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OKLAHOMA UNIV NORMAN DEPT OF CHEMISTRY
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VARIABLE-TEMPERATURE TIN-119M MOESSBAUER STUDY OF THE CHLORODIN--ETC(U)
JUN 81 K C MOLLOY, J J ZUCKERMAN, G DOMAZETIS N00014-77-C-0432

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OFFICE OF NAVAL RESEARCH
Contract N00014-77-C-0432
Task No. NR 053-636
TECHNICAL REPORT No. 31

AD A102483

Variable-Temperature Tin-119m Mössbauer Study of the
Chlorodimethyltin(IV) Derivatives of
L-Cysteine and DL-Penicillamine

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AUG 5 1981
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Prepared for Publication

in

Inorganic Chemistry

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SECURITY CLASSIFICATION OF THIS PAGE (When Data Entered)

REPORT DOCUMENTATION PAGE		READ INSTRUCTIONS BEFORE COMPLETING FORM
1. REPORT NUMBER TR-31	2. GOVT ACCESSION NO. AD-A202 483	3. RECIPIENT'S CATALOG NUMBER
4. TITLE (and Subtitle) Variable-Temperature Tin-119m Mössbauer Study of the Chlorodimethyltin(IV) Derivatives of L-Cysteine and DL-Penicillamine		5. TYPE OF REPORT & PERIOD COVERED
7. AUTHOR(s) K.C. Molloy and J.J. Zuckerman G. Domazetis E.D. James		6. PERFORMING ORG. REPORT NUMBER
9. PERFORMING ORGANIZATION NAME AND ADDRESS Department of Chemistry University of Oklahoma Norman, OK 73019		8. CONTRACT OR GRANT NUMBER(s) N00014-77-C-0432 NSF-74E78-26584
11. CONTROLLING OFFICE NAME AND ADDRESS Office of Naval Research Department of the Navy Arlington, Virginia 22217		10. PROGRAM ELEMENT, PROJECT, TASK AREA & WORK UNIT NUMBERS NR 053-636
14. MONITORING AGENCY NAME & ADDRESS (if different from Controlling Office) (9) Technical Rpts		12. REPORT DATE 15 June, 1981
16. DISTRIBUTION STATEMENT (of this Report) Approved for Public Release, Distribution Unlimited		13. NUMBER OF PAGES 13
17. DISTRIBUTION STATEMENT (of the abstract entered in Block 20, if different from Report) Prepared for Publication in Inorganica Chimica Acta		15. SECURITY CLASS. (of this report) Unclassified
18. SUPPLEMENTARY NOTES		
19. KEY WORDS (Continue on reverse side if necessary and identify by block number) Organotins, Cysteine, Penicillamine, Mössbauer spectroscopy, Hydrogen bonding, Variable-Temperature Mössbauer, Tin		
20. ABSTRACT (Continue on reverse side if necessary and identify by block number) The question of intermolecular association in the chlorodimethyltin(IV) derivatives of L-cysteine (a monohydrate) and DL-penicillamine is settled by variable-temperature tin-119m Mössbauer data between 77 and 140 and 150 K, respectively. The slopes of the logs of the areas under the resonance are $-1.58 \times 10^{-2} \text{ K}^{-1}$ and $-1.46 \times 10^{-2} \text{ K}^{-1}$, respectively, considerably lower than those for monomeric lattices. These data rule out an extended association through bridging amino nitrogen atoms or carboxyate oxygen atoms, and specify a lattice in which small		

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S/N 0102-LF-014-6601

SECURITY CLASSIFICATION OF THIS PAGE (When Data Entered)

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20. small molecular or oligomeric units are associated by an extensive network of hydrogen bonds. An alternative ionic structure of $\text{Cl}(\text{CH}_3)_2\text{SnSCR}_2\text{CH}(\text{NH}_3^+)\text{COO}^-$ units where $\text{R}=\text{H}, \text{CH}_3$ cannot be ruled out on the basis of infrared data [$\nu(\text{OH})=3450 \text{ cm}^{-1}$].

