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RIGOROUS MATHEMATICAL MODELING OF ADSORPTION SYSTEM WITH ELECTROTHERMAL REGENERATION OF THE USED ADSORBENT

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Final Performance Report for Phase 3

**3-D COMSOL Multiphysics™ modeling of an adsorber with four parallel
annular, radial-flow cartridges**

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PROJECT OVERVIEW

OBJECTIVES

The general objective of the project is fundamental mathematical modeling of a complex TSA system with electrothermal desorption step, with adsorbers assembled of one or more cartridge-type, radial flow fixed beds and with in-vessel condensation of the desorbed vapor.

During the first stage of the project, a 1-D mathematical model of a single cartridge without condensation was developed in Matlab and used for simulation. During the second stage of the project, the complete TSA system, for adsorbers with one and two cartridges, was modeled, using a combination of Femlab and Matlab software. 2-D Femlab models that describe adsorption, electrothermal desorption and electrothermal desorption accompanied with condensation of the desorbed vapor were built and integrated using Matlab, in order to describe and simulate complete TSA cycles.

The objective of this, third phase of the project, is to develop mathematical models of another complex configuration, with four parallel annular, radial-flow cartridge-type adsorbers. The analysis of this configuration shows that 3-D models have to be used in order to describe it. The models for adsorption, electrothermal desorption and electrothermal desorption accompanied with condensation of the desorbed vapor are developed again, by using COMSOL Multiphysics (Femlab) 3-D modeling tools.

STATUS OF EFFORT

Three COMSOL Multiphysics 3-D models were built, corresponding the three stages of the TSA process with electrothermal desorption taking place in a 4-cartridge adsorber: adsorption, desorption without condensation and desorption accompanied with condensation at the adsorber wall. The models incorporate seven application modes defined in a number of subdomains. This results with high demands regarding the computer memory and computer time needed for their solution. In order to reduce these demands, the models used for simulations were defined for one quarter of the 4-cartridge adsorber, assuming that all four cartridges are identical. Also, the solution of our models did not converge directly. A three-step iterative solution procedure was developed, leading to the solution of the complete 3-D models. Following this procedure the three stages of the TSA process were simulated for a chosen set of parameters.

ACCOMPLISHMENTS /NEW FINDINGS

The accomplishments of this project can be perceived from two aspects:

1. The developed models represent a mathematical tool which can be used for simulation of the investigated TSA system, its analysis, predicting of the concentration and temperature profiles in the system and its optimization. This was the main objective of this project.
2. The developed solution procedure in COMSOL Multiphysics can be used for solving of other complex 3-D models.

PERSONNEL SUPPORTED

Danijela Antov-Bozalo, Graduate student

Nikola Nikačević, Graduate student

Dr. Menka Petkovska, Professor

PUBLICATIONS

1. M. Petkovska, D. Antov and P. Sullivan, "Electrothermal desorption in an annular - radial flow – ACFC adsorber – Mathematical modeling", *Adsorption*, **11**, 585-590, 2005. (also presented at the FOA-8, 8th International Conference on Fundamentals of Adsorption (FOA8), Sedona, USA, May 23-28, 2004)
2. Petkovska M., Antov-Bozalo D., Marković A. i Sullivan P., "Femlab/Matlab model for complex electric swing adsorption (ESA) system with in-vessel condensation", *Femlab Conference*, Oktobar 2005, Stockholm, Proceedings, 117-122
3. Petkovska M., Antov-Bozalo D., Marković A. i Sullivan P. "Electrical swing adsorption (ESA) regenerative filter modeling and simulation", *2005 Scientific Conference for Chemical and Biological Defense Research*, Novembar 2005, Timonium, SAD, Proceedings
4. Antov-Bozalo D., Sullivan P., Petkovska M., "Adsorption and Electrothermal Desorption in Annular - Radial Flow Adsorber – Mathematical Modeling using FEMLAB", 1st South East European Congress of Chemical Engineering – SEECHE 1, Sept. 2005, Book of Abstracts, 63

In preparation: Two journal and two conference papers

NEW DISCOVERYS, INVENTIONS AND PATENTS:

None

DETAILED REPORT

1. Introduction

The idea about regeneration of used adsorbents by direct heating of the adsorbent particles by passing electric current through them (Joule effect), was first published in the 1970s [1]. Desorption process based on this principle was later named electrothermal desorption [2]. It was recognized as a prospective way to perform desorption steps of TSA cycles. At the same time, fibrous activated carbon was recognized as a very convenient adsorbent form for its realization. Electrothermal desorption has some advantages over conventional methods, regarding adsorption kinetics and dynamics [3, 4] and energy efficiency [2, 5]. Some industrial applications of electrothermal desorption have been reported recently [6, 7]. Nevertheless, the development of mathematical description of processes based on electrothermal desorption hasn't been following the development of the process itself, so far.

A new TSA process with electrothermal desorption step, based on adsorbers assembled of one or more annular, cartridge-type, fixed-beds, with in-vessel condensation, has been presented recently [5, 8, 9]. Its schematic representation is given in Figure 1.

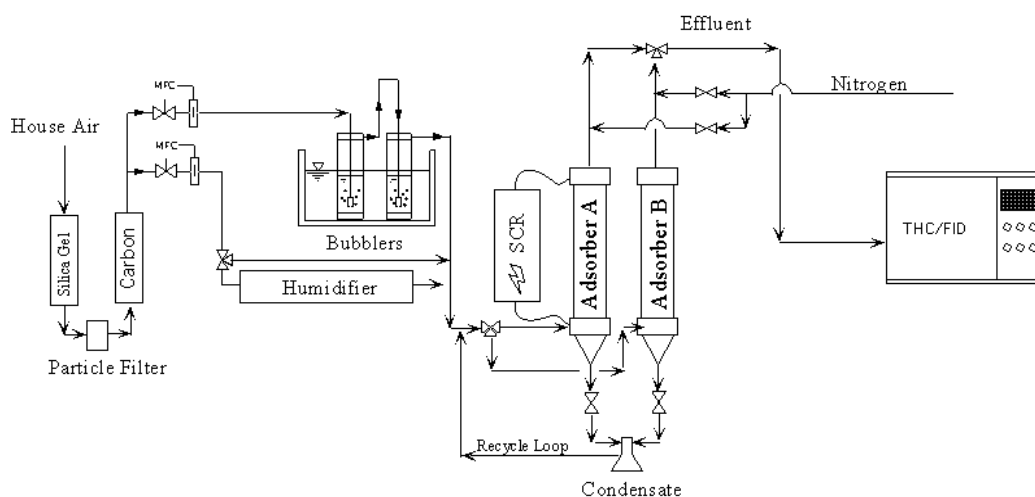


Figure 1. Overall schematic representation of the ACFC adsorption – rapid electrothermal desorption system (from Ref. [5])

The aim of this project is to develop rigorous mathematical models of this system for different adsorber configurations. These models could be used for simulation, analysis and optimization of the TSA process developed by Sullivan [5].

1.1. Description of the system

The adsorbers of the TSA system shown in Figure 1 are composed of one or more annular, radial-flow cartridge-type adsorbent beds. A single cartridge is schematically shown in Figure 2. It is formed as a cylindrical roll of activated carbon fiber cloth (ACFC), spirally coiled around a porous central pipe. The gas flow through the adsorber is in the radial direction. During the desorption step, electric current is passed through the activated carbon cloth in the axial direction, causing heat generation, heating of the adsorbent and desorption.

The cartridges are situated in a cylindrical adsorber vessel. During desorption, the hot gas stream rich with the desorbed vapor comes into contact with the cold vessel wall and the vapor condenses. In that way, the adsorber vessel serves as a passive condenser. The bottom outlet of the adsorber vessel is funnel-shaped, in order to facilitate draining of the condensed liquid.

The structure of the adsorbers with one, two and four cartridges is shown in Figure 3. In principle, the flow directions during the adsorption and desorption steps can be different, and the flow-rates during desorption are considerably lower than during adsorption. The flow directions shown in Figure 3 correspond to the desorption step.

