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THE SYNTHESIS OF
TRIS(PENTAMETHYLCYCLOPENTA-DIENYL)GALLIUM(III)(U) STATE
UNIV OF NEW YORK AT BUFFALO DEPT OF CHEMISTRY
O T BEACHLEY ET AL. 29 JUL 86 TR-16

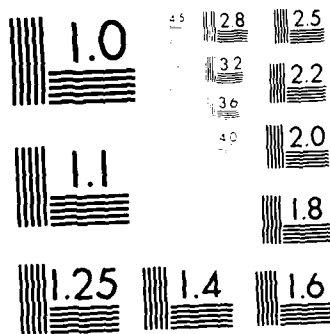
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OFFICE OF NAVAL RESEARCH
Contract N-00014-78-C-0562
Task No. NR 053-686
TECHNICAL REPORT NO. 16

The Synthesis of Tris(pentamethylcyclopentadienyl)gallium(III)

by

O. T. Beachley, Jr. and R. B. Hallock

Prepared for Publication

in

Organometallics

State University of New York at Buffalo
Department of Chemistry
Buffalo, New York 14214

29, July 1986

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REPORT DOCUMENTATION PAGE		READ INSTRUCTIONS BEFORE COMPLETING FORM
1. REPORT NUMBER 16	2. GOVT ACCESSION NO.	3. RECIPIENT'S CATALOG NUMBER Technical Report
4. TITLE (and Subtitle) The Synthesis of Tris(pentamethylcyclopentadienyl)gallium(III)		5. TYPE OF REPORT & PERIOD COVERED
		6. PERFORMING ORG. REPORT NUMBER
7. AUTHOR(s) O. T. Beachley, Jr. and R. B. Hallock		8. CONTRACT OR GRANT NUMBER(s) N-0014-78C-0562
9. PERFORMING ORGANIZATION NAME AND ADDRESS Department of Chemistry State University of New York at Buffalo Buffalo, New York 14214		10. PROGRAM ELEMENT, PROJECT, TASK AREA & WORK UNIT NUMBERS NR-053-686
11. CONTROLLING OFFICE NAME AND ADDRESS Office of Naval Research Department of the Navy Arlington, VA 22117		12. REPORT DATE July 29, 1986
		13. NUMBER OF PAGES 10
14. MONITORING AGENCY NAME & ADDRESS (if different from Controlling Office)		15. SECURITY CLASS. (of this report) Unclassified
		15a. DECLASSIFICATION/DOWNGRADING SCHEDULE
16. DISTRIBUTION STATEMENT (of this Report) Approval for Public Release, Distribution Unlimited		
17. DISTRIBUTION STATEMENT (of the abstract entered in Block 20, if different from Report) Prepared for Publication in Organometallics		
18. SUPPLEMENTARY NOTES		
19. KEY WORDS (Continue on reverse side if necessary and identify by block number) Organogallium Chemistry Synthesis Main-Group Chemistry Pentamethylcyclopentadienyl Chemistry		
20. ABSTRACT (Continue on reverse side if necessary and identify by block number) The previously unknown and elusive compound, $\text{Ga}(\text{C}_5\text{Me}_5)_3$, has been prepared by several synthetic routes and has been characterized by elemental analyses, IR and ^1H NMR data. The compound has been prepared from GaCl_3 and excess $\text{Na}(\text{C}_5\text{Me}_5)$ in THF solution. A second general route to $\text{Ga}(\text{C}_5\text{Me}_5)_3$, the reaction of $\text{Ga}(\text{C}_5\text{Me}_5)_n\text{Cl}_{3-n}$ ($n=1,2$) with sodium naphthalenide in THF solution, has also been realized. Heating of the product mixture at 50-60°C under vacuum leads to the complete removal of the reaction solvent, THF, an observation which suggests that $\text{Ga}(\text{C}_5\text{Me}_5)_3$ is a very weak Lewis acid. The title compound exists as a		

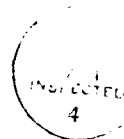
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S/N 0102-014-6601

(Contribution from the Department of Chemistry State University of New York
at Buffalo, Buffalo, New York 14214)



