

AD-A105 377

PENNSYLVANIA STATE UNIV UNIVERSITY PARK DEPT OF CHEMISTRY F/G 7/4
THE IMAGE POTENTIAL AND ION TRAJECTORIES IN SIMS.(U)
SEP 81 R A GIBBS, S P HOLLAND, K E FOLEY

N00014-80-C-0491

NL

UNCLASSIFIED

For 1
AD
A105377

END
DATE
FILMED
O B
DTIC

AD A105377

OFFICE OF NAVAL RESEARCH

Contract N00014-80-C0491 ✓

Task No. NR-051-744

TECHNICAL REPORT NO. 2 ✓

The Image Potential and Ion Trajectories in SIMS

by

R.A. Gibbs, S.P. Holland, K.E. Foley,

B.J. Garrison and N. Winograd

Prepared for Publication

in

Physical Review B

The Pennsylvania State University
Department of Chemistry
University Park, PA 16802-6389

September 1981

Reproduction in whole or in part is permitted for
any purpose of the United State Government.

This document has been approved for public release
and sale; its distribution is unlimited.

FILE COPY

81 10 2 119

REPORT DOCUMENTATION PAGE		READ INSTRUCTIONS BEFORE COMPLETING FORM
1. REPORT NUMBER	2. GOVT ACCESSION NO.	3. RECIPIENT'S CATALOG NUMBER
	AD-A105377	
4. TITLE (and Subtitle)	5. TYPE OF REPORT & PERIOD COVERED	
(6) The Image Potential and Ion Trajectories in SIMS.	(9) Technical Report.	
7. AUTHOR(s)	8. CONTRACT OR GRANT NUMBER(s)	
(10) R.A./Gibbs, S.P./Holland, K.E./Foley! B.J./Garrison and N. Winograd	(15) N00014-80-C0491	
9. PERFORMING ORGANIZATION NAME AND ADDRESS	10. PROGRAM ELEMENT, PROJECT, TASK AREA & WORK UNIT NUMBERS	
Chemistry Department The Pennsylvania State University University Park, PA 16802-6389	NR-051-744	
11. CONTROLLING OFFICE NAME AND ADDRESS	12. REPORT DATE	
	(11) September 1981	
14. MONITORING AGENCY NAME & ADDRESS (if different from Controlling Office)	13. NUMBER OF PAGES	
(12) 20	9	
	15. SECURITY CLASS. (of this report)	
	Unclassified	
	15a. DECLASSIFICATION/DOWNGRADING SCHEDULE	
16. DISTRIBUTION STATEMENT (of this Report)		
This document has been approved for public release and sale; its distribution is unlimited.		
17. DISTRIBUTION STATEMENT (of abstract entered in Block 20, if different from Report)		
18. SUPPLEMENTARY NOTES		
To be published in Physical Review B - Rapid Communications		
19. KEY WORDS (Continue on reverse side if necessary and identify by block number)		
ion bombardment, surface science, SIMS		
20. ABSTRACT (Continue on reverse side if necessary and identify by block number)		
The angle and energy distributions of Ni^+ ions ejected from ion bombarded $Ni(001)_{c(2 \times 2)}-CO$ are shown to be in excellent agreement with classical trajectory calculations for Ni atoms if the calculations are corrected for the presence of an image force. Two important consequences of this observation are that the ionization probability, R^+, is nearly isotropic and that it is only (over)		

DD FORM 1 JAN 73 1473

400343 JAN
SECURITY CLASSIFICATION OF THIS PAGE (When Data Entered)

20. weakly dependent on particle velocity. These constraints impose severe restrictions on proposed ionization theories.

The Image Potential and Ion Trajectories in SIMS

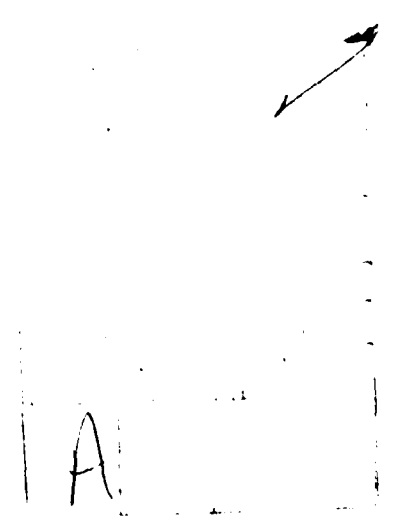
R. A. Gibbs,^(a) S. P. Holland,^(b) K. E. Foley,
B. J. Garrison^(c) and N. Winograd

Department of Chemistry
The Pennsylvania State University
University Park, PA 16802

(a) Present Address: Westhollow Research Center
Shell Development Company
Houston, Texas 77001

(b) Present Address: IBM
General Technology Division
Essex Junction, Vermont 05452

(c) Alfred P. Sloan Research Fellow



Abstract

The angle and energy distributions of Ni^+ ions ejected from ion bombarded $\text{Ni}(001)\sqrt{2}\times\sqrt{2}\text{-CO}$ are shown to be in excellent agreement with classical trajectory calculations for Ni atoms if the calculations are corrected for the presence of an image force. Two important consequences of this observation are that the ionization probability, R^+ , is nearly isotropic and that it is only weakly dependent on particle velocity. These constraints impose severe restrictions on proposed ionization theories.

