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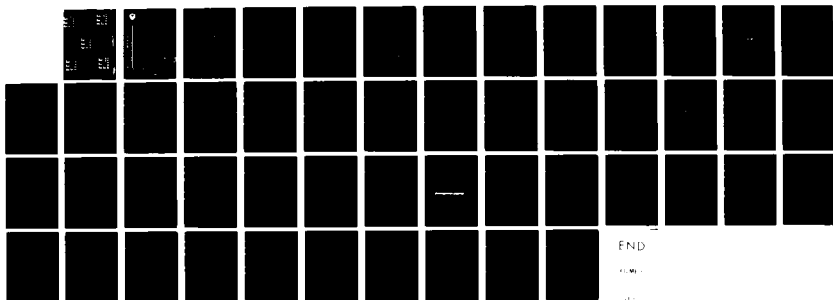
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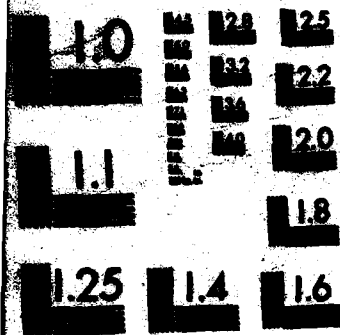
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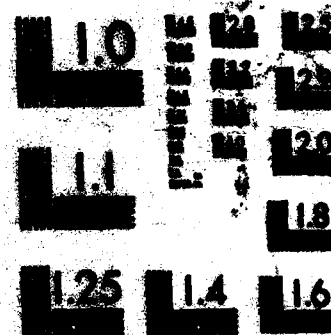
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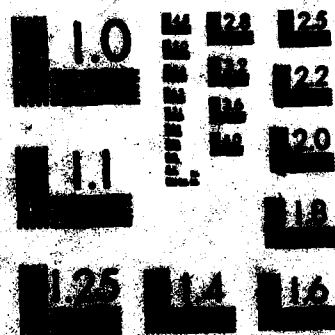




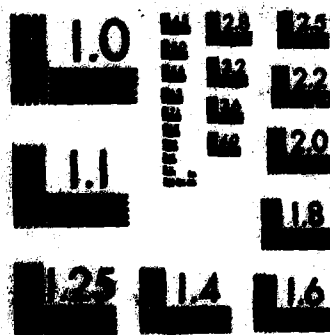
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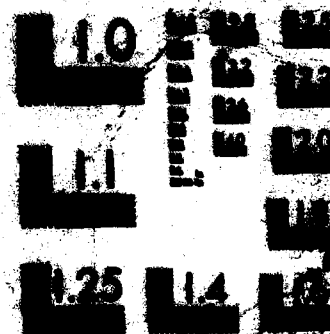
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Research and Development Technical Report

DELET-TR-0420-3

STUDIES LEADING TO THE DEVELOPMENT OF HIGH-RATE LITHIUM-SULFURYL CHLORIDE BATTERY TECHNOLOGY

John C. Hall and Mark Koch
Gould Inc.
Gould Research Center
49 Gould Center
Rolling Meadows, Illinois 60008

October 1982

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The overall aim of the program is an examination of the viability of an active electrolyte lithium-sulfuryl chloride battery system. The specific objectives are:

- o Quantify the stability of lithium in sulfuryl-chloride based electrolyte.
- o Develop means to stabilize in the lithium anode in sulfuryl chloride based electrolyte to meet Army storage requirements.
- o Establish the performance limits of $\text{Li}/\text{SO}_2\text{Cl}_2$ cells with respect to cathode and electrolyte composition.

During the third quarter we have

- o Developed an improved cell design which permit long term storage tests.
- o Conducted stability tests for sulfuryl chloride electrolyte with added chlorine and bromine.
- o Characterized the passivating film in these electrolyte with complex plane impedance.
- o Characterized the performance of cells built with added chlorine and added bromine at temperatures down to -20°C and current densities up to $30\text{mA}/\text{cm}^2$.

We have found that

- o Added bromine has an adverse effect on stability of lithium in sulfuryl chloride cells.
- o By elimination of solder seals a long term lithium corrosion rate in cells without additives of 5 mils per year is feasible.
- o Cells with added chlorine deliver greater coulombic capacities with, however, lower voltages than cells with added bromine at -20°C .
- o Cells with added bromine discharge approximately 500mV above cells with added chlorine at current densities of 20 to 30 mA/cm^2 .

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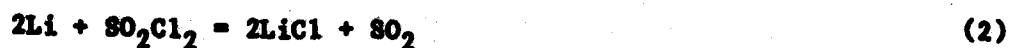
I. Introduction

Batteries built with an inorganic catholytes (e.g., thionyl chloride and sulfuryl chloride) are very attractive for military applications on the basis of their high demonstrated energy density, and high discharge rate capability when compared with organic electrolyte lithium cells. By far the greatest attention has been given to the lithium-thionyl chloride (Li/SOCl₂) system. Cells and batteries have been demonstrated with energy densities in excess of 600 Wh/kg (1).

Despite the impressive levels of performance demonstrated with Li/SOCl₂ wide spread deployment has not yet occurred principally due to unresolved safety issues. One of the probable causes for explosion of Li/SOCl₂ cells during overheating either as a result of short circuiting or an external heat source is the reaction between lithium and sulfur (2), which is formed by cell discharge according to the discharge reaction (3)



In this regard Li/SO₂Cl₂ may be a fundamentally safer system as no sulfur is formed in the discharge reaction (4)



Development of potentially safer Li/SO₂Cl₂ cells for military applications will require the demonstration of cells with a long shelf life and performance which at least approaches that of Li/SOCl₂. The former requirement is principally a function of controlling the rate of lithium corrosion. The latter requirement entails optimisation of the anode with respect to voltage delay and the cathode with respect to voltage and capacity in the range of 1 to 20 mA/cm² at temperatures down to -40°C.

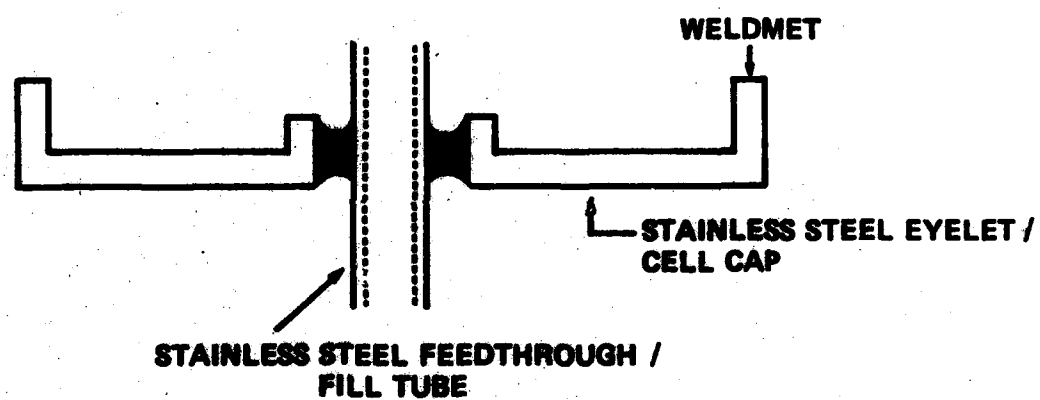
The scope of the present contract encompasses investigation and improvement of both anode stability and the performance of both electrodes.

Previous work in the first and second quarters of the contract has established

- An intrinsically low rate of lithium corrosion (less than 1 mil per year, mpy) in sulfuryl chloride electrolyte
- A much higher in cell corrosion rate (100 to 50 mpy) which we ascribe to impurities in the components of cell construction
- Substantial voltage delay after storage which is substantially reduced by use of a bromine additive to the electrolyte
- Superior room temperature positive electrode operating voltage and capacity with added bromine and to a lesser extent added chlorine
- The conductivity and density of sulfuryl chloride electrolyte.

Work in the third quarter described herein has focused on lowering the in cell corrosion rate, establishing the preferred cathode composition and determining the performance characteristics of $\text{Li}/\text{SO}_2\text{Cl}_2$ cells at temperature of -20°C and below.