The Synthesis of Tris(pentamethylcyclopentadienyl)gallium(III)

by

O. T. Beachley, Jr.* and R. B. Hallock

Abstract

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The previously unknown and elusive compound, $\text{Ga}(\text{C}_5\text{Me}_5)_3$, has been prepared by several synthetic routes and has been characterized by elemental analyses, IR and ^1H NMR data. The compound has been prepared from GaCl_3 and excess $\text{Na}(\text{C}_5\text{Me}_5)$ in THF solution. A second general route to $\text{Ga}(\text{C}_5\text{Me}_5)_3$, the reaction of $\text{Ga}(\text{C}_5\text{Me}_5)_n\text{Cl}_{3-n}$ ($n=1,2$) with sodium naphthalenide in THF solution, has also been realized. Heating of the product mixture at $50-60^\circ\text{C}$ under vacuum leads to the complete removal of the reaction solvent, THF, an observation which suggests that $\text{Ga}(\text{C}_5\text{Me}_5)_3$ is a very weak Lewis acid. The title compound exists as a pentane and ether soluble colorless crystalline solid with a melting point of $168-170^\circ\text{C}$.

Introduction

The study of cyclopentadienyl and pentamethylcyclopentadienyl compounds is part of our ongoing investigations of the chemistry of compounds of group 13 elements in typical and low oxidation states. The most generally useful synthetic route to the desired compounds utilizes metathetical reactions. The compound $\text{Ga}(\text{C}_5\text{H}_5)_3$ was prepared from GaCl_3 and excess LiC_5H_5 in diethyl ether¹. An X-ray structural study revealed a molecule with a simple trigonal planar gallium atom and cyclopentadienyl rings exhibiting η^1 coordination. In contrast, attempts to prepare $\text{Ga}(\text{C}_5\text{Me}_5)_3$ from GaCl_3 and $\text{Li}(\text{C}_5\text{Me}_5)$ proved futile as only $\text{Ga}(\text{C}_5\text{Me}_5)_2\text{Cl}$ was formed². Regardless of the amount of excess $\text{Li}(\text{C}_5\text{Me}_5)$ used, or whether Et_2O or THF was used as the solvent, the only gallium(III) product was $\text{Ga}(\text{C}_5\text{Me}_5)_2\text{Cl}$. Cyclopentadienylindium(I) has also been readily prepared by a metathetical reaction³ between InCl and LiC_5H_5 . Similarly, $\text{In}(\text{C}_5\text{Me}_5)$ has been prepared and characterized by an X-ray structural study⁴. This new indium(I) compound exists as an apparent octahedral cluster with indium(I) atoms on the interior and η^5 -pentamethylcyclopentadienyl groups on the exterior. Since no simple gallium(I) halogen compounds are available as starting materials for the possible preparation of organogallium(I) compounds by metathesis reactions, reduction reactions of organogallium(III) halogen compounds form the basis of a possible synthetic route. In this paper we report the results of our studies of reduction reactions of $\text{Ga}(\text{C}_5\text{Me}_5)\text{Cl}_2$ and $\text{Ga}(\text{C}_5\text{Me}_5)_2\text{Cl}$ using sodium naphthalenide in THF solution. The isolable products from these reactions are $\text{Ga}(\text{C}_5\text{Me}_5)_3$, naphthalene, and an insoluble gray solid indicative of gallium metal and sodium chloride. Tris(pentamethylcyclopentadienyl)gallium(III) has also been independently prepared from GaCl_3 and excess $\text{Na}(\text{C}_5\text{Me}_5)$ in THF solution.

Experimental Section

General Data. All compounds described in this investigation were extremely sensitive to oxygen and moisture and were manipulated in a standard vacuum line or a purified argon atmosphere. The starting materials $\text{Ga}(\text{C}_5\text{Me}_5)_2\text{Cl}$ and $\text{Ga}(\text{C}_5\text{Me}_5)\text{Cl}_2$ were prepared as previously described². The solvents were purified by conventional means (THF-refluxing Na/benzophenone) and were distilled immediately prior to use. Analyses were performed by Schwarzkopf Microanalytical Laboratory, Woodside, NY. Infrared spectra of Nujol mulls between CsI plates were recorded by means of a Perkin-Elmer Model 683 spectrometer. Absorption intensities are reported with abbreviations w(weak), m(medium), s(strong), vs(very strong), and sh(shoulder). The ^1H NMR spectra were recorded at 90 MHz by using a Varian Model EM-390 spectrometer. Chemical shifts are reported in δ units (ppm) and are referenced to SiMe_4 as δ 0.00 and benzene as δ 7.13.

