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SIMULATION OF EDGE EFFECTS IN ELECTROANALYTICAL EXPERIMENTS BY --ETC(11)

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Simulation of Edge Effects in Electroanalytical Experiments.
Part I. Two Dimensional Collocation and Theory for Chrono-
amperometry.

By

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Prepared for Publication
in

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The details for accurate simulations of edge effects experienced at disc electrodes in chronoamperometric experiments is given.		

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Simulation of edge-effects in electroanalytical experiments by orthogonal collocation. Part 1. Two-dimensional collocation and theory for chronoamperometry

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A two-dimensional model to describe edge-effects at planar circular disc electrodes has been developed. Two-dimensional orthogonal collocation is used to discretize the corresponding partial differential equations. The equations for the simulation of a chronoamperometric experiment are derived.

BERND SPEISER et STANLEY PONS, *Can. J. Chim.* 60, 1352 (1982)

On a mis au point une méthode bidimensionnelle pour décrire les effets d'arête au niveau des électrodes à disques circulaires plans. On a utilisé une discrétisation orthogonale pour les équations différentielles partielles correspondantes. On a dérivé les équations de la simulation d'une expérience de chronoamperométrie.

(Traduit par le journal)

Introduction

One common simplification in the theoretical treatment of electrode processes of planar disc electrodes is the assumption of semi-infinite linear diffusion. This results in the reduction of the three-dimensional system to a one-dimensional model, where the distance x from the electrode surface is the only coordinate to be considered. Consequently, the three-dimensional form of Fick's second law

$$[1] \quad \frac{\partial c}{\partial t} = D \left[\frac{\partial^2 c}{\partial x^2} + \frac{\partial^2 c}{\partial r^2} + \frac{\partial^2 c}{\partial z^2} \right]$$

reduces to the simpler differential equation

$$[2] \quad \frac{\partial c}{\partial t} = D \frac{\partial^2 c}{\partial x^2}$$

and the mathematical and computational efforts decrease considerably. The assumption of linear diffusion to a planar electrode might imply an infinite disc without edges. For real systems, however, this of course is not possible.

Thus, measurements at finite disc electrodes should deviate from the simple one-dimensional theory. In fact, it has been shown for different electroanalytical techniques such as chronoamperometry and chronopotentiometry (1) that the diffusion current is higher than expected from the theoretical calculations based on eq. [2], i.e. the Cottrell equation (2) for chronoamperometry. For cyclic voltammograms (3) the peak separation is larger than the theoretical values given by Nicholson and Shain (4). A number of papers have accounted for these deviations as being due to additional diffusion from directions other than perpendicular to the disc electrode (see e.g. refs. 5,

6). In a system with several coupled charge transfer and chemical reactions dramatic changes of the overall voltammogram have been attributed to edge effects (7).

Several attempts have been made to calculate the effects of edge diffusion and to simulate model systems without assuming one-dimensional diffusion (5, 6). Recently, Heinze published a two-dimensional model based on the finite difference method (8, 9), while Oldham (10) and Aoki and Osteryoung (11) developed analytical solutions for the simple reversible mechanism.

These calculations showed, in agreement with experiments, that edge diffusion rates increase with increasing time and decreasing radius of the disc electrode. The deviation from pure Cottrell behavior is claimed to be between 9 and 10% for typical conditions: diffusion coefficient = 10^{-5} cm²/s, electrode radius 0.1 cm, electrolysis time 2.65 s (5, 9–11). All methods presented thus far, however, demonstrate results for only the simplest mechanisms (e.g. reversible electron transfer case). Extensions apparently require extensive computational effort.

Orthogonal collocation has been shown to provide a powerful method to simulate responses of several electroanalytical techniques to a variety of mechanisms (12–18). Up to now all these calculations have been performed using one-dimensional diffusion models. We want to report the construction of a two-dimensional model for the diffusion to and from a planar disc electrode and the discretization of the resulting second order partial differential equations to give a set of simultaneous first order ordinary differential equations.

The comparison of the results of this new model with those already presented is the subject of a

separate report now in preparation. Extensions to other mechanisms and electroanalytical experiments will also be considered.

The two-dimensional model

Consider a planar, circular disc of conducting material embedded into an insulator (Fig. 1). The radius of the disc is r_0 . The radius r_0 of the insulator may go to infinity. From the symmetry of the problem, it follows that we may use two coordinates x (perpendicular to the disc) and r (along the radius vector of the disc). We are required to calculate the concentration c of all species as a function of x , r and the time t , where c is given by second order partial differential equations of the type (9)

$$[3] \quad \frac{\partial c}{\partial t} = D \left[\frac{\partial^2 c}{\partial x^2} + \frac{1}{r} \frac{\partial c}{\partial r} + \frac{\partial^2 c}{\partial r^2} \right] - f(c)$$

and $f(c)$ is determined by the kinetics of the system investigated. D is the diffusion coefficient.

