

REPORT DOCUMENTATION PAGE

Form Approved
OMB No. 0704-0188

Public reporting burden for this collection of information is estimated to average 1 hour per response, including the time for reviewing instructions, searching existing data sources, gathering and maintaining the data needed, and completing and reviewing this collection of information. Send comments regarding this burden estimate or any other aspect of this collection of information, including suggestions for reducing this burden to Department of Defense, Washington Headquarters Services, Directorate for Information Operations and Reports (0704-0188), 1215 Jefferson Davis Highway, Suite 1204, Arlington, VA 22202-4302. Respondents should be aware that notwithstanding any other provision of law, no person shall be subject to any penalty for failing to comply with a collection of information if it does not display a currently valid OMB control number. PLEASE DO NOT RETURN YOUR FORM TO THE ABOVE ADDRESS.

1. REPORT DATE (DD-MM-YYYY)		2. REPORT TYPE Technical Papers		3. DATES COVERED (From - To)	
4. TITLE AND SUBTITLE				5a. CONTRACT NUMBER	
				5b. GRANT NUMBER	
				5c. PROGRAM ELEMENT NUMBER	
6. AUTHOR(S)				5d. PROJECT NUMBER 2303	
				5e. TASK NUMBER m2c8	
				5f. WORK UNIT NUMBER	
7. PERFORMING ORGANIZATION NAME(S) AND ADDRESS(ES) Air Force Research Laboratory (AFMC) AFRL/PRS 5 Pollux Drive Edwards AFB CA 93524-7048				8. PERFORMING ORGANIZATION REPORT	
9. SPONSORING / MONITORING AGENCY NAME(S) AND ADDRESS(ES) Air Force Research Laboratory (AFMC) AFRL/PRS 5 Pollux Drive Edwards AFB CA 93524-7048				10. SPONSOR/MONITOR'S ACRONYM(S)	
				11. SPONSOR/MONITOR'S NUMBER(S)	
12. DISTRIBUTION / AVAILABILITY STATEMENT Approved for public release; distribution unlimited.					
13. SUPPLEMENTARY NOTES					
14. ABSTRACT					
15. SUBJECT TERMS					
16. SECURITY CLASSIFICATION OF:			17. LIMITATION OF ABSTRACT A	18. NUMBER OF PAGES	19a. NAME OF RESPONSIBLE PERSON Leilani Richardson
a. REPORT Unclassified	b. ABSTRACT Unclassified	c. THIS PAGE Unclassified			19b. TELEPHONE NUMBER (include area code) (661) 275-5015

62 1122030 seminate items are enclosed

18

MEMORANDUM FOR PR (Contractor/In-House Publication)

FROM: PROI (TI) (STINFO)

19 Jun 2000

SUBJECT: Authorization for Release of Technical Information, Control Number: **AFRL-PR-ED-TP-2000-136**
E. Lork, R. Mews, D. Viets (University of Bremen, Germany); P. Watson, T. Borrmann (USC); A. Vij, J. Boatz, K. Christe (ERC), "The Structure of the SO_2F^- Anion, a Problem Case"

Readership of Inorganic Chemistry
(2000) (Submission Deadline: none given)

(Statement A)

1. This request has been reviewed by the Foreign Disclosure Office for: a.) appropriateness of distribution statement, b.) military/national critical technology, c.) export controls or distribution restrictions, d.) appropriateness for release to a foreign nation, and e.) technical sensitivity and/or economic sensitivity.

Comments: _____

Signature _____ Date _____

2. This request has been reviewed by the Public Affairs Office for: a.) appropriateness for public release and/or b) possible higher headquarters review.

Comments: _____

Signature _____ Date _____

3. This request has been reviewed by the STINFO for: a.) changes if approved as amended, b.) appropriateness of distribution statement, c.) military/national critical technology, d.) economic sensitivity, e.) parallel review completed if required, and f.) format and completion of meeting clearance form if required

Comments: _____

Signature _____ Date _____

4. This request has been reviewed by PR for: a.) technical accuracy, b.) appropriateness for audience, c.) appropriateness of distribution statement, d.) technical sensitivity and economic sensitivity, e.) military/national critical technology, and f.) data rights and patentability

Comments: _____

APPROVED/APPROVED AS AMENDED/DISAPPROVED

LESLIE. S. PERKINS, Ph.D (Date)
Staff Scientist
Propulsion Directorate

SYNOPSIS

The controversy between crystallographers and theoreticians concerning the structure of SO_2F^- has been resolved in favor of the theoreticians. Crystal structure determinations of the TAS, TAOS, TMA, and K salts of SO_2F^- show that the discrepancies between theory and experiment were due to severe oxygen/fluorine disorder in the crystals. Corrections for the disorder effects lead to a geometry that is in excellent agreement with the theoretical predictions and the observed vibrational spectra. The length and weakness of the S-F bond is confirmed by *ab initio* calculations and a normal coordinate analysis.

Enno Lork, Rüdiger Mews,* Detlef Viets,
Paul G. Watson, Tobias Borrmann,
Ashwani Vij, Jerry A. Boatz, Karl O. Christe*

Inorg. Chem. 2000, 39,

The Structure of the SO_2F^- Anion,
a Problem Case



20021122 030

DISTRIBUTION STATEMENT A:
Approved for Public Release -
Distribution Unlimited

The Structure of the SO_2F^- Anion, a Problem Case¹

Enno Lork,[†] Rüdiger Mews,^{*†} Detlef Viets,[†] Paul G. Watson,[†] Tobias Borrmann,[†]
Ashwani Vij,[‡] Jerry A. Boatz,[‡] Karl O. Christe^{*‡,§}

Department of Chemistry, University of Bremen,
Leobener Strasse NW2, P.O.Box 330440, D- 28334 Bremen, Germany,
Air Force Research Laboratory, Edwards Air Force Base, California 93524, and
Loker Hydrocarbon Research Institute, University of Southern California,
Los Angeles, California 90089-1661

Abstract

Recently, room-temperature crystal structures of SO_2F^- in its K^+ and Rb^+ salts were published in *Z. Anorg. Allg. Chem.* **1999**, 625, 385 and claimed to represent the first reliable geometries for SO_2F^- . However, their almost identical S-O and S-F bond lengths and O-S-O and O-S-F bond angles are in sharp contrast to the results from theoretical calculations. To clarify this discrepancy, the new $[(\text{CH}_3)_2\text{N}]_3\text{SO}^+$ and the known $[\text{N}(\text{CH}_3)_4]^+$, $[(\text{CH}_3)_2\text{N}]_3\text{S}^+$ and K^+ salts of SO_2F^- were prepared and their crystal structures studied at low temperatures. Furthermore, the results from previous RHF and MP2 calculations were confirmed at the RHF, B3LYP and CCSD(T) levels of theory using different basis sets. It is shown that all the SO_2F^- salts studied so far exhibit varying degrees of oxygen-fluorine and, in some cases, oxygen site disorders, with $[(\text{CH}_3)_2\text{N}]_3\text{SO}^+ \text{SO}_2\text{F}^-$ at 113 °K showing the least disorder with $r(\text{SF})-r(\text{SO}) = 17 \text{ pm}$ and $\angle(\text{OSO})-\angle(\text{FSO}) = 6^\circ$. Refinement of the disorder occupancy factors and extrapolation of the observed bond distances for zero disorder resulted in a geometry very close to that predicted by theory. The correctness of the theoretical predictions for SO_2F^- is further supported by the good agreement between the calculated and the experimentally observed

vibrational frequencies and their comparison with those of isoelectronic ClO_2F . A normal coordinate analysis of SO_2F^- confirms the weakness of the S-F bond with a stretching force constant of only 1.63 mdyn/Å and shows that there is no highly characteristic S-F stretching mode. The S-F stretch strongly couples with the SO_2 deformation modes and is concentrated in the two lowest a' frequencies.

Introduction

The SO_2F^- anion has been known since 1953 from the work of Seel and coworkers.²⁻⁵ Its vibrational spectra were studied by Seel and Boudier,⁶ Paetzold and Aurich,⁷ Robinson and coworkers,⁸ Moock and coworkers,⁹ and Kornath and coworkers¹⁰ and correctly assigned with the help of ab initio calculations in terms of a C_s symmetry structure with a predicted geometry of $r(\text{SF}) = 170$ pm, $r(\text{SO}) = 146$ pm, $\angle(\text{OSO}) = 113.2^\circ$, and $\angle(\text{OSF}) = 100.6^\circ$.¹⁰ The agreement between observed and calculated vibrational frequencies was good and supported the calculated geometry. Furthermore, the calculated geometry of SO_2F^- is similar to that experimentally observed for isoelectronic ClO_2F .¹¹

The geometry, predicted from the theoretical calculation and supported by the vibrational analysis¹⁰ is in conflict with the results from three recent X-ray diffraction studies. Zhu and coworkers reported the crystal structure of $\text{Ph}_3\text{PCF}_2\text{H}^+\text{SO}_2\text{F}^-$.¹² In this structure, the SO_2F^- anion is both fluorine-oxygen and oxygen-site disordered and its geometry was listed as $r(\text{S-F}) = 151.6(6)$ pm, $r(\text{S-O1}) = 143.6$ pm, $r(\text{S-O2}) = 141(1)$ pm,

$\angle(\text{F-S-O1}) = 110.5(5)^\circ$. The second structure was obtained by Kuhn and coworkers for 2-fluoro-1,3-diisopropyl-4,5-dimethylimidazolium fluorosulfite.^{13,14} In this structure, the SO_2F^- anion is similarly disordered and, therefore, the authors considered a discussion of the SO_2F^- geometry inappropriate.¹³ In the most recent study by Kessler and Jansen,¹⁵ the room temperature structures of KSO_2F and RbSO_2F were given with $r(\text{SF}) = 159.1(2)$ pm, $r(\text{SO}) = 152.6(2)$ pm, $\angle(\text{OSO}) = 104.9(2)^\circ$, and $\angle(\text{OSF}) = 102.8(1)^\circ$ for the K^+ salt. Possible disorder was ignored and the obtained structural parameters were taken as the correct geometry of isolated SO_2F^- , concluding that the previously published,^{10,14} theoretically calculated structures were invalid. This conclusion, however, raises serious questions. Above all, how could a geometry which deviates by 18 pm for $r(\text{SF})$ and by 8° for $\angle(\text{OSO})$ from the true structure duplicate well the observed vibrational frequencies?

In this paper we wish to report the preparation of the $(\text{Me}_2\text{N})_3\text{S}^+$ (TAS),¹⁶ $(\text{Me}_2\text{N})_3\text{SO}^+$ (TAOS), $\text{N}(\text{CH}_3)_4^+$ (TMA),^{3,4} and K^+ salts of SO_2F^- and their low-temperature crystal-structures. It is shown that the true geometry of SO_2F^- is close to the calculated ones and that the varying degrees of deviation, exhibited in the crystal structures of the different salts, are due to the propensity of SO_2F^- to undergo oxygen-fluorine and, in some cases, oxygen-site disorders.