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$$-1.46 \pm 0.1 / 100 / K$$

Abstract

The question of intermolecular association in the chlorodimethyltin(IV) derivatives of L-cysteine (a monohydrate) and DL-penicillamine is settled by variable-temperature tin-119m Mössbauer data between 77 and 140 or 150 K, respectively. The slopes of the natural logs of the areas under the resonance are $-1.58 \pm 0.1 \times 10^{-2} \text{ K}^{-1}$ and $-1.46 \pm 0.1 \times 10^{-2} \text{ K}^{-1}$, respectively, considerably lower than those for monomeric lattices. These data rule out an extended association through bridging amino nitrogen atoms or carboxyate oxygen atoms, and suggest a lattice in which small molecular or oligomeric units are associated by an extensive network of hydrogen bonds. An alternative ionic structure of $\text{Cl}(\text{CH}_3)_2\text{SnSCR}_2\text{CH}(\text{NH}_3^+)\text{COO}^-$ units where $\text{R}=\text{H}$ can be ruled out on the basis of infrared data [$\nu(\text{OH})=3450/\text{cm}^{-1}$].

$$-1.53 \pm 0.1 / 100 / K$$



INTRODUCTION

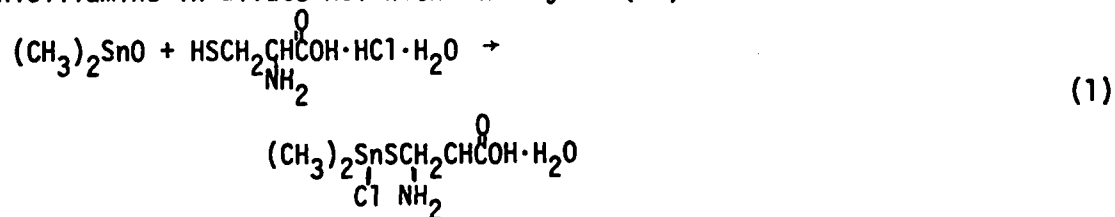
Organotin compounds play an important role in biology and the environment [2,3] and are themselves powerful biocides [4]. We have been pursuing studies in the synthesis [4-6] and structural characterization [8] of organotin(IV) derivatives of amino acids to help to elucidate their mode of action, and have extended these studies to some biologically relevant thio ligands [9-13]. We report in this paper the results from a variable-temperature tin-119m Mössbauer study of the L-cysteine monohydrate and DL-penicillamine derivatives of chlorodimethyltin(IV) which distinguishes between the several structural possibilities for these solids.

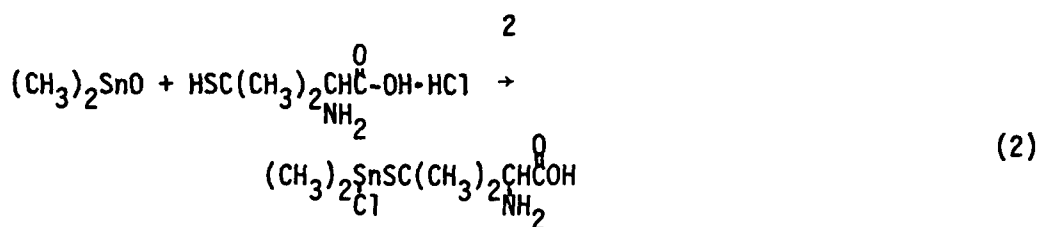
EXPERIMENTAL

Tin-119m Mössbauer spectra were recorded on a Ranger Engineering constant-acceleration spectrometer equipped with an NaI scintillation counter using $\text{Ca}^{119}\text{SnO}_3$ (New England Nuclear Corp.) as the source and $\text{Ca}^{119}\text{SnO}_3$ and β -tin as standard reference materials for velocity calibration at room temperature. Standard, nonlinear, least-squares techniques were used to fit the data to Lorentzian curves. The Ranger Engineering variable-temperature, liquid-nitrogen dewar and controller used in these studies are regulated by a variable-bridge, silicon-controlled-rectifier circuit and is accurate to ± 1 K.