Ion desorption from solids is a general phenomenon which can be induced by photon, electron or particle bombardment. The mechanism of ion formation has been well elucidated when the desorption is stimulated by photon (1) or electron fields (2), although the processes which affect ionization during collisions between atoms in a solid remain quite speculative. This theoretical input is critically needed to interpret experimental results from a variety of ion scattering experiments including secondary ion mass spectrometry (SIMS). In this paper, we examine the ionization problem by performing the first detailed experimental measurements of the angular and energy distributions of ions ejected due to ion bombardment of a well-defined surface. As a model system we have chosen $\text{Ni}(001)_{\text{c}}(2 \times 2)\text{-CO}$ bombarded by 1000 eV Ar^+ at normal incidence since the original surface geometry of CO has been determined by LEED (3) and since the presence of CO enhances the observed Ni^+ yield by more than 4 orders of magnitude over the clean $\text{Ni}(001)$ surface (4). The results are in semi-quantitative agreement with classical dynamical calculations of this system for the neutral particles if their trajectories are modified by the inclusion of an appropriate image force. This agreement provides a convincing basis for the classical dynamics model and suggests that the probability of ionization of the neutral atoms is isotropic and nearly independent of the ejected particle velocity. The latter conclusion supports recent theoretical efforts aimed toward providing physical insight into the ionization process (5) and indicates that Auger neutralization is not a dominant mechanism.

The angle-resolved SIMS measurements are performed using a specially designed UHV chamber such that the quadrupole mass filter (Riber AQX156)

can be rotated with respect to the primary ion beam. This feature is achieved using a set of three 56cm differentially pumped teflon seals. A manipulator for the crystal then provides azimuthal rotation, heating to 1300°K, cooling to 175°K, and vertical translation to a LEED apparatus located on a different level in the chamber. The secondary ions are then angle and energy selected using a 90° spherical electric sector in front of the mass analyzer. The calculated polar angle resolution, based on the size of the apertures and the field-free distance of the sample from the lens, is estimated to be $\pm 7^\circ$. The azimuthal angle resolution $\Delta\phi$ can be approximated from the polar angle, θ , since $\Delta\phi \approx \Delta\theta / \sin\theta$. Using a zoom lens arrangement, the bandpass of the analyzer can be varied from approximately 2 eV for recording energy spectra to nearly 15 eV for obtaining maximum sensitivity. Details of the apparatus will be available elsewhere (6). The total ion flux during the experiment is kept below 10^{13} ions/cm² to avoid significantly altering the surface structure. The cut and polished Ni crystal was cleaned in vacuo in the standard fashion. The clean surface was then exposed to 2 L of CO to obtain a c(2x2) overlayer geometry.

The calculated energy and angular distributions were determined using a previously described classical dynamics procedure (7). The initial nuclear coordinates of the adsorbed CO were taken from LEED measurements (3). These results indicate that the CO is in a linear bonded site with the carbon atom 1.76 Å above the nickel atom. The exact form and parameters of the interaction potential have been published elsewhere (7,8). Comparison between the calculated distributions and the measured ones required the computation of approximately 1400 Ar⁺ ion impacts for both clean Ni(001) and Ni(001)c(2x2)-CO to obtain sufficient statistical accuracy.

It has generally been observed that energy distributions of ejected secondary ions are considerably broader than that of the neutrals. For example, for 1 keV Ar^+ on Cu the ion energy distribution peaks at 4 eV and tails off as $E^{-0.5}$ while the neutral distribution peaks at 3 eV and tails off as roughly E^{-2} (9). Many explanations have been proposed to explain this effect, although most focus on the possibility that the primary phenomenon controlling the ionization probability is the Auger reneutralization rate. In this case, the ionization probability R^+ is a function of ion velocity as (10):

$$R^+ \propto \exp(-A/av_{\perp}) \quad (1)$$

where A is the Auger transition rate at the surface, a is a critical distance and $v_{\perp}$ is the perpendicular velocity component of the ejecting ion. Values of A/a from 2.5×10^5 to 2×10^6 cm/sec have been reported (11). On the other hand, a recent quantum mechanical model of ionization predicts an R^+ which is only a weak function of $v_{\perp}$ (5).