Experimental Section

Standard inert atmosphere techniques were used for the manipulation of all reagents and reaction products. Infrared spectra were recorded on a Nicolet DX-55-FT-IR spectrometer using Nujol/Kel-F mulls, NMR-spectra on a Bruker WP80SY and referenced to Me_4Si (^1H) and CFCI_3 (^{19}F). The NMR samples, dissolved in liquid SO_2 , were contained in sealed 5 mm glass tubes. Commercial grade solvents (MeCN , Et_2O) were dried and purified by employing conventional procedures.¹⁷ Commercial SO_2 was stored over P_4O_{10} at room temperature. TAOSF was obtained, similarly to TASF , from OSF_4 and $\text{Me}_3\text{SiNMe}_2$,¹⁸ TASF ^{16,19} and TMAF ²⁰ were prepared as described, and the fluorosulfites were synthesized according to the method reported for $\text{TAS}^+\text{SO}_2\text{F}^-$.¹⁶ The spectroscopic data of $\text{TAS}^+\text{SO}_2\text{F}^-$ (1) and $\text{TMA}^+\text{SO}_2\text{F}^-$ (3) agreed with those reported in the literature.^{3,4,16}

try to keep
10/10
 $\text{TAOS}^+\text{SO}_2\text{F}^-$ (2) was obtained by dissolving $\text{TAOS}^+\text{Me}_3\text{SiF}_2^-$ (1.7 mmol) in 10 mL SO_2 at $-30\text{ }^\circ\text{C}$. The solution was stirred for 30 min at this temperature, then all volatile material was removed under vacuum. 2 (0.45 g, 100% yield) remained as a colorless solid, mp $239\text{ }^\circ\text{C}$ (dec).
remove space

Salts 1–3 were recrystallised from $\text{MeCN}/\text{Et}_2\text{O}$ mixtures at $-40\text{ }^\circ\text{C}$ to give single crystals suitable for X-ray structure determinations.

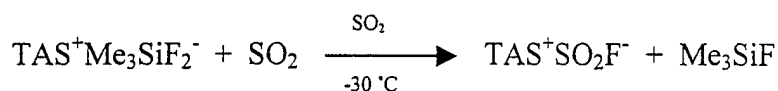
Crystal Structure Determinations. The crystals were mounted on thin glass fibers using the oil-drop technique (Kel-F oil). The data were collected on a Siemens P4 diffractometer using $\text{Mo-K}\alpha$ radiation ($\lambda = 71.073\text{ pm}$) at the given temperatures. The

structures were solved by direct methods. Hydrogen atoms were located from difference electron density maps and refined isotropically.

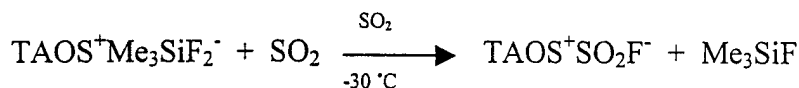
Computational Methods. The geometries and harmonic vibrational frequencies of SOF_2 , SO_2F_2 , and SO_2F^- were calculated using restricted Hartree-Fock (RHF), density functional theory (using the B3LYP hybrid functional²¹), and coupled-cluster methods (CCSD(T)²²). The 6-311+G(2d)²³ and the augmented correlation-consistent polarized valence double- and triple-zeta basis sets²⁴ were used. All calculations were performed using GAMESS²⁵ and GAUSSIAN 94.²⁶

Results and Discussion

Synthesis and Properties of $\text{TAOS}^+\text{SO}_2\text{F}^-$. TAS fluoride, $(\text{Me}_2\text{N})_3\text{S}^+\text{Me}_3\text{SiF}_2^-$,¹⁹ is an excellent fluoride ion donor which readily transfers a fluoride ion to the more acidic SO_2 molecule.¹⁶



Replacement of TASF by TAOSF, $(\text{Me}_2\text{N})_3\text{SO}^+\text{Me}_3\text{SiF}_2^-$, results in the corresponding $\text{TAOS}^+\text{SO}_2\text{F}^-$ salt in quantitative yield.



The new $\text{TAOS}^+\text{SO}_2\text{F}^-$ salt is a colorless solid that is stable at room temperature and melts at 239 °C with decomposition. It was characterized by its crystal structure (see below).

Crystal Structures of $\text{TAS}^+\text{SO}_2\text{F}^-$ (1), $\text{TAOS}^+\text{SO}_2\text{F}^-$ (2), $\text{TMA}^+\text{SO}_2\text{F}^-$ (3), and $\text{K}^+\text{SO}_2\text{F}^-$ (4). The crystal structures of 1–4 were determined at 173 °K and those of 2 and 4 also at 113 °K. The crystal and structure refinement data and the bond distances and angles of the SO_2F^- anions are given in Tables 1 and 2, respectively. The molecular units with atom labeling, the packing diagrams and the closest anion-cations interactions are shown in Figures ^{remove extra space} 1–7. As can be seen, compounds 1–4 are predominantly ionic, containing isolated SO_2F^- anions.

The refinement of the $\text{TAS}^+\text{SO}_2\text{F}^-$ (1) structure was not trivial. Initial structure refinements in the space group Pnma put the atoms S, F and O(2) on special positions located on a mirror plane, while O(1) was site-disordered occupying split positions off the plane with occupancy factors of one-half. In addition, the similarity of the S-O(2) and S-F bond lengths was also indicative of fluorine/oxygen disorder. This solution resulted in an apparent unusually short S-O(1) bond of ~140 pm, about 5 pm shorter than our theoretical predictions for $r(\text{S-O})$, with large thermal motions around the SO(2)F atoms. Since the theoretical calculations also predict that the S-F bonds in SO_2F^- are much longer than the S-O bonds, oxygen/fluorine disorder could only lengthen but not shorten the S-O bonds. The same argument applies to libration corrections. Therefore, this refinement model does not result in a plausible structure for SO_2F^- . The shape of the thermal ellipsoids of the atoms in the SO_2F^- anion suggested that these atoms might be disordered across the crystallographic mirror plane. Consequently, the positions of all

atoms of SO_2F^- were fixed off this plane, resulting in two sets of disordered anions with half occupancy (see Figure 1). The resulting refinement gave a plausible value of 146.5(3) pm for the S-O(2) bond distance, indicating that O(2) is not affected by oxygen/fluorine disorder, and resulted in slightly improved R factors. However, the O(1) and F positions are disordered and, therefore, the S-O and S-F bond distances and their bond angles could not be uniquely determined in this manner. The disorder was first modeled with equal O(1)/F occupancy factors. However, it was found that the occupancy factors cannot be equal because the two resulting S-O(1)/F bond lengths differed by 3.7 pm. The shortcomings of this model are also apparent from the sum of the OSO and OSF bond angles which total 331.6° instead of the theoretically predicted 314° . To obtain the correct occupancy factors which can account for the different S-O(1)/F bond lengths, the sums of the individual occupancy factors were restrained to their required totals (1.0 and 1.0), and the variable individual occupancy factors were refined to give the best fit with the observed data. The occupancy factors of 2-4 were refined in the same manner, and the results are summarized in Table 2.

The above analysis indicates that the disorder of the SO_2F^- anion involves a symmetry plane defined by the S and O(2) atoms and the free valence electron pair on S, with O(1) and F being located off this plane. Therefore, only O(1) and F are affected by this disorder and the actual bond length of the S-O bond is well determined by the observed S-O(2) distance.

The packing diagram of $\text{TAS}^+ \text{SO}_2\text{F}^-$ (Fig. 2) shows a three-dimensional network of intermolecular H...O/F contacts. Out of the four potential binding sites in the disordered SO_2F^- anion, the F/O atoms form a pseudo ten-membered ring (Fig. 1) with a

H1A...F/O distance of 250(3) pm. We use the term "pseudo ring", because one side of the ring structure is always left open due to the disorder-induced half-occupancy. These pseudo ring structures are then connected to the others in the same plane *via* the O2/O2A contacts i.e., H1C...O2 and H2C...O2 at 258(3) and 266(2) pm, respectively. This network is then linked to the other networks via bonding from O1/F1 at 253(3) pm.

The disorder of the SO_2F^- anion in the **TAOS salt** is less complicated than that in the TAS salt because it exhibits only the O(1)/F disorder with respect to the mirror plane formed by O(2), S, and the free valence electron pair on sulfur, but not the additional disorder with respect to the crystallographic mirror plane. Furthermore, from all the presently known SO_2F^- structures, that of the TAOS salt at 113 °K (Figures 3 and 4) exhibits the least amount of oxygen/fluorine disorder, resulting in the largest differences between the apparent S-F and S-O bond lengths and O-S-O and F-S-O bond angles. The degree of O/F disorder also decreases with decreasing temperature, as was established by carrying out the structure determinations at 173 and 113 °K. As in the case of the TAS salt, the S-O(2) bond, which lies on the molecular symmetry plane, is not affected by disorder, resulting in an S-O(2) bond length of 146.8(3) pm which is, within experimental error, identical to that found for the unique oxygen in the TAS salt and, therefore, should be representative of the true S-O bond length in ordered SO_2F^- . Furthermore, the sum of the OSO and OSF bond angles amounts to 313.9 ° which is in excellent agreement with the value of 314 °, predicted from the *ab initio* calculations. The facts that $r(\text{S-O}[\text{1}])$ and $r(\text{S-F})$ in the 113 °K structure of $\text{TAOS}^+\text{SO}_2\text{F}^-$, obtained by refinement assuming an ordered structure, are still somewhat longer and shorter, respectively, than the theoretical

predictions for free SO_2F^- and that $r(\text{S}-\text{O}(1))$ is 1.6 pm longer than $r(\text{S}-\text{O}(2))$, indicates that even in this case there is still some remaining disorder of O(1) and fluorine.

Due to the absence of strong disorder in $\text{TAOS}^+\text{SO}_2\text{F}^-$ at 113 °K, its crystal packing (Fig. 4) can be discussed more meaningfully. The CH_3 group of C3 bridges to two SO_2F^- anions via the $\text{F1}\dots\text{H3A}-\text{C3}-\text{H3C}\dots\text{O1}$ links, forming a zig-zag chain ($\text{F1}\dots\text{H3A} = 243(3)$ pm, $\text{O1}\dots\text{H3C} = 255(4)$ pm). These chains then form a 3-dimensional network by the formation of an intra-molecular bifurcated fluorine contact ($\text{H1A}\dots\text{F1}$, 246(4) pm). These three contacts are the shortest long-range contact distances²⁷ in the crystal lattice.