Preparation of $\text{Na}(\text{C}_5\text{Me}_5)$. Finely divided sodium (1.02 g, 44.3 mmol), which had been isolated by removing the toluene from a 30% Na/toluene dispersion by vacuum distillation, was combined neat with 5.41 g (39.7 mmol) of $\text{C}_5\text{Me}_5\text{H}$ in a reaction tube in an Ar filled drybox. Upon mixing, an exothermic reaction occurred and vigorous bubbling was observed. After removal of the reaction tube from the drybox, solvent (THF-30 mL) was vacuum distilled into the reaction tube. The mixture was then stirred for 4 h at 60-80°C to insure complete reaction, although lack of further hydrogen evolution showed that the reaction had gone to completion in the drybox in the absence of solvent. After attempts to remove excess sodium dispersion from the THF solution of $\text{Na}(\text{C}_5\text{Me}_5)$ by filtration proved unsuccessful due to clogging of the frit by the finely divided sodium, the supernatant brown solution was carefully pipetted away from the fine gray sodium powder. A final

filtration gave a clear brown solution from which THF was removed by vacuum distillation. The resulting brown solid was washed twice with 25 mL of pentane to remove a yellow, pentane soluble oil from the insoluble, brown product, $\text{Na}(\text{C}_5\text{Me}_5)$. Complete removal of THF from the product was insured by evacuation overnight at 75-80°C. The isolated yield of purified $\text{Na}(\text{C}_5\text{Me}_5)$ was 1.6 g (10 mmol, 25% yield).

$\text{Na}(\text{C}_5\text{Me}_5)$. ^1H NMR (d_8 -THF solution, δ): 1.94 (s). IR (Nujol mull, cm^{-1}): 2727 (s), 2671 (w,sh), 1580 (m,br), 1257 (m), 1150 (m), 1110 (m,sh), 1079 (m,br), 1043 (m,sh), 1012 (m,sh), 910 (w,sh), 870 (vw), 800 (w,sh), 787 (w), 762 (vw), 735 (w,sh), 716 (w), 640 (vw), 458 (vw), 397 (w), 328 (vs), 270 (m,sh).

Preparation of $\text{Ga}(\text{C}_5\text{Me}_5)_3$. A solution of 0.536 g (3.04 mmol) of freshly sublimed GaCl_3 in 40 mL of THF was added with stirring at room temperature to a THF (40 mL) solution of $\text{Na}(\text{C}_5\text{Me}_5)$ (1.540 g, 9.733 mmol). The reaction mixture was stirred at room temperature overnight and the THF was completely removed, first by vacuum distillation, then by evacuating the system at 50-60°C for 2 h. Pentane (70 mL) was vacuum distilled onto the products and the mixture was filtered to separate a yellow solution of $\text{Ga}(\text{C}_5\text{Me}_5)_3$ from solid NaCl and unreacted $\text{Na}(\text{C}_5\text{Me}_5)$. Removal of the pentane by vacuum distillation left 0.742 g (1.56 mmol, 51.3% yield) of impure $\text{Ga}(\text{C}_5\text{Me}_5)_3$ as a yellow solid. Further purification was possible by washing the product with 5 mL portions of very cold (about -78°C) pentane. However, this process reduced the yield of pure $\text{Ga}(\text{C}_5\text{Me}_5)_3$ to 0.400 g (0.841 mmol, 27.7% yield). Colorless crystals of $\text{Ga}(\text{C}_5\text{Me}_5)_3$ for a possible X-ray structural study were grown by slow sublimation of a sample of $\text{Ga}(\text{C}_5\text{Me}_5)_3$ in an evacuated sealed tube at 100-105°C.