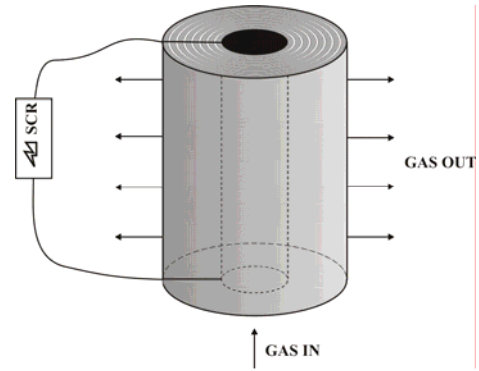


Figure 2. Schematic representation of the annular - radial flow - ACFC adsorbent bed

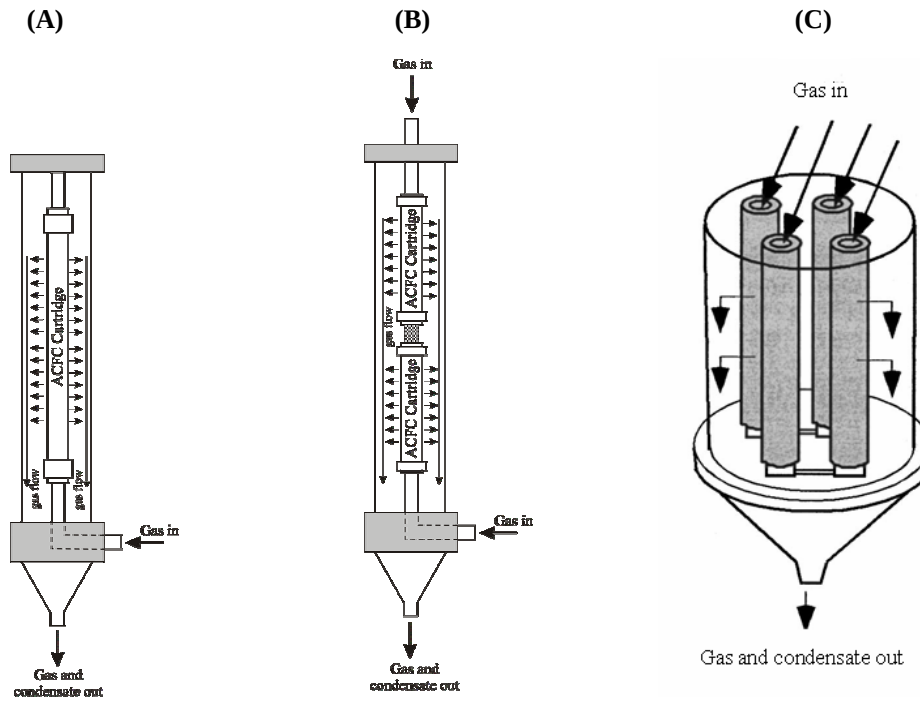


Figure 3. Adsorbers with one (A), two (B) and four (C) annular, radial-flow cartridges

1.2. Brief survey of the results of the first two phases of the project

Our efforts during the first phase of the project were focused on modeling of a single-cartridge adsorber, without condensation. Two deterministic models were postulated: the first one assuming uniform adsorbent density throughout the adsorbent bed, and the second one taking into account the structure of the bed made of layers of ACFC. Both models were obtained as complex sets of coupled nonlinear PDEs, ODEs and algebraic equations. Numerical procedures for solving the model equations were established and used for simulation of adsorption, electroresistive heating, electrothermal desorption and consecutive adsorption-desorption, using Matlab software. Similar results were obtained with the two models. The results of this phase are presented in the Final Performance Report for project FA8655-03-1-3010 [10].

In the second phase of the project, models for two types of adsorbers, one with only one (Figure 3A), and the other with two cartridges (Figure 3B), have been developed. For each adsorber type, three models were built, in order to describe three stages of a complete TSA cycle: adsorption, electrothermal desorption and electrothermal desorption with in-vessel condensation. These models were built using Femlab, a specialized software tool for modeling of complex systems with complex geometry. The models were defined in 2-D axially symmetrical geometry, with seven application modes. In order to describe the complete TSA cycles, the models for the three stages of the TSA process were integrated, by using a combination of Femlab and Matlab. The models were successfully used for simulation of separate stages of the process and of the complete TSA cycles, as well as for their optimization. The results of this phase are presented in the Final Performance Report for project FA8655-04-1-3053 [11].

