

AD-A121 612

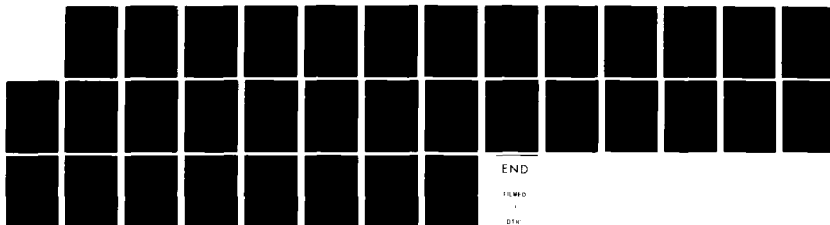
ROLE OF AMINES IN ADHESION OF POLYBUTADIENE TO GLASS
SUBSTRATES II REACTI. (U) AKRON UNIV OH INST OF POLYMER
SCIENCE Y ECKSTEIN ET AL. OCT 82 TR-23
N00014-76-C-0408

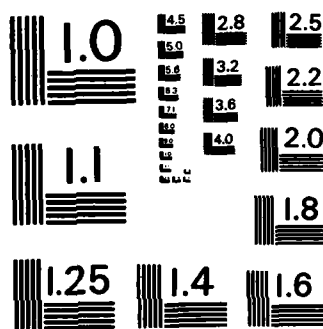
1/1

UNCLASSIFIED

F/G 9/1

NL





MICROCOPY RESOLUTION TEST CHART
NATIONAL BUREAU OF STANDARDS-1963-A

12

SECURITY CLASSIFICATION OF THIS PAGE (When Data Entered)

AD A121612

REPORT DOCUMENTATION PAGE		READ INSTRUCTIONS BEFORE COMPLETING FORM
1. REPORT NUMBER Technical Report No. 23	2. GOVT ACCESSION NO. AD-A121612	3. RECIPIENT'S CATALOG NUMBER
4. TITLE (and Subtitle) Role of Amines in Adhesion of Polybutadiene to Glass Substrates. II. Reactions of Amines with Triethylsilanol and/or Fumed Silica		5. TYPE OF REPORT & PERIOD COVERED Technical Report
7. AUTHOR(s) Y. Eckstein and P. Dreyfuss		6. PERFORMING ORG. REPORT NUMBER
9. PERFORMING ORGANIZATION NAME AND ADDRESS Institute of Polymer Science The University of Akron Akron, Ohio 44325		8. CONTRACT OR GRANT NUMBER(s) N00014-76-C-0408
11. CONTROLLING OFFICE NAME AND ADDRESS Office of Naval Research, Power Program Arlington, VA 22217		10. PROGRAM ELEMENT, PROJECT, TASK AREA & WORK UNIT NUMBERS NR 092-555
14. MONITORING AGENCY NAME & ADDRESS (if different from Controlling Office)		12. REPORT DATE October, 1982
		13. NUMBER OF PAGES 32
		15. SECURITY CLASS. (of this report) Unclassified
		16a. DECLASSIFICATION/DOWNGRADING SCHEDULE
16. DISTRIBUTION STATEMENT (of this Report) According to attached distribution list. Approved for public release; distribution unrestricted		
17. DISTRIBUTION STATEMENT (of the abstract entered in Block 20, if different from Report)		
18. SUPPLEMENTARY NOTES Accepted for publication in J. Adhesion		
19. KEY WORDS (Continue on reverse side if necessary and identify by block number) amines, silanol, fumed silica, structure of amine, carbamate adhesion, polybutadiene, glass		
20. ABSTRACT (Continue on reverse side if necessary and identify by block number) → This paper is part of a larger investigation of the role of amines in the adhesion of polybutadiene to glass substrates. It describes near infrared, infrared and nuclear magnetic resonance spectroscopy studies of the interaction of amines with silanol groups usually under ambient conditions but sometimes when heated in air. Additional supporting evidence was obtained from gas liquid chromatography, mass spectrometry and elemental analyses. ← in order to assure a sufficiently high concentration of silanol groups, triethylsilanol		

DTIC
NOV 19 1982

DTIC FILE COPY

OFFICE OF NAVAL RESEARCH

Contract N00014-76-C-0408

Project NR 092-555

Technical Report No. 23

ROLE OF AMINES IN ADHESION OF POLYBUTADIENE TO GLASS SUBSTRATES.

II. REACTIONS OF AMINES WITH TRIETHYLSILANOL AND/OR FUMED SILICA

by

Y. Eckstein and P. Dreyfuss

Institute of Polymer Science
The University of Akron
Akron, Ohio 44325

October, 1982



A

Reproduction in whole or in part is permitted for any purpose of
the United States Government

Approved for Public Release; Distribution Unrestricted

INTRODUCTION

The goal of this study is to define the role of amines in general, and of 3-aminopropyltriethoxysilane in particular, in the mechanism of adhesion of elastomers to inorganic substrates. The first paper in the series described reactions of amines during peroxide curing of polybutadiene and the effect of low concentrations of amines (comparable to the peroxide concentration) on mechanical properties of the elastomer (1). When aromatic amines or piperazine were added to the elastomer, maximum stress at break and minimum swelling ratios were observed when the number of alkoxide radicals generated was equal to the number of hydrogen atoms on the nitrogen of the amines. The greatest effect of aliphatic amines also occurred when the concentration of primary amino groups was equal to the concentration of peroxide; but the behavior was more complex than that of the aromatic amines and piperazine. These effects were attributed to the formation of chemical bonds between the amines and the polybutadiene. This paper examines model reactions aimed at better defining reactions that occur between amines or aminosilanes and a glass surface.

It is generally agreed that chemical bonds, in the form of a polysiloxane coating, form when silanol groups in a glass surface react with a triethoxysilane under appropriate conditions. However, reactions between amines and the inorganic substrate are also possible. A 1971 infrared study by Kaas and Kardos (2) indicated not only that on treatment with 3-aminopropyltrimethoxysilane the silanol groups on fused silica undergo condensation as expected but

also that hydrogen bonds are formed between the amino groups of the silane compound and the surface hydroxyls. Protonated amine groups were also suggested in a more recent study of the structure of films formed when 3-aminopropyltrimethoxysilane is absorbed into alumina (3) or on the surface of sapphire crystals (4). Hence a study of reactions of amines with triethylsilanol and/or fused silica was included in this project. The present study supports the earlier conclusions that reactions do occur and shows that the mechanism of interaction between amines and silanol groups is strongly influenced by the structure of the amine; in particular, aliphatic amines behave quite differently from aromatic amines. Absorption of carbon dioxide by aliphatic amines appears to be catalyzed by silanols and the adducts formed are more reactive than the uncomplexed amines.

EXPERIMENTAL

Materials

Many of the materials used were described in Part I of this series (1). Cab-O-Sil[®], Grade M5 Fumed Silica from Cabot Corporation, was used as received. Triethylsilanol and the Cab-O-Sil were used as model compounds of a glass surface. Additional silanes included 3-aminopropyldiethoxytethoxysilane, $\text{NH}_2(\text{CH}_2)_3\text{Si}(\text{OEt})\text{Me}_2$, and 3-aminopropyldiethylmethoxysilane, $\text{NH}_2(\text{CH}_2)_3\text{SiMeEt}_2$, from Petrarch Systems, Inc. These silanes were selected for study because they are not capable of undergoing the polysiloxane forming condensation reaction that occurs readily with 3-aminopropyltriethoxysilane,

$\text{NH}_2(\text{CH}_2)_3\text{Si}(\text{OEt})_3$. *n*-Propylamine from Eastman was distilled before use. 1,12-Dodecane-diamine, 97%, (M.P. 62-65°C) and pyrrolidine, 99%, (b.p. 87-88°C), both from Aldrich Chemical Co. were used as received. Carbon tetrachloride, Spectroanalyzed grade, and picric acid, Reagent ACS grade, from Fisher Scientific Co. were used as received. CO was used as powdered dry ice. Deuterium oxide, 99.7% minimum isotopic purity, was obtained from Merck-Sharp and Dohme of Canada Limited.

Preparation of Aliphatic Amine Precipitates

These salts were prepared by mixing equal volumes of the amine and silanol(S) and heating gently at 50°C. After an hour or so a white precipitate formed. The solid was filtered off, washed with acetone to remove any remaining amine or silanol, and dried. Alternatively, the precipitates were prepared by mixing the amine with Cab-O-Sil and heating in an open vessel covered with a Teflon sheet.