Ga(C₅Me₅)₃-Colorless solid; mp 168-170°C, melts to a yellow liquid; ¹H NMR (benzene solution, δ): 1.73; IR (Nujol mull, cm⁻¹): 2720 (m), 1613 (m), 1300 (w,sh), 1236 (vs), 1136 (m), 1125 (m), 1052 (m), 1042 (m,sh), 968 (w,sh), 940 (m), 815 (m), 786 (m), 716 (w,sh), 698 (vs), 685 (m,sh), 644 (w), 592 (m), 560 (m), 518 (w), 412 (s), 292 (vs), 270 (m), 210 (m).

Reaction of Ga(C₅Me₅)₂Cl with Na/Naphthalene. A solution of 3.0 mmol of sodium naphthalenide in 40 mL of THF was added at -78°C to a solution of 1.127 g (3.00 mmol) of Ga(C₅Me₅)₂Cl in 35 mL of THF. After stirring at -78°C for 30 minutes, then at room temperature for 1 h, the THF was removed and 50 mL of pentane was added by vacuum distillation. The mixture was filtered to separate a gray precipitate from a light yellow solution. After removal of pentane, 0.240 g of a gray precipitate, indicative of gallium metal plus NaCl (97.9% yield), was isolated. The pentane soluble, yellow solid was evacuated at 70-80°C overnight to insure quantitative removal of naphthalene, and then was washed several times with 5 mL portions of cold (about -78°C) pentane to yield 0.310 g (0.652 mmol, 32.6% yield) of Ga(C₅Me₅)₃. The identity of the product was confirmed by its mp (165-170°C), ¹H NMR and IR spectroscopic data and C, H analysis (Calcd: C, 75.79; H, 9.54; Found: C, 75.98; H, 9.75).

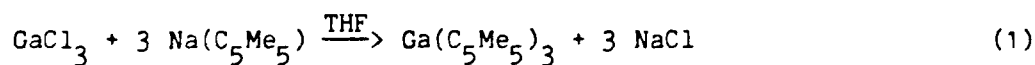
Reaction of Ga(C₅Me₅)Cl₂ with Na/Naphthalene. A solution of 2.62 mmol of sodium naphthalenide in 50 mL of THF was added at -78°C to a solution containing 0.360 g (1.31 mmol) of Ga(C₅Me₅)Cl₂ in 25 mL of THF. After stirring for 1 h at -78°C, the mixture was warmed to room temperature and the THF was removed by vacuum distillation. Then, 50 mL of pentane was added to the reaction mixture. Filtration and removal of pentane produced 0.185 g of a gray insoluble solid, indicative of gallium metal and NaCl (86.4% yield). The pentane soluble material was a pasty, yellow solid which

was evacuated overnight at 60°C to insure quantitative removal of naphthalene. The resulting material was shown to be slightly impure $\text{Ga}(\text{C}_5\text{Me}_5)_3$ by comparison of its mp (148-161°C), ^1H NMR and IR spectroscopic data of that with known samples of $\text{Ga}(\text{C}_5\text{Me}_5)_3$. Qualitative analysis tests showed the lack of a AgCl precipitate when the gallium containing product reacted with HNO_3 and AgNO_3 , as well as the absence of sodium as indicated by a flame test.

Reaction of $\text{Ga}(\text{C}_5\text{Me}_5)_3$ with Na/Naphthalene. A solution of sodium naphthalenide (1.04 mmol in 25 mL of THF) was added at -20°C to a solution containing 0.247 g (0.520 mmol) of $\text{Ga}(\text{C}_5\text{Me}_5)_3$ in 25 mL of THF. A black precipitate formed immediately upon mixing. The mixture was stirred for 3 h at room temperature, was filtered, and then the THF was removed by vacuum distillation. The THF soluble solid was evacuated overnight at 65°C to insure quantitative removal of naphthalene and then was washed 6 times with 25 mL portions of pentane. After removal of pentane under vacuum, the remaining THF soluble, pentane insoluble solid was isolated and identified by ^1H NMR and IR spectroscopic data as $\text{Na}(\text{C}_5\text{Me}_5)$ (0.168 g, 1.06 mmol). The pentane soluble solid was shown to be $\text{Ga}(\text{C}_5\text{Me}_5)_3$ by IR spectroscopic data.