1.3. Outline of the third phase of the project

The third phase of the project was focused on mathematical modeling of the four-cartridge adsorber, presented in Figure 3C. This configuration has been described in detail in Refs. 10 and 12. This modeling is performed using COMSOL Multiphysics software (previously Femlab). Owing to the complex configuration of this adsorber design, 3-D models have to be used, leading to a number of issues regarding computer memory requirements and convergence problems.

Similarly to the 2-D models developed in the second phase, three 3-D models are needed, to describe three steps of the complete cycle:

- adsorption,
- desorption without condensation,
- desorption accompanied with condensation.

The following application modes were incorporated in these models: Non-isothermal flow, Brinkman Equations, Convection and Diffusion, Diffusion, Convection and Conduction, Heat transfer by Conduction and Conductive Media DC.

Solving of these very complex 3-D models is a rather demanding and difficult task. It assumes optimization of the model structure and geometry definition, the generated mesh and the choice of the solver, in regard to the used hardware, computer time and convergence of the solution.

All these issues were attended in this phase of the project.

2. Modeling of the four-cartridge adsorber, using COMSOL Multiphysics

Three 3-D models were built, for the three steps of the complete TSA cycle:

- Model_A3 – for adsorption
- Model_D3 – for desorption without condensation
- Model_DC3 – for desorption with condensation

Model assumptions:

The same assumptions as for the 2-D models derived in the second phase of the project [11] were used in setting up the 3-D models:

- The adsorbent beds are treated as homogeneous, with uniform adsorbent porosity and density.
- The mass and heat transfer resistances on the particle scale are neglected.
- The fluid phase is treated as an ideal gas mixture of the inert and the adsorbate.
- All physical parameters and coefficients are considered as constants.
- The electric resistivity of the ACFC adsorbent is temperature dependent. Linear temperature dependence is assumed, based on experimental results reported in Ref. [5].
- The electric power during electrothermal desorption is supplied under constant voltage conditions.
- The condensation at the adsorber wall is dropwise. This assumption is based on experimental observations reported in Ref. [5].
- The volume of the condensed liquid is neglected, i.e., it is assumed that the liquid drops don't influence the gas flow.
- The heat resistance and heat capacity of the adsorber wall are neglected, so that the wall temperature is equal to the temperature of the environment.
- Initially, the adsorbate concentrations in both phases are in equilibrium and uniform throughout the adsorbent bed. The temperatures of both phases are initially equal and uniform throughout the adsorbent bed.

2.1. Definition of the model geometry

In the second phase of the project we built Femlab models for adsorbers with one and two cartridges and we used 2-D axially symmetrical space [11]. Nevertheless, the four-cartridge adsorber shown in Figure 3C is not axially symmetrical, so it has to be modeled in 3-D space. In COMSOL Multiphysics, the only option for 3-D modeling is to use the Descartes coordinate system (x, y, z).

The gas-flow and electric current directions in the four-cartridge adsorber, as defined in Ref. [12], are shown in Figure 4. This figure shows that the gas-flow directions differ during adsorption and desorption. During adsorption, two-by-two cartridges work in parallel: one pair of cartridges serves for introducing the feed stream, and the other for removal of the gas stream from the adsorber. The two pairs of cartridges are connected in series. From the modeling point of view, the cartridges used for introducing the gas stream into the adsorber are identical, as well as the cartridges used for withdrawing the gas stream from the adsorber, but the two pairs of cartridges are different. On the other hand, during desorption, all four cartridges work in parallel, with gas inlet from the top ends of the central tubes and gas outlet through the bottom end of the outer tube in which all four cartridges are situated (the adsorber vessel). The condensed liquid during desorption leaves the adsorber through the same outlet. From the modeling point of view, all four cartridges are identical, regarding momentum, mass and heat transfer. The electric current flows are also shown in Figure 4. For the design shown

in this figure, all four cartridges are connected in series. From the modeling point of view, all four cartridges are different, regarding the electric current flow.