Usually within 18-48 hrs. significant quantities of the amine salt collected on the sheet and solidified on cooling. It was collected, washed and dried.

Preparation of Adducts of Amines and Carbon Dioxide

Powdered excess dry ice was added to neat liquid amine in a test tube and stirred with a spatula to achieve a uniform mixture. Heat was evolved and a white cloud formed above the amine. The liquid quickly became a uniform white powder. The powder was dried on a filter paper, usually in air but sometimes in a vacuum oven. The method of drying did not seem to affect the properties of the product.

Preparation of Picrates

Picrates were formed from the various compounds following the procedure given in Shriner and Fuson (5).

Near-Infrared Studies

Near-infrared spectra of neat compounds, of compounds diluted with CCl_4 or of reaction products in CCl_4 were recorded using a Cary 17 Spectrophotometer. Reaction mixtures were prepared at room temperature by adding the required amount of amine or amine containing silane and silanol to CCl_4 in a volumetric flask. Usually equimolar quantities of each reagent at a concentration of $\sim 10^{-3}$ mol per 10 ml CCl_4 (0.1M) were used and the spectra were recorded within 2-6 hrs. of mixing the reagents. On standing the amines react with CCl_4 and dark brown solutions are formed (6,7).

Infrared Studies

Infrared spectra were recorded on a Perkin Elmer 521 Grating Infrared Spectrophotometer or on a Perkin Elmer 337 Grating Infrared Spectrometer. The state of the material being examined determined whether a silver chloride cell, a gas cell with 10 cm path length and sodium chloride windows, a KBr pellet, or a cell consisting of sodium chloride windows alone was selected as sample holder.

NMR Studies

^1H NMR spectra of neat compounds or of reaction products in CCl_4 or D_2O were obtained with the aid of a Varian T-60 NMR Spectrometer. ^{13}C NMR spectra were recorded on a Varian Associates CFT-20. The instrumental conditions were $\text{PW} = 19 \mu \text{ sec}$, $\text{SW} = 4000 \text{ Hz}$, $\text{PD} = 1.0 \text{ sec}$, $\text{AT} = 0.5 \text{ sec}$, $\text{T} = +27^\circ\text{C}$.

Other Studies

Gas liquid chromatograms were obtained using a Hewlett Packard 5150 Chromatograph fitted with an SE 30 column, a flame ionization detector and a Perkin-Elmer Model 149B printing integrator. The injection port temperature was 220°C. The program used for separating components consisted of 4 min at 60°C followed by heating at 60°C/min. to 100°C and then 8 min. at 160°C. The mass spectrum of the solid formed by reacting triethylsilanol and 1,4-diaminobutane was recorded on a Finnegan Mass Spectrometer Model 1015S-L. Melting points were taken in capillary tubes in a MEL-TEMP Laboratory Device. Analyses were performed by Galbraith Laboratories, Inc.

RESULTS

Aromatic Amines

Phenylenediamine (PDA) was reacted with triethylsilanol (S) in CCl₄ and the reaction products were examined by near-infrared spectroscopy. PDA is solid at room temperature and is only slightly soluble in CCl₄. However, upon addition of the silanol, the amine dissolved and a small amount of water soluble white precipitate formed. Near-infrared spectra of the amine and of the silanol alone and of the mixture of PDA and S are presented in Figure 1. After addition of the S, all the spectral bands showed broadening; the -NH₂ band at about 1.48 μ m was shifted to longer wavelength; and large quantities of water (broad band at about 1.87 μ m) formed. Broadening of spectral bands suggests the presence of hydrogen bonds,

while the shift of the -NH_2 band to lower energies (longer wavelength) is characteristic of an NH_3R^+ type species.

Aliphatic Amines and Silanes Containing n-Propylamino Groups

Triethylsilanol (S) was reacted with n-propylamine (PA), 1,4-diaminobutane (AB), 3-aminopropyltriethoxysilane (AS), 3-aminopropyldiethylethoxysilane (AMEOS) and 3-aminopropyldiethylmethylsilane (AEMS). Additional experiments were carried out with Cab-O-Sil and some of the same aliphatic amine containing compounds.

Upon addition of S to PA, AB, AS, AMEOS or AEMS a white cloud appeared above the solution. The cloud gave a basic reaction with wet litmus paper and was similar in appearance to the cloud that forms from ammonia and hydrochloric acid. Heat was evolved in all cases and the reaction was more vigorous with the aliphatic amines than with the silanes. Attempts were made to isolate any gaseous products and the products isolated were studied by infrared, nuclear magnetic resonance and mass spectroscopy and by gas liquid chromatography. Indications from the various methods for the presence of materials other than starting materials were contradictory and inconclusive. At present, we believe that the cloud contains a mixture of products possibly including ammonia, cyclic compounds, amine and ethanol and certainly including carbon dioxide and an amine derivative. The exact composition of the cloud varies with the starting reagents. For example, Young et al. have shown that polymerized AS, and presumably by analogy the monomer since the monomer almost always contains some oligomers, is not completely stable and releases ammonia

at both room and elevated temperature (8); also we have found ethanol in the vapors above a mixture of AS and S.

The Solution. Figure 2 shows the near infrared spectra of neat AS and of solutions of AS + S and of AS + S in CCl_4 . The general features of the spectra are common to all the aliphatic amines and aminosilanes studied. As seen in the spectra, addition of the silanol resulted in broadening of all spectral bands and a slight shift to longer wavelengths of the NH_2 -group peak at about $1.525\ \mu\text{m}$ ($1525\ \text{nm}$). The doublet at about $2.0\ \mu\text{m}$ ($2000\ \text{nm}$) is also characteristic of the NH_2 -group, and its spectral distribution varies with the relative concentrations of AS and S in the reaction mixture. Dependence of the spectral bands on concentration and broadening of the spectral bands is characteristic of the presence of hydrogen bonds (9). The shift of the $1.525\ \mu\text{m}$ band to lower energies indicates the presence of protonated amine.

The intensities of peaks for the initial compounds and for the reaction products are summarized in Table IV. The intensity, I , was calculated from the relationship $I = A/C$, where A denotes the absorbance and C the molar concentration of the compounds responsible for the specific band. Upon addition of the silanol to any of the aminosilanes or amines a small decrease in intensity of the NH - peak at about $1.525\ \mu\text{m}$ and a larger decrease in the intensity of the $-\text{OH}$ peak of the silanol at $1.383\ \mu\text{m}$ were observed. If the aminosilane contained ethoxy groups, ethanol, with a peak maximum at $1.408\ \mu\text{m}$, was among the reaction products. The magnitude

of the decrease of the intensities of the 1.525 μm and 1.383 μm peaks seems to be dependent on the initial concentration of the silanol. Usually, small amounts of H_2O are also found among the reaction products.

Heating the reaction mixture of AS and S to 55°C and 95°C caused a continuous increase with temperature in the height of the ethanol peak at 1.408 μm and a continuous decrease with temperature in the height of the silanol peak at 1.383 μm . The NH_2 -peak at 1.525 μm decreased only slightly. (See Figure 3.) These results suggest that heat is driving the following reaction between AS and S to the right:



Alternatively, a mixture of symmetrical and unsymmetrical siloxanes and polysiloxanes may form. Gas liquid chromatography of the reaction mixture of AS and S, measured 24 hrs after mixing, showed the disappearance of a peak at a retention time of 10.2 min., characteristic of AS. A large peak at long retention time, ~24 min., appeared. This is consistent with formation of the condensation product indicated in equation (1).

Precipitates. Addition of S to AB resulted also in the formation of a white precipitate, which had the characteristics of a salt; namely, the precipitate dissolved in water and did not dissolve in most organic liquids. Further investigation showed that a similar precipitate could be obtained on gently heating a mixture of AB and Cab-O-Sil. Liquid droplets collected on the

Teflon sheet covering the vessel containing the mixture and solidified on cooling. The evidence to be presented below indicates that the two solids were identical. Similar solids were collected above mixtures of Cab-O-Sil with 1,12-diaminododecane or with AS (a liquid). Unlike the AB salt, the products from 1,12-diaminododecane and AS did not dissolve in water. Therefore, further efforts to characterize the salts formed when amines interacted with -SiOH groups were concentrated on the products from AB.