Results and Discussion

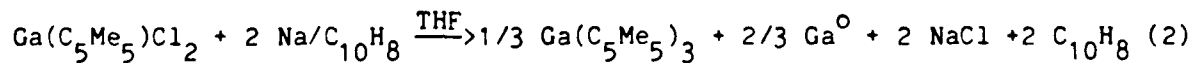
The previously unknown and elusive compound $\text{Ga}(\text{C}_5\text{Me}_5)_3$ has been prepared from GaCl_3 and excess $\text{Na}(\text{C}_5\text{Me}_5)$ in THF solution and has been characterized by melting point, elemental analyses, IR and ^1H NMR spectroscopic data. Since the initial attempts² to prepare $\text{Ga}(\text{C}_5\text{Me}_5)_3$ from GaCl_3 and excess $\text{Li}(\text{C}_5\text{Me}_5)$ led to the formation of only $\text{Ga}(\text{C}_5\text{Me}_5)_2\text{Cl}$, experiments using the more reactive compound $\text{Na}(\text{C}_5\text{Me}_5)$ in THF were undertaken. The observation that $\text{Na}(\text{C}_5\text{Me}_5)$ can be used to prepare $\text{Ga}(\text{C}_5\text{Me}_5)_3$ whereas $\text{Li}(\text{C}_5\text{Me}_5)$ replaces only a maximum of two chlorine



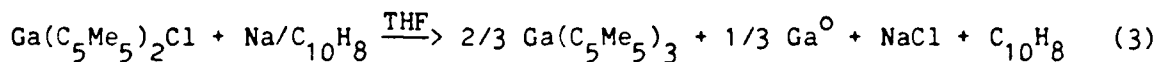
atoms on gallium suggests that the higher reactivity of $\text{Na}(\text{C}_5\text{Me}_5)$ might be related to its increased tendency to exist in solution as simple ionic species rather than associated molecules.⁵ When $\text{Ga}(\text{C}_5\text{Me}_5)_3$ is initially isolated from the reaction mixture, it is contaminated by a yellow material. Most of the unknown yellow impurity can be removed by washing the product with cold (-78°C) pentane. However, this process significantly lowers the yield of the final product. Additional purification can be achieved by sublimation at $100-105^\circ\text{C}$. Colorless crystals of $\text{Ga}(\text{C}_5\text{Me}_5)_3$ have a well-defined, reversible melting point of $168-170^\circ\text{C}$. The compound is soluble in aprotic solvents and both diethyl ether and THF can be removed completely. These observations suggest that $\text{Ga}(\text{C}_5\text{Me}_5)_3$ is a very weak Lewis acid. It is also of interest that $\text{Ga}(\text{C}_5\text{Me}_5)_3$ has enhanced thermal stability compared to $\text{Ga}(\text{C}_5\text{H}_5)_3$, a compound¹ which decomposes at 45°C . It is regrettable that single crystals of $\text{Ga}(\text{C}_5\text{Me}_5)_3$, which were grown by sublimation at $100-105^\circ\text{C}$ for an X-ray structural study, proved unsuitable.

The reactions of pentamethylcyclopentadienylgallium(III) derivatives including $\text{Ga}(\text{C}_5\text{Me}_5)\text{Cl}_2$, $\text{Ga}(\text{C}_5\text{Me}_5)_2\text{Cl}$, and $\text{Ga}(\text{C}_5\text{Me}_5)_3$ with sodium-naphthalenide have been investigated as possible routes to a low oxidation state gallium derivative. The pentamethylcyclopentadienyl ligand might be expected to be an ideal ligand for the stabilization of a low oxidation state gallium compound. Furthermore, alkali metal reduction reactions have been used previously to prepare group 13 compounds which have been described as low oxidation state compounds. The reduction of AlEt_3 using Na/naphthalene ⁶ has been reported to yield $\text{NaAlEt}_2 \cdot \text{naphthalene}$. The reaction of iBu_2AlCl with potassium pellets^{7,8} in hexane yields $\text{Al}_2(\text{iBu})_4$, a species with an apparent Al-Al bond. A related reaction to form $\text{K}_2\text{Al}_2(\text{iBu})_6$ has also been reported⁹. In gallium chemistry, $\text{Na}_2\text{Ga}_2\text{Me}_6$ has been described as the initial product¹⁰ from a GaMe_3 -Na reaction in liquid NH_3 . Similarly, the reaction¹⁰ of GaMe_2Cl with Na in liquid NH_3 has been reported to produce $(\text{Me}_2\text{GaNH}_2)_2$, NaCl and H_2 .