(2031)

Figure 1 Design of Solderless Feedthrough/fill tube/Cell Cap

TABLE 1

**Design Characteristics for Second Generation Cells
employed in Anode Stability Studies**

- 304 stainless steel case with integral glass to metal seal
- Positive grounded
- 90% Shawinigan Black, 10% PTFE cathode, 0.02" thick, nickel Exmet supported
- 0.002" negative nickel sheet current collector
- 0.01" lithium foil pressed to both faces of the current collector
- 0.004" porous Teflon insulator between the anode and the case
- Cell TIG welded

corrosion rate shortly after fill (approximately 20 hours) is about 25% of our previous best data reported in last quarter. However, the data do not show the previously observed exponential decrease in corrosion with time. Rather the corrosion rate appears to be linearly decreasing and has reached a value of 32 mpy 175 hours after cell activation.

This value is close to what was previously observed with a soldered feedthrough. This suggests in the case of the bromine additive at least that the long term corrosion rate may be intrinsic to the chemistry and not the level of purity.

The data in Figure 2 also may explain why cells with added bromine show less voltage delay after storage. If the lithium surface continues to undergo a high rate of corrosion then the formation of a thick passivating film may be inhibited which would result in less voltage delay.

In Figure 3 corrosion data for a $\text{Li}/\text{SO}_2\text{Cl}_2$ cell without any additive is given. In this case the data do show a clear exponential decay. Although the initial corrosion rate is slightly higher than that shown in Figure 2, by 40 hours the corrosion rate has fallen to 32 mpy year whereas with the bromine additive the corrosion rate is 126 mpy.

The data in Figure 3 show an approximate three-fold decrease in lithium corrosion when compared to results reported in the first quarterly report. If the trend in Figure 3 continues we expect the long term (e.g. 500 hour) corrosion rate to fall below 5 mpy.

The markedly different behavior between cells with and without added bromine suggests that LiBr may be disrupting the LiCl lattice in the latter case and result in higher corrosion rate. This conclusion, however, probably over simplifies the situation as ampule experiments have shown a lower corrosion rate for lithium in the bromine containing electrolyte. Thus trace impurities may still be a problem in both $\text{Li}/\text{SO}_2\text{Cl}_2 + \text{Br}_2$ and unadulterated $\text{Li}/\text{SO}_2\text{Cl}_2$.

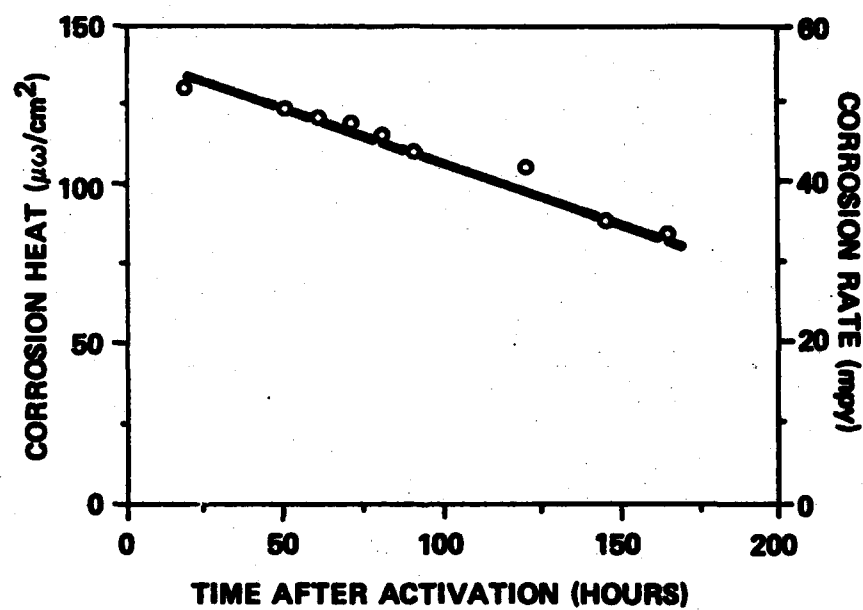


Figure 2 Li/SO₂Cl₂+Br₂ Cell Corrosion Rate Data
 o 1.25 " Ø Cell, 1" Ø Electrodes
 o Floating Double Faced Li Electrode
 o 25°C, Positive SS Case

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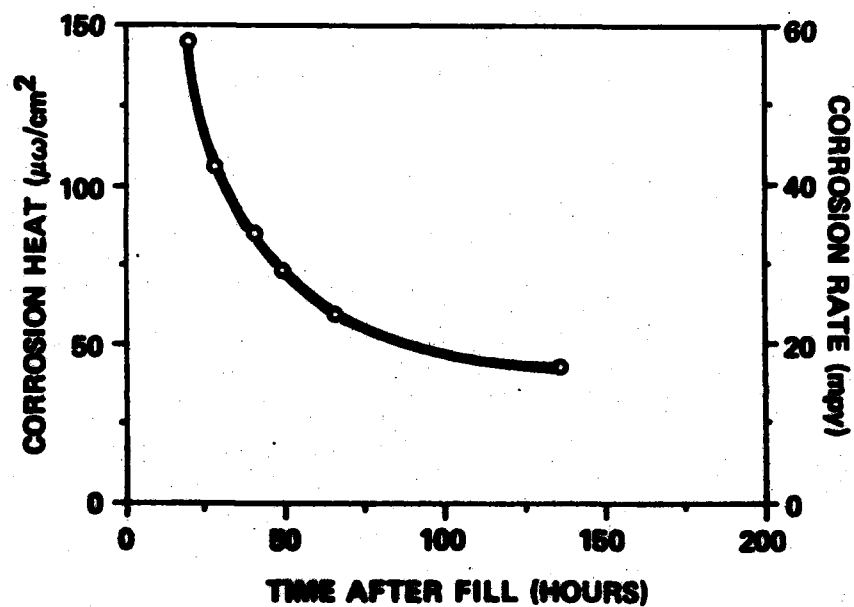


Figure 3 Lithium Corrosion in a $\text{Li}/\text{SO}_2\text{Cl}_2$ Cell
without Additives
 o 25°C
 o 1.25 "D" cell, 1" $\odot$ Electrodes
 o Double faced Anode
 o Positive SS case

(2088)

Finally we are measuring the rate of lithium corrosion in a cell with added chlorine. Data over the first 135 hours of testing are summarized in Figure 4. Qualitatively the behavior of this cell is very similar to that of a Li/SO₂Cl₂ cell without additives. The corrosion rate at any time is slightly higher with added chlorine but in both cases a clear exponential decay is evident.

We would expect the similar data trends shown in Figure 3 and 4. Chlorine will always be present in sulfuryl chloride due to thermal decomposition



With added chlorine no unexpected products should form in the passivating crystal lattice. The slightly higher corrosion rate with chlorine addition probably indicates electrolyte contamination during chlorine addition.

The heat generation of all three cells is still being monitored. Now that we have apparently solved the feedthrough problem these cells will provide baseline data for improvements in anode stability.

II.2 Film Formation Studies

Complex plane impedance analysis on the three cells built for anode corrosion measurements is being simultaneously carried out. In Figures 5 through 8 the impedance spectra for the Li/SO₂Cl₂ + Br₂ cell are given at intervals between 20 and 163 hours after activation. As has been previously observed the frequency at the maximum of the semicircle decreases with storage time and there is clear evidence by 163 hours of storage for the formation of a second lower frequency semicircle.

As has been previously observed in cells without additives both the film thickness and resistance increase with storage time. In Figure 9 both film resistance and thickness changes during storage are given; both continue to increase after 163 hours of storage.