The near infrared and ^1H NMR spectra of the precipitate from S and AB indicated that the compound did not contain any groups derived from the S. Other properties are compared in Table II with those of the starting material and some plausible products (10-14). The carbon dioxide adduct was selected for comparison because AB is known to absorb CO_2 (10,12,14), although properties of the adduct are not reported. Pyrrolidine, $\text{CH}_2\text{CH}_2\text{CH}_2\text{CH}_2\text{NH}$, and its derivatives were selected for comparison because it has been reported that heating the dihydrochloride of AB below 150°C gives pyrrolidine (10,12). Since -SiOH groups are known to be acidic and the near infrared studies indicate that an ammonium type salt is formed when AB and S are mixed, a similar reaction could be possible. Moreover, pyrrolidine is formed during some methods of synthesis of AB (12).

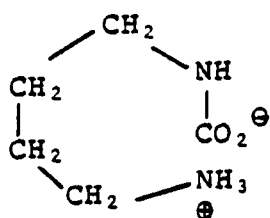
The infrared spectra of the salts obtained from AB in the presence of S or Cab-O-Sil were different from that of AB but closely resembled that of the carbon dioxide adduct of AB. (Compare Figures 4 and 5). The salient features of the IR spectra

of the salts consist of a singlet at 3325 cm^{-1} instead of the doublet in the free amine, a broad strong "ammonium band" (15) in the range 3200 to 1800 cm^{-1} , and a weak amide-like shoulder at 1620 cm^{-1} . (See Figure 4). As can be seen in Figure 5, the infrared spectra of the solids obtained from 1,12-diaminododecane and AS showed similar features. Also, as with the AB salt (Table II) the melting temperatures of these other solids were substantially higher than those of the starting amines, 140°C compared to 64°C for the diamine and $\sim 395^{\circ}\text{C}$ for the AS. The melting behavior of the solids isolated from the AS and Cab-O-Sil mixture was the same as that of a carbon dioxide adduct prepared from the same silane. A solid was not isolated from the Teflon sheet above the PA and Cab-O-Sil mixture but it was demonstrated that PA also forms an adduct with carbon dioxide (M.P. = $70-72^{\circ}\text{C}$). The adduct from PA is soluble in water.

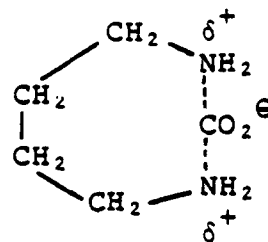
The carbonyl peak in the infrared spectrum of the salt from AB and S or Cab-O-Sil was not well resolved but in the ^{13}C NMR spectrum shown in Figure 6, the carbonyl peak was clearly visible. The ^{13}C NMR spectrum, in fact, revealed 2 carbonyl peaks, 4 kinds of C-N bonds and 4 kinds of C-C bonds. Comparison with spectra of the authentic compounds allowed one of the carbonyls, one C-N bond and one C-C bond to be attributed to the carbon dioxide adduct of AB. Another pair of bonds, one C-C and one C-N could be associated with AB itself. The carbon dioxide salt clearly is $\sim 67\%$ dissociated in D_2O . The other peaks could arise from a different

amine, whose structure is unknown at the present time. Those peaks definitely cannot be attributed to pyrrolidine, a likely side product (12), since they are at different fields from those that arise from the authentic compounds. Whatever the unknown structure, it arises in the course of preparation or isolation of the salt, because the same sample of diamine was used to prepare the CO_2 adduct used for comparison.

From all the above data we conclude that the salt isolated from AB plus S or AB plus Cab-O-Sil was a 1:1 carbon dioxide adduct of AB, probably having the structure:



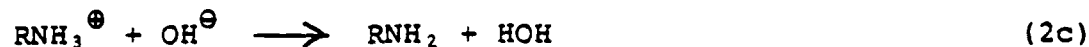
or



Fragments observed by mass spectrometry were in agreement with the proposed composition. The other solids are linear carbamates of the structure RNHCO_2H or $\text{RNHCO}_2^- \text{H}_3\text{O}^+$ (11).

It is noteworthy that the same AB salt forms when AB is heated in air or when AB stands in an open vessel in air, but its formation is much slower in the absence of $-\text{SiOH}$ groups or added dry ice.

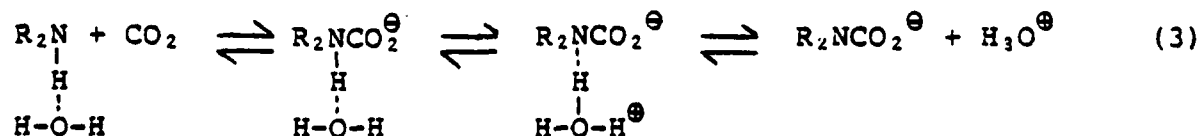
Then the $\text{OH}^\ominus$ reacts with the protonated amine to the extent needed to satisfy the equilibrium required by the concentration of ions in the solution:



These reactions probably occur with all amines but are dominant only with aromatic amines and certain cyclic amines like piperazine.

Aliphatic Amines

The reaction of aliphatic amines with $-\text{SiOH}$ groups is more complex than the corresponding reaction with aromatic amines. The aliphatic amines can undergo all the reactions already described for aromatic amines but probably do so to a lesser extent. Several features of the reaction are different: (1) a white cloud forms above the solution immediately upon mixing an aliphatic amine and a compound containing $-\text{SiOH}$ groups, (2) heat is evolved, (3) the $-\text{NH}_2$ group concentration definitely decreases while that of the aromatic amine hardly changes (4) only traces of water are present in the reaction products and (5) precipitates form. The main mechanism seems to proceed through an active intermediate formed at the earliest stage of the reaction. The evidence indicates that the active intermediate is an adduct of carbon dioxide and amine, a carbamate, probably formed as suggested in Caplow's mechanism (11):



An -SiOH group could facilitate formation of the required protonated amine in the first step but would not form part of the final salt because of subsequent siloxane formation as described above in the section on aromatic amines. It appears that -SiOH groups or at least the silanol or silica facilitate the formation of the CO_2 adduct of the amines also. It is possible they do so by formation of a reactive ammonium intermediate which dissolves readily in the silanol or complexes firmly with silica. It is also likely that the reaction proceeds so well because silicon compounds absorb large quantities of air including the CO_2 . (We have observed that it is almost impossible to keep a sample of silicones degassed. After thorough degassing and storage on a high vacuum line, the silicones become resaturated with air within seconds after opening to the atmosphere.) Thus the silicon compounds may merely be serving as a medium for bringing substantial amounts of amine and of CO_2 into favorable contact for reaction.

Carbamate formation is probably facilitated also by other fillers containing -OH groups. Moreover, if the relative ease of formation of picrates from the carbon dioxide adducts compared to their ease of formation from amines is a valid indication of the general reactivity of carbamates, the presence of the amines in the form of carbamates greatly enhances their reactivity in subsequent reactions.

Finally, it should be mentioned that although no evidence for the formation of pyrrolidine in reactions involving AB and S or Cab-O-Sil was found, the formation of pyrrolidine during curing at elevated temperatures such as 150°C cannot be definitely excluded. These model studies were carried out mostly at room temperature and were always below 100°C.

ACKNOWLEDGEMENT

This work forms part of a program of research on the adhesion of elastomers supported by a research grant from the Office of Naval Research. The authors are grateful to H. J. Harwood and M. P. Dreyfuss for helpful discussions regarding the salt from 1,4-butanediamine and to E. R. Santee, Jr. for taking the ^{13}C NMR spectra and assisting in the interpretation of all the NMR spectra.

REFERENCES

1. Y. Eckstein and P. Dreyfuss, J. Adhesion, 13, 303 (1982).
2. R. L. Kaas and J. L. Kardos, Polym. Eng. and Sci., 11, 11 (1971).
3. C. S. Park, S. H. Lee and N. H. Sung in "Polym. Sci. and Technol.", 12B, L. H. Lee, Ed., Plenum Press, N.Y., N.Y., 1980, p. 757.
4. F. J. Boerio, C. A. Gosselin and Y. W. Williams, Abstracts 4th Annual Meeting of the Adhesion Society, Savannah, Georgia, Feb. 23-25, 1981, p. 7.
5. R. L. Shriner and R. C. Fuson, "The Systematic Identification of Organic Compounds", John Wiley & Sons, Inc., N.Y. (1948), p. 180.