Refinement of the 173 °K structure of $\text{TAOS}^+\text{SO}_2\text{F}^-$, assuming no disorder, suffered from shortcomings similar to those previously reported in the literature for the $\text{Ph}_3\text{PCF}_2\text{H}^+$ salt¹² and provided no useful information.

In the $\text{TMA}^+\text{SO}_2\text{F}^-$ salt (see Figures 5 and 6 and Tables 1 and 2) at 173 °K, the SO_2F^- anion shows, as in the TAOS^+ salt, only O(1)/F disorder with respect to the mirror plane formed by O(2), S, and the free valence electron pair on sulfur, resulting again in a unique S-O(2) distance of 147.0(2) pm, which is almost identical to those found for the TAS^+ and TAOS^+ salts. Refinements, carried out assuming equal occupancy factors for O(1) and F, resulted again in different bond lengths for the two S-O(1)/F bonds, demonstrating the need for refining the occupancy factors.

The packing diagram for $\text{TMA}^+\text{SO}_2\text{F}^-$ is shown in Figure 6 and exhibits bifurcated bridges between O(2) and two hydrogen atoms from the same TMA cation, and between one of the O(1)/F atoms and hydrogen atoms from two different TMA cations, with the other O(1)/F atom forming a single bridge to a TMA cation. The bifurcated bridging is

similar to that observed for $\text{TAOS}^+\text{SO}_2\text{F}^-$, resulting in 16-membered rings that contain two cations and two anions and are three-dimensionally cross-linked to other rings. The two shortest of the many close distances are $\text{H2C}\dots\text{O1/F1}$ at 235(2) pm and $\text{H1B}\dots\text{O2}$ at 243(2) pm.

The **low-temperature (113 °K) structure of $\text{K}^+\text{SO}_2\text{F}^-$** is shown in Figure 7. This structure also exhibits oxygen/fluorine disorder but, contrary to the TAS, TAOS, and TMA salt structures, both oxygen atoms participate in the disorder. Therefore, this structure does not provide a unique S-O bond distance. The differences between the KSO_2F and the TAS, TAOS and TMA salt structures suggest a different kind of disorder. In the TAS, TAOS, and TMA salts, which exhibit significant anion-cation interactions through hydrogen bridges, the disorder is related to a symmetry plane through O(2), S, and the free valence electron pair on sulfur, while in the K^+ salt, which exhibits weaker anion-cation bridging, it appears to involve positional disorder around a pseudo three-fold axis along sulfur and its free valence electron pair. In the positionally disordered structure of KSO_2F , the occupancy factors for the three O/F positions are not equal and one bond is considerably longer than the other two, requiring again a refinement of the occupancy factors. Except for the expected temperature effects, our low-temperature structure of $\text{K}^+\text{SO}_2\text{F}^-$ is essentially identical to that previously found at room temperature,¹⁵ although the longer S-F/O bond is about 4 pm longer at the lower temperature, indicating a decrease in disorder with decreasing temperature.

As can be seen from Figure 7, the potassium ion is surrounded by 6 SO_2F^- anions. Out of these, 3 form bidentate and 3 form monodentate bridges to K^+ . Out of the 3 bidentate SO_2F^- groups, one is bridged through the two atoms with the higher oxygen

occupancy factors, designated as O/F; the other two are bridged through one atom with the higher oxygen occupancy factor (O/F) and one atom with the higher fluorine occupancy factor, designated as F/O. Out of the 3 monodentate SO_2F^- groups, 2 are bridged through O/F atoms and one through F/O atoms, resulting in an unusual coordination number of 9 around K^+ .

The above structure analyses resulted for the TAS, TAOS, and TMA salts in a well defined S-O(2) distance of about 147 pm that is in excellent agreement with the theoretical predictions (see below). However in all these structures, the observed S-F distances are too short due to O(1)/F disorder with varying occupancy factors. Since the observed bond lengths should be a function of the corresponding occupancy factors, one should be able to extrapolate to a 100 % occupancy factor and thus obtain an estimate for the individual S-F bond length in SO_2F^- . This was done in Figure 8, which shows that the length of the S-F bond should exceed 170 pm, in agreement with the theoretical predictions. The previous conclusion¹⁵ that the geometry derived from the strongly disordered room temperature structure of KSO_2F represents the first reliable geometry for SO_2F^- is obviously flawed. The pronounced tendency of SO_2F^- to undergo disorder can be attributed to its pyramidal shape, the similar size and electronegativity of its ligands, and packing arrangements that favor disorder.

Computational Results

Molecular vibrations occur on a very fast time scale and, therefore, are not affected by disorder phenomena in the crystal. Furthermore, the vibrational frequencies of a molecule are uniquely determined by its geometry. Therefore, good agreement

between observed and calculated frequencies can be obtained only with the correct geometry,²² and not, as previously implied,¹⁵ with an incorrect geometry. Good quality theoretical calculations generally yield relatively accurate bond angles.²² Depending on the level of theory and the quality of the basis set used, the bond distances are generally also accurate within a few pm. When comparing calculated and observed frequencies, it should be kept in mind that the calculated values are harmonic frequencies while the observed ones are anharmonic values, ^{frequencies? or values?} thus giving rise to small discrepancies. Furthermore, different physical states can cause minor deviations. The calculated values are for the isolated free species in the gas phase, while observed spectra are frequently for solids or liquids. However, these effects are generally small, and systematic over- or under- estimation of bond lengths and vibrational frequencies at certain levels of theory can be corrected by using scaling factors.

Reliable experimental vibrational spectra for SO_2F^- have recently been published by Kornath and coworkers and were shown to be in reasonable agreement with the values calculated at the uncorrelated RHF/6-31 +G* level of theory.¹⁰ The calculated geometry was also in good agreement with values¹⁴ previously obtained at the RHF/6-31+G(3df) and MP2/6-31+G(3df) levels of theory. We have calculated the geometry and vibrational frequencies of SO_2F^- at the uncorrelated RHF, the correlated CCSD(T), and the density functional B3LYP levels of theory using 6-311+G(2d),²¹ aug-cc-pvdz, and aug-cc-pvtz²² basis sets. To judge the reliability of these calculations, the geometries and vibrational frequencies of the closely related and well known SOF_2 ^{28,29} and SO_2F_2 ^{30,31} molecules were also calculated.

The results from these calculations are summarized in Tables 3 and 4, respectively, and show that the bond angles of free SO_2F^- change only little with the method of calculation and the basis set. Values of $\angle(\text{FSO}) = 100.5^\circ$ and $\angle(\text{OSO}) = 113^\circ$ should be close to the actual values. The calculated S-O bond length range in SO_2F^- is also narrow, and this bond is predicted to be about 147 pm which is in excellent agreement with the value of 146.8(3) pm observed in one of our crystal structure studies. The S-F bond length, however, is more sensitive to the level of theory and quality of the basis set chosen and is predicted to fall into the range of 170-176 pm, which is not too different from the minimum bond length of 165.7 pm, derived from our crystal structure studies. This relatively wide range is not surprising since the S-F bond in SO_2F^- is highly polar (see normal coordinate analysis), and the calculations for highly polar bonds are very sensitive to correlation and basis set polarization functions. Based on our predictions (see Table 3), the difference between $r(\text{SF})$ and $r(\text{SO})$ should range from 23 to 29 pm, which is in marked contrast to the 6 pm, required if the published¹⁵ crystal structures of KSO_2F and RbSO_2F were representative of the free SO_2F^- anion. Tables 3 and 4 also show that even the RHF calculations with a good basis set give surprisingly good results for these sulfur oxofluorides. For OSF_2 , which has the same C_s symmetry as SO_2F^- , only the best coupled cluster calculation, CCSD(T)/aug-cc-pvtz, gave a better result, i.e., a smaller average frequency deviation and smaller scaling factors, than the RHF/6-311+G(2d) calculation.

Additional evidence for the existence of a long and highly polar S-F bond in SO_2F^- comes from the structure¹¹ and vibrational spectra³² of isoelectronic ClO_2F . The observed difference between the bond lengths of Cl-F and Cl=O is 27.4 pm,¹¹ and the

structure is very similar to that predicted for free gaseous SO_2F^- (see Table 3). Further support for the geometry, predicted by us in Table 3 for free SO_2F^- , comes from a recent study of $\text{N}(\text{CH}_3)_4^+\text{SO}_2\text{CN}^-$.³³ The observed vibrational spectra were found to be in good agreement with those calculated at the RHF/6-31+G* level for a minimum energy structure with $r(\text{SO}) = 146.7$ pm, $r(\text{SC}) = 190.8$ pm, $\angle(\text{OSO}) = 114.0^\circ$, and $\angle(\text{OSC}) = 100.8^\circ$. These structural parameters are very similar to those predicted by us for SO_2F^- in Table 3.

Normal Coordinate Analysis

The assignment and force constant of the S-F stretching mode of SO_2F^- are of significant interest. If indeed the geometry proposed in reference 15 for SO_2F^- in its K^+ and Rb^+ salts with $r(\text{SF}) = 159$ pm were correct, its S-F stretching mode and force constant should be comparable to those found for SF_2 [$r_e(\text{SF}) = 159.2$ pm, $\nu_{\text{as}} \text{SF}_2 = 813.0 \text{ cm}^{-1}$, $\nu_{\text{sym}} \text{SF}_2 = 838.5 \text{ cm}^{-1}$, $f_r = 4.72 \text{ mdyne/\AA}$].^{34,35} If, on the other hand, $r(\text{SF})$ in SO_2F^- falls into the range of 170-176 pm, its stretching mode and force constant should be much lower.

The assignment of the S-F stretching mode in SO_2F^- has been controversial from the very beginning. Paetzold and Aurich published the infrared and Raman spectra of KSO_2F in 1965 and assigned the 595 cm^{-1} band to the SO_2 scissoring mode and the 496 cm^{-1} band to the S-F stretching mode.⁷ They attributed the surprisingly low S-F stretching frequency to a highly polar S-F bond and to fluorine bridging between the SO_2F^- ions. Shortly afterwards, Seel and Boudier reported the infrared spectra of KSO_2F , RbSO_2F and CsSO_2F . Since the frequencies of the bands at about 595 and 495 cm^{-1} changed only little for the different cations, they correctly concluded that the SO_2F^- anions should not

be significantly associated. In view of the lack of significant fluorine bridging which could have explained the low S-F stretching frequency, they proposed to assign the 595 cm^{-1} band to the S-F stretching mode.⁶ In a 1990 publication on $\text{NH}_4^+\text{SO}_2\text{F}^-$, Moock and coworkers followed⁹ the assignments of Paetzold and Aurich,⁷ whereas in 1997 Kornath and coworkers followed¹⁰ that of Seel and Boudier.⁶ In the most recent paper,¹⁵ Kessler and Jansen believed to have confirmed the assignments of Seel and Boudier, because they observed isolated SO_2F^- ions in the crystal structures of KSO_2F and RbSO_2F and falsely assumed that Kornath's *ab initio* calculations had established not only the observed frequencies, but also their identities. The identification of fundamental vibrations within a given symmetry species, however, requires the knowledge of the potential energy distribution (PED) obtainable through a standard normal coordinate analysis.

