings of 3.15 Å, followed by a longer		

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The Synthesis of 1,1'-Bicobaltocene Salts of Tetracyano-p-quinodimethanide and the Structure of
 1,1'-Bicobaltocene[Co(III)Co(III)][TCNQ]₃

By Chak-po Lau, Phirtu Singh, Susan J. Cline, Reginald Seiders, Maurice Brookhart,
 Wayne E. Marsh, Derek J. Hodgson, and William E. Hatfield*

ABSTRACT

The reaction of 1,1'-bicobaltocene[Co(III)Co(III)](PF₆)₂ or the analogous mixed valence compound [Co(III)Co(II)] with excess [(C₂H₅)₃NH][TCNQ]₂ and TCNQ yields the complex salt 1,1'-bicobaltocene[Co(III)Co(III)][TCNQ]₃. The complex salt crystallizes in the triclinic system, space group $P\bar{1}$ with unit cell dimensions $a = 7.475(4)$, $b = 10.113(5)$, and $c = 14.306(3)$ Å, and $\alpha = 92.51(3)^\circ$, $\beta = 98.84(3)^\circ$, and $\gamma = 92.64(4)^\circ$. The conventional and weighted R factors for 1979 reflections for which I was greater than $3\sigma(I)$ are $R = 0.047$ and $R_w = 0.044$. The TCNQ units stack to form an alternating chain with two short, adjacent interplanar spacings of 3.15 Å, followed by a longer interplanar spacing of 3.45 Å. The interplanar spacing of the fulvalene moieties in the cation is 3.270 Å, and the cobalt-carbon bond distances range from 2.012(5) to 2.041(5) Å. The complex salt is diamagnetic and the electrical conductivity of a pressed pellet sample is $0.03 \text{ ohm}^{-1} \text{ cm}^{-1}$.

INTRODUCTION

The relatively high electrical conductivities of pressed pellet samples of the mixed valence compound 1,1'-biferrocene[Fe(II)Fe(III)][TCNQ]₂^{1,2} where [TCNQ]₂⁻ is the tetracyano-p-quinodimethanide radical anion dimer, continues to attract attention³ since the mechanism of the conductivity process is unknown. The analogous 1,1'-bicobaltocene salts of TCNQ are of interest in this regard since it is well known that bicobaltocene is much more easily oxidized than the iron compound, and a different degree of charge transfer, TCNQ band filling, and electrical conductivities may be anticipated in bicobaltocene salts of TCNQ. Results of our studies of the TCNQ salts of 1,1'-bicobaltocene are reported in this Article.

EXPERIMENTAL SECTION

Synthesis of Salts of Bicobaltocene.—The hexafluorophosphate salt of bicobaltocene(III,III) was prepared by the method of Davison and Smart⁴ and the orange product was purified by elution with acetone from a Florisil column. The purified copper-red compound gave two well separated, equally intense triplets at $\tau = 5.5$ and 7.5 in the NMR spectrum.

Tetracyano-p-quinodimethane (TCNQ) was obtained from Pfaltz and Bauer, Inc. and purified by recrystallization from tetrahydrofuran. The lithium salt of TCNQ was prepared by adding a boiling solution of lithium iodide in acetonitrile to a boiling solution of TCNQ in acetonitrile. The reaction mixture was allowed to stand for 4 hours and the purple solid was separated from the brown solution. The purple solid was washed with acetonitrile until the washing was bright green, and then it was washed with a large volume of diethyl ether.

The triethylammonium salt of TCNQ was prepared by adding triethylamine to a very hot solution of TCNQ in THF. After several hours, black shiny crystals of $[\text{Et}_3\text{NH}][\text{TCNQ}]_2$ separated from the intensely green colored solution. The black crystals were collected, washed with THF, and purified by recrystallization from acetonitrile.

The mixed valence salt 1,1'-bicobaltocene $[\text{Co(III)Co(II)}](\text{PF}_6)$ was prepared by the method described by Smart⁵ by reducing 1,1'-bicobaltocene $[\text{Co(III)Co(III)}](\text{PF}_6)_2$ in acetonitrile with hydrazine hydrate. The initial dark orange solution turned green immediately, and dark green microcrystals separated upon cooling. The product was collected, washed with absolute ethanol, anhydrous diethyl ether, and dried in vacuum at room temperature.

The new compound 1,1'-bicobaltocene $[\text{Co(III)Co(III)}][\text{TCNQ}]_3$ resulted from the reaction of the mixed valence hexafluorophosphate salt with a mixture of $[\text{Et}_3\text{NH}][\text{TCNQ}]_2$ and neutral TCNQ. However, it was later found that this compound is more conveniently prepared from 1,1'-bicobaltocene $[\text{Co(III)Co(III)}](\text{PF}_6)_2$, and this procedure is given here. $[\text{Et}_3\text{NH}][\text{TCNQ}]_2$ (0.262 g) and TCNQ (1.046 g) were dissolved in a minimum amount of hot acetonitrile (about 200 mL) in a closed flask under nitrogen, and a solution of the cobalt(III,III) hexafluorophosphate compound (0.170 g) in acetonitrile (20 mL) was added dropwise through a dropping funnel. After the addition of cobalt compound, the mixture was refluxed for an hour. The black precipitate which had formed was collected, and washed thoroughly with distilled water until the washing was colorless. The precipitate was then washed with acetonitrile until the washing was no longer yellow and dried in a vacuum at room temperature. The yield was 0.225 g (89%). Analyses: Calc'd for $\text{C}_{56}\text{Co}_2\text{H}_{28}\text{N}_{12}$ 68.16% C; 2.86% H; and 17.04% N. Found 68.02% C; 2.92% H; and 17.15% N.

The simple salt 1,1'-bicobaltocene $[\text{Co(III)Co(III)}][\text{TCNQ}]_2$ was prepared by the reaction of 1,1'-bicobaltocene $[\text{Co(III)Co(III)}](\text{PF}_6)_2$ with a stoichiometric amount of LiTCNQ in refluxing

acetonitrile.

Physical Characterizations.—Electron paramagnetic resonance (EPR) spectra were recorded using a Varian E-3 spectrometer, and infrared spectra were collected using a Beckmann IR-4250 spectrophotometer. Electrical conductivities of pressed pellet samples were measured using four point contacts as necessary for the van der Pauw method⁶. A constant current of 10 microamperes was supplied by a Keithley Model 227 constant current source, and the voltage drop was measured with a Keithley Model 180 nanovoltmeter. Magnetic susceptibilities were measured as a function of temperature using procedures which have been described previously.⁷ The data were corrected for the diamagnetism of the constituent atoms.⁸

X-ray Crystallographic Studies.—Deep red crystals of 1,1'-biscobaltocene[Co(III)Co(III)](TCNQ)₃ were obtained by slowly cooling a hot saturated solution of the compound in dimethylformamide. The crystals were thin and did not diffract well. A crystal of dimensions (0.30 x 0.20 x 0.06) mm was mounted on a glass fiber and used for all data collection at 19°C on an Enraf-Nonius CAD-4 computer controlled diffractometer equipped with a molybdenum X-ray tube and a graphite monochromator. The cell constants were determined by a least squares refinement of the setting angle of 25 high angle reflections. Pertinent crystallographic data are presented in Table I. Data reduction and all other calculations, except those for the drawings, were done with the CAD-4 SDP set of programs on a PDP 11/34 computer. Drawings of the structure were made using ORTEP-II.⁹ The structure was solved by locating the cobalt atom from an origin-removed, sharpened Patterson function, and the remaining atoms were located from successive difference Fourier maps. The refinement was carried out by the full matrix least squares procedure with the function minimized being

$$\sum w \{ |F_o| - |F_c| \}^2$$

The weights, w , were taken to be

$$4F_o^2 / \sigma(F_o^2)$$

All atoms except hydrogens were refined with anisotropic temperature factors. The hydrogen atoms were refined with isotropic temperature factors. The final discrepancy indices

$$R = \sum ||F_o| - |F_c|| / \sum |F_o|$$

and

$$R_w = \left\{ \sum w \left(|F_o| - |F_c| \right)^2 / \sum w |F_o|^2 \right\}^{1/2}$$

are 0.047 and 0.044, respectively. The largest shift error in the last cycle of refinement was 0.28 and the estimated standard deviation of an observation of unit weight was 1.3. The positional parameters are given in Table II. A list of structure factors and a table of thermal parameters are deposited as supplementary material.