6. F. M. Unger, Dissertation Abstracts 29B, 2821-B (1969).
Ph.D. Thesis, Georgetown, University, 1968.
7. W. J. Lautenberger, Dissertation Abstracts 29B, 140-B (1968).
Ph.D. Thesis, University of Pennsylvania, 1967.
8. F. M. Young, T. G. Roskos and H. R. Radloff, SPI, 17th Conference, Section 14-C, 1 (1962).
9. R. T. Conley in "Infrared Spectroscopy", Allyn Bacon, Inc. Boston, 1966, p. 119.
10. Beilsteins Handbuch der Organischen Chemie, 4)1), 264.
11. M. Caplow, J. Am. Chem. Soc., 90, 6795 (1968).
12. Berichte der Deutschen Chemischen Gesellschaft, 19, 780 (1886),
and 20, 442 (1887).
13. The Merck Index, 9th Edition, M. Windholz, Ed., Merck & Co., Inc., Rahway, N.J., 1976, No. 7802.
14. Dictionary of Organic Compounds, 4th Edition, J. R. A. Pollock and R. Stevens, Eds., Oxford University Press, 1965, pp 2809, 2827.
15. K. Nakanshi and P. H. Solomon, "Infrared Absorption Spectroscopy", 2nd Edition, Holden-Day, Inc., San Francisco, 1977, pp 33-44.

TABLE I

Near-Infrared Spectra of Amines Alone
and of Amines plus Silanol Mixtures

reaction mixture ^(a)	Intensities			
	NH ₂ ~1.5 μ m	$\equiv$ SiOH 1.383 μ m	EtOH ^(b) 1.408 μ m	H ₂ O 1.87 μ m
S	-	0.97	-	-
PA	0.47	-	-	-
AB	1.06	-	-	-
AS	0.80	-	-	-
AMEOS	0.68	-	0.27	-
AEMS	0.44	-	-	-
PA + S	0.43	0.44	-	traces
AB + S	0.95	0.3	-	-
AS + S	0.69	0.62	0.21	traces
AMEOS + S	0.54	0.73	0.41	traces
AMEOS + 2S	0.49	0.575	0.15	0.06
AMEOS + 3S	0.41	0.53	0.1	0.03
AEMS + S	0.41	0.72	-	0.04

- a) S = Et₃SiOH, PA = CH₃CH₂CH₂NH₂, AB = NH₂(CH₂)₄NH₂,
 AS = NH₂(CH₂)₃Si(OEt)₃, AMEOS = NH₂(CH₂)₃Si(OEt)Me₂,
 AEMS = NH₂(CH₂)₃Si MeEt₂. The reaction mixture was prepared in
 CCl₄ from equimolar concentrations of the amine and of the silanol.
 2S or 3S indicate that the concentration of S was twice or three
 times that of AS.
- b) The intensities were calculated relative to the concentration of
 silanol.

TABLE II

Comparison of Properties of Salt of AB with Properties of Known Materials

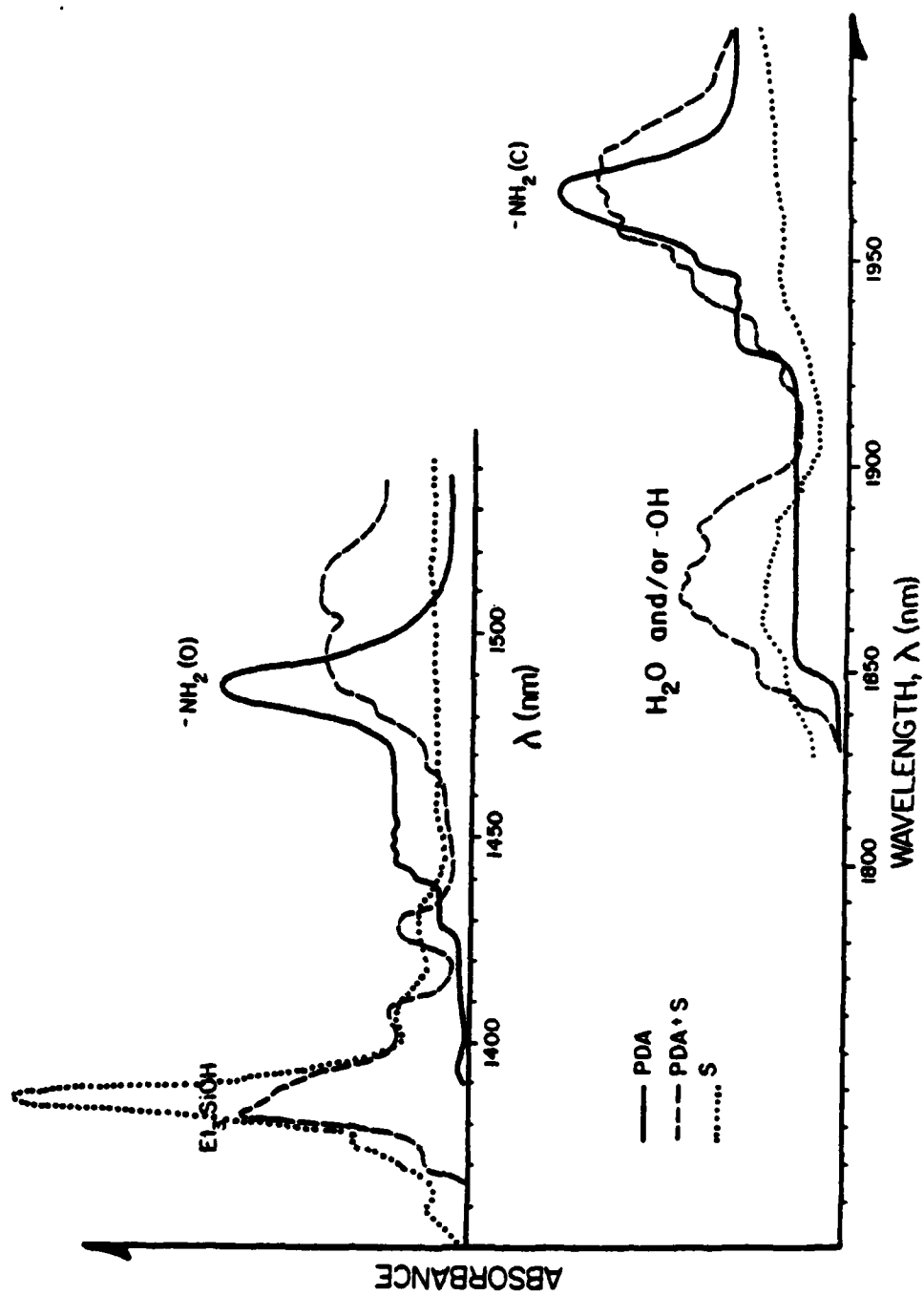
Property	Material			
	Salt	AB	AB·CO ₂	Pyrrolidine Pyrrolidine·CO ₂
M.P., °C	sublimes 167	27	sublimes 167	67.75
Picrate	Y,d257	Y,d265 ^b	Y, 256	Y,112 ^b
		Y,d263 ^c		Y,112 ^d
		Y, 250-55 ^d		
Elemental Analysis				
	45.84 ^b	54.50 ^e	45.44 ^e	67.55 ^e
%C				52.16 ^e
%H	9.61 ^b	13.72 ^e	9.15 ^e	12.75 ^e
%N	21.40 ^b	31.78 ^e	21.20 ^e	19.69 ^e
%O	23.01 ^b	-	24.21 ^e	12.17 ^e
%Si	0.38 ^b	-	-	27.79 ^e