RESULTS AND DISCUSSION

The products from the reaction of L-cysteine hydrochloride hydrate and DL-penicillamine in dilute HCl with dimethyltin(IV) oxide contain sulfur-tin bonds:





The former is obtained in quantitative yield as a monohydrate, the latter in 90% yield anhydrous. Titration of the cysteine derivative with sodium hydroxide removes the carboxylate and amino protons stepwise, and the absence of uv bands assignable to the cysteinate and chlorodimethyltin ions rules out a simple ionic formulation in aqueous solution. Infrared absorptions for the Sn-S bond near 400 cm^{-1} with the corresponding absence of the $\nu(\text{SH})$ mode, together with mass spectral fragments such as $[\text{Sn-SCH}_2\text{CHNH}_2]^+$ lend corroborative evidence for the sulfur bonding [9-11].

The structures of the solids are of interest because of the possibilities for oligomeric or polymeric behavior, as found in the organotin derivatives of the simple amino acids [5-8]. In both products the $\nu(\text{C=O})$ modes are reduced in frequency, presumably because of association to tin or strong hydrogen bonding. The $|^1J(^{119}\text{Sn}-^{13}\text{C})|$ and $|^2J(^{119}\text{Sn}-\text{C}-^1\text{H})|$ coupling constants in D_2O solution are indicative of higher coordinated tin atoms in that solvent [9-11].

Mössbauer spectroscopy can provide useful information in such cases. The Isomer Shift and Quadrupole Splitting data for these thiol derivatives [13] are consistent with five-coordinated structures, and the spectrum of the penicillamine derivative - being observable at ambient temperatures - is suggestive of an associated lattice. Such association can arise either from intermolecular association through bridging amino groups, as in solid trimethyltin(IV) glycinate [8], bridging carboxylate groups as in the organotin acetates [14,15], or through intermolecular hydrogen bonding which is also present in the structure of the

solid glycinate where it stabilizes the lattice. Indeed, the crystal in that case propagates along the direction of the interchain hydrogen bonds and not along the propagation axis of the polymer chain, indicating that hydrogen bonding makes a more important contribution to the lattice energy than does the coordinate-covalent bonding to tin [8]. In contrast, the crystal structure of the ethyl ester of the cysteine derivative contains a monomeric, five-coordinated, trigonal bipyramidal tin atom bonded to sulfur and chelated by the amino group^[11,12] which is also intramolecularly hydrogen bonded to the carbonyl. Thus the question of the lattice structure adopted in these species is a complex one and requires further investigation.

Variable temperature Mössbauer studies can distinguish these possibilities. The Mössbauer recoil-free fraction, f , reflects the mean square displacement of the tin atom $\langle \chi^2 \rangle$ from its equilibrium position:

$$f = \exp \left[- \frac{\langle \chi^2 \rangle}{\lambda^2} \right] \quad (3)$$

where λ is the wavelength of the tin-119m γ -ray divided by 2π . By the Debye model, for thin absorbers, f is linearly related to the area under the resonance, A_T , whose temperature dependence is given by:

$$A_T \left\{ f = \exp \left[- \frac{6E_R T}{k\theta_D^2} \right]; \text{ for } T \geq \frac{\theta_D}{2} \right. \quad (4)$$

where E_R is the recoil energy, and θ_D is a temperature which for an ideal monatomic solid is equivalent to the Debye temperature. The more tightly held are the tin atoms in a lattice, the slower will f - and hence A_T - decrease with increasing temperature. For methyltin derivatives, lattices of monomers exhibit slopes of $\log_e A_T$ vs. T of ca. $-1.8 \times 10^{-2} \text{ K}^{-1}$, regardless of the coordination number at tin. Weak hydrogen bonding reduces this value to ca. $-1.7 \times 10^{-2} \text{ K}^{-1}$, while complex networks of hydrogen bonding reduces it still more to ca. $-1.2 \times 10^{-2} \text{ K}^{-1}$. Strongly hydrogen bonded lattices and solids in which one- and higher dimension association is present exhibit slopes of ca. $-0.9 \times 10^{-2} \text{ K}^{-1}$ [16,17]. A

particularly low value is given by tin(II) oxide at $-0.23 \times 10^{-2} \text{ K}^{-1}$ [16,18-20], and tetramethyltin is at the other end of the methyltin scale at $-2.76 \times 10^{-2} \text{ K}^{-1}$ [21].