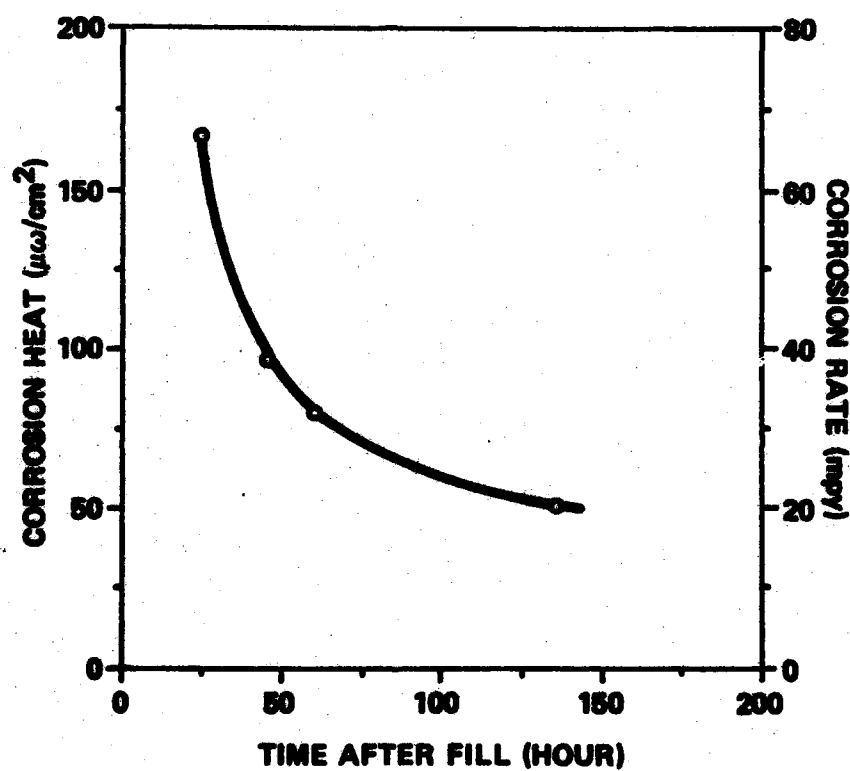


Figure 4 Lithium Corrosion in a $\text{Li}/\text{SO}_2\text{Cl}_2$ Cell with Added Cl_2

- o 25°C
- o 1.25 "Ø cell, 1" Ø Electrodes
- o Double faced Anode
- o Positive SS Case

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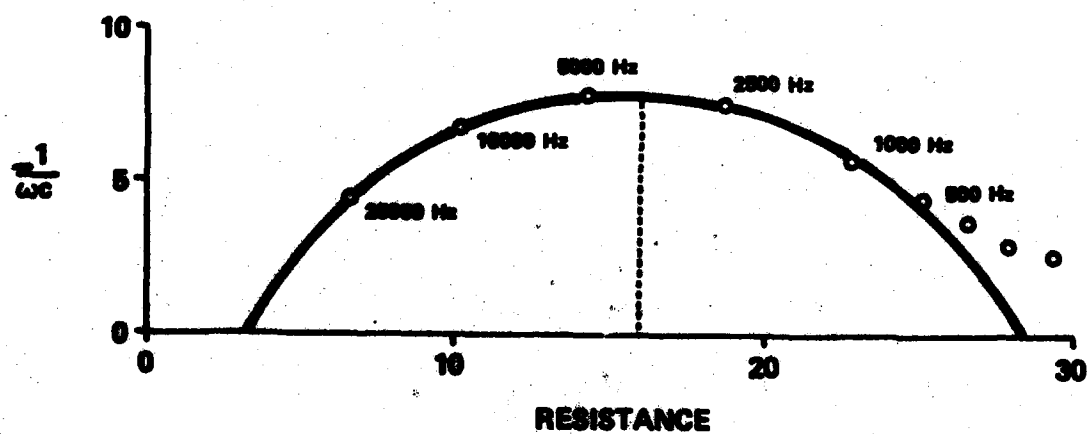


Figure 5 Complex Plane Impedance Spectra of $\text{Li}/\text{SO}_2\text{Cl}_2+\text{Br}_2$
 20 hours after activation
 $R_{\text{film}} = 25 \text{ ohms}, 13 \times 10^7 \text{ ohm cm}$
 $C_{\text{film}} = 5.1 \mu \text{ F}$
 Thickness = 19.1

(3085)

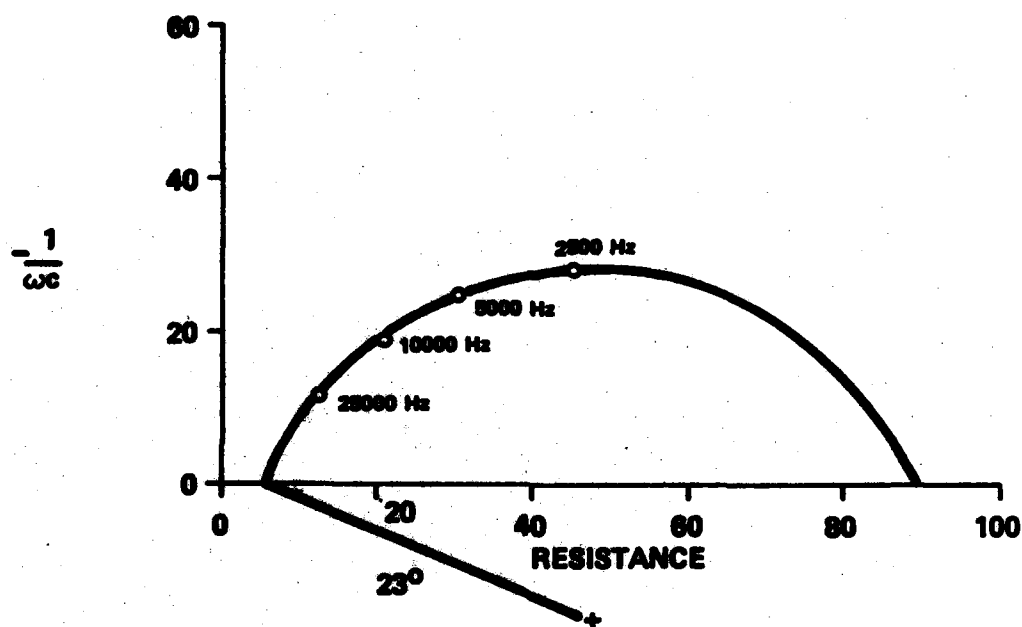


Figure 6 Complex Plane Impedance Spectrum of $\text{Li}/\text{SO}_2\text{Cl}_2+\text{Br}_2$
 68 hrs. after fill
 $R_{\text{film}} = 88 \text{ ohms}, 25 \times 10^8 \text{ ohm-cm}$
 $C_{\text{film}} = 2.83 \mu \text{ F}$
 Thickness = $345 \text{ \AA}$

(2069)

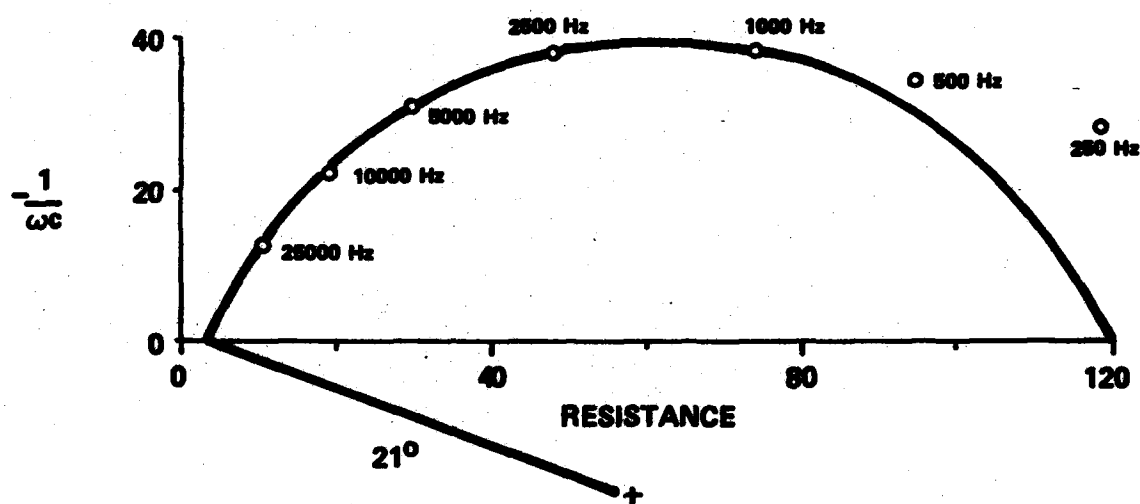


Figure 7 Complex Plane Impedance Spectrum of $\text{Li}/\text{SO}_2\text{Cl}_2+\text{Br}_2$
 144 hrs. after fill
 $R_{\text{film}} = 117 \text{ ohms}, 28 \times 10^7 \text{ ohm-cm}$
 $C_{\text{film}} = 2.3 \mu \text{ f}$
 Thickness = $420 \text{ \AA}$