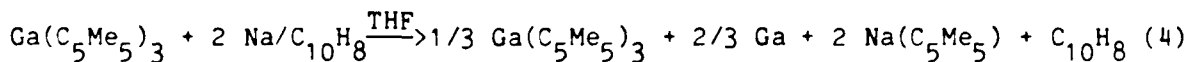
In the pentamethylcyclopentadienylgallium(III)-sodium naphthalenide reactions the only isolated gallium products were $\text{Ga}(\text{C}_5\text{Me}_5)_3$ and gallium metal. When $\text{Ga}(\text{C}_5\text{Me}_5)\text{Cl}_2$ was reacted with $\text{Na/C}_{10}\text{H}_8$ in a 0.50 mole ratio, an insoluble gray solid consistent with an 86% yield (based on sodium) of NaCl and gallium metal was isolated (equation 2). The identity of the



product as $\text{Ga}(\text{C}_5\text{Me}_5)_3$ was verified by a comparison of the melting point, ^1H NMR and IR spectral data with that of the known compound. The reaction of $\text{Ga}(\text{C}_5\text{Me}_5)\text{Cl}$ (equation 3) occurred similarly.



When $\text{Ga}(\text{C}_5\text{Me}_5)_3$ was reacted with $\text{Na}/\text{C}_{10}\text{H}_8$ in a 0.50 mole ratio (equation 4), the products were gallium metal and $\text{Na}(\text{C}_5\text{Me}_5)$. It is of interest that



$\text{NaGa}(\text{C}_5\text{Me}_5)_4$ is not formed. Apparently, the steric bulk of the C_5Me_5 ligand reduces either the Lewis acidity of $\text{Ga}(\text{C}_5\text{Me}_5)_3$ or the Lewis basicity of C_5Me_5^- to preclude the formation of $\text{NaGa}(\text{C}_5\text{Me}_5)_4$. Thus, the reactions of $\text{Ga}(\text{C}_5\text{Me}_5)_n\text{Cl}_{3-n}$ ($n = 3, 2, 1$) with $\text{Na}/\text{naphthalene}$ occur as previously observed for other group 13 alkyl halide - sodium reactions except that MR_3 is formed instead of NaMR_4 when excess sodium is used. Additional reactions will have to be investigated for us to realize the successful synthesis of an organogallium(I) derivative.

Acknowledgements. This work was supported in part by the National Science Foundation through Grant CHE-81-3316 and by the Office of Naval Research. We also thank Professor J. L. Atwood and his students at University of Alabama, University, Alabama for attempting the X-ray structural study of $\text{Ga}(\text{C}_5\text{Me}_5)_3$.

References

1. Beachley, O. T., Jr.; Getman, T. D.; Kirss, R. U.; Hallock, R. B.; Hunter, W. E.; Atwood, J. L. *Organometallics* 1985, 4, 751.
2. Beachley, O. T., Jr.; Hallock, R. B.; Zhang, H. M.; Atwood, J. L. *Organometallics* 1985, 4, 1675.
3. Peppe, C.; Tuck, D. G.; Victoriano, L. *J.C.S. Dalton* 1981, 2592.
4. Beachley, O. T., Jr.; Churchill, M. R.; Fettingner, J. C.; Pazik, J. C., Victoriano, L. *J. Amer. Chem. Soc.* (Submitted).
5. Coates, G. E.; Green, M. L. H.; Wade, K. "Organometallic Compounds" Vol. 1, Methuen & Co. Ltd. 1967.
6. Boehm, H. P.; Schneider, M.; Wistuba, H. *Angew. Chem. Int. Ed. Engl.* 1965, 4, 600.
7. Hoberg, H.; Krause, S. *Angew. Chem. Int. Ed. Engl.* 1976, 15, 694.
8. Miller, M. A.; Schram, E. P. *Organometallics* 1985, 4, 1362.
9. Hoberg, H.; Krause, S. *Angew. Chem. Int. Ed. Engl.* 1978, 17, 949.
10. Krause, C. A.; Toonder, F. E. *J. Amer. Chem. Soc.* 1933, 55, 3547.

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