RESULTS

Description of the structure of 1,1'-bicobaltocene[Co(III)Co(III)](TCNQ)₂.—The bond distances and angles in the fulvalene (C₅H₄—C₅H₄) moiety are shown in Figure 1. The central bond, C19—C24, of 1.465(6) Å is somewhat longer than the carbon—carbon bond distances in the rings, the latter having a mean value of 1.420 Å. The fulvalene ring bonds thus have more double bond character than the central bond connecting the two rings. The mean value of the ten ring angles in the fulvalene moiety is 108.0°. Each five membered ring of the fulvalene moiety is planar (Table III, planes VI and VII). The two five membered rings of the fulvalene moieties are, however, slightly noncoplanar (Table III, plane V). The dihedral angle between the mean planes of the two five-membered rings is 1.3°. Two cobalt atoms related by an inversion center are sandwiched between two centrosymmetrically related fulvalene molecules thus forming the complex cation, bicobaltocene, as shown in Figure 2. There is complete overlap of the corresponding atoms of the two fulvalene moieties in the cation. The distance between the planes of the fulvalene moieties is 3.270 Å and that between the two cobalt atoms is 3.912(1) Å. The distances

between the cobalt atom and the carbon atoms of fulvalene range from 2.012(5) Å to 2.041(5) Å.

The bond distances and the bond angles for the TCNQ moiety are given in Figure 3(a) for TCNQ-1 and in Figure 3(b) for TCNQ-2. The latter TCNQ entity lies on a crystallographic inversion center. These values are consistent with the usual quinonoid structure for this molecule. The estimated standard deviations in the bond distances are unfortunately too high (~ 0.007 Å) for any reasonable estimate to be made of the amount negative charge associated with the TCNQ molecule.¹⁰ The six atom ring of both TCNQ-1 and TCNQ-2 are very nearly planar (Table III, planes II and III, respectively) but the TCNQ unit as a whole is slightly nonplanar in both cases (Table III, planes I and IV).

The packing of the ions in the crystal lattice is shown in two projections; along the b -axis in Figure 4, and along the a -axis in Figure 5. It can be seen that the TCNQ anions form parallel stacks approximately along the b -axis at $c = \pm 1/2$. The bicobaltocene cations are situated in between the TCNQ anion stacks, and these cations do not form parallel planar stacks. All of the TCNQ anions along a stack are parallel to each other, the dihedral angle between the least squares planes of TCNQ-1 and TCNQ-2 (Table III, planes II and III, respectively) being only 0.2° . The stacking distances between two centrosymmetrically related TCNQ-1 anions is 3.45 Å, which is a normal distance for the van der Waals separation between such units. The stacking distance between TCNQ-1 and TCNQ-2 anions, however, is 3.15 Å, which is an extremely short interplanar separation in stacks of TCNQ anions. The modes of molecular overlap, as shown in Figures 6(a) and 6(b), are the so-called "ring external bond" type.¹⁰ There is more direct atomic overlap between two centrosymmetrically related TCNQ-1 anions, Figure 6(a), than between TCNQ-1 and TCNQ-2 anions, Figure 6(b), which may explain partly the smaller stacking distance (3.15 Å) in the latter case.

There are a number of short contacts between atoms in the two TCNQ units in the stack. The shortest distance is 3.17 Å between C1 in TCNQ-1 and C16 in TCNQ-2. Owing to the large stacking distance of 3.45 Å between the two centrosymmetric TCNQ-1 anions, there are no short contacts between atoms in the two units. In addition to those mentioned above, there are only two other nominally short contacts: C9-C27 = 3.36 Å and N6-HC20 = 2.45 Å.

The TCNQ stack exhibits an unusual alternation in the the interplanar stacking distances. Essentially the stack may be considered to be a chain of trimers with two short interplanar distances of 3.15 Å followed by a longer interplanar distance of 3.45 Å. A crude model for the electronic structure of the compound involves dicationic bicobaltocene[Co(III)Co(III)]²⁺ and dianionic [TCNQ]₃²⁻ units. However, as discussed below, the magnetic and electrical conductivity properties of the compound argue against such a simple localized description of the electronic structure.

Magnetic and electrical properties.—The magnetic properties of all of the compounds prepared in this study were measured with a vibrating sample magnetometer. All of the compounds were diamagnetic, although 1,1'-bicobaltocene[Co(III)Co(III)][TCNQ]₃ exhibited a sharp singlet at $g = 2.0033$ in the EPR spectrum.

The room temperature electrical conductivity of a pressed pellet sample of 1,1'-bicobaltocene[Co(III)Co(III)][TCNQ]₃ was measured using four point van der Pauw technique and found to be $0.03 \text{ ohm}^{-1} \text{ cm}^{-1}$. The room temperature conductivity of the simple salt 1,1'-bicobaltocene[Co(III)Co(III)][TCNQ]₂ was several orders of magnitude smaller, being $5.0 \times 10^{-6} \text{ ohm}^{-1} \text{ cm}^{-1}$.

DISCUSSION

The reaction of 1,1'-bicobaltocene[Co(III)Co(III)](PF₆)₂ with excess [Et₃NH][TCNQ]₂ and TCNQ yields the complex salt 1,1'-bicobaltocene[Co(III)Co(III)][TCNQ]₃. The formation of this complex salt was not expected, since the reaction had been designed to form a compound with a TCNQ to cobalt ratio of two. The complex salt also results from the reaction of the mixed valence compound 1,1'-bicobaltocene[Co(II)Co(II)](PF₆) with excess [Et₃NH][TCNQ]₂ and TCNQ. In this latter reaction apparently the "cobalt(II) center" is oxidized to cobalt(III) by the TCNQ. Attempts to prepare salts containing the mixed valence cobaltocene have not been successful.

The TCNQ units stack in the solid state to give an alternating chain which may be viewed as being formed by the stacking of [TCNQ]₃²⁻ units. The repeat distance in the trimer is 3.15 Å and the distance between trimers is 3.45 Å. It is of value to compare the distance between the centers of TCNQⁿ⁻ in electrically conducting TCNQⁿ⁻ chain systems. The TCNQ-TCNQ distance is 3.091(2) Å in [TTF][TCNQ] at, and below, the metal-insulator transition of approximately 60 K,¹² while the separation is 3.168(2) Å at 295 K.¹³ TCNQ-TCNQ distances in other uniformly spaced cation TCNQ chain compounds include the following: the interplanar stacking distance is 3.26 Å in [TMTSF][TCNQ] (where TMTSF is the cation of 4,4',5,5'-tetramethyl- $\Delta^{2,2'}$ -bis-1,3-diselenole)¹⁴ and [NMP][TCNQ] (where NMP is the cation of N-methylphenazine),¹⁵ 3.22 Å in [quinolinium][TCNQ],¹⁶ 3.20 Å in [HMTSF][TCNQ] (where HMTSF is the cation of $\Delta^{2,2'}$ -bis(4,5-trimethylene)-1,3-diselenole),¹⁷ and 3.45 Å in [1,1'-dimethylferrocenium][TCNQ]₂.¹¹ Neutral TCNQ also crystallizes with a chain structure. In the neutral compound the TCNQ-TCNQ interplanar spacing is 3.45 Å.¹⁸

A vast array of alternately spaced TCNQ chains exist. The simplest of these are a series

of compounds in which there is pairwise alternation of the interplanar separations. Alternating interplanar spacings of 3.15 and 3.27 Å have been observed in $[\text{MEM}][\text{TCNQ}]_2$ (where MEM is the cation of N-methyl-N-ethylmorpholine),¹⁹ while alternating interplanar spacings of 3.162(5) and 3.423(10) Å have been determined in $[\text{Fe}(\text{C}_5\text{H}_4)_2(\text{CH}_2)_3][\text{TCNQ}]_2$.²⁰ The cation in this latter compound is the trimethyleneferrocenium ion. The alternating interplanar spacings are 3.149 and 3.484 Å in the chains found in $[\text{Rb}][\text{TCNQ}]$.²¹ It is tempting to think of these latter compounds with large differences in the interplanar separations as being chains of non-interacting, tightly bound dimers. However, the semiconducting nature of the electrical conductivity of these materials indicates that there is some interaction between the dimers.