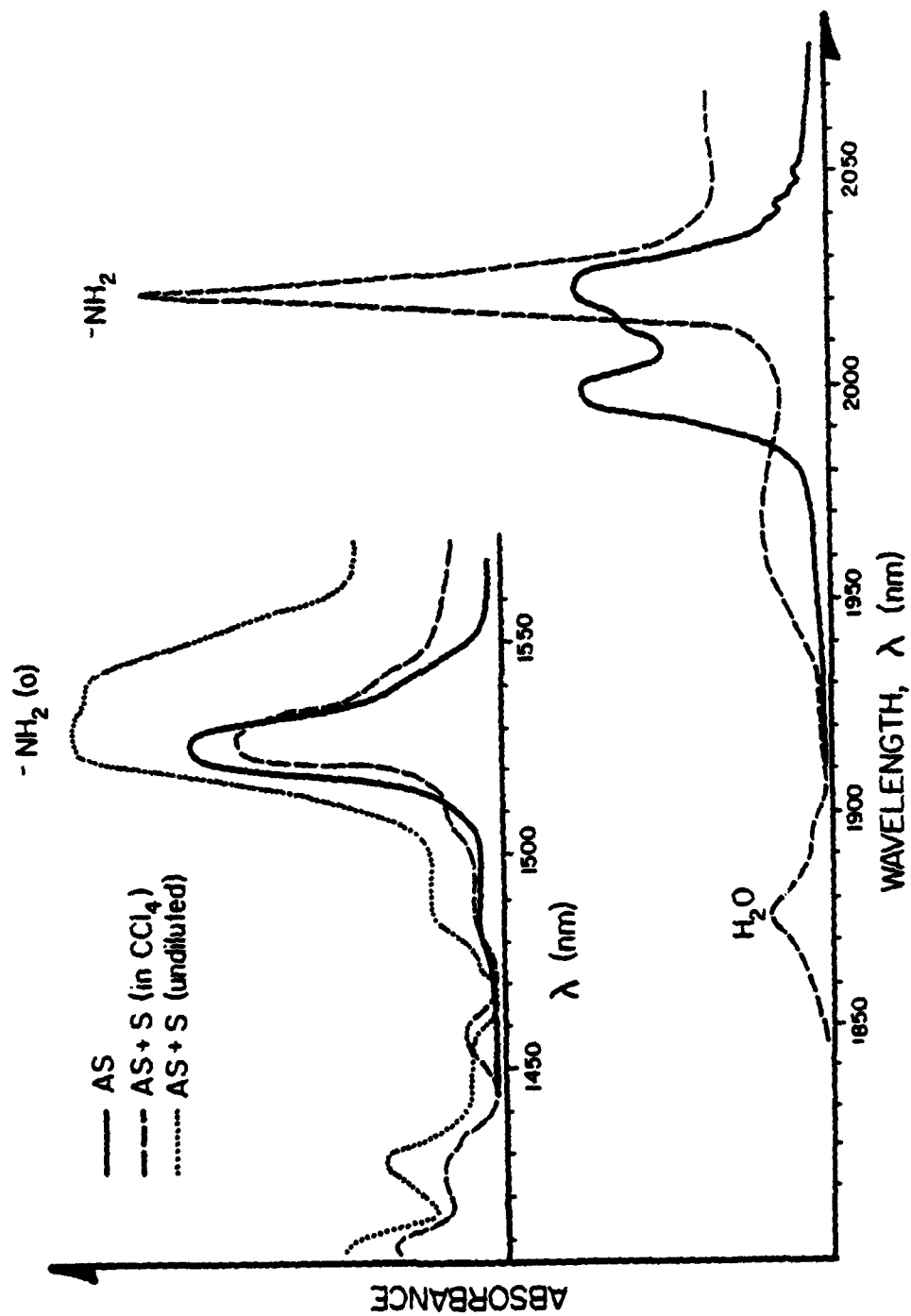
18

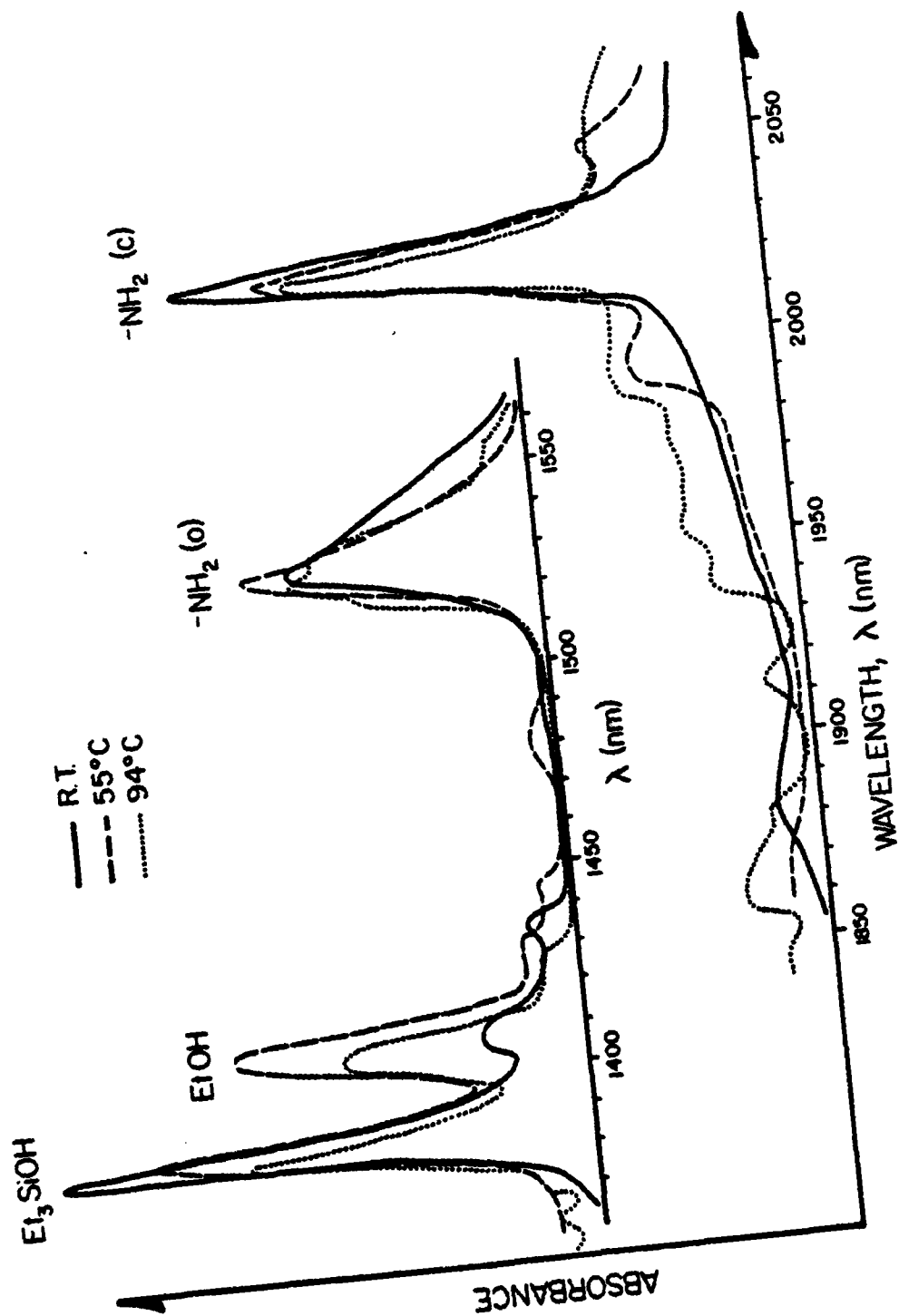
^a A liquid with B.P. = 87-88°C^b Observed^c Reference 7^d Reference 11^e Calculated for AB, AB·CO₂, Pyrrolidine, and Pyrrolidine·CO using the formulas C₄H₁₂N₂, C₅H₁₂N₂O₂, C₄H₉N and C₅H₉NO₂, respectively