Let us consider an EC-mechanism (4). There the differential equations describing the concentration of species A, B, and C are given by [5]–[7].



$$[5] \quad \frac{\partial c_A}{\partial t} = D_A \left[\frac{\partial^2 c_A}{\partial x^2} + \frac{1}{r} \frac{\partial c_A}{\partial r} + \frac{\partial^2 c_A}{\partial r^2} \right]$$

$$[6] \quad \frac{\partial c_B}{\partial t} = D_B \left[\frac{\partial^2 c_B}{\partial x^2} + \frac{1}{r} \frac{\partial c_B}{\partial r} + \frac{\partial^2 c_B}{\partial r^2} \right] - k_3 c_B$$

$$[7] \quad \frac{\partial c_C}{\partial t} = D_C \left[\frac{\partial^2 c_C}{\partial x^2} + \frac{1}{r} \frac{\partial c_C}{\partial r} + \frac{\partial^2 c_C}{\partial r^2} \right] + k_3 c_B$$

For $k_1 = 0$ we deal with a simple reversible electron transfer. The electrode reaction in eq. [4] is drawn as an oxidation but the extension to reductions is immediate. For simplification we may assume that the diffusion coefficients are equal:

$$[8] \quad D_A = D_B = D_C$$

This, however, is not necessary for orthogonal collocation simulations.

Depending on the electroanalytical method we wish to simulate for the mechanism chosen, certain initial and boundary conditions are imposed upon the solution of eq. [3]. Let us consider a chronoamperometric experiment, where species A, the substrate, is present at a homogeneous concentration c_A^0 at the beginning of the experiment. The product

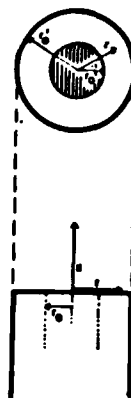


Fig. 1. Geometry of the two-dimensional model.

species B and C may not be present at $t = 0$ (initial conditions [9], [10]):

$$[9] \quad c_B(x, r, 0) = c_C(x, r, 0) = 0$$

$$[10] \quad c_A(x, r, 0) = c_A^0(x, r, 0) = 0$$

At the boundary $x \rightarrow \infty$, i.e. at a distance x far from the electrode where the concentrations are not changed during the experiment:

$$[11] \quad c_A(\infty, r, t) = c_A^0$$

$$[12] \quad c_B(\infty, r, t) = c_C(\infty, r, t) = 0$$

The same condition may hold for a distance r far from the electrode radially:

$$[13] \quad c_A(x, \infty, t) = c_A^0$$

$$[14] \quad c_B(x, \infty, t) = c_C(x, \infty, t) = 0$$

At the electrode surface the Nernst equation has to be fulfilled for c_A and c_B (eq. [14a]). If we step to a potential E at least 200 mV past the formal potential $E_{f,A}$ the concentration of the substrate A at the electrode is immediately forced to zero and [14a] may be replaced by [14b]. Additional conditions at

points in the x and r directions, respectively

$$[55] \quad c(x_i, r_k, t) = \sum_{j=1}^{N_x+1} \sum_{l=1}^{N_r+1} \lambda_{j,l} x_j^{i-1} r_l^{k-1}$$

The concentrations at the collocation points may be the elements of a concentration matrix $\Gamma_{i,k}$

$$[56] \quad \Gamma_{i,k} = X_{i,j} \cdot A_{j,l} \cdot R_{l,k}$$

which is calculated by multiplication of a matrix $X_{i,j}$, containing the x_j^{i-1} as elements, a matrix $A_{j,l}$, containing the elements $\lambda_{j,l}$, and a matrix $R_{l,k}$, containing the elements r_l^{k-1} . The superscript T denotes a transposed matrix.

Now we can calculate the first and second derivatives of the concentration at the collocation points with respect to x and r respectively:

$$[57] \quad \Gamma_{i,k}^x = X_{i,j}^x \cdot A_{j,l} \cdot R_{l,k}$$

$$[58] \quad \Gamma_{i,k}^r = X_{i,j} \cdot A_{j,l} \cdot R_{l,k}^r$$

$$[59] \quad \Gamma_{i,k}^{xx} = X_{i,j}^{xx} \cdot A_{j,l} \cdot R_{l,k}$$

$$[60] \quad \Gamma_{i,k}^{rr} = X_{i,j} \cdot A_{j,l} \cdot R_{l,k}^{rr}$$

where the superscript $x(i)$ denotes the first and accordingly $r(i)$ characterizes the second derivative with respect to $x(r)$.