The results from our normal coordinate analysis of SO_2F^- are summarized in Table 5. The unscaled CCSD(T)/aug-cc-pvtz data were used because they duplicate the observed frequencies well and require by far the smallest scaling factors. The PED of Table 5 contains a big surprise. It shows that ν_2 , ν_3 , and ν_4 are all strongly mixed and that S2, the symmetry coordinate of the S-F stretch, contributes only 20 % to ν_2 , and 40 and 39 % to ν_3 and ν_4 , respectively. Thus, the S-F stretch is distributed over all three fundamental vibrations and is concentrated in ν_3 and ν_4 . An inspection of the signs in the PED reveals the following coupling effects. 580 cm^{-1} : in phase coupling of δ_{sym} FSO_2 (45%) with δ_{sciss} SO_2 (31%), and their out of phase coupling with νSF (20%); 506 cm^{-1} : in phase coupling of δ_{sciss} SO_2 (56%) with νSF (40%), and their out of phase coupling with δ_{sym} FSO_2 (3%); 365 cm^{-1} : in phase coupling of δ_{sym} FSO_2 (51%) with νSF (39%), and

their out of phase coupling with δ_{sciss} SO₂ (10%). This analysis demonstrates the irrelevance and fallacies of arguments over the assignment of certain modes without the benefits of a normal coordinate analysis. As so often, nature disregards our desire to paint simple black and white pictures of complex issues.

A second surprising result from the normal coordinate analysis is the very small value of 1.6 mdyn/Å for the SF stretching force constant. Normal, predominantly covalent S-F bonds exhibit values ranging from about 4.5 to 5.4 mdyn/Å.^{36,37} The low value of the SF stretching force constant demonstrates the high polarity of the S-F bond in SO₂F⁻. Furthermore, the 39% contribution of S-F stretching to the low frequency 365 cm⁻¹ mode provides a dissociative mode with a low activation energy barrier toward the loss of a fluoride ion.

Conclusions

- (i) The experimental crystal structures of the TMA⁺, TAS⁺, TAOS⁺ and K⁺ salts of SO₂F⁻ show that all the crystal structures reported so far for SO₂F⁻ suffer from severe oxygen/fluorine disorder and that the given geometries do not reflect the true structure of SO₂F⁻. In the structures of the TMA⁺, TAS⁺, and TAOS⁺ salts, one oxygen atom is ordered and its S-O distance of 147 pm is well defined. The S-F distances can be estimated to equal or exceed 170 pm by refining the occupancy factors in the O/F disorders and by an extrapolation of a correlation between the occupancy factors and the observed bond lengths to 100 % occupancy.

Supporting Information Available: Tables of structure determination summaries, atomic coordinates, bond lengths and angles, anisotropic displacement parameters, and hydrogen coordinates of (1), (2), (3) and (4) in CIF format. This material is available free of charge via the Internet at <http://pubs.acs.org>.

- (14) Maulitz, A. H.; Boese, R.; Kuhn, N. J. *J. Mol. Struct.* **1995**, 333, 227.
- (15) Kessler, U.; Jansen, M. *Z. Anorg. Allg. Chem.* **1999**, 625, 385.
- (16) Heilemann, W.; Mews, R. *Chem. Ber.* **1988**, 121, 461.
- (17) Perrin, D. D.; Armageo, W. L. F. *Purification of Laboratory Chemicals*, 3rd ed.; Pergamon, London, 1988.
- (18) Wessel, J., Dissertation, University of Bremen, Germany, 1995. Wessel, J.; Lork, E.; Behrens, U.; Borrmann, T.; Stohrer, W.-D.; Mews, R., to be published.
- (19) Middleton, W. J. U. S. Pat. 3 940 402 (1976), and *Org. Synth.* **1985**, 64, 221.
- (20) Christe, K. O.; Wilson, W. W.; Wilson, R. D.; Bau, R.; Feng, J. *J. Am. Chem. Soc.* **1990**, 112, 7619.
- (21) Becke, A.D. *J. Chem. Phys.* **1993**, 98, 5648.
- (22) See, for example: Bartlett, R. J.; Stanton, J. F., *Applications of Post-Hartree-Fock Methods: A Tutorial*, in *Reviews of Computational Chemistry*, Vol. V. Lipkowitz, A.; Boyd D. B.; eds. VCH Publishers, Inc. New York, 1994.
- (23) a) Krishnan, R.; Binkley J. S.; Seeger, R.; Pople J.A. *J. Chem. Phys.* **1980**, 72, 650.
- b) McLean, A. D.; Chandler, G. S. *J. Chem. Phys.* **1980**, 72, 5639.
- c) Frisch, M. J.; Pople, J. A.; Binkley, J. S. *J. Chem. Phys.* **1984**, 80, 3265.
- d) Clark, T.; Chandrasekhar, J.; Spitznagel, G. W.; Schleyer, P. von R. *J. Comput. Chem.* **1983**, 4, 294.
- (24) a) Dunning Jr., T. H. *J. Chem. Phys.* **1989**, 90, 1007.
- b) Kendall, R. A.; Dunning Jr., T. H.; Harrison R. J. *J. Chem. Phys.* **1992**, 96, 6796.
- c) Woon, D. E.; Dunning Jr., T. H. *J. Chem. Phys.* **1993**, 98, 1358.

- (25) Schmidt, M. W.; Baldrige, K. K.; Boatz, J. A.; Elbert, S. T.; Gordon, M. S.; Jensen, J. J.; Koseki, S.; Matsunaga, N.; Nguyen, K. A.; Su, S.; Windus, T. L.; Dupuis, M.; Montgomery, J. A. *J. Comput. Chem.* **1993**, *14*, 1347.
- (26) Gaussian 94, Revision D.4, Frisch, M. J.; Trucks, G. W.; Schlegel, H. B.; Gill, P. M. W.; Johnson, B. G.; Robb, M. A.; Cheeseman, J. R.; Keith, T.; Petersson, G. A.; Montgomery, J. A.; Raghavachari, K.; Al-Laham, M. A.; Zakrzewski, V. G.; Ortiz, J. V.; Foresman, J. B.; Cioslowski, J.; Stefanov, B. B.; Nanayakkara, A.; Challacombe, M.; Peng, C. Y.; Ayala, P. Y.; Chen, W.; Wong, M. W.; Andres, J. L.; Replogge, E. S.; Gomperts, R.; Martin, R. L.; Fox, D. J.; Binkley, J. S.; Defrees, D. J.; Baker, J.; Stewart, J. P.; Head-Gordon, M.; Gonzales, C.; Pople, J. A. Gaussian, Inc. Pittsburg PA, 1995
- (27) Bondi, A. *J. Phys. Chem.* **1964**, *68*, 441.
- (28) Hargittai, I.; Miglhoff, F. C. *J. Mol. Struct.* **1973**, *16*, 69.
- (29) Shimanouchi, T. *J. Phys. Chem. Ref. Data* **1977**, *6*, 1021.
- (30) Lide, D. R.; Mann, D. E.; Fristom, R. M. *J. Chem. Phys.* **1957**, *26*, 734.
- (31) a) Shimanouchi, T. *J. Phys. Chem. Ref. Data* **1977**, *6*, 1036. b) Sumodi, A. J.; Pace, E. L. *Spectrochim. Acta*, **1972**, *28A*, 1129. c) Nolin, C.; Tremblay, J.; Savoie, R. *J. Raman Spectrosc.* **1974**, *2*, 71. d) Sportouch, S.; Clark, R. J. H.; Gaufres, R. *J. Raman Spectrosc.* **1974**, *2*, 153.
- (32) Smith, D. F.; Begun, G. M.; Fletcher, W. H. *Spectrochim. Acta*, **1964**, *20*, 1763.
- (33) Kornath, A.; Blecher, O.; Ludwig, R. *J. Am. Chem. Soc.* **1999**, *121*, 4019.
- (34) a) Kirchhoff, W. H.; Johnson, D. R.; Powell, F. X. *J. Mol. Spectrosc.* **1973**, *48*, 1157. b) Endo, Y.; Saito, S.; Hirota, E.; Chikaraisli, T. *J. Mol. Spectrosc.* **1979**, *377*, 222.

- (35) Deroche, J. C.; Buerger, H.; Schulz, P.; Willner, H. *J. Mol. Spectrosc.* **1981**, *89*, 269.
- (36) Haas, A.; Willner, H. *Spectrochim. Acta* **1979**, *35A*, 953.
- (37) Siebert, H. "Anwendungen der Schwingungsspektroskopie in der Anorganischen Chemie; Anorganische und Allgemeine Chemie in Einzeldarstellungen, VII," Springer Verlag, Berlin, 1966.

Diagram Captions

Figure 1: Crystal structure, numbering scheme, and hydrogen bridging of $\text{TAS}^+\text{SO}_2\text{F}^-$ (1), showing the two disordered positions of SO_2F^- . The displacement ellipsoids are drawn at the 50 % probability level, and the hydrogen atoms were located from difference electron density maps.

Figure 2: Packing diagram and hydrogen bridging of $\text{TAS}^+\text{SO}_2\text{F}^-$ (1).

Figure 3: Crystal structure, numbering scheme, and hydrogen bridging of $\text{TAOS}^+\text{SO}_2\text{F}^-$ (2) at 113 °K. The displacement ellipsoids are drawn at the 50 % probability level, and the hydrogen atoms were located from difference electron density maps.

Figure 4: Packing diagram and hydrogen bridging of $\text{TAOS}^+\text{SO}_2\text{F}^-$.

Figure 5: Crystal structure and numbering scheme of $\text{TMA}^+\text{SO}_2\text{F}^-$ (3). The displacement ellipsoids are drawn at the 50 % probability level, and the hydrogen atoms were located from difference electron density maps.

Figure 6: Packing diagram and hydrogen bridging of $\text{TMA}^+\text{SO}_2\text{F}^-$.

Figure 7: Crystal structure at 113 °K and numbering scheme of $\text{K}^+\text{SO}_2\text{F}^-$ (4), showing the coordination of 6 SO_2F^- anions around the K^+ cation.

Figure 8: Plot of the observed S-O/F and S-F/O bond lengths of (1)-(4) in pm as a function of their O/F occupancy factors.