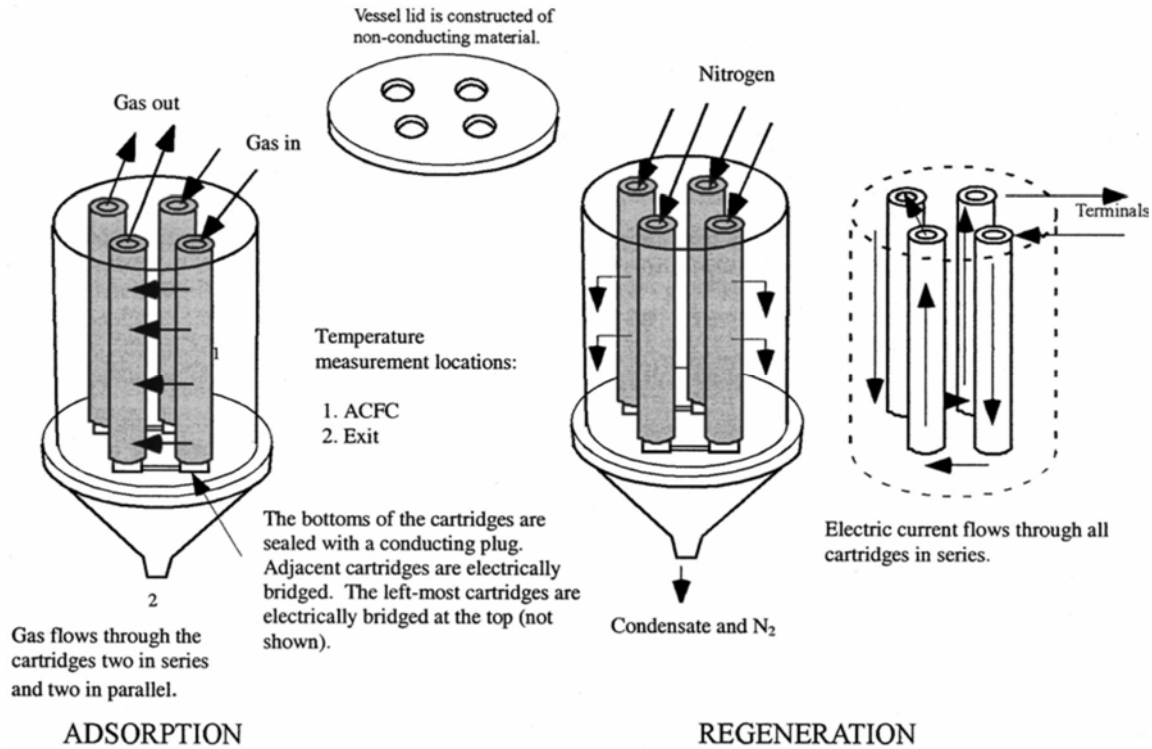


Figure 4. Four-cartridge adsorber [12]

In principle, the complete adsorber geometry with four cartridges should be used, in order to model the adsorber in which all four cartridges are different (e.g. because of different electric current flows through them). Nevertheless, 3-D models are very complex and very demanding regarding computer memory and computer time. For that reason, we explored two other options: to model a four-cartridge adsorber in which two cartridges can be considered as equal, and a four-cartridge adsorber in which all four cartridges are equal.

2.1.1. A 3-D model of the whole adsorber

The most general and most complex geometry is the one representing the whole adsorber with 4 cartridges. It is shown in Figure 5, in 3-D perspective and in the 2-D x-y plane. The complete adsorber consists of 9 sub-domains (shown in the 2-D view in Fig. 5):

- Sub-domains EXT1, EXT2, EXT3 and EXT4 correspond to the adsorbent beds of the 4 cartridges;
- Sub-domain EXT5 corresponds to the adsorber space around the 4 cartridges;
- Sub-domains EXT6, EXT7, EXT8 and EXT9 correspond to the central tubes of the 4 cartridges.