From eq. [56] it follows that

$$[61] \quad A_{j,l} = X_{i,j}^{-1} \cdot \Gamma_{i,k} \cdot (R_{l,k}^T)^{-1}$$

After inserting [61] into [57] we arrive at

$$[62] \quad \Gamma_{i,k}^x = X_{i,j}^x \cdot X_{i,j}^{-1} \cdot \Gamma_{i,k} = A_{i,j} \cdot \Gamma_{i,k}$$

where the first derivative of the concentration with respect to x is expressed as a function of the concentration itself. In analogy eqs. [61] and [58] give

$$[63] \quad \Gamma_{i,k}^r = X_{i,j}^r \cdot X_{i,j}^{-1} \cdot \Gamma_{i,k} = B_{i,j} \cdot \Gamma_{i,k}$$

From [61] and [59] we get

$$[64] \quad \Gamma_{i,k}^{xx} = \Gamma_{i,k} \cdot (B_{i,j}^T)^{-1} \cdot B_{i,j}^{xx} = \Gamma_{i,k} \cdot (B_{i,j}^T)^{-1} \cdot B_{i,j}^{xx} \\ = \Gamma_{i,k} \cdot (B_{i,j} \cdot B_{i,j}^T)^{-1} \cdot B_{i,j}^{xx}$$

or

$$[65] \quad \Gamma_{i,k}^{xx} = \Gamma_{i,k} \cdot C_{i,j}$$

Analogous from [61] and [60]

$$[66] \quad \Gamma_{i,k}^{rr} = \Gamma_{i,k} \cdot (B_{i,j}^T \cdot B_{i,j})^{-1} \cdot B_{i,j}^{rr} = \Gamma_{i,k} \cdot D_{i,j}$$

We may calculate the matrices containing the derivatives of the concentration at the collocation points from the concentration matrix once the matrices $A_{j,l}$, $B_{i,j}$, $C_{i,j}$, and $D_{i,j}$ have been determined. Returning to summation expression we

can write

$$[67] \quad \left(\frac{\partial c}{\partial x} \right)_{i,k} = \sum_{j=1}^{N_x+1} A_{i,j} c(x_j, r_k, t)$$

$$[68] \quad \left(\frac{\partial^2 c}{\partial x^2} \right)_{i,k} = \sum_{j=1}^{N_x+1} C_{i,j} c(x_j, r_k, t)$$

$$[69] \quad \left(\frac{\partial c}{\partial r} \right)_{i,k} = \sum_{l=1}^{N_r+1} B_{i,l} c(x_i, r_l, t)$$

and

$$[70] \quad \left(\frac{\partial^2 c}{\partial r^2} \right)_{i,k} = \sum_{l=1}^{N_r+1} D_{i,l} c(x_i, r_l, t)$$

As has been pointed out by Whiting and Carr (12) for the one-dimensional case, the matrices $A_{i,j}$, $B_{i,j}$, $C_{i,j}$, and $D_{i,j}$ depend solely on the type of polynomial used as trial function and may be calculated from the zeroes of this polynomial. This has to be done only once. Using the values of the elements of these matrices we are able to express the differential quotients of c with respect to x and r at all collocation points.

For linear diffusion problems it has been shown that Legendre polynomials are most suitable (12). Pons (18) used Chebyshev polynomials for diffusion under spherical geometry. It may be noted, however, that the choice of the trial function does not have a large effect on the accuracy of the solutions for higher order approximations. In general, the best trial function is determined by the symmetry of the problem.

In this work, Legendre polynomials have been used for discretization along the x -coordinate. The collocation in the direction of the r -coordinate has been performed using different types of functions including Legendre, Chebyshev, and polynomials constructed from even powers of r , which is suggested by the symmetry of the problem around $r = 0$ (19).

Discretization of the equations for chronoamperometry

We can substitute the summation expressions [67]–[70] for the partial differential quotients in eqs. [24]–[28] and [30]–[32]. This yields expressions for the time derivative of the concentration of species A, B, and C at all the collocation points as a function of the concentrations at the collocation points. If we substitute the boundary conditions [36]–[41], [42b] (i.e. we assume a step to a potential far behind E^0) and [47]–[49], we arrive at

$$[71] \quad \left. \frac{dc_A}{dt} \right|_{i,k} = -\beta \left[B_{i,j} c_A(x_j, r_k, T) + \sum_{l=1}^{N_r+1} B_{i,l} c_A(x_i, r_l, T) \right] \\ + \left[E_{i,j} c_A(x_i, 0, T) + E_{i,j+1} c_A(x_{j+1}, 1, T) + \sum_{l=1}^{N_r+1} E_{i,l} c_A(x_i, r_l, T) \right]$$