Alternating chains of TCNQ which contain tetrameric units are found in $[(\text{C}_2\text{H}_5)_3\text{NH}][\text{TCNQ}]_2$,²² the isomorphous methyltriphenylarsonium and methyltriphenylphosphonium salts of $[\text{TCNQ}]_2^-$,²³ the N-(n-propyl)-quinolinium salt of $[\text{TCNQ}]_2^-$,²⁴ the N,N'-dibenzyl-4,4'-bipyridinium salt of $[\text{TCNQ}]_4^{2-}$,²⁵ and the 1,3,3-trimethyl-2-(p-N-methyl-N-p-chloroethylstyryl)indolinium salt of $[\text{TCNQ}]_2^-$, although the spacings in the last compound are consistent with a pair of dimers.²⁶ In the triethylammonium compound the TCNQ-TCNQ interplanar separations do not differ greatly and range from 3.24 to 3.32 Å. In the other three compounds, however, there are much larger differences between the interplanar spacings within the tetrameric unit and those between tetrameric units. For example, the interplanar spacings within the tetrameric units in [methyltriphenylphosphonium] $[\text{TCNQ}]_2$ are 3.20 Å while the spacing between tetrameric units is 3.58 Å.²³

The alternating chains which exist in the 5:1 compounds formed between 1,2-bis(1-benzyl-4-pyridyl)ethane or 1,2-bis(1-benzyl-4-pyridinio)ethylene and TCNQ ^{27,28} may be thought of as stacks of pentamers with TCNQ-TCNQ spacings within the pentameric unit being 3.23-3.25 Å and the TCNQ-TCNQ spacings between pentamers being 3.36-3.41 Å. The alternating chain

compounds have electrical conductivities on the order of $0.03 \text{ ohm}^{-1} \text{ cm}^{-1}$, values which are comparable to that observed for $[\text{bicobaltocene}][\text{TCNQ}]_3$.

Trimeric units similar to those found in $[\text{bicobaltocene}][\text{TCNQ}]_2$ have also been found in $\text{Cs}_2[\text{TCNQ}]_3$ ²⁹ and $(\text{morpholinium})_2[\text{TCNQ}]_3$ ³⁰ although the alternation spacings in these two compounds are very similar, being 3.22–3.26 Å in the former and 3.24–3.25 in the latter. The alternation spacings in $[\text{bicobaltocene}][\text{TCNQ}]_3$ vary much more widely than these two examples.

The interplanar TCNQ–TCNQ spacings in discrete dimers are germane to this discussion. The interplanar spacing is 3.26 Å in $[(\text{CH}_3)_3\text{NH}]_2[\text{TCNQ}]_3$,³¹ 3.12 Å in $[\text{benzimidazolium}][\text{TCNQ}]$,³² and 3.10 Å in $[\text{Nb}_3(\mu\text{-Cl})_6(\text{C}_6(\text{CH}_3)_6)_3][(\text{TCNQ})_2]$.³³ The intratrimer interplanar spacing of 3.15 Å in $[\text{bicobaltocene}][\text{TCNQ}]_3$ is among the shortest TCNQ–TCNQ contacts that has yet been observed.

There has been considerable discussion concerning the charge distribution within TCNQ clusters based on differences in bond distances, and attention is usually focussed on the exocyclic carbon–carbon double bond. In neutral TCNQ the exocyclic carbon–carbon bond distance is 1.374(4) Å,³⁴ and the bond distance increases in the radical anion. Depending on the counterion, the exocyclic bond distance in TCNQ^- ranges from 1.401(6)³⁵ to 1.418 Å,³⁶ and the bond distance in $\text{TCNQ}^{1/2}$, that is, in $[\text{TCNQ}]_2^-$, is usually taken to be about 1.395.¹⁰ It is clear that caution must be exercised in the analysis of these data since the estimated standard deviations are within the range of the calculated differences in bond distances. Our structural data do not permit a conclusion concerning the charge distribution.

A first approximation to the bonding in the TCNQ chain involves the trimeric unit and the lowest unoccupied molecular orbital on each TCNQ molecule. Assuming D_{2h} symmetry for TCNQ,³³ the LUMO is B_{2g} , and in the trimeric unit of idealized point symmetry C_{2h} the three LUMOs combine to

give one bonding level, one antibonding level, and an intermediate level which is largely nonbonding. The two electrons associated with the charge transfer in the formulation $[\text{TCNQ}]_3^{2-}$ are accommodated in the bonding level, thereby accounting for the short interplanar spacing and the diamagnetism of the compound. This picture is incomplete since the intertrimer, interplanar TCNQ-TCNQ separation is relatively small, being 3.45 Å, and a refined description of the electronic structure must take interactions between trimers in the chain into consideration. The interactions between the trimers lead to the relatively high electrical conductivity of this semiconductor.

ACKNOWLEDGEMENTS

This research was supported in part by the Office of Naval Research.

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FIGURE CAPTIONS

Figure 1. Bond distances and bond angles in the fulvalene moiety. The estimated standard deviations in the bond angles are 0.5° . The view is approximately normal to the molecular plane with atoms drawn as thermal ellipsoids at the 40% probability level. Hydrogen and cobalt atoms are drawn as open circles of arbitrary size.

Figure 2. A view of the centrosymmetric 1,1'-bicobaltocene[Co(III)Co(III)]²⁺ cation with the atoms drawn as thermal ellipsoids at the 40% probability level. Hydrogen atoms and cobalt atoms are drawn as open circles of arbitrary size.

Figure 3. Bond distances and bond angles (a) for TCNQ-1, and (b) for TCNQ-2. The estimated standard deviations in the bond distances, if not shown are 0.007 Å. Thermal ellipsoids are drawn at the 40% level except for hydrogens which are drawn as circles of arbitrary size.

Figure 4. A view of the crystal packing along the *b*-axis. The origin is in the center of the figure with *a* going up and *c* to the right. The TCNQ-2 molecules are shown with unshaded bonds.

Figure 5. A view of the crystal packing along the *a* axis. The origin is in the center of the figure with *b* going down and *c* to the right.

Figure 6. Modes of molecular overlap in pairs of TCNQ. a.) Overlap between centrosymmetrically related TCNQ-1 anions. b.) Overlap between TCNQ-1 and TCNQ-2 anions.

TABLE 1. Crystallographic Data.

Formula: $\text{Co}_2\text{C}_{56}\text{H}_{28}\text{N}_{12}$

FW = 986.79

 $\lambda(\text{MoK}\alpha) = 0.70926 \text{ \AA}$ $\underline{a} = 7.475 (4) \text{ \AA}$ $\underline{\alpha} = 92.51 (3)^\circ$ $\underline{b} = 10.113 (5)$ $\underline{\beta} = 98.84 (3)$ $\underline{c} = 14.306 (3)$ $\underline{\gamma} = 92.64 (4)$ $U = 1066.1 \text{ \AA}^3$ Triclinic, $P\bar{1}$ $D_c = 1.54 \text{ g cm}^{-3}$ $Z = 1$ $D_o = 1.54 \text{ g cm}^{-3}$ by flotation in $\text{CHBr}_3/\text{CHCl}_3$.Intensity data collected by ω - θ scan between $2 < \theta < 30^\circ$.Number of reflections with $I > 3\sigma = 1979$. $R = 0.047$ $R_w = 0.044$