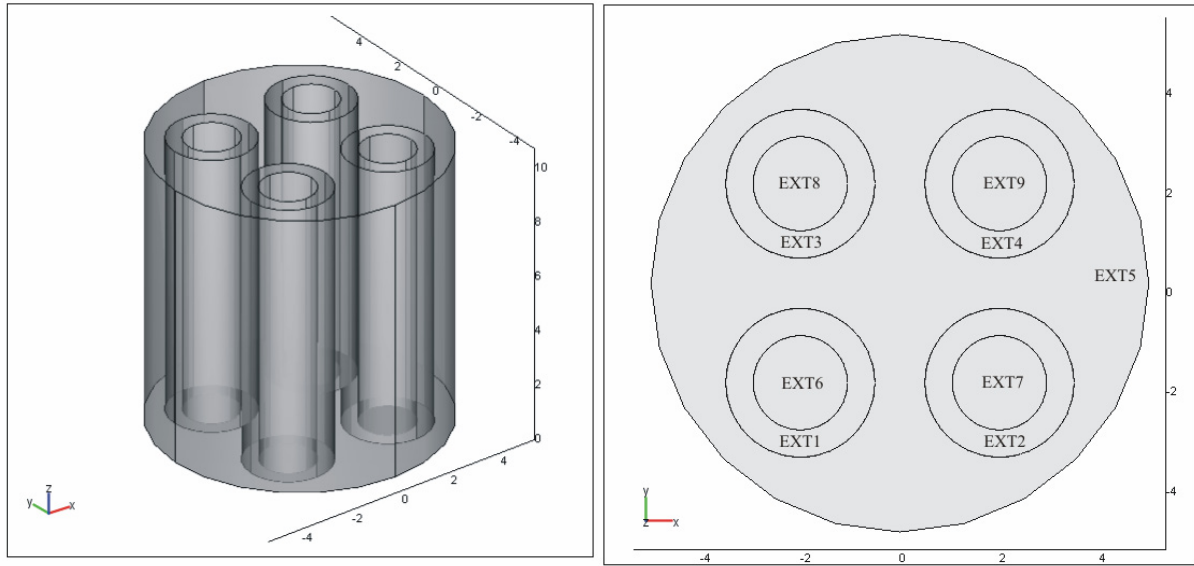


Figure 5. Geometry definition for the whole four-cartridge adsorber: 3-D view (left) and 2-D x-y view (right)

If “coarser mesh” is chosen as a mesh parameter, with all other default parameter settings, the mesh corresponding to this geometry contains **24481** 3-D elements.

2.1.2. A 3-D model of one half of the adsorber

If the adsorber contains two pairs of identical cartridges, the model geometry could be simplified by defining only one half of the adsorber and assuming that the other half is completely symmetrical to it. This geometry is shown in Figure 6.