Our calculated energy distributions for Ni atoms ejected from $\text{Ni}(001)\sqrt{2} \times \sqrt{2}$ -CO are quite different from the rather broad energy distributions found experimentally for Ni^+ ions as shown in Figure 1. We can obtain reasonable agreement between calculated and experimental distributions at a polar angle of 30° by correcting the calculated trajectories using Eq. 1 and an A/a value of 1.1×10^6 cm/sec, although computed energy distributions at other angles then deviate rather dramatically from experimental curves. As we shall see later, this equation poorly predicts polar and azimuthal angle Ni^+ ion intensities.

In contrast, incorporation of a simple image force allows quantitative agreement between the calculations of neutral atom trajectories and the

experimentally measured ion distributions. To overcome the image force and escape the solid, the ion must perform work, E_{image} . If it is assumed that the particle instantaneously becomes an ion, E_{image} is given by

$$E_{\text{image}} = \frac{e^2}{4z} = \frac{3.6\text{eV}}{a_0} \quad (2)$$

where a_0 is the height in Å of the particle above the jellium step-edge at the instant of ionization. If θ_n is the angle of the atom's velocity vector at this point, the ion emerges with a final direction given by

$$\theta_i = \tan^{-1} \left\{ \frac{E_0 \sin^2 \theta_n}{E_0 \cos^2 \theta_n - E_{\text{image}}} \right\}^{1/2} \quad (3)$$

where θ_i is the corrected polar ejection angle as measured from the surface normal of the leaving ion, and E_0 is the kinetic energy of the neutral particle ejected at polar angle θ_n . The corrected theoretical energy distribution using $E_{\text{image}} = 3.6 \text{ eV}$ is also shown in Figure 1. The agreement with experiment over all polar angles is quite good. Note that in this comparison we choose a specific value of E_{image} without explicitly assuming the functional form given in Eq. 2.

A further test of the relevance of this correction can also be developed by comparing neutral and ionic trajectories of Ni at various polar and azimuthal angles. As shown in Figure 2, a selection of both low and high kinetic energy particles produces a maximum Ni^+ signal at a polar angle of about 50° . The calculations produce a distribution similar in shape, which peaks at significantly lower angles. Correction of the neutral trajectories using Eq. 3 with $E_{\text{image}} = 3.6 \text{ eV}$ however, provides quantitative agreement in both energy regimes. An E_{image} value other than $3.6 \pm 0.3 \text{ eV}$ yields a poor fit to the data in Figure 2 and Figure 1. It is

also important to note that correction of the calculated trajectories using Eq. 1 shifts the polar distributions significantly closer to the normal than the experimental data.

Information contained in azimuthal angular spectra is more sensitive to surface structure than either the polar or energy distributions (12,13). The azimuthal spectra obtained at large polar angles should be strongly influenced by any image force, since the image force acts to bend particles originally ejected at smaller polar angles into the detector. In Figure 3, the angular spectra obtained for 3 ± 3 eV Ni^+ ions ejected from $\text{Ni}(001)\sqrt{2} \times \sqrt{2}$ -CO are shown at $\theta = 30^\circ, 45^\circ, 60^\circ$ and 70° . Predicted neutral and image force corrected distributions are also shown, again derived assuming the same E_{image} value of 3.6 eV. Although the magnitude of the measured anisotropy is slightly smaller than calculated at 45° and 30° , the level of agreement is quite remarkable. Note that the calculated distributions are unaffected by the incorporation of Eq. 1. It is of interest that if the CO is placed in other bonding geometries, that poor agreement with experiment is found. These types of angle-resolved experiments, then, should be valuable aids in the analysis of unknown surface structures.

It is clear that the classical dynamics calculations can provide an accurate description of ion trajectories with simple inclusion of an appropriate image force. The agreement between theory and experiment for Ni trajectories is excellent for energy distributions at various polar and azimuthal angles. Similar levels of agreement are found for Ni_2^+ and NiCO^+ trajectories, although statistical fluctuations in the theoretical development so far preclude detailed comparisons. Unfortunately, the CO^+ ion is not experimentally observed.