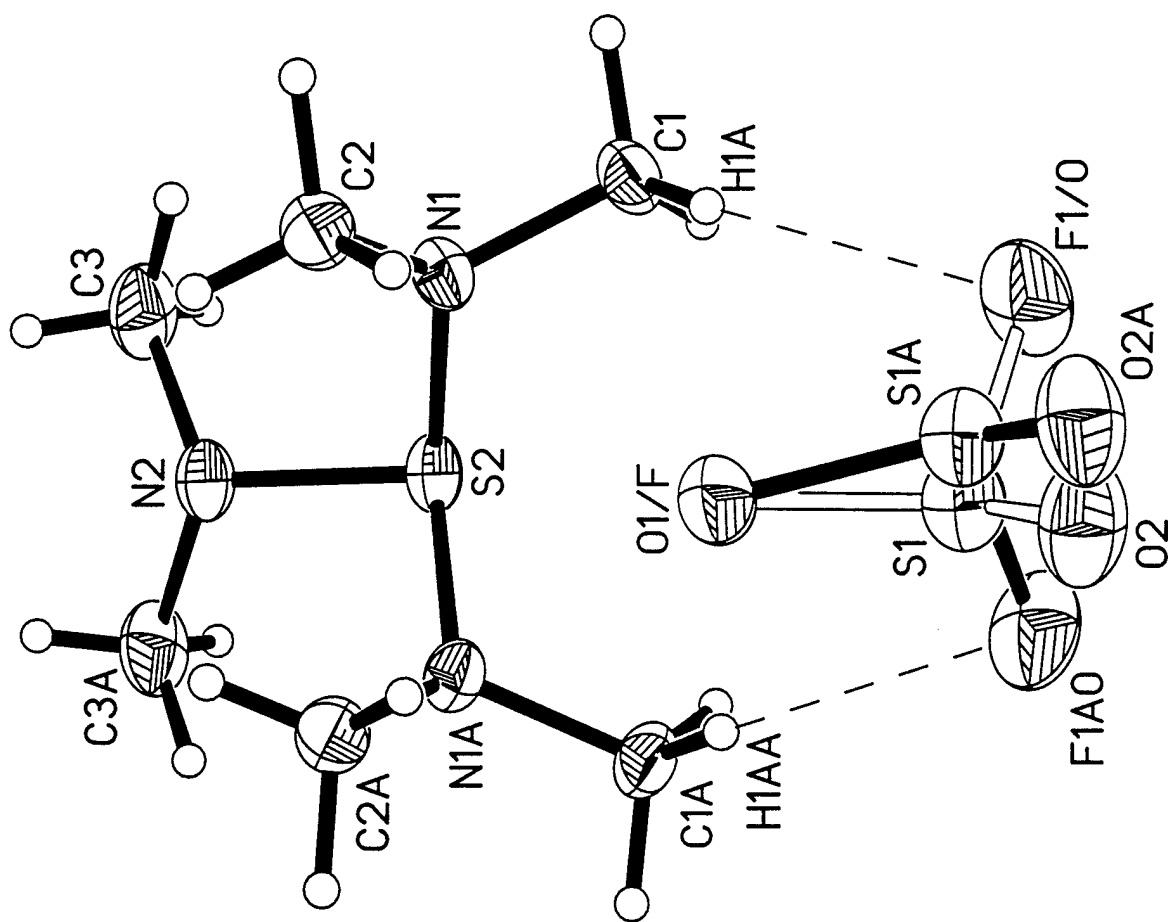
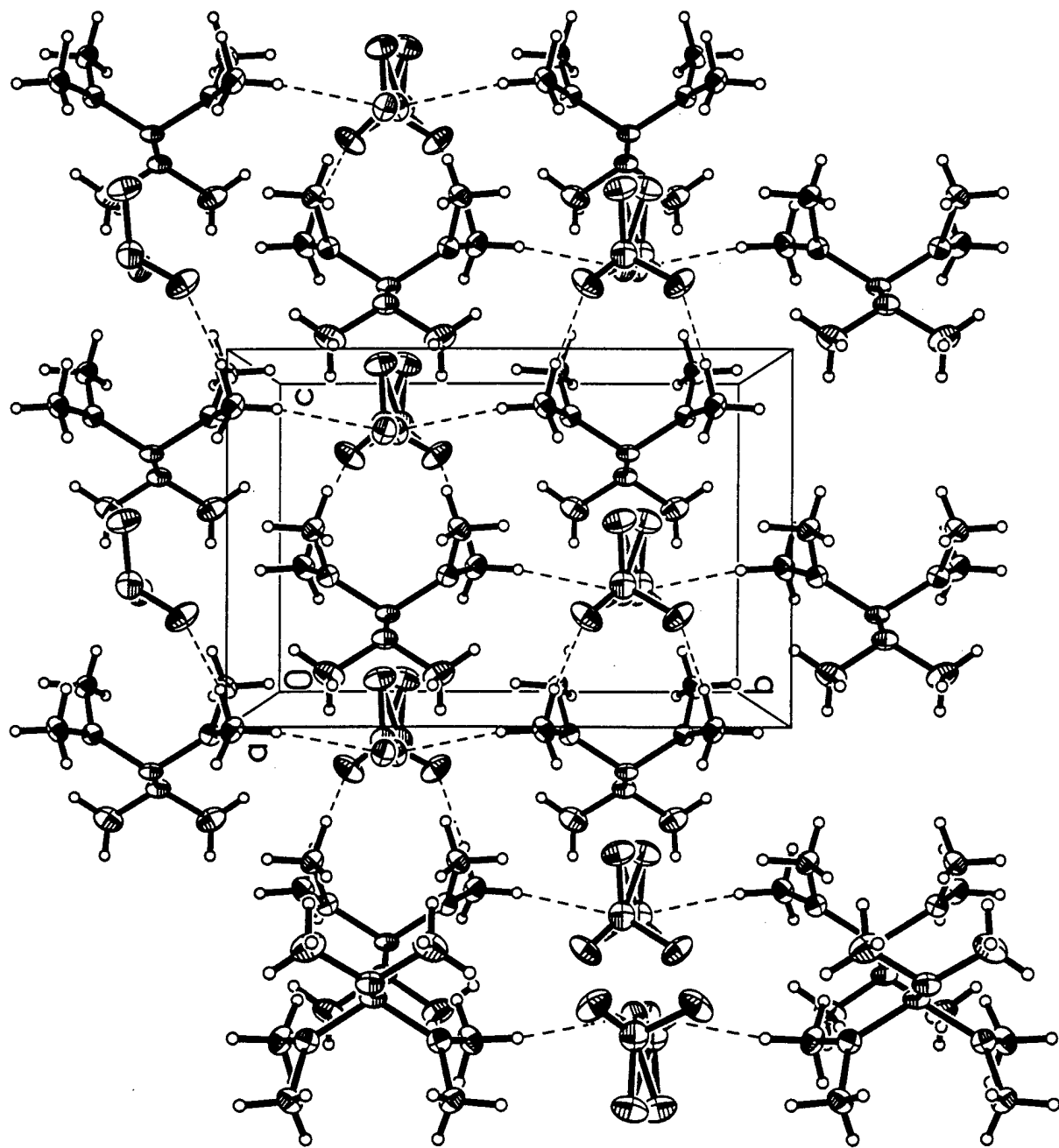
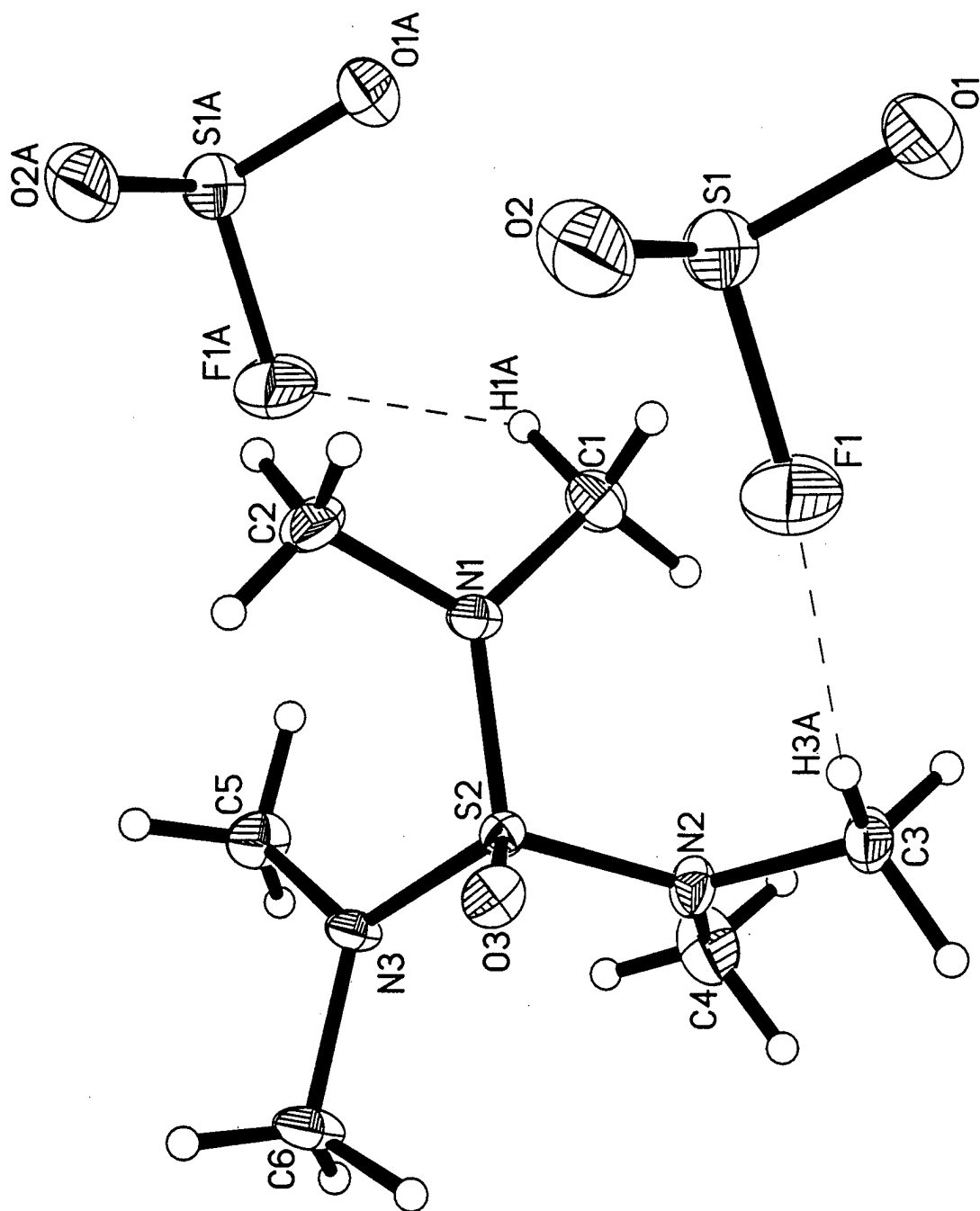
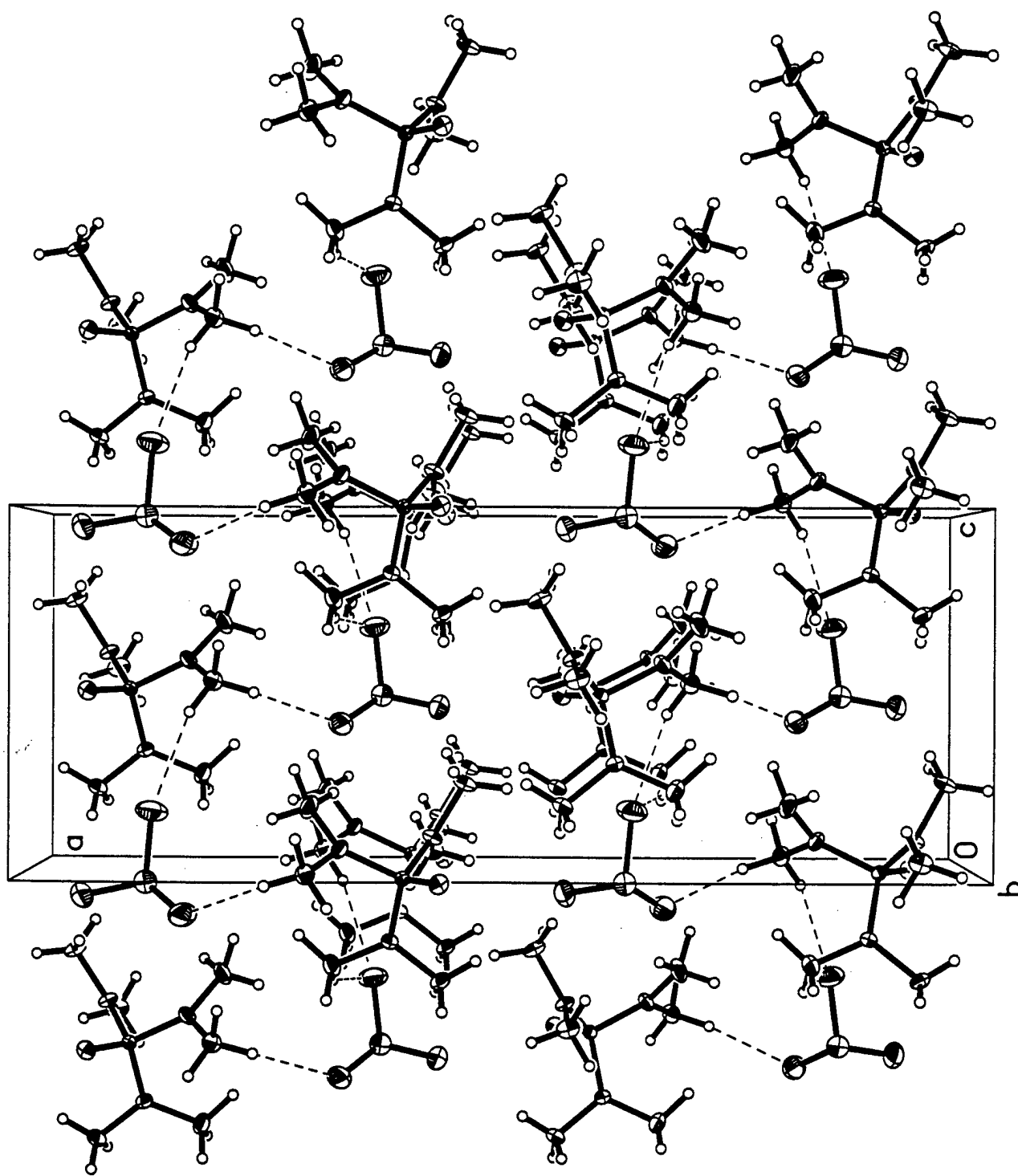