$$[72] \quad \left. \frac{dc_B}{dt} \right|_{i,k} = -\beta \left[B_{i,j} c_B(x_j, r_k, T) + \sum_{l=1}^{N_r+1} B_{i,l} c_B(x_i, r_l, T) \right] \\ + \left[E_{i,j} c_B(x_i, 0, T) + E_{i,j+1} c_B(x_{j+1}, 1, T) + \sum_{l=1}^{N_r+1} E_{i,l} c_B(x_i, r_l, T) \right] - \alpha c_B$$

$$[73] \quad \left. \frac{dc_C}{dt} \right|_{i,k} = -\beta \left[B_{i,j} c_C(x_j, r_k, T) + \sum_{l=1}^{N_r+1} B_{i,l} c_C(x_i, r_l, T) \right] \\ + \left[E_{i,j} c_C(x_i, 0, T) + E_{i,j+1} c_C(x_{j+1}, 1, T) + \sum_{l=1}^{N_r+1} E_{i,l} c_C(x_i, r_l, T) \right] + \alpha c_B$$

$$[74] \quad \left. \frac{dc_A}{dt} \right|_{i,k} = -\beta \left[B_{i,j} c_A(x_j, r_k, T) + B_{i,j+1} c_A(x_{j+1}, r_{j+1}, T) + \sum_{l=1}^{N_r+1} B_{i,l} c_A(x_i, r_l, T) \right] \\ + \beta' \left[F_{i,j} c_A(x_i, 0, T) + F_{i,j+1} c_A(x_{j+1}, r_{j+1}, T) + \sum_{l=1}^{N_r+1} F_{i,l} c_A(x_i, r_l, T) \right]$$

$$[75] \quad \left. \frac{dc_B}{dt} \right|_{i,k} = -\beta \left[B_{i,j} c_B(x_j, r_k, T) + \sum_{l=1}^{N_r+1} B_{i,l} c_B(x_i, r_l, T) \right] \\ + \beta' \left[F_{i,j} c_B(x_i, 0, T) + \sum_{l=1}^{N_r+1} F_{i,l} c_B(x_i, r_l, T) \right] - \alpha c_B$$

$$[76] \quad \left. \frac{dc_C}{dt} \right|_{i,k} = -\beta \left[B_{i,j} c_C(x_j, r_k, T) + \sum_{l=1}^{N_r+1} B_{i,l} c_C(x_i, r_l, T) \right] \\ + \beta' \left[F_{i,j} c_C(x_i, 0, T) + \sum_{l=1}^{N_r+1} F_{i,l} c_C(x_i, r_l, T) \right] + \alpha c_B$$

where

$$[77] \quad E_{i,j} = \frac{C_{i,j} + R_{i,j} D_{i,j}}{R_{i,j}}$$

and

$$[78] \quad F_{i,j} = \frac{C_{i,j}}{\sqrt{\beta' + R_{i,j}}} + D_{i,j}$$

The primed terms, except T' and β' , correspond to the outer functions. Equations [71]–[76] contain $9N_x + 2N_r + 3N_r$ unknowns, i.e. $c_A(x_i, 0, T)$, $c_A(x_{j+1}, 1, T)$, $c_B(x_i, 0, T)$, $c_B(x_{j+1}, 1, T)$, $c_C(x_i, 0, T)$, $c_C(x_{j+1}, 1, T)$, $c_A(x_i, r_l, T)$, $c_B(x_i, r_l, T)$, $c_C(x_i, r_l, T)$ for all x_i ; $c_A(0, R_k, T)$, $c_B(0, R_k, T)$, $c_C(0, R_k, T)$ for all R_k ; $c_A(0, R_k, T)$, $c_B(0, R_k, T)$, $c_C(0, R_k, T)$ for all R_k which may be computed from the boundary conditions as is shown below.

Condition [43] yields

$$[79] \quad \sum_{j=1}^{N_x+1} A_{i,j} c_A(x_j, r_k, T) = - \sum_{j=1}^{N_x+1} A_{i,j} c_B(x_j, r_k, T)$$

Expanding the sums, introducing [36], [37], and [42b], and solving for the concentration of species B at the electrode surface ($x_i = 0$) gives:

$$[80] \quad c_h(0, R_i, T) = -\frac{1}{A_{i,1}} \left[A_{i,0,1} + \sum_{j=1}^{N+1} A_{i,j} c_h(X_j, R_i, T) + c_h(X_i, R_i, T) \right]$$

Condition [44] is changed to

$$\sum_{j=1}^{N+1} A_{i,j} c_h(X_j, R_i, T) = 0$$

which gives, after introduction of [37],

$$[81] \quad c_h(0, R_i, T) = -\frac{1}{A_{i,1}} \sum_{j=1}^{N+1} A_{i,j} c_h(X_j, R_i, T)$$