(2068)

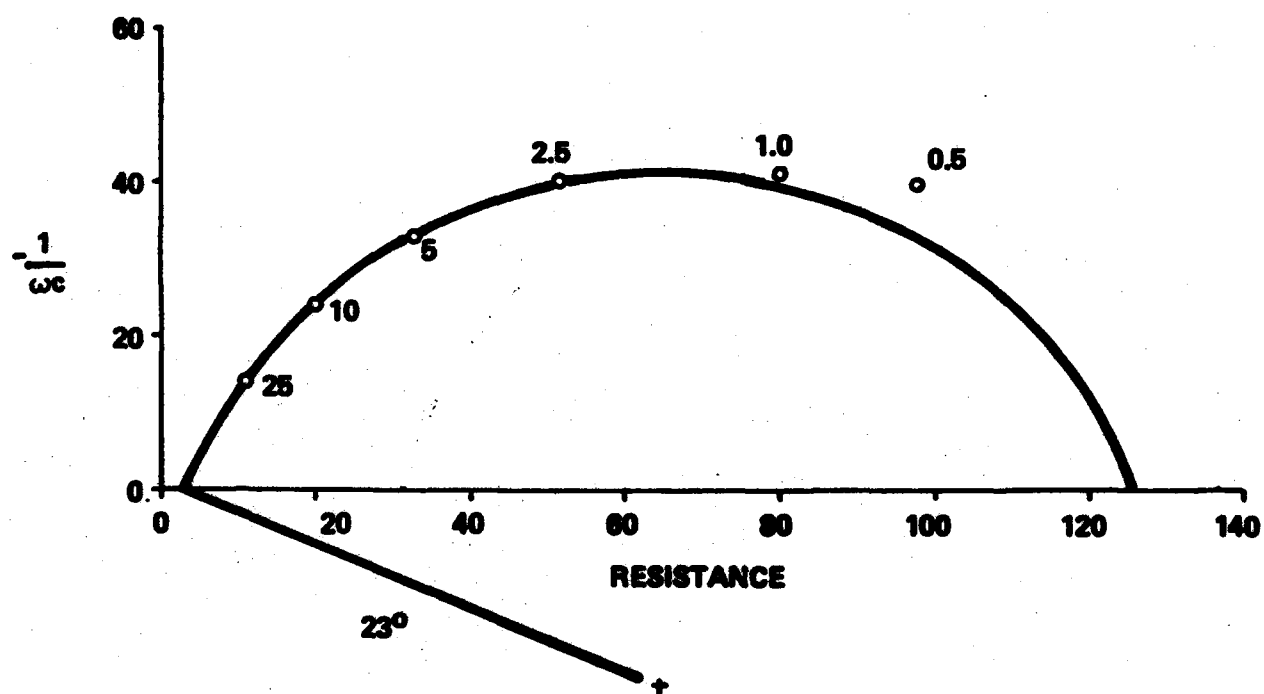


Figure 8 Complex Plane Impedance Spectrum of $\text{Li}/\text{SO}_2\text{Cl}_2+\text{Br}_2$ 163 hrs. after fill

$R_{\text{film}} = 123 \, \Omega, 24 \times 10^7 \, \Omega\text{-cm}$

$C_{\text{film}} = 2.0 \, \mu \text{ F}$

Thickness = $510 \, \Omega$

(2067)

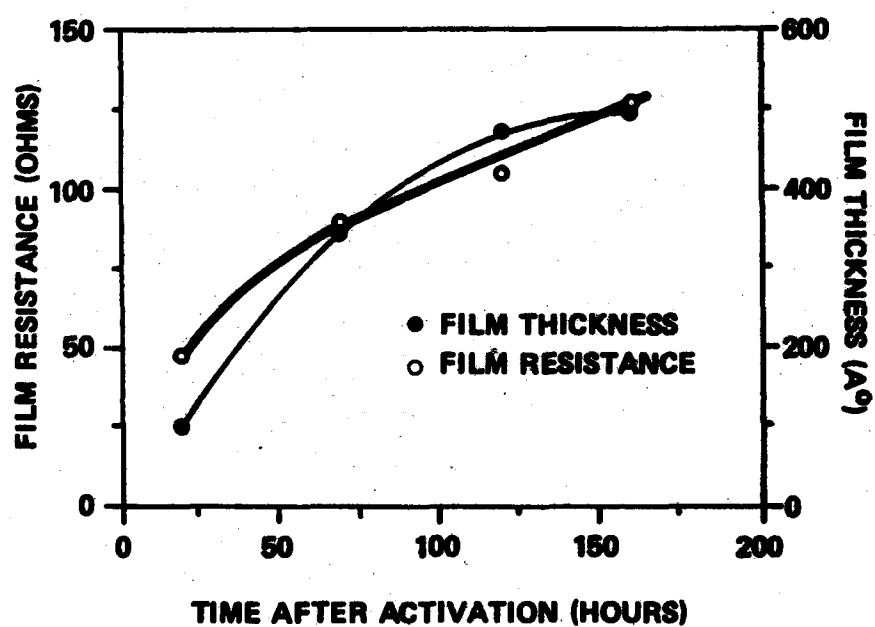


Figure 9 Change in Film Resistance and Thickness
During Storage of $\text{Li/SO}_2\text{Cl}_2+\text{Br}_2$

(2058)

The film resistivity is in fact considerably greater than that previously reported to $\text{Li}/\text{SO}_2\text{Cl}_2$ without additives. As can be seen from the numerical data in Figure 5-8 after the initial 20 hour measurement the resistivity is essentially constant at approximately 25×10^7 ohm-cm. This compares with the previously measured film resistance 8.2×10^7 ohm-cm after 412 hours of storage of a cell with a solder seal and no additive. It is probable that the lower film resistance previously measured indicates film doping with heavy metal ions which would also be expected to increase the corrosion rate.

A final observation from the data in Figures 5-8 is that the high frequency intercept of the semicircle with the real axis is essentially invariant (approximately 3 ohms) between 20 and 161 hours of storage. Previous solder sealed cells without additives had shown an increase of this intercept with storage time. This was attributed to an accumulation of non-adherent LiCl crystals adjacent to the anode increasing circuit resistance and hence the high frequency intercept. The constant intercept observed in the present work is most probably a result of the lower corrosion rate. It is also possible that the corrosion products remain adherent to the lithium possibly affording greater anode protection.

In Figures 10-12 complex plane impedance spectra for the cell with no additives are given. The data cover the period between 2 and 44 hours after activation. The following features are noted with regards to the data.

- Film resistivity, as was the case with $\text{Li}/\text{SO}_2\text{Cl}_2 + \text{Br}_2$, is considerably greater than observed in the first quarter work.
- Film thickness is much lower for equivalent storage periods than observed in the first quarter.
- The high frequency intercept appears to be constant with storage time.