TABLE 2. Fractional Coordinates

<u>Atom</u>	<u>X</u>	<u>Y</u>	<u>Z</u>
Co	0.4856(1)	0.31552(7)	0.03517(5)
N1	-0.0394(7)	-0.0695(5)	0.1081(3)
N2	0.5127(7)	-0.0185(5)	0.2632(4)
N3	0.1159(7)	-0.2340(5)	0.7597(3)
N4	-0.4221(6)	-0.2674(5)	0.6035(3)
N5	0.4788(6)	0.6325(5)	0.3323(3)
N6	-0.0536(7)	0.5786(5)	0.1739(3)
C1	0.1116(7)	-0.1163(5)	0.3477(3)
C2	-0.0746(7)	-0.1451(5)	0.3510(4)
C3	-0.1397(7)	-0.1781(5)	0.4298(3)
C4	-0.0167(7)	-0.1845(5)	0.5167(3)
C5	0.1715(7)	-0.1550(5)	0.5157(3)
C6	0.2314(7)	-0.1214(5)	0.4350(3)
C7	0.1744(7)	-0.0795(5)	0.2656(4)
C8	0.0552(8)	-0.0738(5)	0.1769(4)
C9	0.3621(8)	-0.0469(5)	0.2645(4)
C10	-0.0843(6)	-0.2179(5)	0.5980(3)
C11	0.0307(7)	-0.2268(5)	0.6881(4)
C12	-0.2725(7)	-0.2450(5)	0.6001(4)
C13	0.0730(6)	0.5359(4)	0.4175(3)
C14	-0.1161(6)	0.5062(4)	0.4153(3)
C15	-0.1863(6)	0.4730(5)	0.4933(3)
C16	0.1449(6)	0.5726(5)	0.3361(3)
C17	0.3310(7)	0.6049(5)	0.3355(3)
C18	0.0316(7)	0.5748(5)	0.2470(3)
C19	0.3607(6)	0.4705(5)	0.0888(3)
C20	0.2306(7)	0.3681(5)	0.0470(4)
C21	0.2689(7)	0.2502(5)	0.0948(4)
C22	0.4215(8)	0.2770(5)	0.1648(3)
C23	0.4810(7)	0.4098(5)	0.1615(3)
C24	0.3709(6)	0.6083(5)	0.0618(3)
C25	0.2544(7)	0.6667(5)	-0.0112(4)
C26	0.3106(7)	0.8019(5)	-0.0133(4)

TABLE 2. (continued)

C27	0.4626(8)	0.8282(5)	0.0586(4)
C28	0.4999(8)	0.7095(5)	0.1053(3)
HC2	-0.155(4)	0.865(3)	0.306(2)
HC3	-0.267(5)	-0.189(4)	0.427(3)
HC5	0.248(5)	-0.159(4)	0.569(3)
HC6	0.347(6)	-0.122(4)	0.437(3)
HC14	-0.193(5)	0.517(4)	0.359(3)
HC15	-0.306(5)	0.453(4)	0.488(3)
HC20	0.143(6)	0.386(4)	-0.005(3)
HC21	0.192(6)	0.161(5)	0.076(3)
HC22	0.482(5)	0.209(4)	0.207(2)
HC23	0.569(5)	0.462(4)	0.200(3)
HC25	0.158(6)	0.615(4)	-0.055(3)
HC26	0.259(5)	0.860(4)	-0.064(3)
HC27	0.555(6)	0.906(5)	0.075(3)
HC28	0.594(5)	0.701(3)	0.156(2)

TABLE 3. Deviations ($\times 10^3 \text{ \AA}$) of Atoms from Least-Squares Planes through Various Moieties

TCNQ Planes					
Atom	I	II	Atom	III ^a	IV ^a
C1	29	7	C13	-3	-1
C2	11	-3	C14	3	-32
C3	8	-2	C15	-3	-40
C4	15	2	C16		-15
C5	22	2	C17		-7
C6	17	-7	C18		8
C7	1		N5		-12
C8	12		N6		21
C9	-23				
C10	5				
C11	14				
C12	-29				
N1	19				
N2	-62				
N3	13				
N4	-52				
Bisfulvalene Planes					
	V	VI	VII		
C19	1	-9	-5 ^b		
C20	-2	6	-23 ^b		
C21	-14	-1	-35 ^b		
C22	-5	-5	-10 ^b		
C23	24	9	27 ^b		
C24	4	-17 ^b	2		
C25	11	-6 ^b	-1		
C26	3	-28 ^b	-1		
C27	-9	-55 ^b	2		
C28	-13	-53 ^b	-2		

^a Each atom in this plane has a corresponding atom related by a center of inversion, the latter being also included in the plane calculation.

^b Atom not included in the plane calculation.

TABLE 4. Thermal Parameters for (BFDco)₂(TCNQ)₃. Form of the Thermal Ellipsoids is

$$\exp[-(\beta_{11}h^2 + \beta_{22}k^2 + \beta_{33}l^2 + \beta_{12}hk + \beta_{13}hl + \beta_{23}kl)].$$