In analogy it may be derived from [45] and [40] or [41]

$$[82] \quad c_h(0, R_i, T) = -\frac{1}{A_{i,1}} \left[A_{i,0,1} + \sum_{j=1}^{N+1} A_{i,j} c_h(X_j, R_i, T) \right]$$

$$[83] \quad c_h(0, R_i, T) = -\frac{1}{A_{i,1}} \sum_{j=1}^{N+1} A_{i,j} c_h(X_j, R_i, T)$$

and

$$[84] \quad c_h(0, R_i, T) = -\frac{1}{A_{i,1}} \sum_{j=1}^{N+1} A_{i,j} c_h(X_j, R_i, T)$$

The condition [46] gives, for species A, B, and C

$$[85] \quad c_h(X_i, 0, T) = -\frac{1}{C_{i,1}} \left[C_{i,0,1} c_h(X_i, 1, T) + \sum_{j=1}^{N+1} C_{i,j} c_h(X_j, R_i, T) \right]$$

$$[86] \quad c_h(X_i, 0, T) = -\frac{1}{C_{i,1}} \left[C_{i,0,1} c_h(X_i, 1, T) + \sum_{j=1}^{N+1} C_{i,j} c_h(X_j, R_i, T) \right]$$

and

$$[87] \quad c_h(X_i, 0, T) = -\frac{1}{C_{i,1}} \left[C_{i,0,1} c_h(X_i, 1, T) + \sum_{j=1}^{N+1} C_{i,j} c_h(X_j, R_i, T) \right]$$

Condition [50] yields

$$[88] \quad \sum_{j=1}^{N+1} C_{i,j} c_h(X_j, R_i, T) = \sum_{j=1}^{N+1} \sqrt{F} C_{i,j} c_h(X_j, R_i, T)$$

which may be expanded and transformed using [40] and [47]

$$[89] \quad [C_{i,0,1} c_h(X_i, 1, T) + C_{i,1,1} c_h(X_i, 0, T)] - \sum_{j=1}^{N+1} C_{i,j} c_h(X_j, R_i, T) + \sum_{j=1}^{N+1} \sqrt{F} C_{i,j} c_h(X_j, R_i, T) + \sqrt{F} C_{i,0,1}$$

From [85] and [89] it follows that

$$[90] \quad c_h(X_i, 1, T) = c_h(X_i, 0, T) = \frac{1}{G_i} \left[\sum_{j=1}^{N+1} H_{i,j} c_h(X_j, R_i, T) + \sum_{j=1}^{N+1} C_{i,j} \sqrt{F} C_{i,j} c_h(X_j, R_i, T) + C_{i,1} \sqrt{F} C_{i,0,1} \right]$$

where

$$[91] \quad G_i = [C_{i,0,1} c_h(X_i, 1, T) - \sqrt{F} C_{i,1} C_{i,0,1} - C_{i,1,1} c_h(X_i, 0, T)]$$

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and

$$[92] \quad H_i = C_{i,0,1} C_{i,1} - C_{i,1} C_{i,0,1}$$

Conditions [46] and [51] ([46] and [52]) are treated similarly using the corresponding boundary expressions. Thus,

$$[93] \quad c_h(X_i, 1, T) = c_h(X_i, 0, T) = \frac{1}{G_i} \left[\sum_{j=1}^{N+1} H_{i,j} c_h(X_j, R_i, T) + \sum_{j=1}^{N+1} C_{i,j} \sqrt{F} C_{i,j} c_h(X_j, R_i, T) \right]$$

and

$$[94] \quad c_h(X_i, 1, T) = c_h(X_i, 0, T) = \frac{1}{G_i} \left[\sum_{j=1}^{N+1} H_{i,j} c_h(X_j, R_i, T) + \sum_{j=1}^{N+1} C_{i,j} \sqrt{F} C_{i,j} c_h(X_j, R_i, T) \right]$$

Finally, the concentrations of species A at $R = 0$ and all X_i may be calculated by introducing [90] into [85]:

$$[95] \quad c_h(X_i, 0, T) = -\frac{1}{G_i} \left[\sum_{j=1}^{N+1} G_{i,j} c_h(X_j, R_i, T) + \sum_{j=1}^{N+1} C_{i,0,1} \sqrt{F} C_{i,j} c_h(X_j, R_i, T) + C_{i,0,1} \sqrt{F} C_{i,0,1} \right]$$

while it holds from [93] and [94] ([94] and [97]), that

$$[96] \quad c_h(X_i, 0, T) = -\frac{1}{G_i} \left[\sum_{j=1}^{N+1} G_{i,j} c_h(X_j, R_i, T) + \sum_{j=1}^{N+1} C_{i,0,1} \sqrt{F} C_{i,j} c_h(X_j, R_i, T) \right]$$

and that

$$[97] \quad c_h(X_i, 0, T) = -\frac{1}{G_i} \left[\sum_{j=1}^{N+1} G_{i,j} c_h(X_j, R_i, T) + \sum_{j=1}^{N+1} C_{i,0,1} \sqrt{F} C_{i,j} c_h(X_j, R_i, T) \right]$$