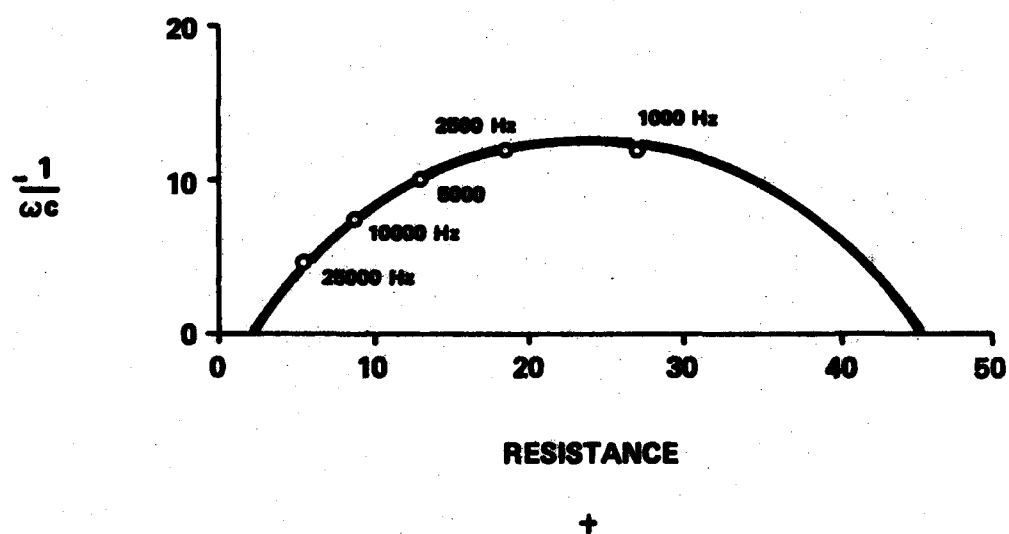


Figure 11 **Complex Plane Impedance Spectrum of $\text{Li}/\text{SO}_2\text{Cl}_2$**
without Additives 19 hrs. after fill
 $R_{\text{film}} = 43 \text{ ohms}, 37 \times 10^7 \text{ ohm cm}$
 $C_{\text{film}} = 8.43 \mu \text{ F}$
Thickness = 116A°

(2063)

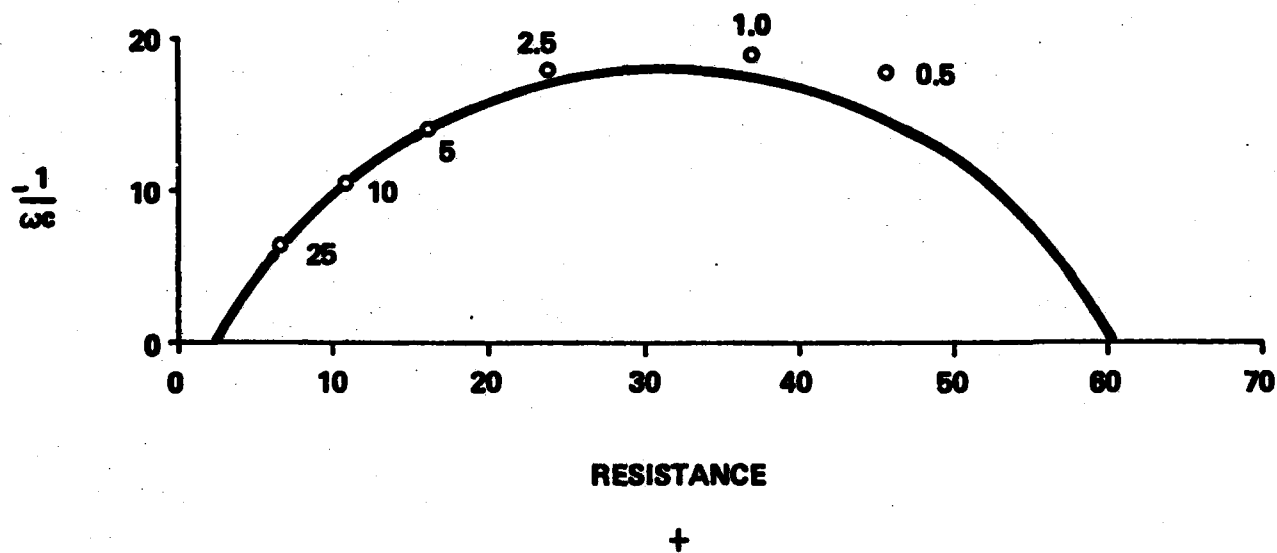


Figure 12 Complex Plane Impedance Spectrum of $\text{Li}/\text{SO}_2\text{Cl}_2$ without Additives 44 hrs. after fill
 $R_{\text{film}} = 58 \text{ ohm}, 25 \times 10^7 \text{ ohm-cm}$
 $C_{\text{film}} = 4.17 \mu \text{ F}$
 Thickness = $230 \text{ \AA}$

(2062)

- The frequency at the maximum of the semicircle decreases with storage time.

The cell has not been on test for a sufficient period to establish long term trends. The initial data are, however, encouraging with regard to improved anode performance as thinner films should result in less voltage delay.

An interesting though not yet fully understood result was obtained in our investigation of $\text{Li/SO}_2\text{Cl}_2$ anode film in the absence of additives. The cell under test is the second one built in this reporting period. The first cell for unknown reasons failed after 20 hours of testing. Prior to failure limited complex plane impedance and calorimetric was obtained and is summarized in Figure 13. What is noteworthy about the data is

- The cell corrosion rate at 20 hours is approximate 75% of that of the second cell.
- The film resistance and thickness are both considerably greater and the film resistivity less.
- The center of the semicircle lies much closer to the real axis than we have previously observed.

The final set of film studies carried out this quarter were on the cell with added chlorine. Complex plane impedance spectra for this cell are given in Figures 14-16 between 2 and 45 hours after activation.

The data differ radically from the cases summarized in Figures 5-8 and 10-12. For a given storage time film thickness is nearly double that of the cell without additives. Film resistivity on the other hand is much lower

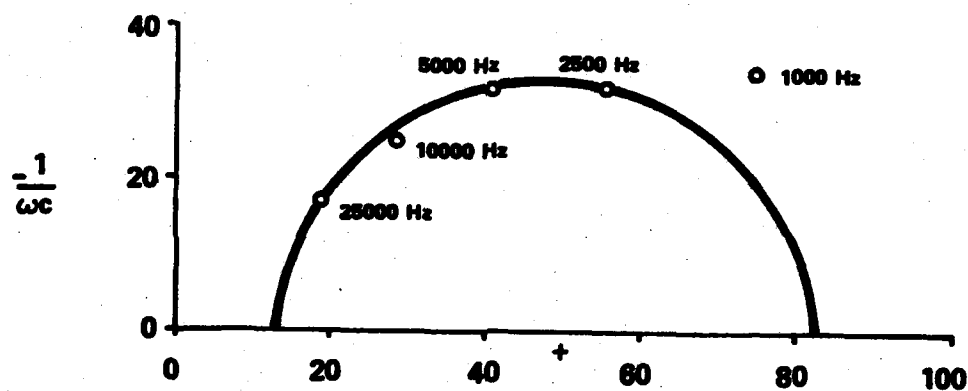


Figure 13 Complex Plane Impedance Spectrum of Li/SO₂Cl₂
Without Additives 20 hrs. after fill

$R_{\text{film}} = 68 \text{ ohms}, 8 \times 10^7 \text{ ohm cm}$

$C_{\text{film}} = 1.21 \mu \text{ F}$

Thickness = 806 Å

Corrosion Rate = 88 $\mu\omega/\text{cm}^2$, 36 mpy

OCV = 3.886V

(2061)

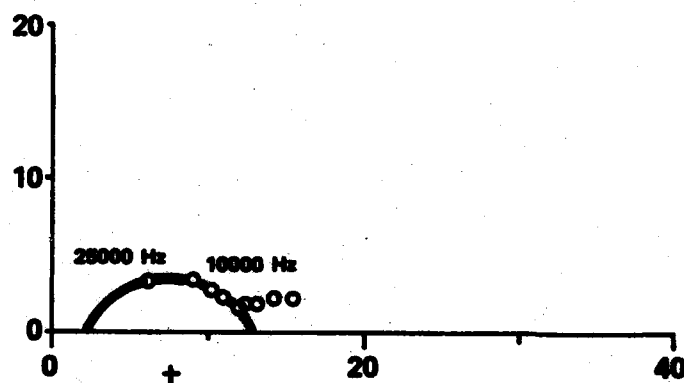


Figure 14 **Complex Impedance Spectrum of $\text{Li}/\text{SO}_2\text{Cl}_2+\text{Cl}_2$**
 2 hrs. after fill
 $R_{\text{film}} = 11 \text{ ohm}, 3 \times 10^7 \text{ ohm cm}$
 $C_{\text{film}} = 2.67 \mu \text{ F}$
 Thickness = $366 \text{ \AA}$

(2066)