If the above arguments concerning the presence of a relatively strong image force are correct, there are a number of constraints placed upon any ionization theory. First, the ion must be formed very close to the surface. The $1.0 \text{ \AA}$ distance we find for $\text{Ni}(001)_{\sqrt{2}\times\sqrt{2}}\text{-CO}$ should be viewed only qualitatively since other factors such as saturation of the image force (14) and the change in partial charge on the ejecting atom with distance (5) are not taken into account with our simple approach. A second constraint is that for $\text{Ni}(001)_{\sqrt{2}\times\sqrt{2}}\text{-CO}$, R^+ is only weakly dependent on v_i and is not a measureable function of ejection angle. These results have only recently been predicted using a quantum mechanical model (5). This model utilizes the electronic density of states to calculate electronic hopping probabilities during the atomic collisions which lead to particle ejection. Although the influence of ejection angle on R^+ has not been fully tested, it is possible that the large number of different ejection mechanisms which are observed cause an averaging of angular anisotropies. At least for $\text{Ni}(001)_{\sqrt{2}\times\sqrt{2}}\text{-CO}$, the Auger neutralization mechanism does not appear to be important.

The authors wish to thank the National Science Foundation, The Office of Naval Research and the Petroleum Research Foundation administered by The American Chemical Society for financial support. One of us (BJG) also acknowledges the A. P. Sloan Foundation for a Research Fellowship and the Camille and Henry Dreyfus Foundation for a grant for newly appointed young faculty.

References

1. M. L. Knotek, V. O. Jones and V. Rehn, Phys. Rev. Lett., 43, 300 (1979).
2. (a) M. L. Knotek and P. J. Feibelman, Phys. Rev. Lett., 40, 964 (1978);
(b) D. Menzel and R. Gomer, J. Chem. Phys., 41, 311 (1964). (c) P. A. Redhead, Can. J. Phys. 42, 886 (1964).
3. (a) M. Passler, A. Ignatiev, F. Jona, D. W. Jepsen and D. M. Marcus, Phys. Rev. Lett. 43, 360 (1979); (b) S. Andersson and J. B. Pendry, Phys. Rev. Lett., 43, 363 (1979).
4. P. H. Dawson and W. Tam, Surface Sci., 91, 153 (1980).
5. Z. Šroubek, K. Ždánský and J. Žavadil, Phys. Rev. Lett., 45, 580 (1980).
6. R. A. Gibbs and N. Winograd, Rev. Sci. Instrum., submitted.
7. N. Winograd, B. J. Garrison and D. E. Harrison, Jr., J. Chem. Phys., 73, 3473 (1980).
8. D. E. Harrison, Jr., P. W. Kelly, B. J. Garrison and N. Winograd, Surface Sci., 76, 311 (1978).
9. R. G. Hart and C. B. Cooper, Surface Sci., 94, 105 (1980).
10. H. D. Hagstrum, J. Vac. Sci. Technol., 12, 7 (1975).
11. K. Wittmaack, in: Inelastic Ion-Surface Collisions, Ed. Tolk, Tully, Heiland and White (Academic Press, N.Y., 1977) p. 153.
12. (a) S. P. Holland, B. J. Garrison and N. Winograd, Phys. Rev. Lett., 43, 220 (1979); (b) S. P. Holland, B. J. Garrison and N. Winograd, Phys. Rev. Lett., 44, 3473 (1980).
13. N. Winograd and B. J. Garrison, Accounts of Chem. Res., 13, 406 (1980).
14. R. E. Dietz, E. G. McRae and R. L. Campbell, Phys. Rev. Lett., 45, 1280 (1980).

Figure Captions

Figure 1 - Energy distributions for Ni ejected from $\text{Ni}(001)\sqrt{2}\times\sqrt{2}\text{-CO}$ due to bombardment by 1 KeV Ar^+ at normal incidence. The polar angle, θ , is defined with respect to the surface normal. The azimuthal angle, ϕ , is along the $\langle 100 \rangle$ direction for all cases. The three curves in each panel are identified as follows: — calculated Ni distribution, ---- calculated Ni distribution with inclusion of the image force and experimental points for Ni^+ .

Figure 2 - Polar angle distributions for Ni ejected from $\text{Ni}(001)\sqrt{2}\times\sqrt{2}\text{-CO}$. Experimental conditions and codings for each curve are the same as for Figure 1. The upper set of curves is recorded for a secondary ion energy of 7 ± 2 eV while the lower set of curves is taken at 22 ± 2 eV.

Figure 3 - Azimuthal angle distributions at various polar angles for Ni ejected from $\text{Ni}(001)\sqrt{2}\times\sqrt{2}\text{-CO}$. The curve codings are the same as for Figure 1. Only those Ni particles with an energy of 3 ± 3 eV are detected. The value of $\phi=0^\circ$ corresponds to $\langle 100 \rangle$ while $\phi=\pm 45^\circ$ corresponds to $\langle 110 \rangle$.