Table 1 lists the slope data for chlorodimethyltin(IV) cysteine between 77 and 140 K which are plotted as $\ln A$ (normalized to the area under the resonance curve at 77 K) vs. temperature in K in Figure 1. The analogous data for the DL-penicillamine derivatives between 77 and 150 K are given in Table 1 and plotted in Figure 2. The slopes are found to be $-1.58 \pm 0.1 \times 10^{-2} \text{ K}^{-1}$ (intercept at $T=0 \text{ K}$, 1.045, correlation coefficient 0.973 for 7 points) and $-1.46 \pm 0.1 \times 10^{-2} \text{ K}^{-1}$ (intercept at $T=0 \text{ K}$, 1.215, correlation coefficient 0.999 for 7 points), respectively.

These values are higher than for trimethyltin(IV) glycinate ($-1.15 \times 10^{-2} \text{ K}^{-1}$) in which both hydrogen bonding and a one-dimensional association through bridging amino group nitrogen atoms are present [8], but considerably lower than those for monomeric lattices [16,17,21].

Against the background of these systematics for methyltin derivatives [22], the magnitudes of the slopes of the plots of A_T vs. temperature for the two chlorodimethyltin(IV) derivatives suggest a lattice in which small molecular or oligomeric units [23] are associated by a relatively extensive network of hydrogen bonds. The slope values serve to rule out an extended association through bridging amino nitrogen atoms as in the trimethyltin(IV) glycinate [8], or by carboxylate oxygen atoms as in the bridged acetates for which many examples are available [14,15], and are consistent with the dimeric, hydrogen-bonded structure originally proposed on the basis of strong ditin fragments in the mass spectra [9,25].

The hydrogen bonding in the chlorodimethyltin(IV) cysteine monohydrate involves both the water of hydration and the $-\text{NH}_2$ group on the basis of the broad absorptions seen at 3450 ($\nu\text{O-H}$) and 3000 cm^{-1} ($\nu\text{N-H}$) in the infrared spectrum [9] and may involve the carbonyl group as in the ethyl ester [11,12].

In the corresponding anhydrous DL-penicillamine derivative which exhibits an ambient temperature Mössbauer spectrum and a shallower slope, the hydrogen bonding can only involve the -NH_2 and carboxylate groups. The strongly lowered infrared $\nu(\text{C=O})$ frequencies in these derivatives (1613 vs,br, 1620 s cm^{-1} , respectively [9]) compared with the monomeric ethyl ester [11,12] (1733 vs cm^{-1} [9]), and the similarity of the $\nu(\text{NH}_2)$ frequencies (3450 s,br, 3450 s,br, respectively vs. 3309 s 3243 s cm^{-1} [9]) specify that carboxylate rather than amino group bridging links the molecular units into the dimer or small oligomer. An alternative structure involving $\text{Cl}(\text{CH}_3)_2\text{SnSCR}_2\text{CH}(\text{NH}_3^+)\text{COO}^-$ units in which $\text{R}=\text{H}$, CH_3 can be ruled out by the presence of $\nu(\text{OH})$ stretches in the infrared spectra of these compounds [9].

Acknowledgement

This work is supported by the U.S. Office of Naval Research and the U.S. National Science Foundation through grant CHE 78-26584 (to J.J.Z.). A Commonwealth of Australia Postgraduate Research Award is acknowledged by G.D.

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- 22 The effect of passing from the trimethyltin (whose slopes we are comparing) to our dimethyltin chlorides is not known, but the monomeric tetraphenyltin and diphenyltin dichloride solids [14,15] have very similar slopes (-1.66 [20,21] and $-1.71 \times 10^{-2} \text{K}^{-1}$ [20]).
- 23 These data do not specify the molecularity of the unit. Attempts to determine the value of n in the formula $[\text{Cl}(\text{CH}_3)_2\text{SnSCH}_2\text{CH}(\text{NH}_2)\text{COOH}]_n$ by use of the effective vibrating mass model [24] proved inconclusive since no bands $< 200 \text{ cm}^{-1}$ were found in the Raman spectrum which could be assigned to the inter-molecular, intra-unit cell modes.
- 24 R. H. Herber and M. F. Leahy, in ref. 4, page 155.
- 25 Very weak ($< 1\%$) ditin fragments in the mass spectrum of the ethyl ester were shown from a thermogram to arise from decomposition at the probe temperature (160°C) [9].