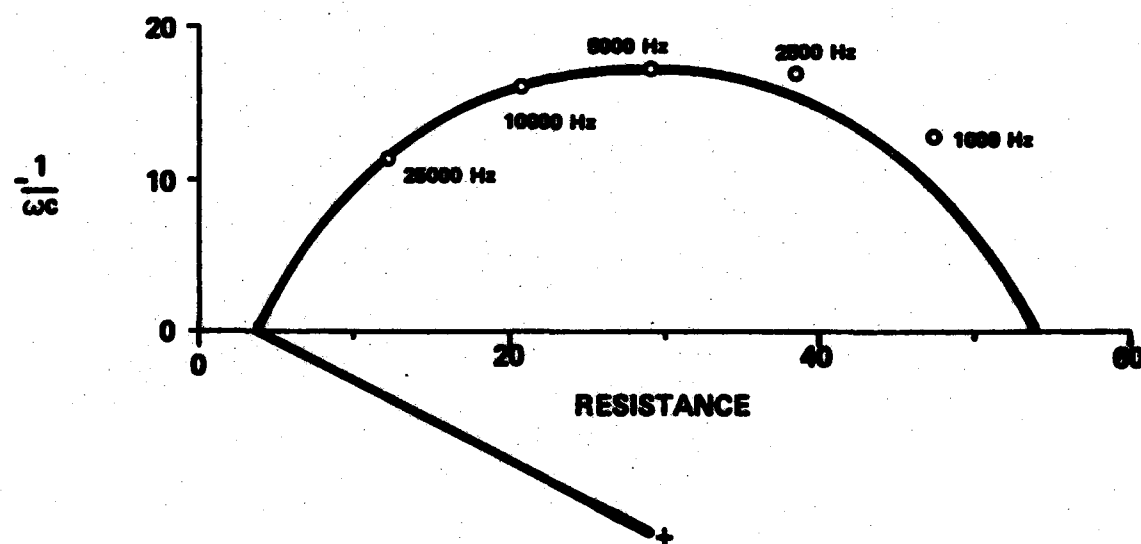


Figure 15 Complex Plane Impedance Spectrum of $\text{Li/SOCl}_2 + \text{Cl}_2$ 19 hrs, after fill

$R_{\text{film}} = 50 \text{ ohms}, 9 \times 10^7 \text{ ohm cm}$

$C_{\text{film}} = 1.83 \mu \text{ F}$

Thickness = $556 \text{ \AA}$

(2080)

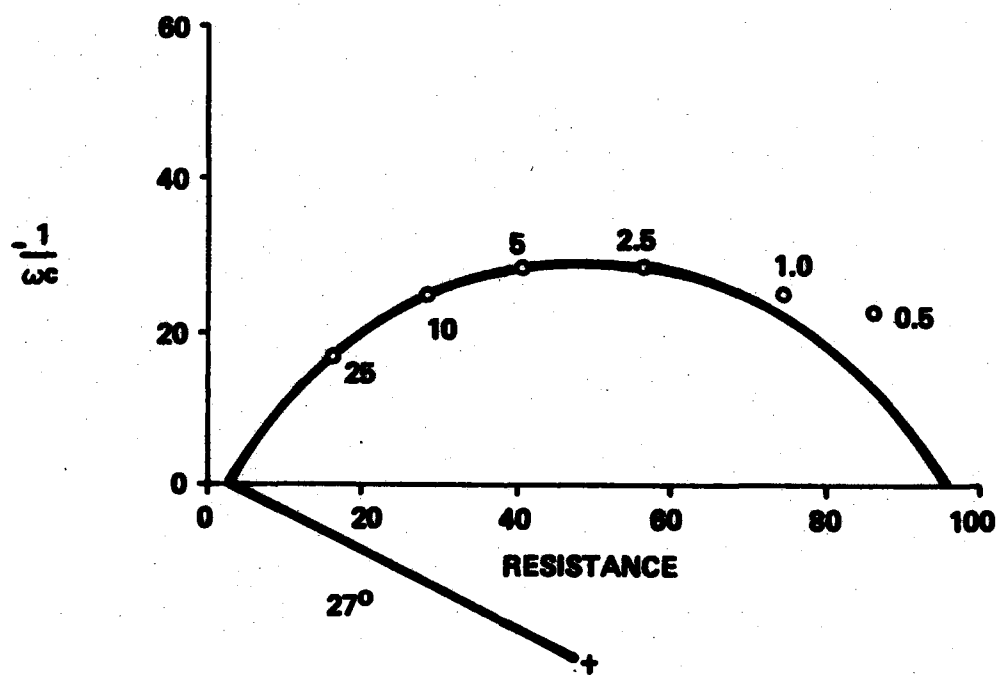


Figure 16 Complex Plane Impedance Spectrum of $\text{Li}/\text{SO}_2\text{Cl}_2+\text{Cl}_2$
 45 hrs. after fill
 $R_{\text{film}} = 93 \text{ ohms}, 13 \times 10^7 \text{ ohm cm}$
 $C_{\text{film}} = 1.45 \mu \text{ F}$
 Thickness = $705 \text{ \AA}$

(2059)

possibly indicating metal atom doping of the film. Further complex plane analysis over a longer time span will be required to determine if these differences persist. In any event the high frequency intercept is not changing in agreement with the lower corrosion rate when compared to earlier work.

III. Cathode Studies

Under the heading of cathode studies three areas of interest were investigated: (i) electrolyte additives, (ii) low temperature performance, (iii) cathode composition.

The test vehicle for these studies consisted of 2 inch hermetically sealed cells with both electrodes floating as shown in Figure 17. The anodes were composed of 20 mil Li pressed on 5 mil nickel mesh; the cathodes were composed of approximately 0.35 g of a Shawinigan Black (SB)/PTFE mixture pressed on 11 mil Ni mesh to a uniform thickness of 0.020 in. Five mil Manninglas was used as the separator. Double distilled SO_2Cl_2 was used in the formulation of all electrolytes.

III.1 Electrolyte Additives

The addition of LiAlCl_3Br to $\text{SO}_2\text{Cl}_2/\text{LiAlCl}_4$ electrolyte was shown in the last quarter to out perform all other electrolyte tested at 10 mA/cm^2 . Cells using the bromide additive were compared to cells using $\text{SO}_2\text{Cl}_2/1.5 \text{ M LiAlCl}_4$ (standard electrolyte); the bromide cells discharged with higher capacities and at higher voltages. A comparison between cells using the standard electrolyte plus Cl_2 and cells using the bromide additive was also made, the results again demonstrated the bromide additives beneficial effects. This quarter additional runs comparing cells w/ Cl_2 added vs. cells w/'Br' added were made at 20 mA/cm^2 and 30 mA/cm^2 .

Eight 2 inch cells as described above were fabricated using 90% SB/10% PTFE as the cathode mix. The electrolytes used were $\text{SO}_2\text{Cl}_2/1.0 \text{ M LiAlCl}_4/0.5 \text{ M LiAlCl}_3\text{Br}$ and $\text{SO}_2\text{Cl}_2/1.5 \text{ M LiAlCl}_4/0.5 \text{ M Cl}_2$. Four cells were discharged at 20 mA/cm^2 , 3 cells at 30 mA/cm^2 (one cell leaked upon filling). The results are seen in Table 2 and Figure 18.