Now we may substitute the unknown terms in eqs. [71]–[76]. From [71], using [90] and [95] it follows

$$[98] \quad \frac{dc_h}{dt} \Big|_{X_i=0} = \theta \left[R_{i,0,1} + \sum_{j=1}^{N+1} R_{i,j} c_h(X_j, R_i, T) \right] + \frac{1}{G_i} \left[I_{i,0} \sqrt{F} C_{i,0,1} + \sum_{j=1}^{N+1} I_{i,j} (\sqrt{F} C_{i,1} - C_{i,0,1} C_{i,j}) - I_{i,1} C_{i,0,1} + I_{i,1} C_{i,0,1} c_h(X_i, R_i, T) + \sum_{j=1}^{N+1} I_{i,j} \sqrt{F} C_{i,j} c_h(X_j, R_i, T) \right]$$

where

$$[99] \quad I_{i,0} = E_{i,0,1} C_{i,1} - E_{i,1} C_{i,0,1}$$

and

$$[100] \quad I_{i,j} = E_{i,j} C_{i,1} - E_{i,1} C_{i,j}$$

Combination of [72] with [80], [93], and [96] gives

$$[101] \quad \frac{dc_h}{dt} \Big|_{X_i=0} = \frac{\theta}{A_{i,1}} \left[\sum_{j=1}^{N+1} H_{i,j} c_h(X_j, R_i, T) - R_{i,1} \left[A_{i,0,1} + \sum_{j=1}^{N+1} A_{i,j} c_h(X_j, R_i, T) \right] + \frac{1}{G_i} \left[\sum_{j=1}^{N+1} I_{i,j} (\sqrt{F} C_{i,1} - C_{i,0,1} C_{i,j}) - I_{i,1} C_{i,0,1} + I_{i,1} C_{i,0,1} c_h(X_i, R_i, T) + \sum_{j=1}^{N+1} I_{i,j} \sqrt{F} C_{i,j} c_h(X_j, R_i, T) \right] - \theta c_h \right]$$

where

$$[102] \quad K_{i,j} = B_{i,j}A_{i,j} - B_{i,j}A_{i,j}$$

For the inner function describing the concentration of species C it holds from [73] with [81], [94] and [97]

$$[103] \quad \frac{dc_i}{dt} \Big|_{R_i} = \frac{\beta}{A_{i,j}} \left[\sum_{j=1}^{N_j} K_{i,j} c_i(X_i, R_i, T) \right] + \frac{1}{G_i} \left[\sum_{j=1}^{N_j} L_{i,j} \sqrt{D} C_{i,j} - C_{i,j+1} - L_{i,j} C_{i,j+1} + L_{i,j} C_{i,j+1} \right] c_i(X_i, R_i, T) + \sum_{j=1}^{N_j} L_{i,j} \sqrt{D} C_{i,j} c_i(X_i, R_i, T) + ac_i$$

The outer functions for all three species are derived following the same scheme. The time derivative of c_i in the outer domain is found from [74], [82], and [90], and in analogy the corresponding equations for species B or C may be calculated from [75], [83], and [93] or [76], [84], and [94].

$$[104] \quad \frac{dc_i}{dt} \Big|_{R_i} = \frac{\beta}{A_{i,j}} \left[K_{i,j} c_i(X_i, R_i, T) \right] + \frac{\beta}{G_i} \left[L_{i,j+1} + \sum_{j=1}^{N_j} F_{i,j} K_{i,j} c_i(X_i, R_i, T) + \sum_{j=1}^{N_j} L_{i,j} c_i(X_i, R_i, T) \right] + ac_i$$

with

$$[105] \quad L_{i,j} = F_{i,j} G_i + C_{i,j} \sqrt{D} C_{i,j} F_{i,j}$$

$$[106] \quad \frac{dc_i}{dt} \Big|_{R_i} = \frac{\beta}{A_{i,j}} \left[\sum_{j=1}^{N_j} K_{i,j} c_i(X_i, R_i, T) \right] + \frac{\beta}{G_i} \left[\sum_{j=1}^{N_j} F_{i,j} K_{i,j} c_i(X_i, R_i, T) + \sum_{j=1}^{N_j} L_{i,j} c_i(X_i, R_i, T) \right] + ac_i$$

$$[107] \quad \frac{dc_i}{dt} \Big|_{R_i} = \frac{\beta}{A_{i,j}} \left[\sum_{j=1}^{N_j} K_{i,j} c_i(X_i, R_i, T) \right] + \frac{\beta}{G_i} \left[\sum_{j=1}^{N_j} F_{i,j} K_{i,j} c_i(X_i, R_i, T) + \sum_{j=1}^{N_j} L_{i,j} c_i(X_i, R_i, T) \right] + ac_i$$

Now the set of six second order partial differential equations [26]–[28] and [30]–[32] has been transformed into a set of $3N_j(N_j + N_s)$ simultaneous ordinary differential equations [90], [101], [103], [104], [106], and [107]. Such systems may be easily integrated by computer programs to give the time dependence of the concentrations of species A, B, and C at the collocation points (X_i, R_i) and (X_i, R_i) . We describe the concentration profiles which develop in front of the electrode by the concentrations at the collocation points.