Table I. Variable Temperature Tin-119m Mössbauer Data for Chlorodimethyltin(IV)
L-Cysteine Monohydrate and Chlorodimethyltin(IV) DL-Penicillamine

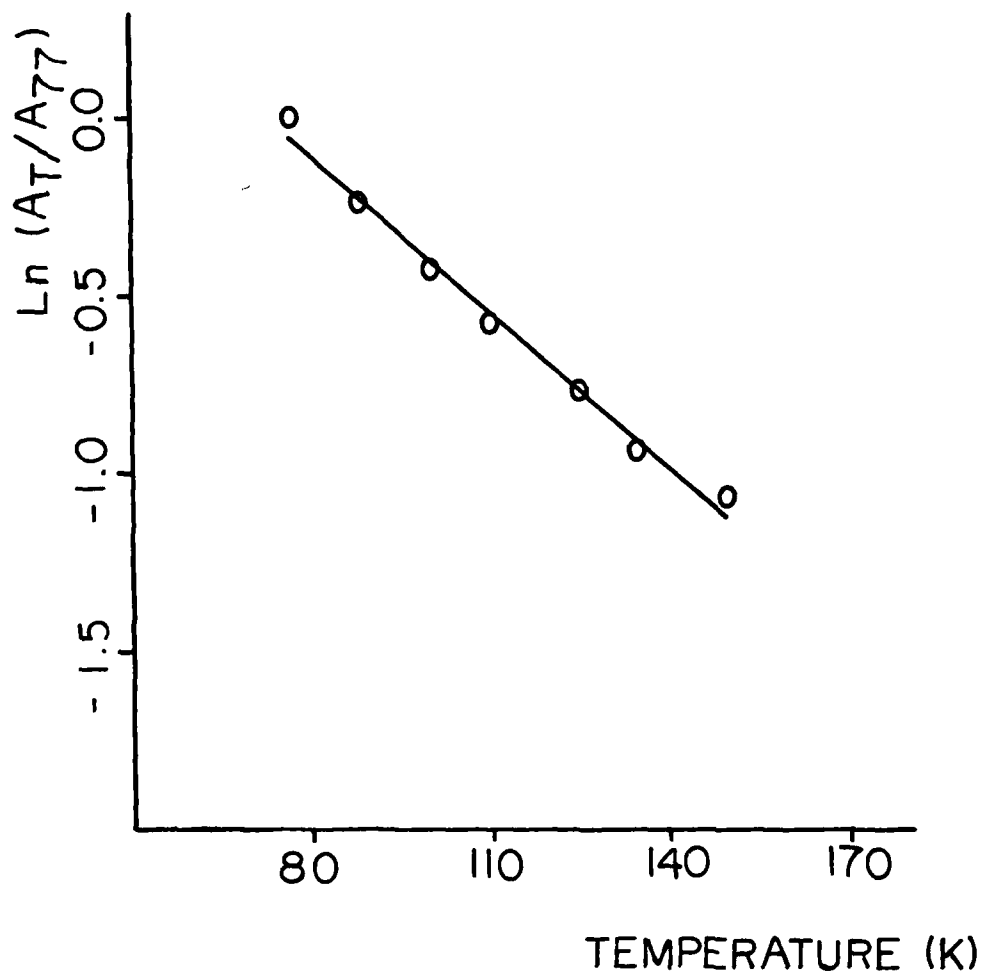
<u>T(K)</u>	<u>L-Cysteine</u>		<u>DL-Penicillamine</u>	
	<u>A(T)^a</u>	<u>log_e [A(T)/A(77)]</u>	<u>A(T)^a</u>	<u>log_e [A(T)/A(77)]</u>
77	0.487	0.000	0.439	0.000
88	0.414	-0.162	0.346	-0.238
100	0.332	-0.383	0.287	-0.425
110	0.286	-0.532	0.244	-0.587
120	0.245	-0.687		
125			0.202	-0.776
130	0.214	-0.822		
135			0.172	-0.937
140	0.179	-1.001		
150			0.151	-1.067