From the table and figure the Cl_2 cells appear to deliver at 20 mA/cm^2 a higher capacity. At 30 mA/cm^2 the difference may be attributable to scatter. The bromide additive cells still operate at a higher voltage, but

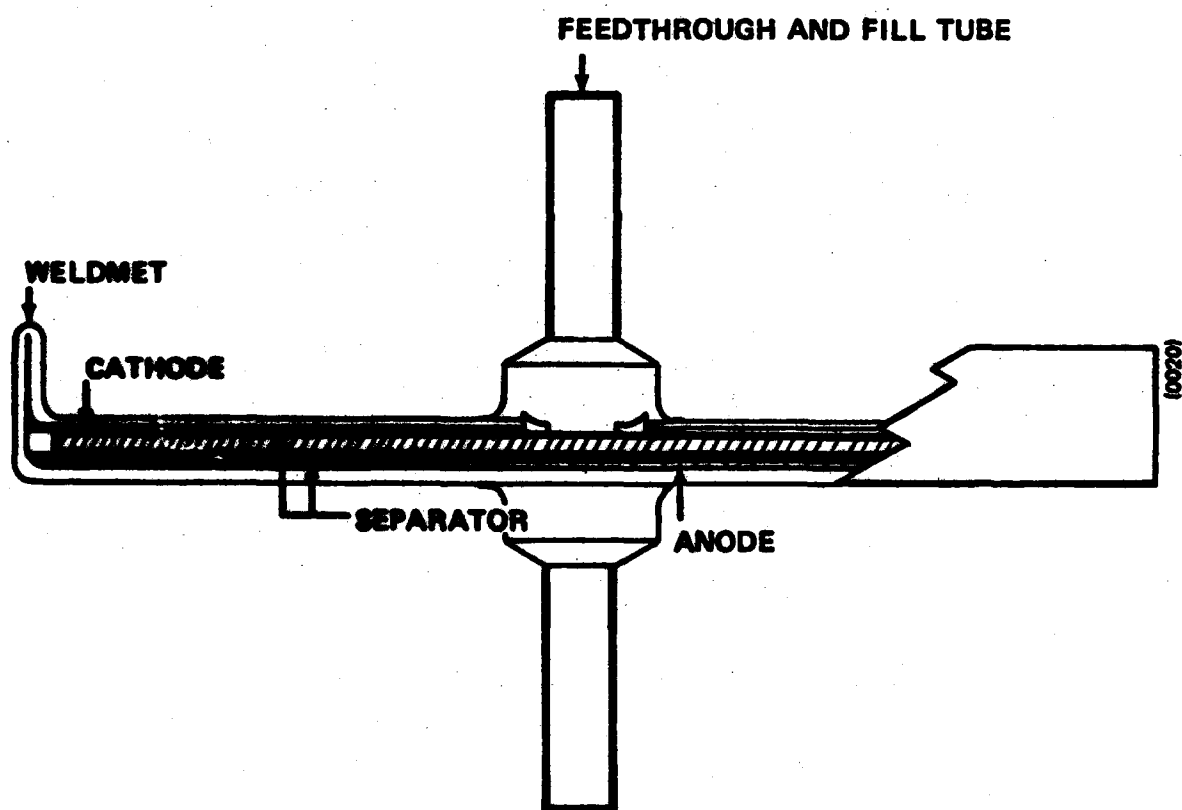


Figure 17 Schematic Cross Section and Description of 50mm diameter Cell used in $\text{Li}/\text{SO}_2\text{Cl}_2$ Studies

CATHODE: 90/10 Shawinigan Black/PTFE 0.35g pressed onto expanded Ni screen

ANODE: 20mil Lithium foil pressed onto expanded Ni screen

SEPARATOR: 5 mil Manniinglas

INSULATOR: 10 mil Telfon sheet

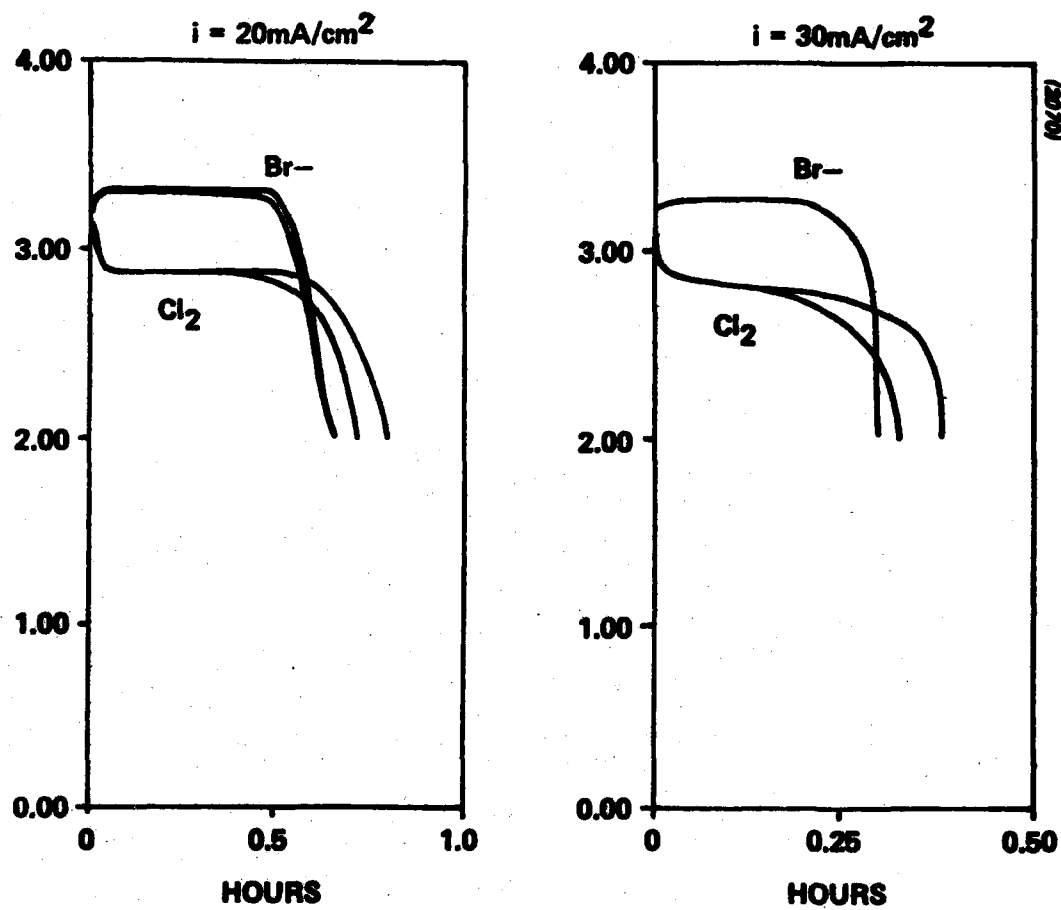


Figure 18 Discharge Curves of $\text{Li}/\text{SO}_2\text{Cl}_2$ Test Cells
Run at 20 and 30 mA/cm^2 with LiAlCl_4Br Added and Cl_2 Added

TABLE 2

Capacities (in Ah) of Li/SO₂Cl₂ Cells
 at 20 mA/cm² and 30 mA/cm² for Br₂ and Cl₂ Additives
 At Room Temperatures

i	SO ₂ Cl ₂	SO ₂ Cl ₂
	1.0 M LiAlCl ₄ /0.5 M LiAlCl ₂ Br ₂	1.5 M LiAlCl ₄ /0.5 M Cl ₂
20 mA/cm ²	0.198	0.240
	0.198	0.216
30 mA/cm ²	0.139	0.144
		0.176

the higher Cl_2 capacities are contrary to the trends observed at 10 mA/cm^2 . At 10 mA/cm^2 the Cl_2 cells behaved identically to the cells using the standard electrolyte (lower capacity than Br_2 cells) except for the higher initial voltage. From this observation one might expect comparable results at 20 and 30 mA/cm^2 . This is not the case as seen in Table 2. Apparently the addition of Cl_2 facilitates more complete utilization of the cathode at 20 mA/cm^2 than the addition of bromide. How Cl_2 more efficiently utilizes the cathode is not known but possibly it may affect the solubility of the reaction products or cause the discharge to proceed via a more efficient mechanism.

A comparison of Cl_2 cells vs Br_2 cells at additional current densities is planned for the coming months.

III.2 Low Temperature Testing

Most of the work presented to date was from discharge data accumulated at 25°C . Cells were stored at 60°C and 40°C , but were discharged after they cooled to room temperature. All of our 25°C discharge data show the superiority of the Br_2 additive compared to Cl_2 or no additive at all.

This quarter discharge data were obtained at -20°C to investigate what effect low temperatures have on cell performance.