We note the structural similarity of the formulae describing the time derivatives of the concentrations. This is one of the advantages of the orthog-

onal collocation method, for if the theory has been derived once it is very easy to express it in a computer language. Furthermore, if more complicated mechanisms are to be simulated, only minor changes in the formulae have to be performed. Also, the extension of the theory to other experimental techniques, i.e. other boundary conditions, is easy. This will be shown in another paper of this series.

We also note that the implementation of the expressions describing species C is not essential to calculate the concentration profiles of A and B, for C is not electroactive and does not react to or with an electroactive intermediate [16]. However, if one

desires, for example, to calculate the spectroelectrochemical response [14] of species C one is interested in the concentration profiles of C as well.

Calculation of the current

In the one-dimensional diffusion model the current i is simply calculated from the flux of the substrate at the electrode surface:

$$[108] \quad i = nFAD \left(\frac{\partial c_A}{\partial x} \right)_{x=0}$$

where A is the electroactive area of the electrode. This implies that the flux of the substrate is uniform over A .

This is clearly not true in the case of a two-dimensional model considering edge effects. Instead, the current is given by

$$[109] \quad i = nFD \int_0^{2\pi} \left(\frac{\partial c_A}{\partial r} \right)_{r=0} dr$$

where the r -dependent flux through an infinitesimally small part $2\pi r dr$ of the electrode is integrated over the electroactive section, $0 \leq r \leq r_0$ of the r -coordinate.

Using the transformation equations [21]–[23], [25], and the definition of β (eq. [29]), eq. [109] yields

$$[110] \quad i = \frac{nFD \int_0^{2\pi} \left(\frac{\partial c_A}{\partial x} \right)_{x=0} dR}{L} = 2nFAC \frac{\sqrt{D}}{\sqrt{t}} \int_0^{2\pi} \left(\frac{\partial c_A}{\partial x} \right)_{x=0} dR$$

The chronoamperometric "constant" $i\sqrt{t}$ is given by

$$[111] \quad i\sqrt{t} = 2nFAC \frac{\sqrt{D}}{\sqrt{t}} \int_0^{2\pi} \left(\frac{\partial c_A}{\partial x} \right)_{x=0} dR$$

Alternatively, we may express this product in terms of the "Cottrell constant" (2)

$$[112] \quad (i\sqrt{t})_C = \frac{nFAC \sqrt{D}}{\sqrt{t}}$$

as

$$[113] \quad \frac{i\sqrt{t}}{(i\sqrt{t})_C} = 2\sqrt{\pi} \int_0^{2\pi} \left(\frac{\partial c_A}{\partial x} \right)_{x=0} dR$$

Since the differential quotients $(\partial c_A / \partial x)$ are given at the collocation points we may calculate the integral exactly by a quadrature formula of the type (6.28, p. 134)

$$[114] \quad \int_0^{2\pi} f(r) w_j(r) dr = \sum_{j=1}^{N_j} Q_j f(r_j)$$

where the Q_j are elements of a vector depending on the trial function chosen.

Thus, from [67]

$$[115] \quad \int_0^{2\pi} R \left(\frac{\partial c_A}{\partial x} \right)_{x=0} dR = \sum_{j=1}^{N_j} Q_j \left[R_j \sum_{i=1}^{N_j} A_{i,j} c_i(X_i, R_i, T) \right]$$

Expanding the sum over j and introducing the boundary conditions [36] and [42b], [115] changes to

$$[116] \quad \int_0^{2\pi} R \left(\frac{\partial c_A}{\partial x} \right)_{x=0} dR = \sum_{j=1}^{N_j} Q_j \left[R_j A_{i,j+1} + R_j \sum_{i=1}^{N_j} A_{i,j} c_i(X_i, R_i, T) \right]$$

The Gaussian weights Q_j for $j=1$ and $j=N_j+2$ are equal to zero, thus [116] reduces to

$$[117] \quad \int_0^{2\pi} R \left(\frac{\partial c_A}{\partial x} \right)_{x=0} dR = \sum_{j=1}^{N_j} Q_j \left[R_j A_{i,j+1} + R_j \sum_{i=1}^{N_j} A_{i,j} c_i(X_i, R_i, T) \right]$$

which is a convenient form for computation of the integral.