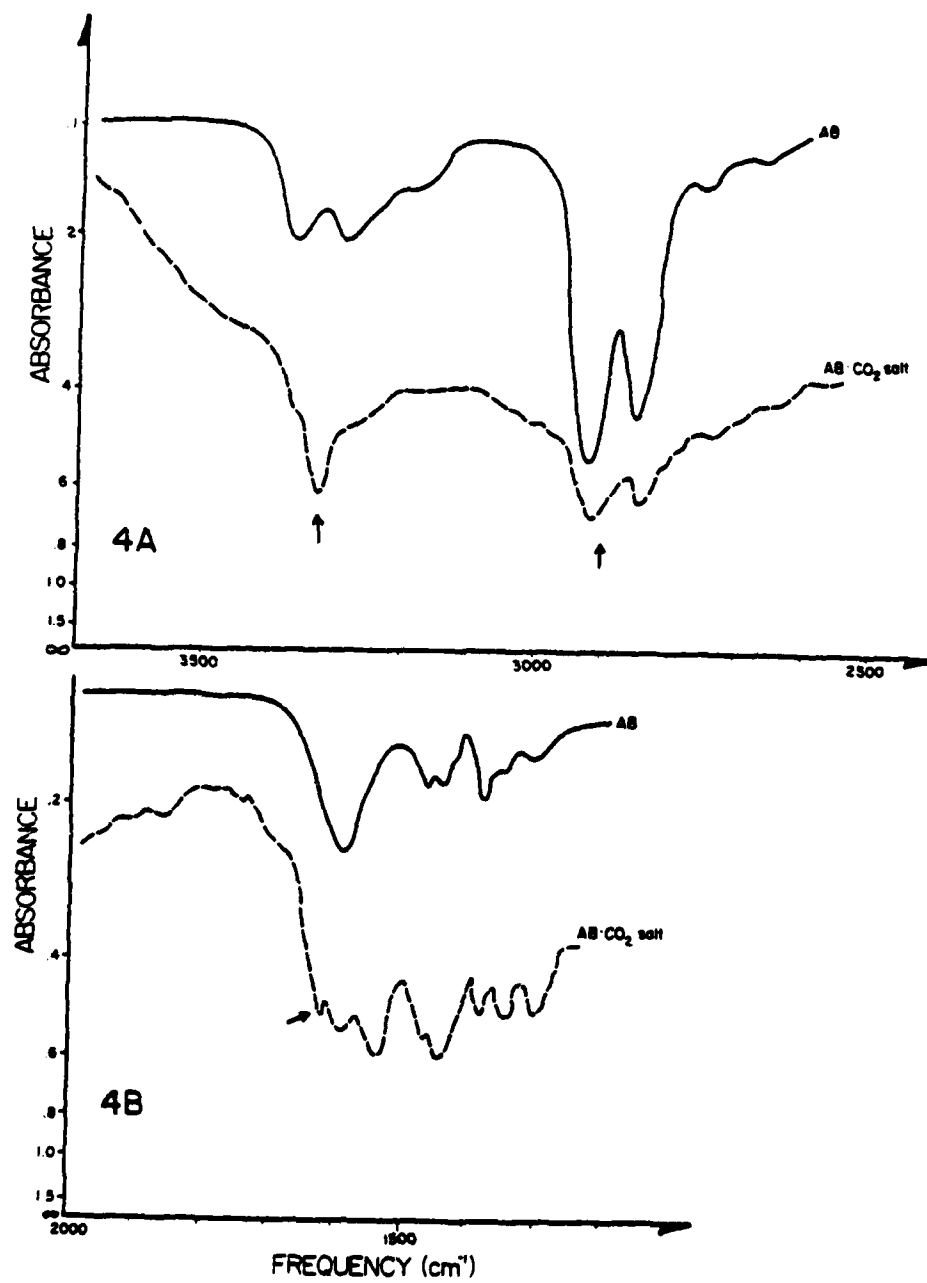
Figure Legends

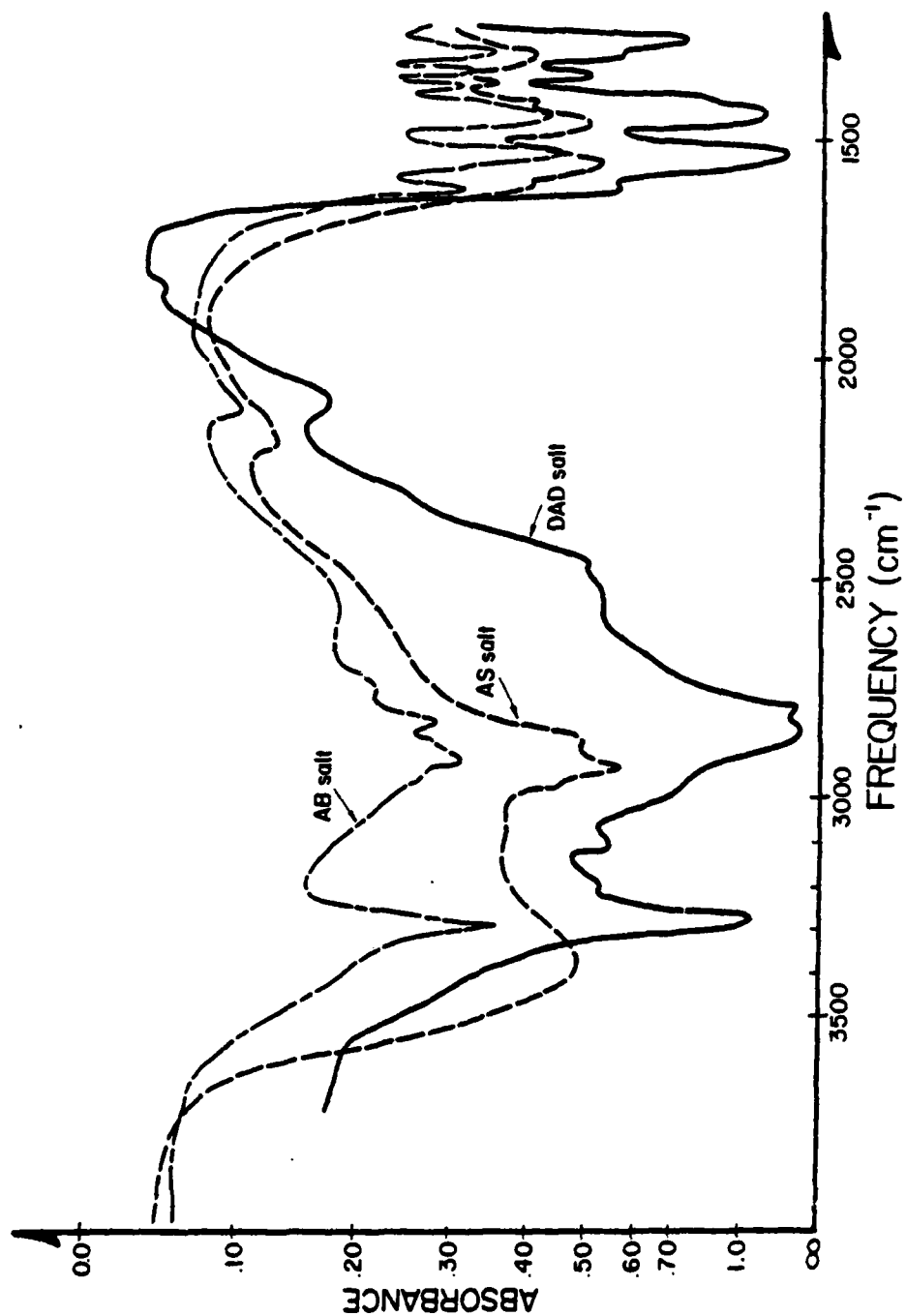
- Figure 1. Near infrared spectra of m-phenylenediamine (PDA) alone, of triethylsilanol(S) and of a mixture of PDA and S. All compounds were dissolved in CCl_4 .
- Figure 2. Near infrared spectra of neat 3-aminopropyltriethoxysilane (AS) and of mixtures of AS and triethoxysilanol(S) both neat and diluted with CCl_4 . The spectrum of AS + S undiluted is not shown in the 1900 to 2050 nm region because it absorbs very strongly in this region.
- Figure 3. Effect of heating on near infrared spectra of reaction mixtures of 3-aminopropyltriethoxysilane(AS) and triethoxysilanol(S).
- Figure 4. Two regions of the infrared spectra of 1,4-diaminobutane(AB) and of its adduct with carbon dioxide($\text{AB} \cdot \text{CO}_2$) taken neat between windows on the Perkin Elmer 521 Grating Infrared Spectrophotometer. Figure 4A shows the change from the doublet to the singlet on formation of the adduct as well as the broad "ammonium band". Figure 4B shows the carbonyl peak. The regions in question are marked with arrows.
- Figure 5. Comparison of infrared spectra of amine salts taken in KBr pellets on the Perkin Elmer 337 Grating Infrared Spectrometer. AB, AS, and DAD are 1,4-diaminobutane, 3-propylaminotriethoxysilane and 1,12-diaminododecane, respectively.
- Figure 6. ^{13}C NMR spectra of AB, $\text{AB} \cdot \text{CO}_2$ and the AB salt.
O = transmitter frequency. TMS = 3.5 external ppm.