Four kinds of electrolyte were investigated in this experiment

$\text{SO}_2\text{Cl}_2/1.5 \text{ M LiAlCl}_4$
 $\text{SO}_2\text{Cl}_2/1.5 \text{ M LiAlCl}_4/0.5 \text{ M Cl}_2$
 $\text{SO}_2\text{Cl}_2/1.0 \text{ M LiAlCl}_4/0.5 \text{ M LiAlCl}_3\text{Br}$
 $\text{SO}_2/\text{Cl}_2/1.5 \text{ M LiAlCl}_3\text{Br}$

The lithium salts were obtained from a commercial laboratory, the Cl_2 was added employing a double gas scrubbing apparatus. Cells were discharged at 10 mA/cm^2 at -20°C and at 25°C for comparison. Since we have no way in this experiment to determine the day to day variability, only comparisons of cells fabricated and run on the same day will be made.

From the results seen in Table 3 one startling result is apparent, at -20°C the bromide additive cells have lower capacities than cells using standard electrolyte. Also since there does appear to be no significant differences between Cl_2 added cells and the standard cells at -20°C , we can imply that the Cl_2 additive cells also have better capacities than the bromide additive cells at -20°C . This trend is opposite to the results evident at 25°C in this experiment and what we have observed in past experiments at 25°C .

The reactions involving bromine additive are more sensitive to low temperature than the reactions taking place in the cell in the presence of Cl_2 , or in the absence of any additives.

In addition two cells were run using $\text{SO}_2\text{Cl}_2/1.5 \text{ M LiAlCl}_3\text{Br}$. The results seen in Table 3 do not appear promising as the capacities in both cells (one at 25°C and one at -20°C) are below normal for this cell chemistry. The operating voltages though in these cells were 100 mV higher than cells using a $\text{LiAlCl}_4/\text{LiAlCl}_3\text{Br}$ mixture and 500 mV higher than standard electrolyte or Cl_2 added electrolyte cells. More testing on the optimum LiAlCl_3Br concentration is planned.

Additional testing comparing these four electrolytes at -20°C is planned, as well as testing at -40°C . If cells using the Br_2 additive run at low temperatures continue to demonstrate poor performance than Cl_2 added and standard electrolyte cells, this additive may prove unacceptable in batteries performing in this environment.

III.3 Cathode Composition Study

Most of the data obtained to date under this contract was from cells using 90% Shawinigan Black/10% PTFE binder as cathode material. This study was devised to compare the 90/10 cathode to the 95/5 cathode. A 2^3 factorial design experiment was developed for this study. The two additional factors

TABLE 3

Capacities (in Ah) of Cells Run
with various Electrolytes at 25°C and -20°C

Run	Temp.	Standard Electrolyte	Cl ₂ Added	Bromide Additive (0.5)	Bromide Additive (1.5)
1	25°C	0.360	0.390 0.447		
2	25°C	0.303		0.441 0.348	0.285
3	-20°C	0.422		0.270 0.315	0.300
4	-20°C	0.353 0.294	0.316 0.341		
5	-20°C	0.388	0.347 0.347		
6	-20°C	0.316 0.338		0.244 0.264	

are current density and the bromide additive. The experimental design and the data obtained are as follows:

l	-	-	-	A.	Cathode composition	+ 95% SB/5% PTFE
a	+	-	-			- 90% SB/10% PTFE
b	-	-	-	B.	Additive	+ w/B ₂ ⁻
ab	+	+	-			- w/o B ₂ ⁻
c	-	-	+	C.	Current density	+ 2 mA/cm ²
ac	+	-	+			- 10 mA/cm ²
abc	+	+	+			

Results

	Capacity (Ah)	Total Watt-Hours
l	0.447	1.36
a	0.804	2.45
b	0.540	1.86
ab	0.930	3.21
c	0.789	2.60
ac	1.010	3.33
bc	0.710	2.56
abc	0.908	3.27

No statistical significance can be assigned to any of the factors until a replicate is performed, but the performance of cells with the 95/5 cathode is much better than those with the 90/10 cathode as shown in the above data. Wherever the 95/5 cathode was used (a plus (+) in column A) the capacity and the total watt-hours are at least 20% higher than its counterpart using

90/10. This data is contrary to previous work using sealed jar cells, in those experiments 90/10 was shown to be better. The other two factors exhibit predictable results. Lower current density = higher capacity; presence of additive = higher capacities.

Additional runs of this experiment are planned to determine if this initial trend in cathode composition is statistically significant.

IV. Future Work

Now that we have developed a long life cell design with at least improved anode stability we plan to focus our attention in the fourth contract quarter on

- Investigation of means of improving anode stability and minimizing voltage delay. Three possible means of doing this are cyanoacrylate coating, vinyl chloride coating and prepassivation of the anode in pure thionyl chloride.
- Reinvestigate design variables which may effect anode stability. Some work in this area was carried out in the second quarter on cells with a solder joint between the case and feedthrough. The instability of this joint requires further study of the proper choice of grounding. In addition the use of a PTFE insulator between the anode and case may not be advisable as these two components react.
- Ampule studies of lithium foils in sulfuryl chloride electrolytes doped with possible cell contaminants such as water and heavy metal ions. The great difference between cell and ampule results to date strongly indicates a contaminant problem. Ampule studies with known contaminant levels will provide an efficient means of identifying which species must be excluded from the cell.

Most of our attention will be directed at resolving the anode problem. We do plan, however, to continue parametric cathode optimization and initiate limited performance characterization in simulated Army use such as a laser designator battery discharge regime.

Finally we will carry out spectroelectrochemical studies of $\text{Li}/\text{SO}_2\text{Cl}_2$ speciation during discharge and reversal. The equipment and techniques for this have already been perfected under a Navy contract at Gould for study of Li/SOCl_2 reactions.

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Newport Laboratory
Newport, RI 02840 (1)

Dr. Leonard Nanis

G207

S.R.I.

Menlo Park, CA 94025

(1)

Dr. Stuart Fordyce

NASA Lewis Research Center

Mail Stop 6-1

21000 Brookpark Road

Cleveland, OH 44135

(1)

Dr. J.J. Auburn, Rm 1A-317

Bell Laboratories

600 Mountain Avenue

Murray Hill, NJ 07974

(1)

Mr. Myles Walsh

ECO Incorporated

P.O. Box 578

Buzzards Bay, MA 02532

(1)

Mr. Harvey Frank

Mail Stop 198-220

Jet Propulsion Laboratory

4800 Oak Grove Drive

Pasadena, CA 91103

(1)

S.B. Brummer

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Newton, MA 02158

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Naval Surface Weapons Center

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(1)

Dr. Frank Bis

Naval Surface Weapons Center

White Oak Laboratory, Code R-33

Silver Spring, MD 20910

(1)

Mr. Eddie T. Seo
Research and Development Division
The Gates Rubber Co.
999 S. Broadway
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(1)

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ONR Chemical Program
Arlington, VA 22217

(1)

Mr. Sidney Gross
Mail Stop 8C-62
Boeing Aerospace Company
P.O. Box 3999
Seattle, WA 98124

(1)

Mr. Lou Adams
Saft Corporation of America
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Valdosta, GA 31601

Dr. H.V. Venkatesetty
Honeywell Technology Center
10701 Lyndale Avenue, South
Bloomington, MN 55420

(1)

Dr. C. Liang
Electrochem Industries (E-I) Inc.
9990 Wehrle Drive
Clarence, NY 14031

(1)

Mr. Aiji Uchiyama
Jet Propulsion Laboratory-M.S. 198-220
4800 Oak Grove Drive
Pasadena, CA 91103

(1)

Dr. C. Schlaikjer
Duracell International
Northwest Industrial Park
Burlington, MA 01803

(1)

Dr. Harvey N. Seiger
Harvey N. Seiger Associates
8 Beacon Hill Drive
Waterford, CT 06385

(1)

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TRW/DSSG
Ballistic Missile Division
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San Bernardino, CA 92402 (1)

Dr. S. P. Wolsky
Ansum Enterprises, Inc.
153 Horseshoe Lane
Centerville, MA 02634 (1)

Dr. John C. Hall
Altus Corporation
1610 Crane Ct.
San Jose, CA 95112 (1)

Dr. Faruq Marikar
Calenese Research Company
86 Morris Avenue
Summit, New Jersey 07901 (1)

K. M. Abraham
EIC, Inc.
Newton, MA 02158 (1)