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Correlation of the hydrogenation catalytic activity of dinuclear complexes of rhodium(I) with the formation of carbonyl-bridged complexes

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Methanolic solutions of the dinuclear cationic complexes of rhodium(I), $[\text{Rh}_2(\mu\text{-X})(\text{CO})_2(\text{PPh}_3)_2(\text{ECN})_2]^+$ ($\text{X} = \text{Cl}, \text{Br}; \text{E} = \text{P}, \text{As}$) and iridium(I), $[\text{Ir}_2(\text{Cl})_2(\text{PPh}_3)_2(\text{PCH}_2\text{PPh}_2)_2]^+$, are each active catalysts for the hydrogenation of alkynes to alkenes, and alkenes to alkanes. Neutral complexes with pseudohalogen ligands have also been investigated but only the cyano-complex $[\text{Rh}_2(\mu\text{-N})_2(\text{CO})_2(\text{PPh}_3)_2(\text{PCH}_2\text{PPh}_2)_2]$, and the arsenic analog, are of significant catalytic activity. Only those complexes which reversibly form bridged complexes with ligands such as CO are also catalytically active.

Soluble complexes with dithiocarbamate ligands are not efficient catalysts under the conditions tested.

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Les complexes cationiques dinucléaires de rhodium(I), $[\text{Rh}_2(\mu\text{-X})(\text{CO})_2(\text{PPh}_3)_2(\text{ECN})_2]^+$ ($\text{X} = \text{Cl}, \text{Br}; \text{E} = \text{P}, \text{As}$), et d'iridium(I), $[\text{Ir}_2(\text{Cl})_2(\text{PPh}_3)_2(\text{PCH}_2\text{PPh}_2)_2]^+$, en solution dans le méthanol, sont des catalyseurs actifs d'hydrogénation des alcynes en alcènes et d'alcènes en alcanes. On a également étudié les complexes neutres ayant des ligands pseudohalogénés, mais seuls les complexes de type cyano $[\text{Rh}_2(\mu\text{-N})_2(\text{CO})_2(\text{PPh}_3)_2(\text{PCH}_2\text{PPh}_2)_2]$ et leurs analogues de type arsène ont une activité catalytique significative. Seuls les complexes qui forment de manière réversible des complexes liés par un pont avec des ligands, tels le CO, sont également des catalyseurs actifs.

Les complexes solubles avec les ligands dithiocarbamates ne sont pas des catalyseurs efficaces dans les conditions étudiées.

(Traduit par le journal)

Introduction

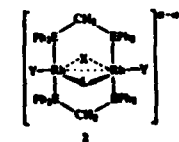
The first dinuclear rhodium complex of the geometry designated "A-frame" (1) $[\text{Rh}_2(\mu\text{-S})(\text{CO})_2(\text{dppe})_2]$ (dppe is $\text{Ph}_2\text{PCH}_2\text{PPh}_2$ and dppm is $\text{Ph}_2\text{PCH}_2\text{AsPh}_2$) was described in 1977 (1). The synthesis and structure of related complexes of rhodium or iridium has since received much attention (2-16). For complexes of rhodium(I), several instances of co-ordination of two-electron donor ligands (L; e.g. CO, SO_2) at the site bridging the metal centers (2) have been reported (2, 3, 6, 9, 12-16). The complex cation, $[\text{Rh}_2(\mu\text{-Cl})(\text{CO})_2(\text{dppe})_2]^+$, forms a 1:1 labile, symmetrical (μ , ^{31}P NMR) complex with diphenylacetylene (6),¹ and similar neutral complexes have also recently been described (11). Co-ordination of an alkyne at the bridging site (2; L = alkyne) is accompanied by lengthening of the carbon-carbon bond such that the product is accurately described as aryl-dimetallated alkynyl complex (11). No complex in which an alkene is co-ordinated at either a bridging or terminal site of an A-frame complex has yet been described.

¹Received February 12, 1982.

²This system has been investigated elsewhere (16) Canada, private communication) and in this laboratory. The differences in spectroscopic parameters between solutions of the cationic and neutral complexes are now considered to be due to differences in conditions. A cationic complex containing bridging $\text{CH}_3\text{XCOXCOCH}_3$ has been well characterized (16).



1 ($\text{E} = \text{P}$ or As , X is no longer ligand of formal charge +). Y is a second two-electron donor ligand, usually CO



The A-frame structure, with the proximity of the two metal centres, has strong potential for dinuclear activation of unsaturated substrates. However, the complex $[\text{Rh}_2(\mu\text{-S})(\text{CO})_2(\text{dppe})_2]$ has been shown to have low activity for the hydrogenation of ethylene (13). Herein I report that those cationic or neutral complexes of rhodium (or iridium) which form carbonyl-bridged complexes (2, L = CO) are also efficient catalysts.

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