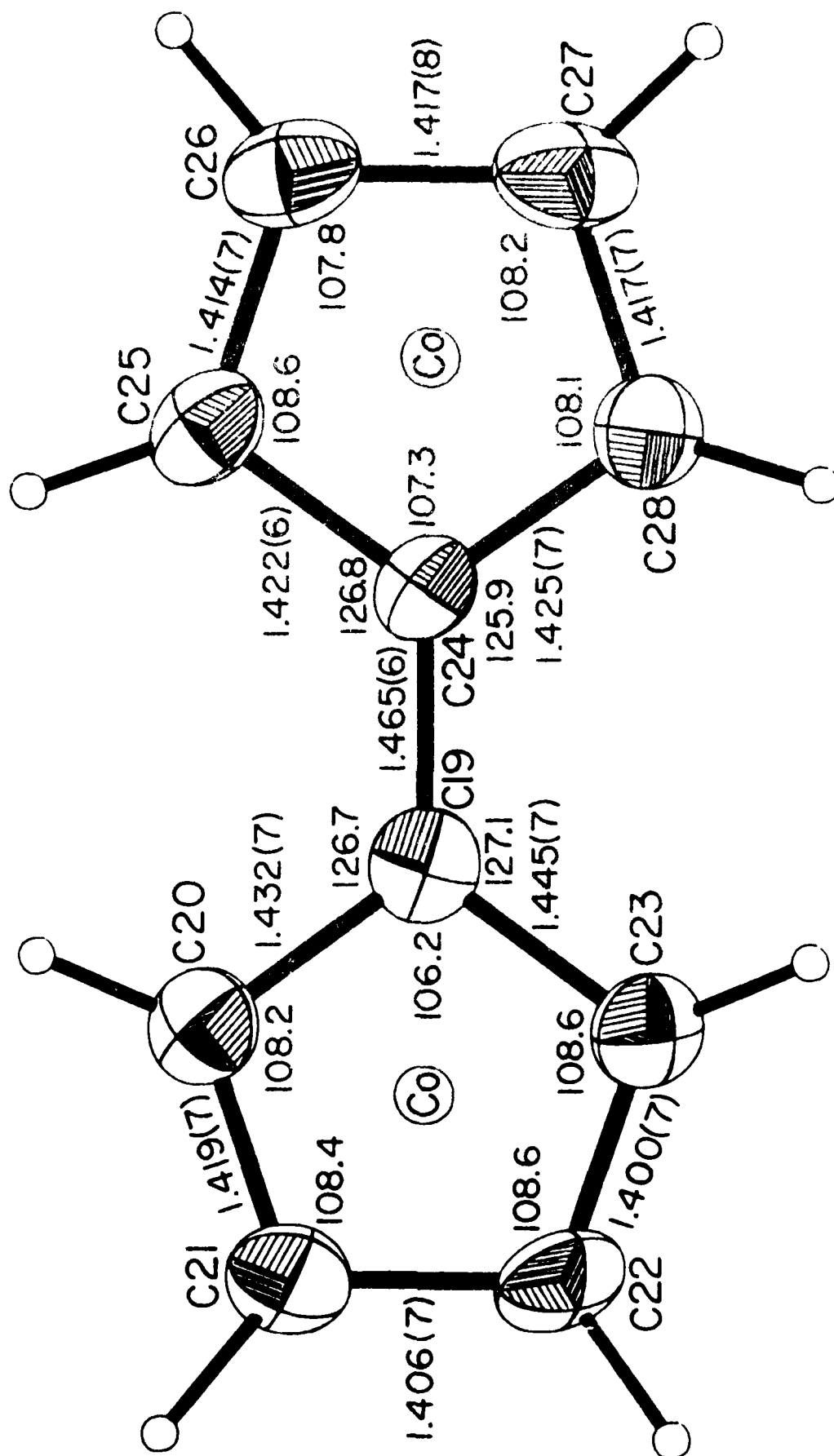
	β_{11}	β_{22}	β_{33}	β_{12}	β_{13}	β_{23}
Co	0.0135(1)	0.00768(7)	0.00326(3)	0.0006(2)	0.0015(1)	0.00161(8)
N1	0.033(1)	0.0170(7)	0.0053(3)	0.003(2)	-0.0007(11)	0.0052(8)
N2	0.026(1)	0.0110(6)	0.0084(3)	-0.001(1)	0.0078(11)	0.0025(8)
N3	0.030(1)	0.0170(7)	0.0052(3)	0.002(2)	-0.0023(11)	-0.0008(8)
N4	0.019(1)	0.0172(7)	0.0064(3)	0.003(1)	0.0037(10)	-0.0022(8)
N5	0.016(1)	0.0128(6)	0.0064(3)	0.001(1)	0.0042(9)	-0.0010(7)
N6	0.027(1)	0.0146(7)	0.0042(3)	-0.002(2)	-0.0027(10)	0.0017(7)
C1	0.016(1)	0.0060(5)	0.0047(3)	0.005(1)	-0.0001(10)	-0.0008(7)
C2	0.017(1)	0.0096(6)	0.0040(3)	0.008(1)	-0.0065(10)	-0.0004(7)
C3	0.013(1)	0.0072(6)	0.0042(3)	0.005(1)	-0.0002(9)	-0.0008(7)
C4	0.015(1)	0.0059(5)	0.0038(3)	0.006(1)	0.0011(10)	-0.0003(7)
C5	0.015(1)	0.0076(6)	0.0037(3)	0.004(1)	-0.0044(10)	-0.0026(7)
C6	0.014(1)	0.0075(6)	0.0045(3)	0.007(1)	0.0006(10)	-0.0006(7)
C7	0.022(1)	0.0070(6)	0.0044(3)	0.003(1)	0.0033(10)	0.0022(7)
C8	0.025(1)	0.0089(6)	0.0048(3)	0.002(2)	0.0040(11)	0.0031(8)
C9	0.025(1)	0.0069(6)	0.0050(3)	-0.000(2)	0.0055(11)	0.0011(7)
C10	0.011(1)	0.0075(6)	0.0044(3)	0.003(1)	0.0006(9)	-0.0023(7)
C11	0.020(1)	0.0090(6)	0.0042(3)	0.001(2)	0.0023(10)	-0.0009(7)
C12	0.020(1)	0.0104(6)	0.0034(3)	0.006(1)	0.0021(10)	-0.0006(7)
C13	0.011(1)	0.0044(5)	0.0041(3)	0.001(1)	-0.0014(9)	-0.0009(6)
C14	0.012(1)	0.0064(5)	0.0032(3)	0.004(1)	-0.0026(9)	0.0014(6)

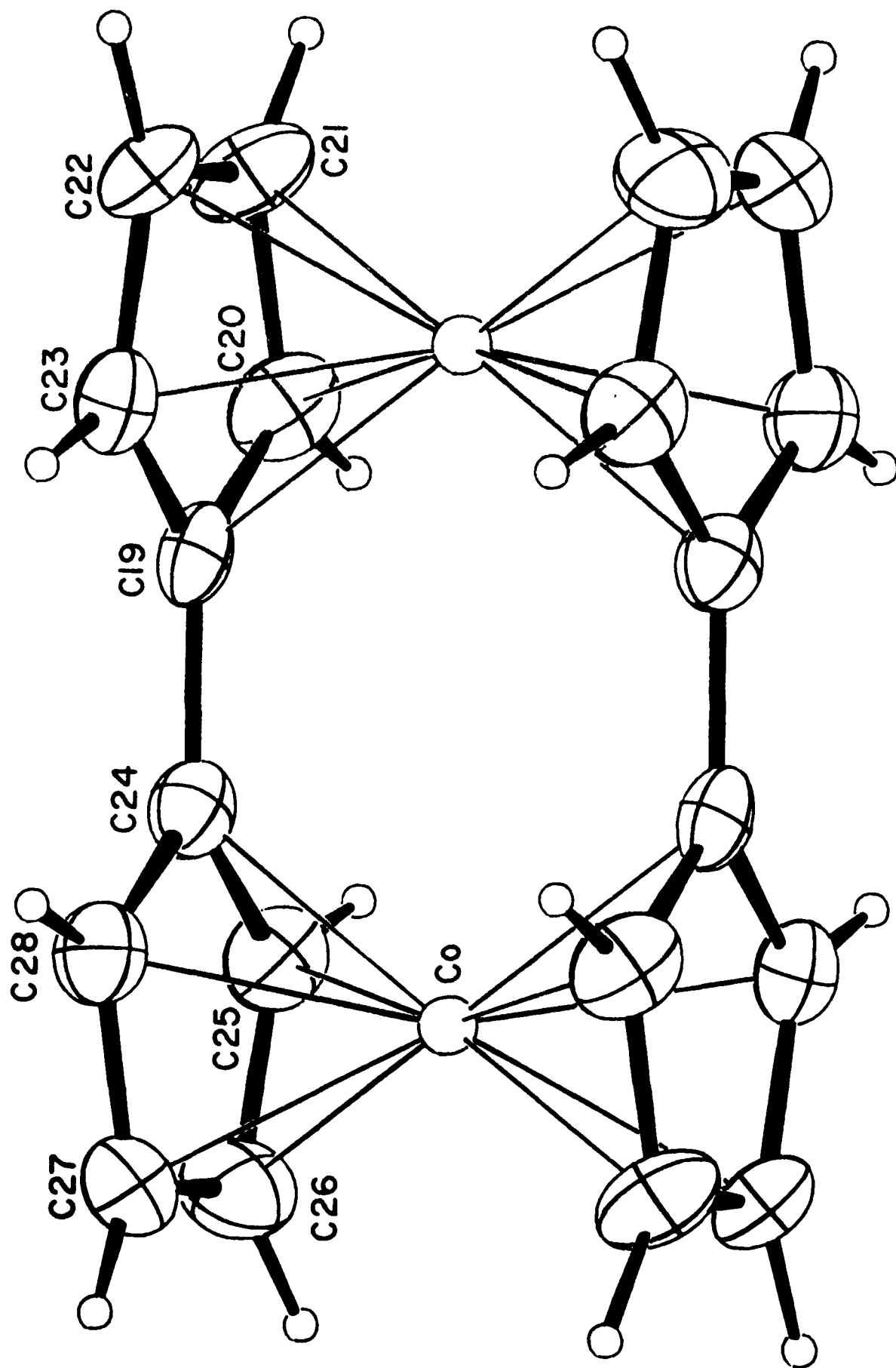
TABLE 4.. (continued)