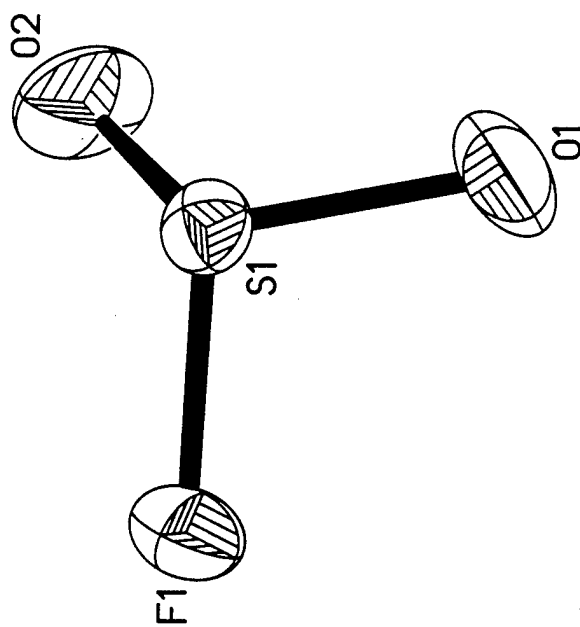
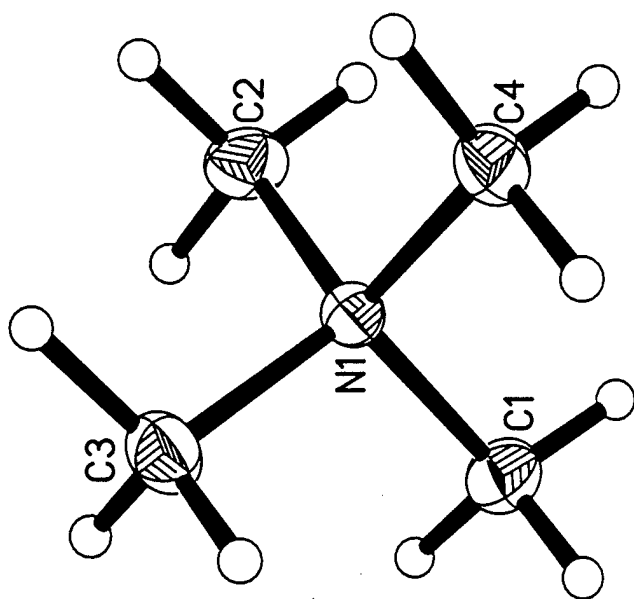
FIG. 1 TAC

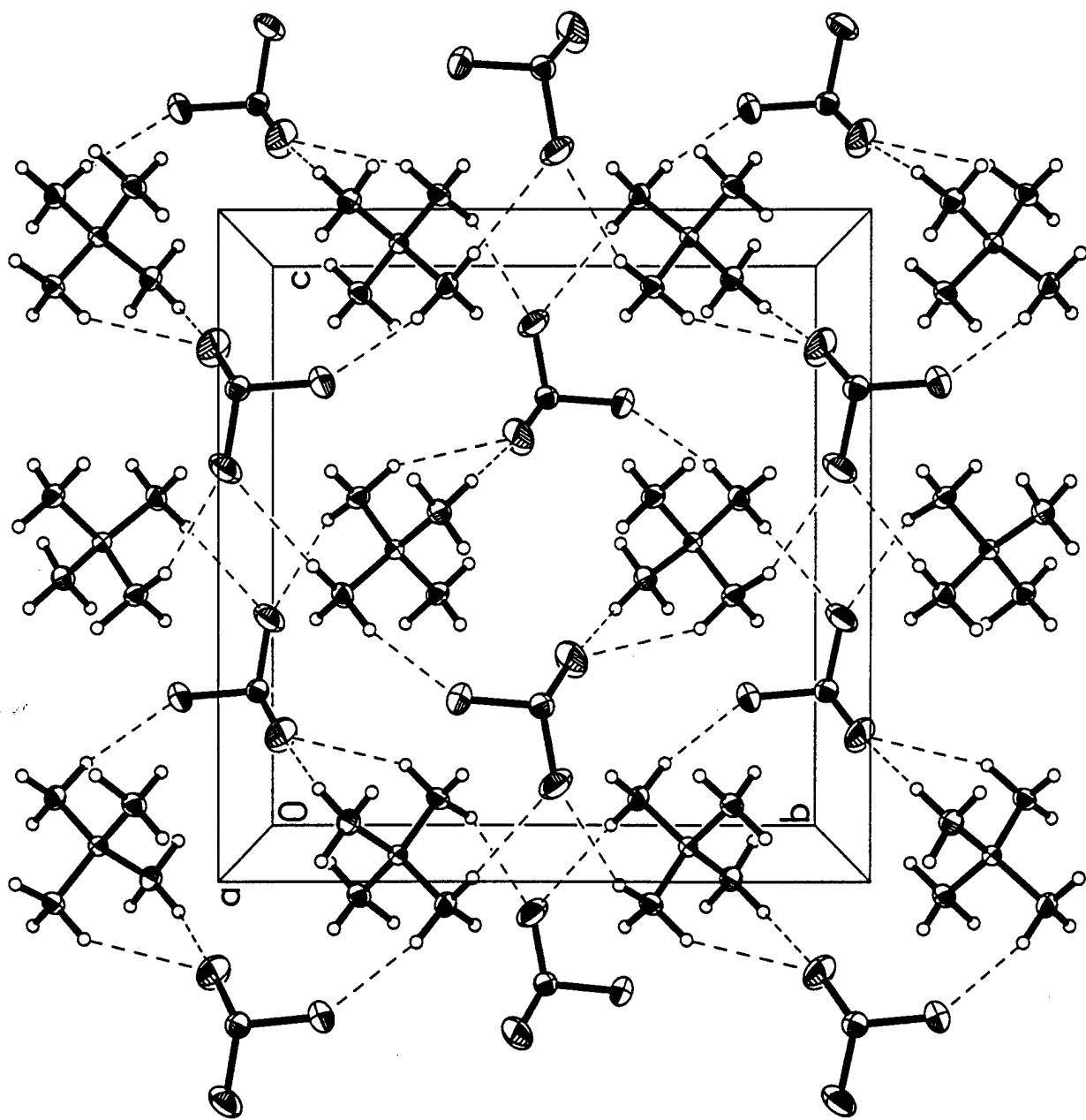
70. 70. 70. 70. 70.