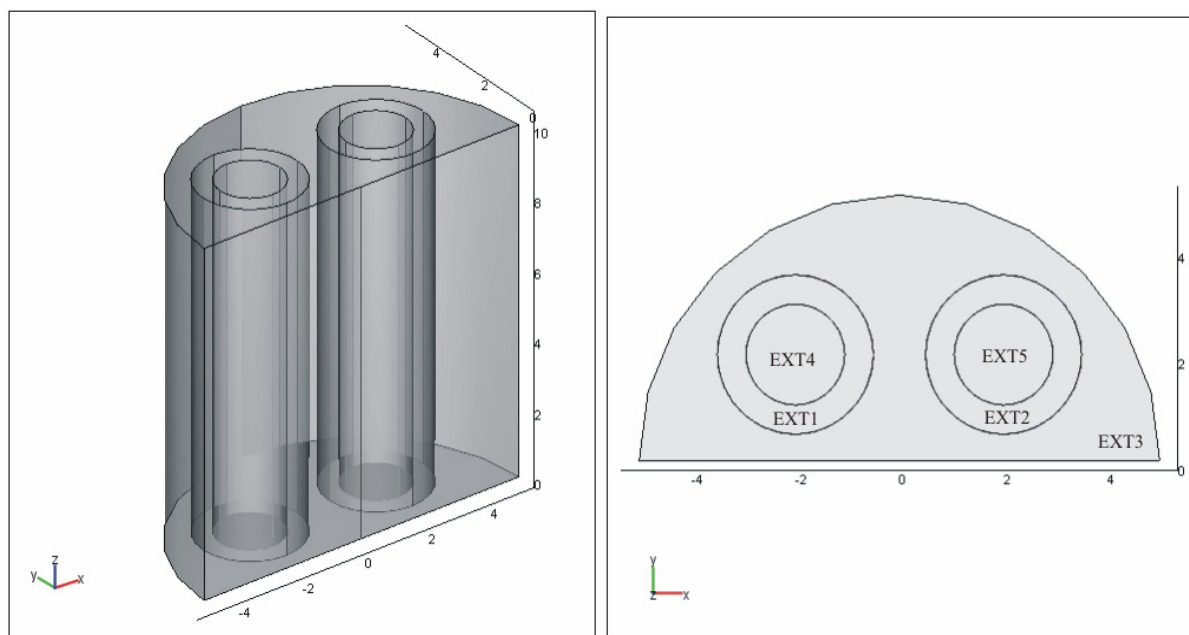


Figure 6. Geometry definition for one half of the four-cartridge adsorber: 3-D view (left) and 2-D x-y view (right)

For this geometry, the number of sub-domains reduces to 5 (shown in the 2-D view in Fig. 6):

- Sub-domains EXT1 and EXT2 correspond to the adsorbent beds of the two cartridges;
- Sub-domain EXT3 corresponds to the space around the cartridges;
- Sub-domains EXT4 and EXT5 correspond to the central tubes of the two cartridges.

For “coarser mesh” setting, the mesh for this geometry consists of **12118** 3-D elements.

2.1.3. A 3-D model of one quarter of the adsorber

If all four cartridges can be considered as identical, the model geometry can be further simplified, by defining only one quarter of the adsorber and assuming symmetry both in the x - and in the y -direction. This geometry is shown in Figure 7 (the main dimensions of the adsorber are also shown).

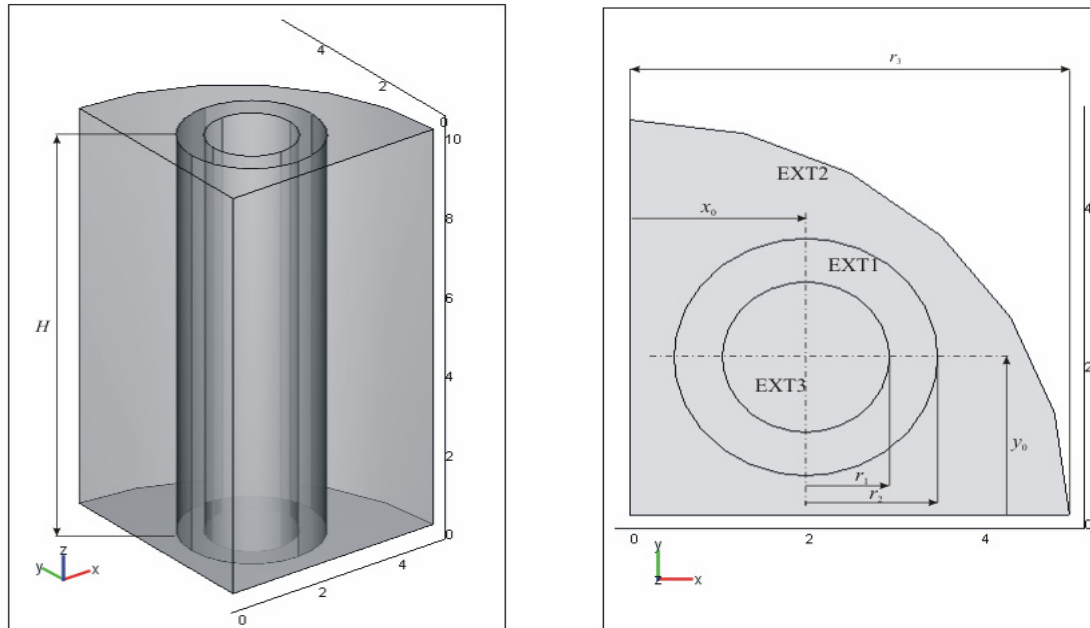


Figure 7. Geometry definition for one quarter of the four-cartridge adsorber: 3-D view (left) and 2-D x - y view (right)

In this case, three sub-domains have to be defined (shown in the 2-D view in Fig. 7):

- Sub-domain EXT1 – the adsorbent bed;
- Sub-domain EXT2 – the space around the cartridges;
- Sub-domain EXT3 – the central tube of the cartridge.