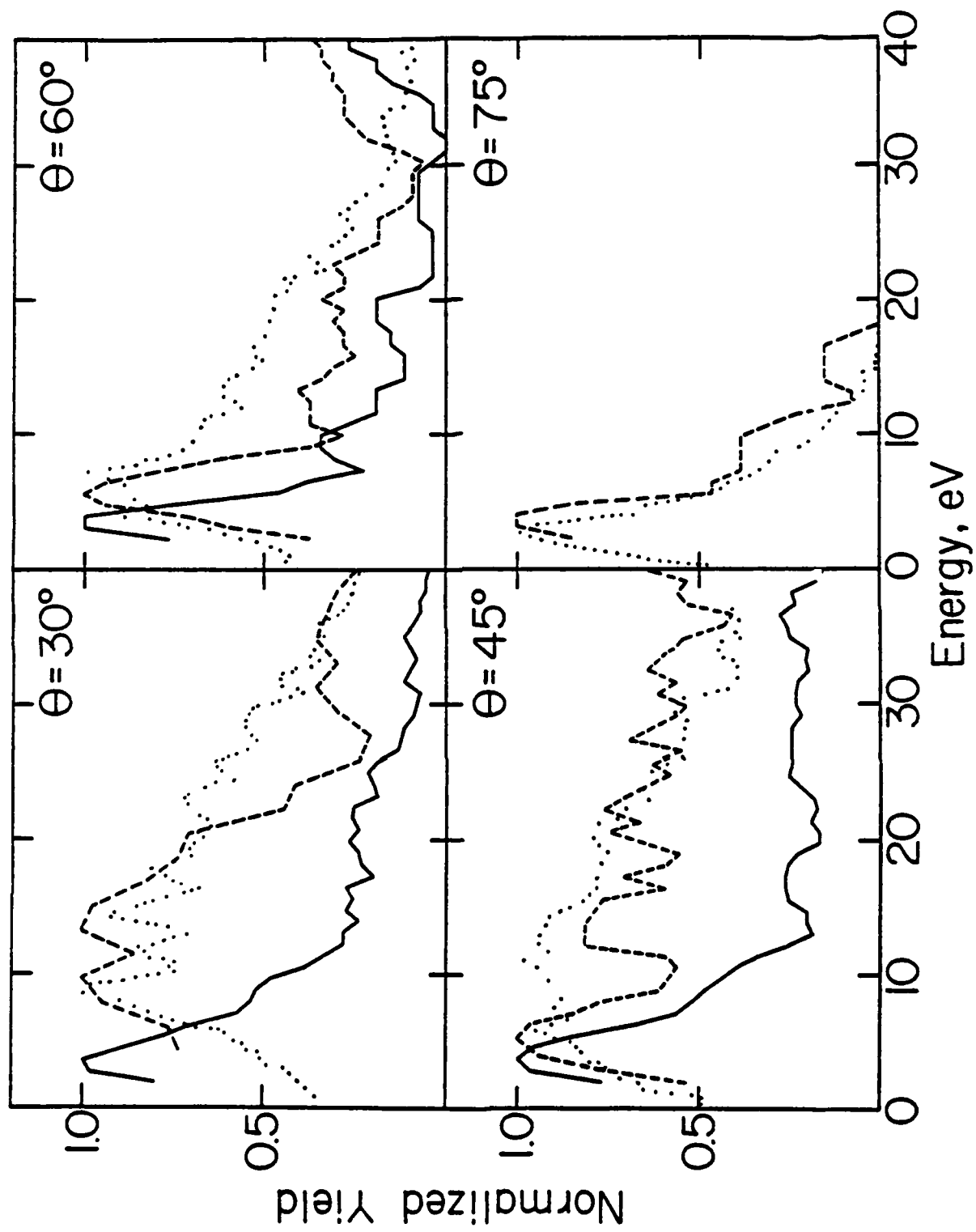


FIGURE 1

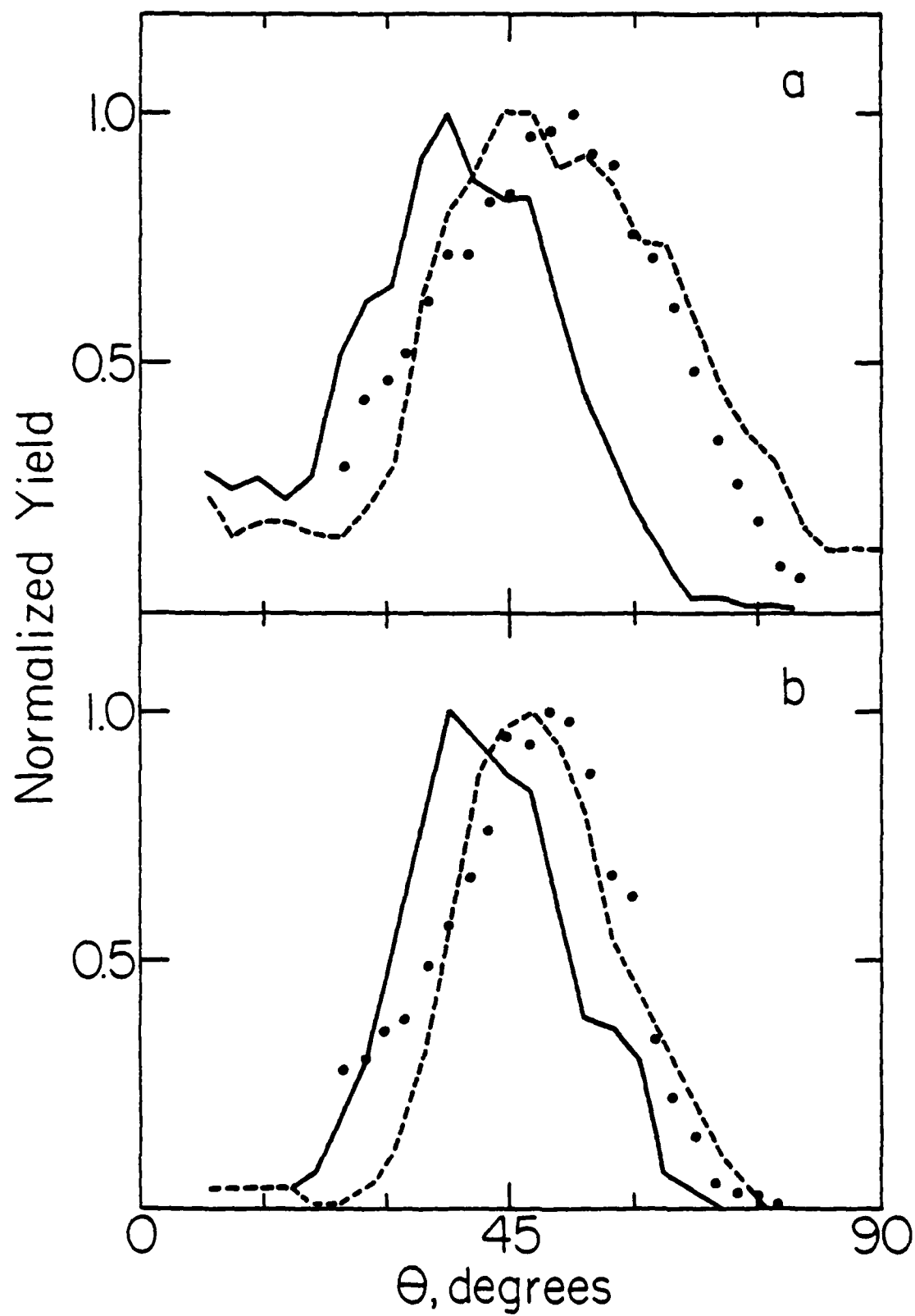


FIGURE 2

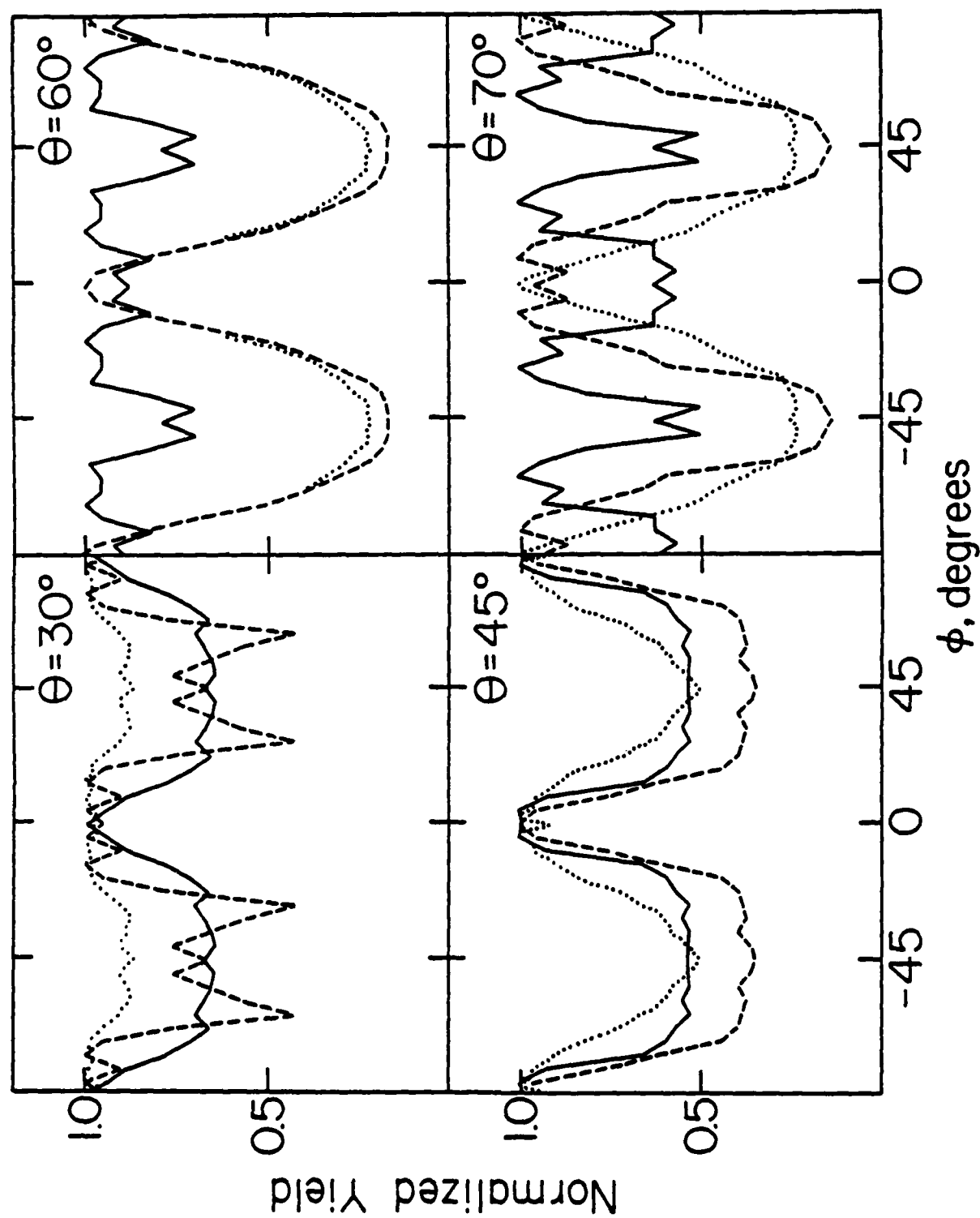


FIGURE 3

TECHNICAL REPORT DISTRIBUTION LIST, GEN

	<u>No. Copies</u>		<u>No. Copies</u>
Office of Naval Research Attn: Code 472 800 North Quincy Street Arlington, Virginia 22217	2	U.S. Army Research Office Attn: CRD-AA-IP P.O. Box 1211 Research Triangle Park, N.C. 27709	1
ONR Western Regional Office Attn: Dr. R. J. Marcus 1030 East Green Street Pasadena, California 91106	1	Naval Ocean Systems Center Attn: Mr. Joe McCartney San Diego, California 92152	1