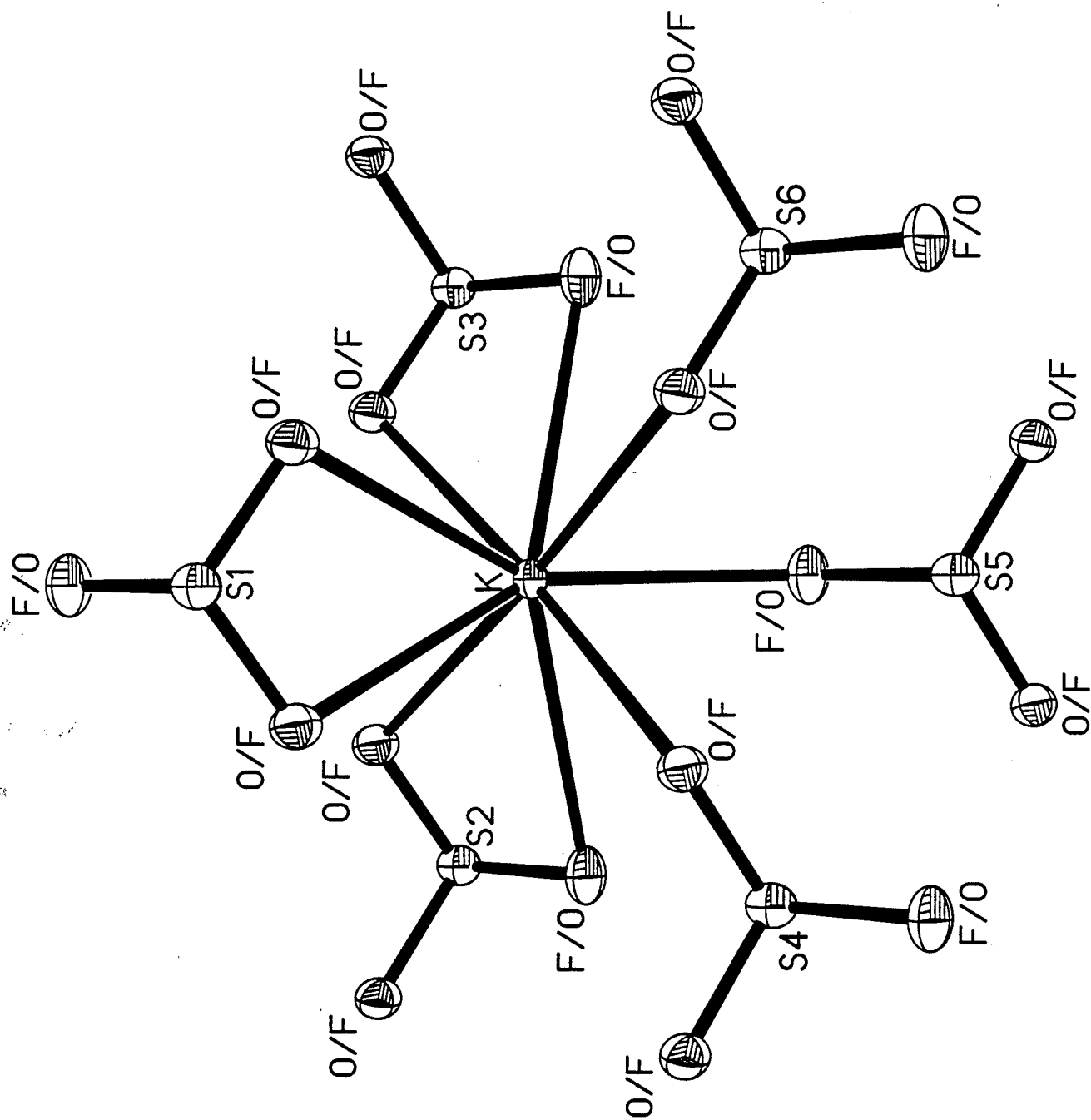


Figure 7 K501

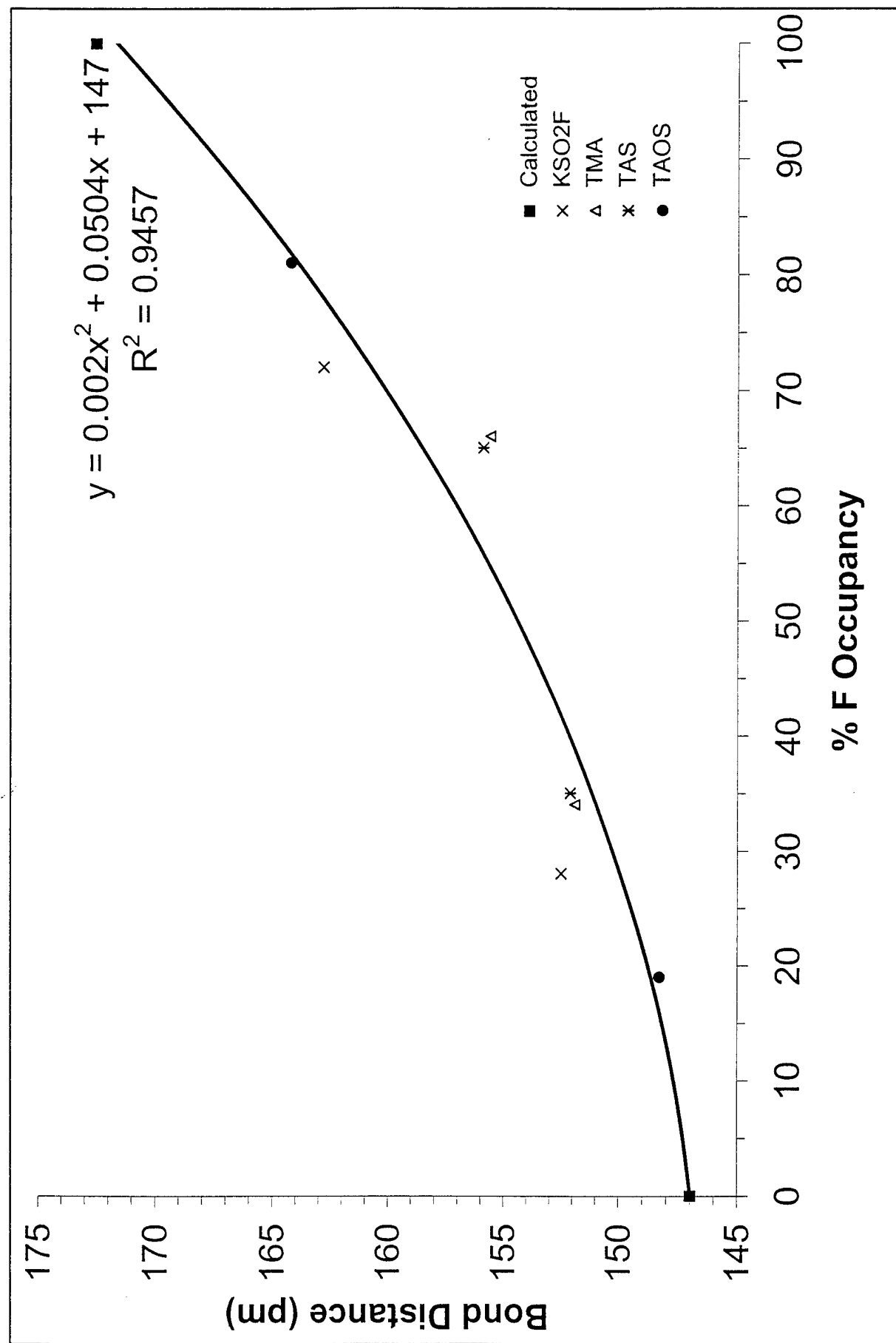


Table 1. Crystal data and structure refinement for (1) - (4)

	1	2	3	4
Empirical formula	C ₆ H ₁₈ FN ₃ O ₃ S ₂	C ₆ H ₁₈ FN ₃ O ₃ S ₂	C ₄ H ₁₂ FN ₂ O ₂ S	FKO ₂ S
M	247.35	263.35	157.21	122.16
T (K)	173(2)	113(2)	173(2)	113(2)
Crystal system	orthorhombic	orthorhombic	orthorhombic	monoclinic
Space group	<i>Pnma</i>	<i>Pna2₁</i>	<i>Pbca</i>	<i>P2₁/m</i>
a (pm)	1469.0(3)	2185.0(2)	1152.0(1)	461.3(1)
b (pm)	1117.4(2)	673.3(1)	1150.5(1)	564.0(1)
c (pm)	733.4(1)	819.4(1)	1162.7(2)	683.0(1)
β (°)				107.06(1)
V (nm ³)	1.2038(4)	1.2055(3)	1.5410(3)	0.16988(5)
Z	4	4	8	2
D _c (Mgm ⁻³)	1.365	1.451	1.355	2.388
Reflections collected	1993	3019	2357	691
Independent	1444	2771	1763	427
R(int)	0.0283	0.0810	0.0468	0.0166
Parameters	122	212	135	33
μ(MoKα), mm ⁻¹	0.438	0.449	0.375	2.000
R ₁	0.0430	0.0434	0.0517	0.0263
wR ₂	0.0961	0.1045	0.1275	0.0748
Flack's parameter		0.01(9)		
Difference electron density e.Å ⁻³	0.412 / -0.257	0.461 / -0.547	0.448 / -0.594	0.600 / -0.407

Table 2. Bond distances and angles of the fluorosulfite anion $\text{FSO}(1)\text{O}(2)^-$ in different salts compared to those calculated for the free gaseous anion. The data for (5)-(8) are literature values which were not corrected O/F mixed occupancies and therefore, do not represent the true structures.

Compound	SO(1)	SO(2) [pm]	SF	O(1)SO(2)	O(1)SF	[°]	O(2)SF	Temp. [K] ^a	ref
1 ^b	152.1(3) [O _F = 0.35]	146.2(3) [O _O = 1.0]	155.9(2) [O _O = 0.35]	109.8(2)	108.8(1)		113.0(1)	173	
2	148.3(3) [O _F = 0.19]	146.8(3) [O _O = 1.0]	164.2(2) [O _O = 0.19]	108.8(2)	102.6(2)		102.4(2)	113	
3	151.9(2) [O _F = 0.34]	147.0(2) [O _O = 1.0]	155.6(2) [O _O = 0.34]	106.9(2)	102.7(1)		104.2(1)	173	
4	152.5(2) [O _F = 0.28]	152.5(2) ^f [O _O = 0.28]	162.8(2) [O _O = 0.28]	105.9(1)	101.9(7)		101.9(7) ^f	113	
Predicted ^g for free gaseous ion	147	147	170-175	113	100.5		100.5		
KSO ₂ F (5)	152.6(2)	152.6(2)	159.1(2)	104.9(2)	102.8(1)		102.8(1)	294	15 ^e
RbSO ₂ F (6)	153.0(6)	153.6(6)	159.8(8)	105.7(5)	102.6(3)		102.6(3)	294	15 ^e
Ph ₃ PCF ₂ H ^c (7)	143.6(0)	141.1(1)	151.6(6)	-	110.5(5)		-	293	12 ^e
F-imid ^d (8)	148.9(11)	146.8(14)	149.6(16)	108.7(7)	104.4(8)		107.4(8)	190	13,14 ^e

^a) Temperature of data collection; ^b) The anion is disordered across a mirror plane, the occupancies of mixed O(1) and F positions (O_O = oxygen occupancy and O_F = fluorine occupancy) are refined; ^c) Disordered; ^d) F-imid = 2-fluoro-1,3-diisopropyl-4,5-imethylimidazolium cation. ^e) Data uncorrected for mixed O/F occupancies; ^f) symmetry generated: x, -y+1/2, z; ^g) Predicted values based on the calculations given in Table 3.