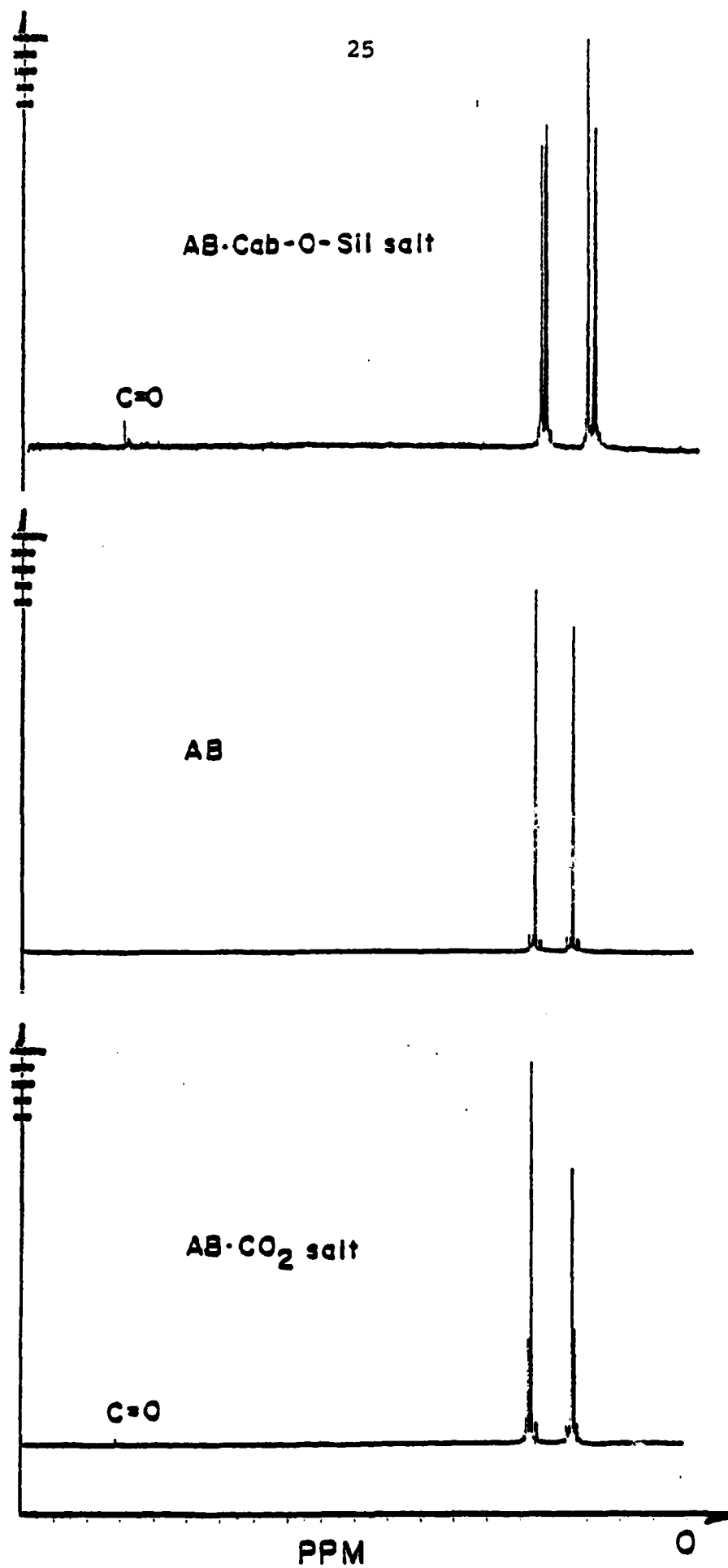












DISTRIBUTION LIST

	<u>No. Copies</u>		<u>No. Copies</u>
Dr. L.V. Schmidt Assistant Secretary of the Navy (R,E, and S) Room 5E 731 Pentagon Washington, D.C. 20350	1	Dr. F. Roberto Code AFRPL MKPA Edwards AFB, CA 93523	1
Dr. A.L. Slafkosky Scientific Advisor Commandant of the Marine Corps Code RD-1 Washington, D.C. 20380	1	Dr. L.H. Caveny Air Force Office of Scientific Research Directorate of Aerospace Sciences Bolling Air Force Base Washington, D.C. 20332	1
Dr. Richard S. Miller Office of Naval Research Code 413 Arlington, VA 22217	10	Mr. Donald L. Ball Air Force Office of Scientific Research Directorate of Chemical Sciences Bolling Air Force Base Washington, D.C. 20332	1
Mr. David Siegel Office of Naval Research Code 260 Arlington, VA 22217	1	Dr. John S. Wilkes, Jr. FJSRL/NC USAF Academy, CO 80840	1
Dr. R.J. Marcus Office of Naval Research Western Office 1030 East Green Street Pasadena, CA 91106	1	Dr. R.L. Lou Aerojet Strategic Propulsion Co. P.O. Box 15699C Sacramento, CA 95813	1
Dr. Larry Peebles Office of Naval Research East Central Regional Office 666 Summer Street, Bldg. 114-D Boston, MA 02210	1	Dr. V.J. Keenan Anal-Syn Lab Inc. P.O. Box 547 Paoli, PA 19301	1
Dr. Phillip A. Miller Office of Naval Research San Francisco Area Office One Hallidie Plaza, Suite 601 San Francisco, CA 94102	1	Dr. Philip Howe Army Ballistic Research Labs ARRADCOM Code ORDAR-BLT Aberdeen Proving Ground, MD 21005	1
Mr. Otto K. Heiney AFATL - DLDL Elgin AFB, FL 32542	1	Mr. L.A. Watermeier Army Ballistic Research Labs ARRADCOM Code ORDAR-BLI Aberdeen Proving Ground, MD 21005	1
Mr. R Geisler ATTN: MKP/MS24 AFRPL Edwards AFB, CA 93523	1	Dr. W.W. Wharton Attn: DR3MZ-RKL Commander U.S. Army Missile Command Redstone Arsenal, AL 35898	1

DISTRIBUTION LIST

	<u>No. Copies</u>		<u>No. Copies</u>
Dr. R.G. Rhoades Commander Army Missile Command DRSMI-R Redstone Arsenal, AL 35898	1	Dr. E.H. Debutts Hercules Inc. Baccus Works P.O. Box 98 Magna, UT 84044	1
Dr. W.D. Stephens Atlantic Research Corp. Pine Ridge Plant 7511 Wellington Rd. Gainesville, VA 22065	1	Dr. James H. Thacher Hercules Inc. Magna Baccus Works P.O. Box 98 Magna, UT 84044	1
Dr. A.W. Barrows Ballistic Research Laboratory USA ARRADCOM DRDAR-BLP Aberdeen Proving Ground, MD 21005	1	Mr. Theodore M. Gilliland Johns Hopkins University APL Chemical Propulsion Info. Agency Johns Hopkins Road Laurel, MD 20810	1
Dr. C.M. Frey Chemical Systems Division P.O. Box 358 Sunnyvale, CA 94086	1	Dr. R. McGuire Lawrence Livermore Laboratory University of California Code L-324 Livermore, CA 94550	1
Professor F. Rodriguez Cornell University School of Chemical Engineering Olin Hall, Ithaca, N.Y. 14853	1	Dr. Jack Linsk Lockheed Missiles & Space Co. P.O. Box 504 Code Org. 83-10, Bldg. 154 Sunnyvale, CA 94088	1
Defense Technical Information Center DTIC-DDA-2 Cameron Station Alexandria, VA 22314	12	Dr. B.G. Craig Los Alamos National Lab P.O. Box 1663 NSP/DOD, MS-245 Los Alamos, NM 87545	1
Dr. Rocco C. Musso Hercules Aerospace Division Hercules Incorporated Alleghany Ballistic Lab P.O. Box 210 Washington, D.C. 21502	1	Dr. R.L. Rabie WX-2, MS-952 Los Alamos National Lab. P.O. Box 1663 Los Alamos NM 37545	1
Dr. Ronald L. Simmons Hercules Inc. Eglin AFATL/DLOL Eglin AFB, FL 32542	1	Dr. R. Rogers Los Alamos Scientific Lab. P.O. Box 1663 Los Alamos, NM 87545	1

DISTRIBUTION LIST

	<u>No. Copies</u>		<u>No. Copies</u>
Mr. R. Brown Naval Air Systems Command Code 330 Washington, D.C. 20361	1	Dr. J. Schnur Naval Research Lab. Code 6510 Washington, D.C. 20375	1
Dr. H. Rosenwasser Naval Air Systems Command AIR-310C Washington, D.C. 20360	1	Mr. R. Beauregard Naval Sea Systems Command SEA 64E Washington, D.C. 20362	1
Mr. B. Sobers Naval Air Systems Command Code 03P25 Washington, D.C. 20360	1	Mr. G. Edwards Naval Sea Systems Command Code 62R3 Washington, D.C. 20362	1
Dr. L.R. Rothstein Assistant Director Naval Explosives Dev. Engineering Dept. Naval Weapons Station Yorktown, VA 23691	1	Mr. John Boyle Materials Branch Naval Ship Engineering Center Philadelphia, PA 19112	1
Dr. Lionel Dickinson Naval Explosive Ordnance Disposal Tech. Center Code D Indian Head, MD 20640	1	Dr. H.G. Adolph Naval Surface Weapons Center Code R11 White Oak Silver Spring, MD 20910	1
Mr. C.L. Adams Naval Ordnance Station Code PM4 Indian Head, MD 20640	1	Dr. T.D. Austin Naval Surface Weapons Center Code R16 Indian Head, MD 20640	1
Mr. S. Mitchell Naval Ordnance Station Code 5253 Indian Head, MD 20640	1	Dr. T. Hall Code R-11 Naval Surface Weapons Center White Oak Laboratory Silver Spring, MD 20910	1
Dr. William Tolles Dean of Research Naval Postgraduate School Monterey, CA 93940	1	Mr. G.L. Mackenzie Naval Surface Weapons Center Code R101 Indian Head, MD 20640	1
Naval Research Lab. Code 6100 Washington, D.C. 20375	1	Dr. K.F. Mueller Naval Surface Weapons Center Code R11 White Oak Silver Spring, MD 20910	1