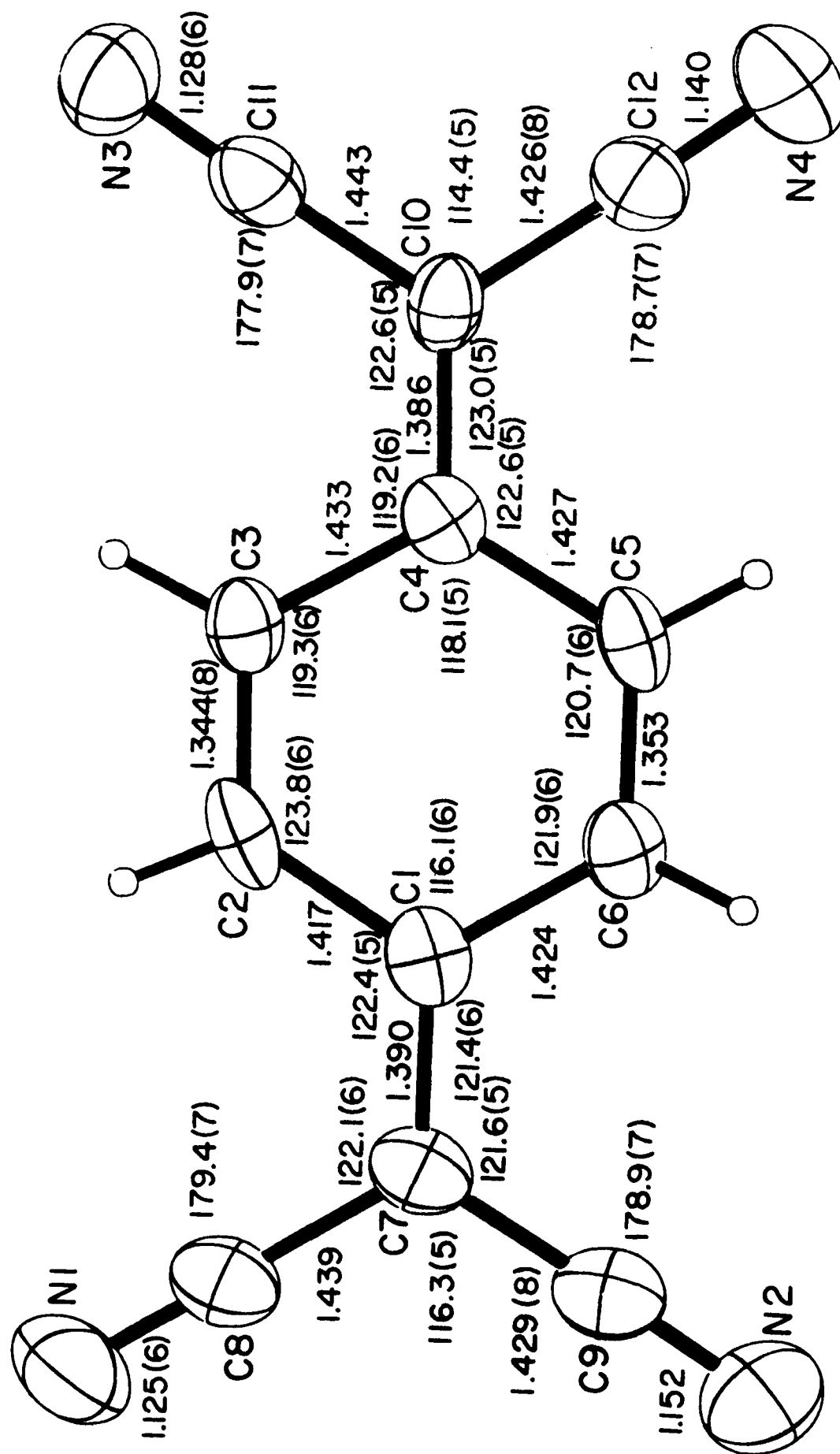
C15	0.011(1)	0.0075(6)	0.0043(3)	0.001(1)	-0.0002(9)	0.0013(7)
C16	0.013(1)	0.0075(6)	0.0036(3)	0.003(1)	0.0008(9)	0.0008(7)
C17	0.016(1)	0.0087(6)	0.0034(3)	0.004(1)	0.0004(10)	0.0006(7)
C18	0.018(1)	0.0073(6)	0.0040(3)	0.000(1)	0.0027(10)	-0.0004(7)
C19	0.013(1)	0.0100(6)	0.0031(2)	0.002(1)	0.0066(8)	0.0008(7)
C20	0.010(1)	0.0113(7)	0.0053(3)	0.000(1)	0.0040(9)	0.0032(8)
C21	0.017(1)	0.0113(7)	0.0057(3)	-0.006(1)	0.0056(10)	0.0060(8)
C22	0.026(1)	0.0097(6)	0.0033(3)	0.003(2)	0.0045(10)	0.0050(7)
C23	0.021(1)	0.0097(6)	0.0027(3)	0.002(2)	0.0028(10)	-0.0000(7)
C24	0.012(1)	0.0087(6)	0.0034(3)	0.005(1)	0.0032(8)	0.0012(7)
C25	0.012(1)	0.0107(6)	0.0049(3)	0.007(1)	0.0041(9)	0.0030(7)
C26	0.020(1)	0.0089(6)	0.0067(4)	0.010(1)	0.0061(11)	0.0026(8)
C27	0.034(2)	0.0075(6)	0.0045(3)	0.001(2)	0.0066(12)	-0.0015(8)
C28	0.024(1)	0.0093(6)	0.0030(3)	-0.001(2)	0.0003(10)	-0.0003(7)
HC2	0.9(8)					
HC3	3.8(11)					
HC5	2.4(9)					
HC6	4.6(12)					
HC14	3.5(10)					
HC15	2.6(10)					
HC20	4.3(11)					
HC21	5.4(13)					
HC22	2.2(9)					
HC23	2.3(9)					
HC25	4.3(11)					
HC26	3.9(11)					