Table 3. Calculated geometries^a of SO₂F⁻ compared to the calculated and observed geometries of SOF₂ and SO₂F₂ and the observed geometry of ClO₂F

SOF ₂	R1(SO)	R2-R1	R2(SF)	<(FSO)	<(FSF)	ref
RHF/6-31 + G*	140.9	16.7	157.6	106.6	92.5	10
RHF/6-311+G(2d)	139.5	15.9	155.4	106.5	92.6	
RHF/aug-cc-pvdz	142.6	16.6	159.2	106.3	92.3	
RHF/aug-cc-pvtz	140.1	14.7	154.8	106.3	93.0	
B3LYP/6-311+G(2d)	142.9	20.4	163.3	106.7	92.8	
B3LYP/aug-cc-pvdz	146.1	20.5	166.6	106.5	92.8	
B3LYP/aug-cc-pvtz	143.4	18.8	162.2	106.4	93.0	
CCSD(T)/6-311+G(2d)	143.1	18.8	161.9	106.7	92.3	
CCSD(T)/aug-cc-pvdz	146.9	19.3	166.2	106.4	92.3	
CCSD(T)/aug-cc-pvtz	143.6	16.7	160.3	106.3	92.5	
obsd.	142.0(3)	16.3	158.3(3)	106.2(2)	92.2(2)	28
SO ₂ F ₂	R1(SO)	R2-R1	R2(SF)	<(OSO)	<(FSF)	ref
RHF/6-311 + G(2d)	138.0	12.6	150.6	124.6	95.6	
RHF/aug-cc-pvdz	141.3	13.3	154.6	125.2	94.8	
RHF/aug-cc-pvtz	138.7	11.7	150.4	124.1	95.9	
B3LYP/6-311+G(2d)	141.6	15.7	157.3	125.6	94.9	
B3LYP/aug-cc-pvdz	145.2	16.4	161.6	126.0	94.3	
B3LYP/aug-cc-pvtz	142.2	14.8	157.0	125.2	95.1	
CCSD(T)/6-311+G(2d)	141.6	14.4	156.0	125.5	94.9	
CCSD(T)/aug-cc-pvdz	145.7	15.4	161.1	126.3	94.0	
CCSD(T)/aug-cc-pvtz	142.1	13.3	155.4	125.3	95.1	
obsd.	140.5	12.5	153.0	124.0	96.1	30
SO ₂ F ⁻	R1(SO)	R2-R1	R2(SF)	<(FSO)	<(OSO)	ref
RHF/6-31 + G*	145.8	24.0	169.8	100.6	113.2	10
RHF/6-31 + G(3df)	143.7	22.7	166.4	100.5	112.9	14
RHF/6-311 + G(2d)	144.1	19.8	163.9	100.4	113.1	
RHF/aug-cc-pvdz	147.8	23.4	171.2	100.6	112.6	
RHF/aug-cc-pvtz	144.8	21.1	165.9	100.5	112.8	
MP2/6-31+G(3df)	147.1	30.9	178.0	100.3	113.8	14
B3LYP/6-311+G(2d)	147.8	37.2	185.0	100.8	113.4	
B3LYP/aug-cc-pvdz	151.7	32.7	184.4	101.1	112.6	
B3LYP/aug-cc-pvtz	148.3	31.9	180.2	100.7	113.0	
CCSD(T)/6-311+G(2d)	147.9	34.5	182.4	100.5	113.6	
CCSD(T)/aug-cc-pvdz	152.4	30.2	182.6	100.7	112.8	
CCSD(T)/aug-cc-pvtz	148.4	27.3	175.7	100.3	115.9	
Predicted	147	23-29	170-176	100.5	113	
ClO ₂ F	R1(ClO)	R2-R1	R2(ClF)	<FCIO	<OCIO	ref
	142.0	27.4	169.4	101.8	115.2	11

a) Bond lengths in pm and angles in degrees.

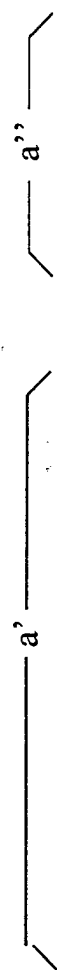
Table 4. Calculated scaled (unscaled) and observed vibrational frequencies of SO₂F₂, SOF₂ and SO₂F

	 a ₁ 			a ₂ 		b ₁ 		b ₂ 		average
SO ₂ F ₂ (C _{2v})	V ₁	V ₂	V ₃	V ₄	V ₅	V ₆	V ₇	V ₈	V ₉	Δ(V _{obsd} - V _{calcd}) ref
Observed	1270	849	553	384	384	1504	544	887	540	31
RHF/6-311+G(2d)	1259(1384)	869(955)	557(612)	384(422)	379(417)	1473(1619)	547(595)	897(986)	541(601)	9.4
RHF/aug-cc-pvdz	1234(1305)	881(932)	561(569)	383(389)	374(380)	1444(1528)	555(551)	913(966)	543(563)	21.5
RHF/aug-cc-pvtz	1246(1393)	876(979)	550(615)	385(426)	380(421)	1453(1624)	558(601)	904(1010)	542(606)	15.9
IR and RA Int. ^a	212.5;13.0, .09	128.2;11.5,0	47.9;1.7,.73	.12;.64;.75	0; .99	384.3; 1.72	44.7; 1.21	276.5; 1.81	42.1; 1.29	
B3LYP/6-311+G(2d)	1304(1226)	823(774)	557(516)	383(355)	383(355)	1553(1460)	546(506)	859(808)	542(502)	16.3
B3LYP/aug-cc-pvdz	1281(1162)	841(763)	560(477)	380(324)	378(322)	1518(1377)	551(466)	875(794)	547(469)	8.4
B3LYP/aug-cc-pvtz	1287(1231)	837(800)	557(520)	384(359)	383(358)	1527(1460)	547(511)	870(832)	541(505)	7.5
CCSD(T)/6-311+G(2d)	1289(1238)	833(800)	555(532)	387(372)	382(366)	1531(1471)	544(521)	876(841)	538(515)	9.0
CCSD(T)/aug-cc-pvdz	1262(1159)	848(779)	558(484)	385(334)	377(327)	1498(1376)	548(474)	893(820)	546(475)	4.9

^a IR and Raman intensities in km mol⁻¹ and Å⁴ amu⁻¹, respectively, and polarization ratios.



SOF ₂ (C ₂)	V ₁	V ₂	V ₃	V ₄	V ₅	V ₆	$\Delta(V_{\text{obsd}} - V_{\text{calcd}})$	ref
Observed	1333	808	530	378	747	393		29
RHF/6-311+G(2d)	1321(1457)	811(894)	533(601)	380(428)	751(828)	389(439)	4.7	
RHF/aug-cc-pvdz	1269(1389)	822(900)	536(565)	376(396)	773(846)	391(412)	19	
RHF/aug-cc-pvtz	1288(1466)	819(932)	531(611)	377(434)	764(870)	390(449)	11	
B3LYP/6-311+G(2d)	1420(1313)	794(734)	535(486)	369(335)	716(662)	399(362)	25	
B3LYP/aug-cc-pvdz	1348(1261)	804(752)	540(465)	364(313)	743(695)	401(345)	5.0	
B3LYP/aug-cc-pvtz	1370(1318)	802(771)	535(501)	370(346)	733(705)	398(373)	12	
CCSD(T)/6-311+G(2d)	1390(1307)	798(750)	532(500)	376(354)	729(685)	392(369)	12.7	
CCSD(T)/aug-cc-pvdz	1313(1241)	809(765)	536(469)	369(324)	757(716)	399(349)	8.7	
CCSD(T)/aug-cc-pvtz	1330(1322)	808(803)	531(517)	378(368)	749(745)	392(381)	1.2	



$\text{SO}_2\text{F}^- (\text{C}_s)$	V_1	V_2	V_3	V_4	V_5	V_6	$\Delta(V_{\text{obsd}} - V_{\text{calcd}})$	ref
Observed	1108	590	497	387	1184	363		10
RHF/6-31+G*	1086	605	536	394	1174	362	15.7	10
RHF/6-311+G(2d)	1098(1219)	597(663)	511(567)	377(418)	1184(1314)	362(402)	6.8	
RHF/aug-cc-pvdz	1088(1163)	611(653)	529(584)	381(421)	1166(1247)	347(383)	18.8	
RHF/aug-cc-pvtz	1088(1227)	608(686)	514(612)	384(458)	1170(1319)	354(422)	13.5	
B3LYP/6-311+G(2d)	1114(1079)	571(553)	554(472)	364(310)	1218(1179)	350(298)	25.3	
B3LYP/aug-cc-pvdz	1117(1034)	578(535)	558(479)	371(318)	1200(1111)	340(292)	22.8	
B3LYP/aug-cc-pvtz	1113(1086)	577(563)	540(490)	370(336)	1206(1177)	351(318)	18.7	
CCSD(T)/6-311+G(2d)	1105(1076)	576(561)	538(475)	367(324)	1215(1183)	355(313)	19.5	
CCSD(T)/aug-cc-pvdz	1102(1017)	589(543)	550(485)	364(330)	1193(1101)	342(302)	17.2	
CCSD(T)/aug-cc-pvtz	1102(1092)	585(580)	525(506)	378(365)	1201(1190)	352(340)	12.7	

Table 5. Unscaled CCSD(T)/aug-cc-pvtz symmetry force constants and potential energy distribution of SO₂F

	Calcd. Freq., cm ⁻¹	Symmetry force constants ^a				PED ^b
		F ₁₁	F ₁₂	F ₃₃	F ₄₄	
a' V ₁	1092	F ₁₁				
		8.520				
V ₂	580	F ₂₂	1.626			96.6(1) 2.9(3) 0.6(4)
V ₃	506	F ₃₃	-0.076	1.913		45.0(4) 31.3(3) 20.4(2) 3.4(1)
V ₄	365	F ₄₄				56.1(3) 40.4(2) 3.4(4)
		-0.007	0.014	0.596	1.924	50.9(4) 39.3(2) 9.8(3)
		F ₅₅	F ₆₆			
a'' V ₅	1190	F ₅₅				99.1(5) 0.8(6)
V ₆	340	F ₆₆	1.049			99.1(6) 0.8(5)
f _{iso} = 8.180						
f _{SF} = 1.626						

^a Stretching force constants in mdyn/Å, deformation constants in mdynÅ/rad², and stretch-bend interaction constants in mdyn/rad.

^b PED in percent; the symmetry coordinates are defined as: (1) V_{sym}SO₂, (2) V SF, (3) δ_{sciss} SO₂, (4) (3) δ_{sym}SO₂, (5) V_{asym}SO₂, (6) δ_{asym} O₂SF