DISTRIBUTION LIST

	<u>No. Copies</u>		<u>No. Copies</u>
Mr. J. Murrin Naval Sea Systems Command Code 62R2 Washington, D.C. 20362	1	Dr. A. Nielsen Naval Weapons Center Code 385 China Lake, CA 93555	1
Dr. D.J. Pastine Naval Surface Weapons Center Code R04 White Oak Silver Spring, MD 20910	1	Dr. R. Reed, Jr. Naval Weapons Center Code 388 China Lake, CA 93555	1
Mr. L. Roslund Naval Surface Weapons Center Code R122 White Oak, Silver Spring MD 20910	1	Dr. L. Smith Naval Weapons Center Code 3205 China Lake, CA 93555	1
Mr. M. Stosz Naval Surface Weapons Center Code R121 White Oak Silver Spring, MD 20910	1	Dr. B. Douda Naval Weapons Support Center Code 5042 Crane, Indiana 47522	1
Dr. E. Zimmet Naval Surface Weapons Center Code R13 White Oak Silver Spring, MD 20910	1	Dr. A. Faulstich Chief of Naval Technology MAT Code 0716 Washington, D.C. 20360	1
Dr. D. R. Derr Naval Weapons Center Code 388 China Lake, CA 93555	1	LCDR J. Walker Chief of Naval Material Office of Naval Technology MAT, Code 0712 Washington, D.C. 20360	1
Mr. Lee N. Gilbert Naval Weapons Center Code 3205 China Lake, CA 93555	1	Mr. Joe McCartney Naval Ocean Systems Center San Diego, CA 92152	1
Dr. E. Martin Naval Weapons Center Code 3858 China Lake, CA 93555	1	Dr. S. Yamamoto Marine Sciences Division Naval Ocean Systems Center San Diego, CA 91232	1
Mr. R. McCarten Naval Weapons Center Code 3272 China Lake, CA 93555	1	Dr. G. Bosmajian Applied Chemistry Division Naval Ship Research & Development Center Annapolis, MD 21401	1
		Dr. H. Shuey Rohn and Haas Company Huntsville, Alabama 35801	1

DISTRIBUTION LIST

	<u>No. Copies</u>		<u>No. Copies</u>
Dr. J.F. Kincaid Strategic Systems Project Office Department of the Navy Room 901 Washington, D.C. 20376	1	Dr. C.W. Vriesen Thiokol Elkton Division P.O. Box 241 Elkton, MD 21921	1
Strategic Systems Project Office Propulsion Unit Code SP2731 Department of the Navy Washington, D.C. 20376	1	Dr. J.C. Hinshaw Thiokol Wasatch Division P.O. Box 524 Brigham City, Utah 84302	1
Mr. E.L. Throckmorton Strategic Systems Project Office Department of the Navy Room 1043 Washington, D.C. 20376	1	U.S. Army Research Office Chemical & Biological Sciences Division P.O. Box 12211 Research Triangle Park NC 27709	1
Dr. D.A. Flanigan Thiokol Huntsville Division Huntsville, Alabama 35807	1	Dr. R.F. Walker USA ARRADCOM DRDAR-LCE Dover, NJ 07801	1
Mr. G.F. Mangum Thiokol Corporation Huntsville Division Huntsville, Alabama 35807	1	Dr. T. Sinden Munitions Directorate Propellants and Explosives Defence Equipment Staff British Embassy 3100 Massachusetts Ave. Washington, D.C. 20008	1
Mr. E.S. Sutton Thiokol Corporation Elkton Division P.O. Box 241 Elkton, MD 21921	1	LTC B. Loving AFROL/LK Edwards AFB, CA 93523	1
Dr. G. Thompson Thiokol Wasatch Division MS 240 P.O. Box 524 Brigham City, UT 84302	1	Professor Alan N. Gent Institute of Polymer Science University of Akron Akron, OH 44325	1
Dr. T.F. Davidson Technical Director Thiokol Corporation Government Systems Group P.O. Box 9253 Ogden, Utah 84409	1	Mr. J. M. Frankle Army Ballistic Research Labs ARRADCOM Code DRDAR-BLI Aberdeen Proving Ground, MD 21005	1

DISTRIBUTION LIST

	<u>No. Copies</u>		<u>No. Copies</u>
Dr. Ingo W. May Army Ballistic Research Labs ARRADCOM Code DRDAR-BLI Aberdeen Proving Ground, MD 21005	1	Dr. J. P. Marshall Dept. 52-35, Bldg. 204/2 Lockheed Missile & Space Co. 3251 Hanover Street Palo Alto, CA 94304	1
Professor N.W. Tschoegl California Institute of Tech Dept. of Chemical Engineering Pasadena, CA 91125	1	Ms. Joan L. Janney Los Alamos National Lab Mail Stop 920 Los Alamos, NM 87545	1
Professor M.D. Nicol University of California Dept. of Chemistry 405 Hilgard Avenue Los Angeles, CA 90024	1	Dr. J. M. Walsh Los Alamos Scientific Lab Los Alamos, NM 87545	1
Professor A. G. Evans University of California Berkeley, CA 94720	1	Professor R. W. Armstrong Univ. of Maryland Department of Mechanical Eng. College Park, MD 20742	1
Professor T. Litovitz Catholic Univ. of America Physics Department 520 Michigan Ave., N.E. Washington, D.C. 20017	1	Prof. Richard A. Reinhardt Naval Postgraduate School Physics & Chemistry Dept. Monterey, CA 93940	1
Professor W. G. Knauss Graduate Aeronautical Lab California Institute of Tech. Pasadena, CA 91125	1	Dr. R. Bernecker Naval Surface Weapons Center Code R13 White Oak, Silver Spring, MD 20910	1
Professor Edward Price Georgia Institute of Tech. School of Aerospace Engin. Atlanta, Georgia 30332	1	Dr. M. J. Kamlet Naval Surface Weapons Center Code R11 White Oak, Silver Spring, MD 20910	1
Dr. Kenneth O. Hartman Hercules Aerospace Division Hercules Incorporated P.O. Box 210 Cumberland, MD 21502	1	Professor J. D. Achenbach Northwestern University Dept. of Civil Engineering Evanston, IL 60201	1
Dr. Thor L. Smith IBM Research Lab D42.282 San Jose, CA 95193	1	Dr. N. L. Basdekas Office of Naval Research Mechanics Program, Code 432 Arlington, VA 22217	1
		Professor Kenneth Kuo Pennsylvania State Univ. Dept. of Mechanical Engineering University Park, PA 16802	1

DISTRIBUTION LIST

	<u>No. Copies</u>		<u>No. Copies</u>
Dr. S. Sheffield Sandia Laboratories Division 2513 P.O. Box 5800 Albuquerque, NM 87185	1	ONR Resident Representative Ohio State University Res. Ctr. 1314 Kinnear Road Columbus, OH 43212	1
Dr. M. Farber Space Sciences, Inc. 135 Maple Avenue Monrovia, CA 91016	1		
Dr. Y. M. Gupta SRI International 333 Ravenswood Avenue Menlo Park, CA 94025	1		
Mr. M. Hill SRI International 333 Ravenswood Avenue Menlo Park, CA 94025	1		
Professor Richard A. Schapery Texas A&M Univ. Dept of Civil Engineering College Station, TX 77843	1		
Dr. Stephen Swanson Univ. of Utah Dept. of Mech. & Industrial Engineering MEB 3008 Salt Lake City, UT 84112	1		
Mr. J. D. Byrd Thiokol Corp. Huntsville Huntsville Div. Huntsville, AL 35807	1		
Professor G. D. Duval Washington State University Dept. of Physics Pullman, WA 99163	1		
Prof. T. Dickinson Washington State University Dept. of Physics Pullman, WA 99163	1		