ONR Eastern Regional Office Attn: Dr. L. H. Peebles Building 114, Section D 666 Summer Street Boston, Massachusetts 02210	1	Naval Weapons Center Attn: Dr. A. B. Amster, Chemistry Division China Lake, California 93555	1
Director, Naval Research Laboratory Attn: Code 6100 Washington, D.C. 20390	1	Naval Civil Engineering Laboratory Attn: Dr. R. W. Drisko Port Hueneme, California 93401	1
The Assistant Secretary of the Navy (RE&S) Department of the Navy Room 4E736, Pentagon Washington, D.C. 20350	1	Department of Physics & Chemistry Naval Postgraduate School Monterey, California 93940	1
Commander, Naval Air Systems Command Attn: Code 3100 (H. Rosenwasser) Department of the Navy Washington, D.C. 20360	1	Scientific Advisor Commandant of the Marine Corps (Code RD-1) Washington, D.C. 20380	1
Defense Technical Information Center Building 5, Cameron Station Alexandria, Virginia 22314	12	Naval Ship Research and Development Center Attn: Dr. G. Bosmajian, Applied Chemistry Division Annapolis, Maryland 21401	1
Dr. Fred Saalfeld Chemistry Division, Code 6100 Naval Research Laboratory Washington, D.C. 20375	1	Naval Ocean Systems Center Attn: Dr. S. Yamamoto, Marine Sciences Division San Diego, California 91232	1
		Mr. John Boyle Materials Branch Naval Ship Engineering Center Philadelphia, Pennsylvania 19112	1