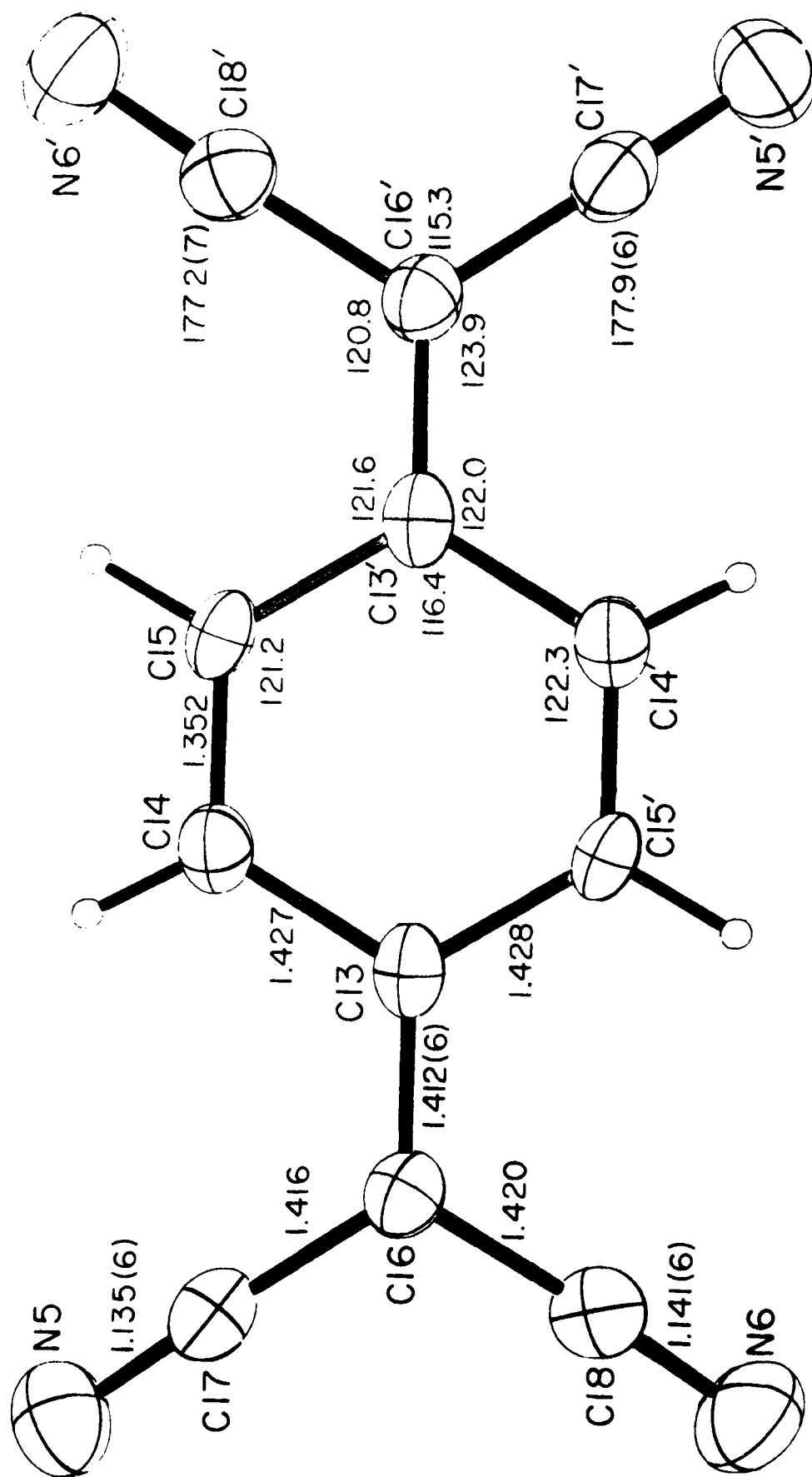
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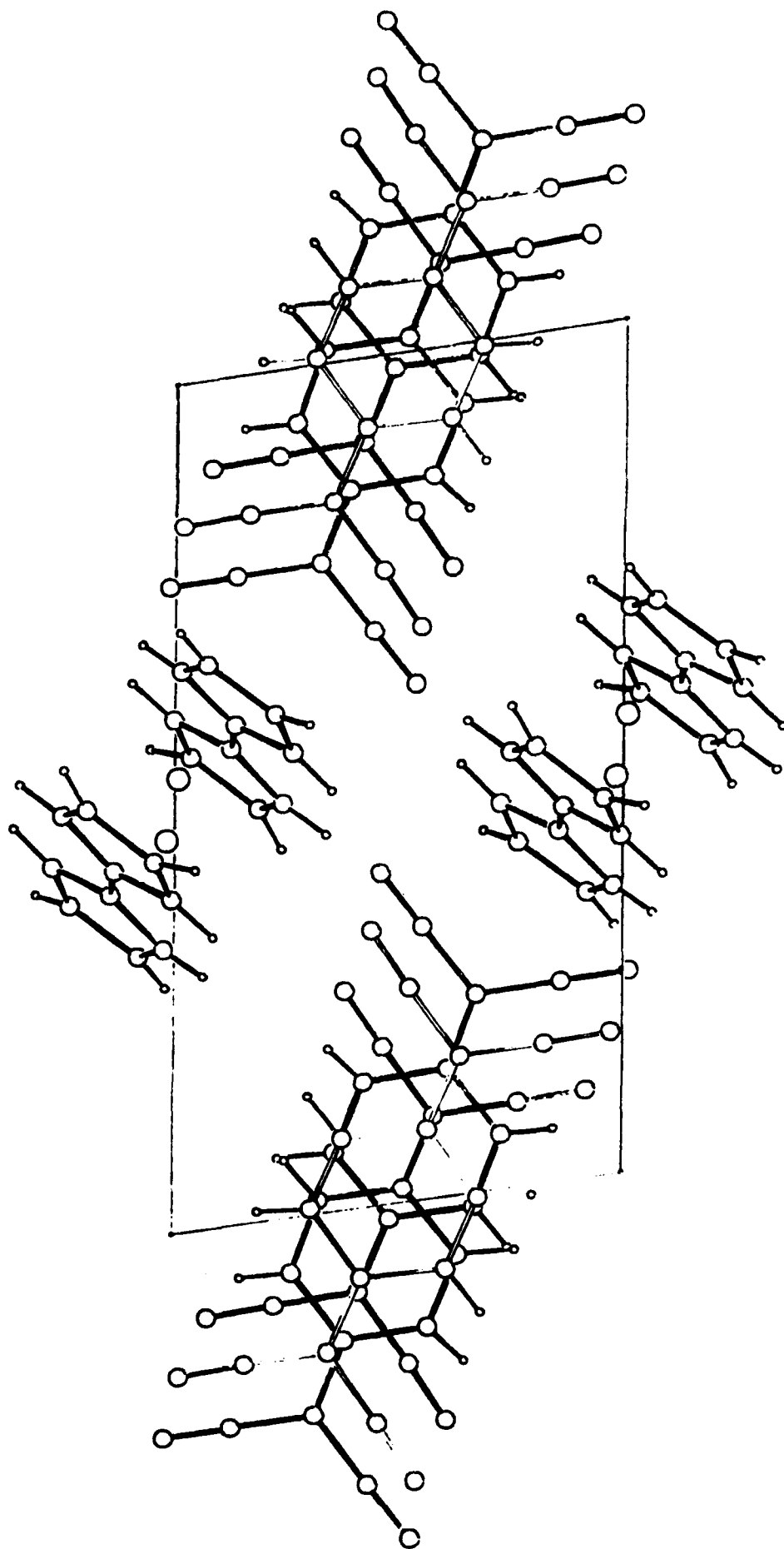
HC27	5.7(13)
HC28	1.8(9)