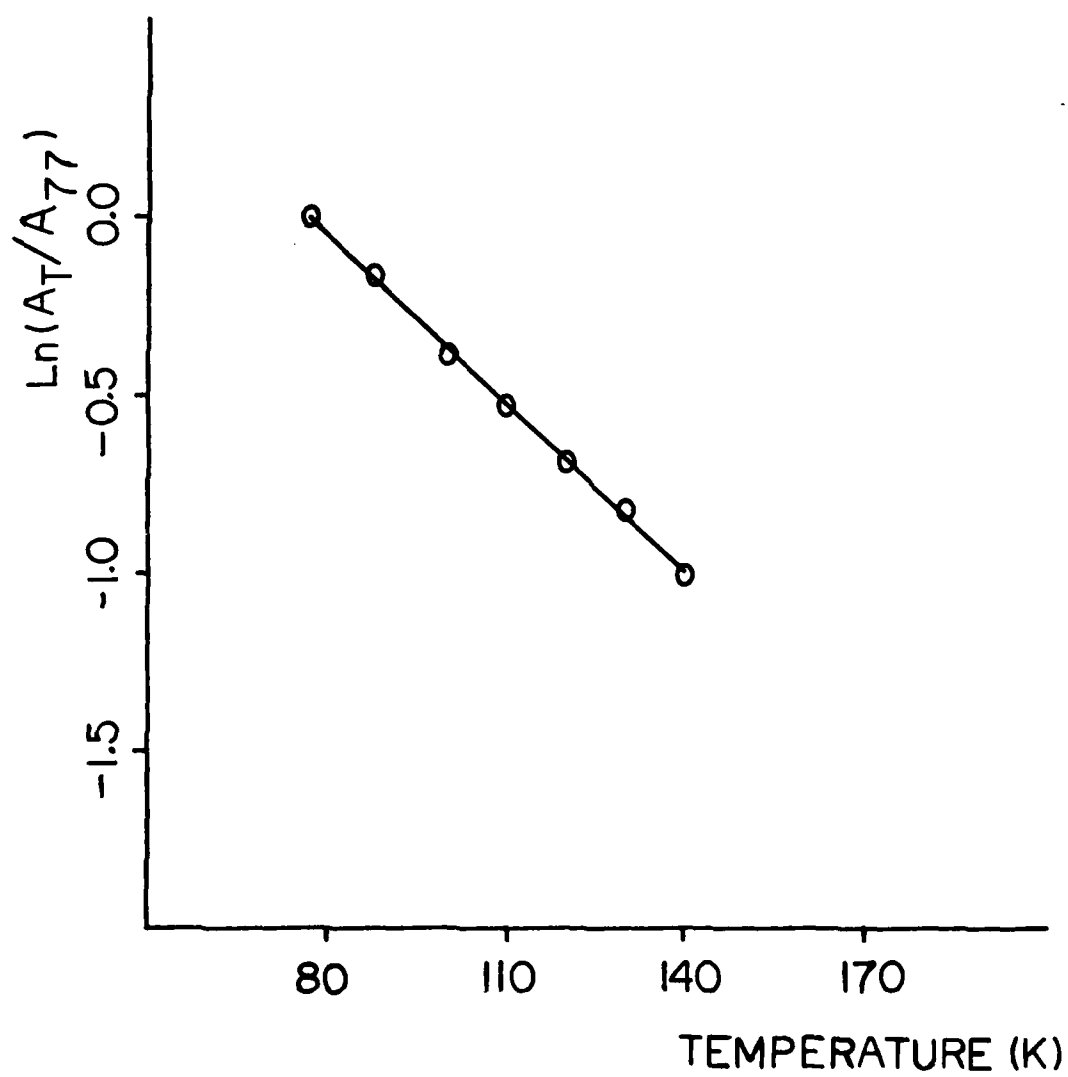
^aA(T) is the sum of the areas under the two wings of the doublet.

Figure Captions

Figure 1. Plot of $\ln A$ (normalized to the area under the resonance curve at 77 K) vs. temperature in K for chlorodimethyltin(IV) L-cysteine hydrate. The slope is $-1.58 \times 10^{-2} \text{ K}^{-1}$.

Figure 2. Plot of $\ln A$ (normalized to the area under the resonance curve at 77 K) vs. temperature in K for chlorodimethyltin(IV) DL-penicillamine. The slope is $-1.46 \times 10^{-2} \text{ K}^{-1}$.





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