TECHNICAL REPORT DISTRIBUTION LIST, 056

	<u>No.</u> <u>Copies</u>		<u>No.</u> <u>Copies</u>
Dr. G. A. Somorjai Department of Chemistry University of California Berkeley, California 94720	1	Dr. C. P. Flynn Department of Physics University of Illinois Urbana, Illinois 61801	1
Dr. L. N. Jarvis Surface Chemistry Division 4555 Overlook Avenue, S.W. Washington, D.C. 20375	1	Dr. W. Kohn Department of Physics University of California (San Diego) LaJolla, California 92037	1
Dr. J. B. Hudson Materials Division Rensselaer Polytechnic Institute Troy, New York 12181	1	Dr. R. L. Park Director, Center of Materials Research University of Maryland College Park, Maryland 20742	1
Dr. John T. Yates Surface Chemistry Section National Bureau of Standards Department of Commerce Washington, D.C. 20234	1	Dr. W. T. Peria Electrical Engineering Department University of Minnesota Minneapolis, Minnesota 55455	1
Dr. Theodore E. Madey Surface Chemistry Section Department of Commerce National Bureau of Standards Washington, D.C. 20234	1	Dr. Chia-wei Woo Department of Physics Northwestern University Evanston, Illinois 60201	1
Dr. J. M. White Department of Chemistry University of Texas Austin, Texas 78712	1	Dr. D. C. Mattis Polytechnic Institute of New York 333 Jay Street Brooklyn, New York 11201	1
Dr. Keith H. Johnson Department of Metallurgy and Materials Science Massachusetts Institute of Technology Cambridge, Massachusetts 02139	1	Dr. Robert M. Hexter Department of Chemistry University of Minnesota Minneapolis, Minnesota 55455	1
Dr. J. E. Demuth IBM Corporation Thomas J. Watson Research Center P.O. Box 218 Yorktown Heights, New York 10598	1	Dr. R. P. Van Duyne Chemistry Department Northwestern University Evanston, Illinois 60201	1

TECHNICAL REPORT DISTRIBUTION LIST, 056

	<u>No. Copies</u>		<u>No. Copies</u>
Dr. S. Sibener Department of Chemistry James Franck Institute 5640 Ellis Avenue Chicago, Illinois 60637	1	Dr. Martin Fleischmann Department of Chemistry Southampton University Southampton 509 5NH Hampshire, England	1
Dr. M. G. Lagally Department of Metallurgical and Mining Engineering University of Wisconsin Madison, Wisconsin 53706	1	Dr. J. Osteryoung Chemistry Department State University of New York at Buffalo Buffalo, New York 14214	1
Dr. Robert Gomer Department of Chemistry James Franck Institute 5640 Ellis Avenue Chicago, Illinois 60637	1	Dr. G. Rubloff I.B.M. Thomas J. Watson Research Center P. O. Box 218 Yorktown Heights, New York 10598	1
Dr. R. G. Wallis Department of Physics University of California, Irvine Irvine, California 92664	1	Dr. J. A. Gardner Department of Physics Oregon State University Corvallis, Oregon 97331	1
Dr. D. Ramaker Chemistry Department George Washington University Washington, D.C. 20052	1	Dr. G. D. Stein Mechanical Engineering Department Northwestern University Evanston, Illinois 60201	1
Dr. P. Hansma Chemistry Department University of California, Santa Barbara Santa Barbara, California 93106	1	Dr. K. G. Spears Chemistry Department Northwestern University Evanston, Illinois 60201	1
Dr. P. Hendra Chemistry Department Southampton University England SO9JNH	1	Dr. R. W. Plummer University of Pennsylvania Department of Physics Philadelphia, Pennsylvania 19104	1
Professor P. Skell Chemistry Department Pennsylvania State University University Park, Pennsylvania 16802	1	Dr. E. Yeager Department of Chemistry Case Western Reserve University Cleveland, Ohio 41106	2
Dr. J. C. Hemminger Chemistry Department University of California, Irvine Irvine, California 92717	1	Professor D. Hercules University of Pittsburgh Chemistry Department Pittsburgh, Pennsylvania 15260	1

TECHNICAL REPORT DISTRIBUTION LIST, 056

No.
Copies

Professor N. Winograd
The Pennsylvania State University
Department of Chemistry
University Park, Pennsylvania 16802 1

Professor T. F. George
The University of Rochester
Chemistry Department
Rochester, New York 14627 1

Professor Dudley R. Herschbach
Harvard College
Office for Research Contracts
1350 Massachusetts Avenue
Cambridge, Massachusetts 02138 1

Professor Horia Metiu
University of California,
Santa Barbara
Chemistry Department
Santa Barbara, California 93106 1

Professor A. Steckl
Rensselaer Polytechnic Institute
Department of Electrical and
Systems Engineering
Integrated Circuits Laboratories
Troy, New York 12181 1

Professor R. D. Archer
University of Massachusetts
Chemistry Department
Amherst, Massachusetts 01003 1

Dr. A. C. Pastor
Hughes Research Laboratories
3011 Malibu Canyon Road
Malibu, California 90265 1