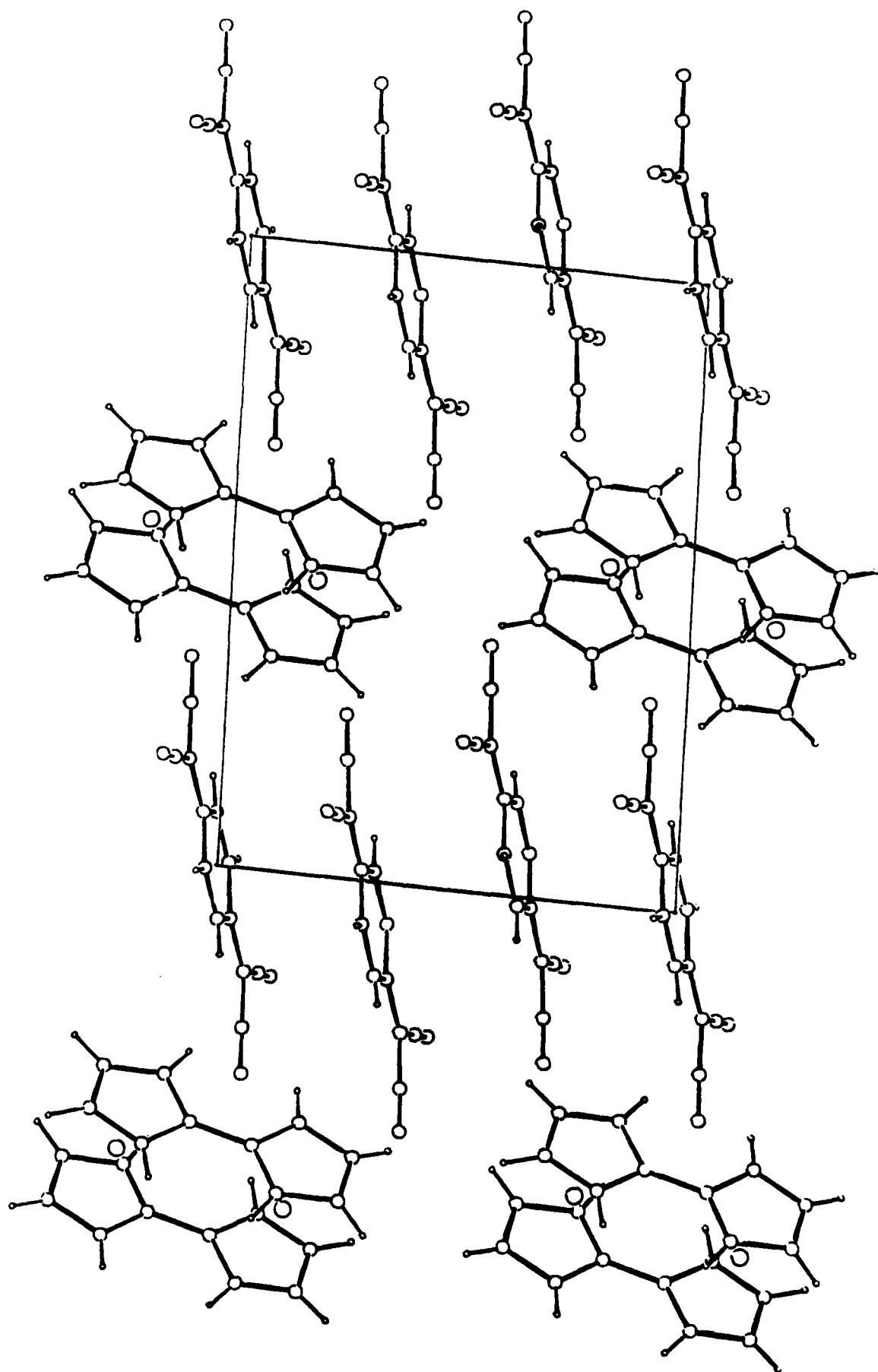


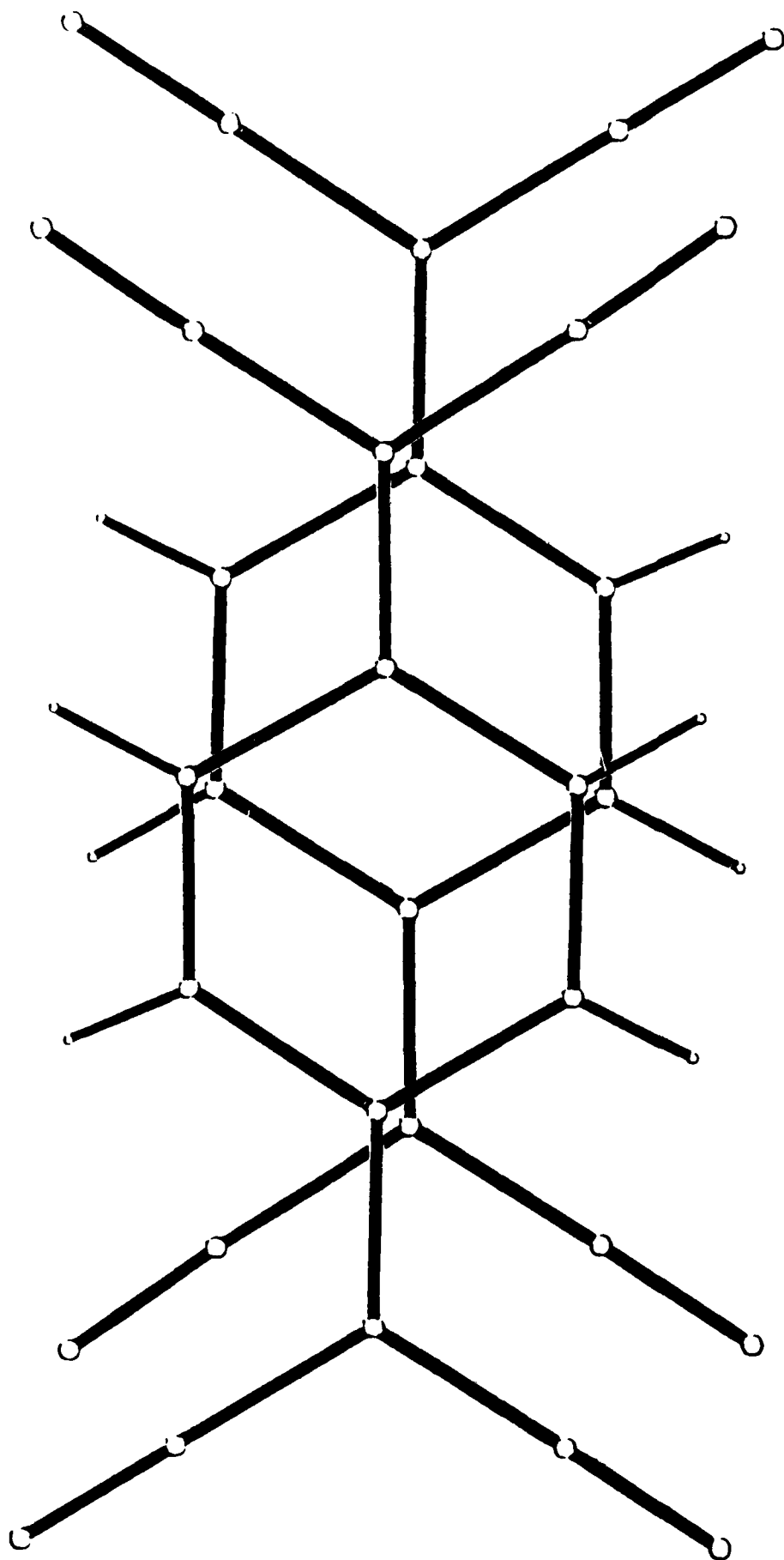


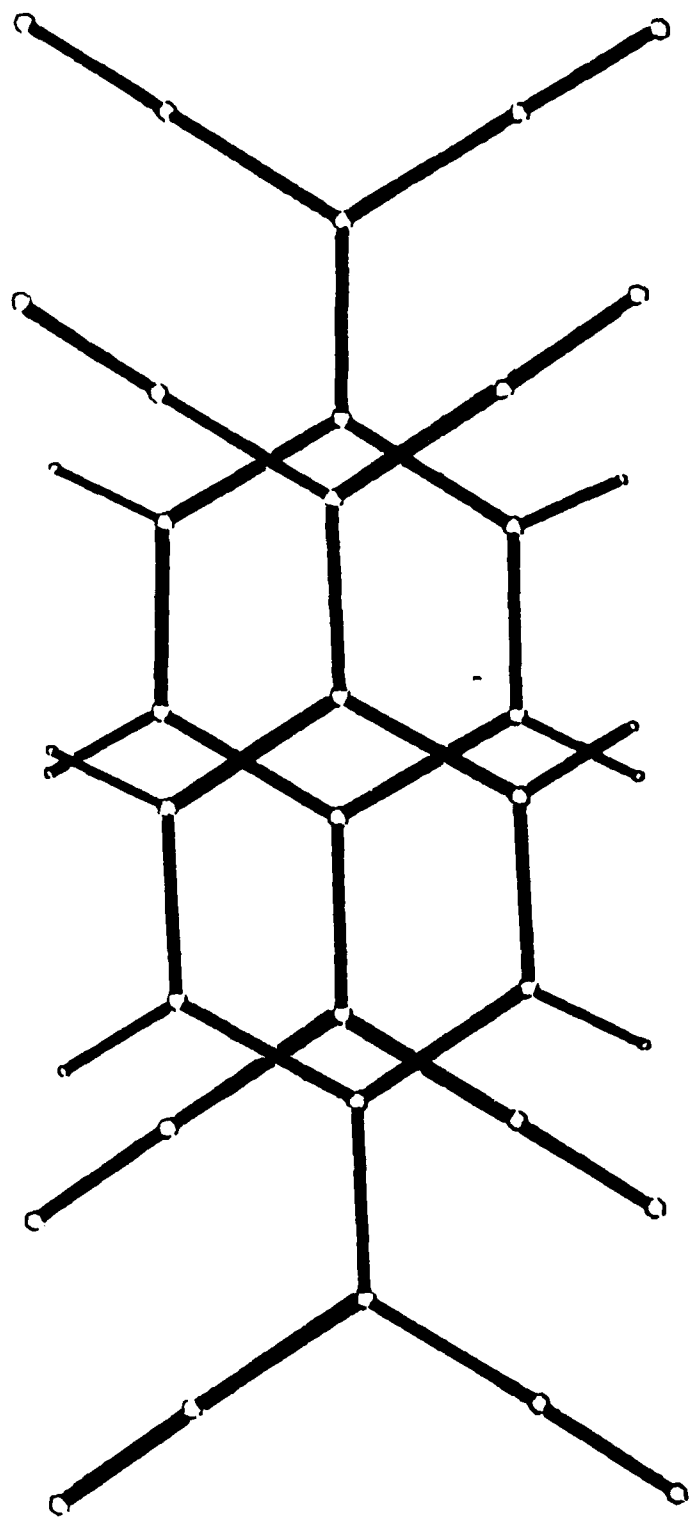












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