The mesh for this geometry, for “coarser mesh” setting, contains **6126** elements.

All three geometries shown, in Figures 5, 6 and 7, were defined for a laboratory scale adsorber of the following dimensions:

- $H=10$ cm (height of the adsorber)
- $r_1=0.95$ cm (radius of the central tubes)
- $r_2= 1.5$ cm (outer radius of the adsorbent beds)
- $r_3=5$ cm (radius of the adsorber vessel)
- $x_0=y_0=2$ cm (position of the cartridge center in the x and y directions)

2.2. The application modes used

For modeling of the four-cartridge adsorber, the same built-in application modes are used, as for the one- or two-cartridge adsorbers (Ref. 11). For the most complex process – electrothermal desorption, in which simultaneous gas flow throw tubes and adsorbent beds, mass and heat transfer and heat generation by the Joule effect are coupled, seven application modes have to be incorporated in the COMSOL Multiphysics model:

1. *Non-isothermal flow* – for defining the momentum balances in the central tubes and outer space;
2. *Brinkman equations* – for defining the momentum balances for the adsorbent beds;
3. *Convection and Conduction* – for defining the heat balances for the gas phase;
4. *Heat transfer by Conduction* – for defining the heat balances for the solid phase;
5. *Conductive Media DC* – for defining heat generation by Joule effect;
6. *Convection and Diffusion* – for defining the mass balances for the gas phase;
7. *Diffusion* – for defining the mass balances for the solid phase.

The application modes are active in the sub-domains according to this scheme:

- *Convection and Conduction* and *Convection and Diffusion* are active in all sub-domains;
- *Non-isothermal flow* is active in sub-domains corresponding to the central tubes of the cartridges and the space around the cartridges;
- *Brinkman equations*, *Heat transfer by Conduction*, *Conductive Media DC* and *Diffusion* are active only in the sub-domains corresponding to the adsorbent beds.

2.3 The underlying equations

In principle, the phenomena taking place in the four-cartridge TSA system are the same as in the one- and two-cartridge systems. Nevertheless, as a result of using 3-D Descartes instead of 2-D radial coordinate system, the model equations are not exactly the same. Although the model equations and their boundary conditions for different steps of the TSA process are somewhat different, as electrothermal desorption involves all phenomena and the corresponding application modes, here is an overview of the PDEs underlying the FEMLAB model for electrothermal desorption:

1) Momentum balances and continuity equations for the inert gas:

For the central tubes and the space around the cartridges:

$$\rho_g \frac{\partial u}{\partial t} + \nabla \left[-2\mu \frac{\partial u}{\partial x} - \mu \left(\frac{\partial u}{\partial y} + \frac{\partial v}{\partial x} \right) - \mu \left(\frac{\partial u}{\partial z} + \frac{\partial w}{\partial x} \right) \right] = - \left(\rho_g \left(u \frac{\partial u}{\partial x} + v \frac{\partial u}{\partial y} + w \frac{\partial u}{\partial z} \right) + \frac{\partial p}{\partial x} \right) \quad (1)$$

$$\rho_g \frac{\partial v}{\partial t} + \nabla \left[-\mu \left(\frac{\partial v}{\partial x} + \frac{\partial u}{\partial y} \right) - 2\mu \frac{\partial v}{\partial y} - \mu \left(\frac{\partial v}{\partial z} + \frac{\partial w}{\partial y} \right) \right] = - \left(\rho_g \left(u \frac{\partial v}{\partial x} + v \frac{\partial v}{\partial y} + w \frac{\partial v}{\partial z} \right) + \frac{\partial p}{\partial y} \right) \quad (2)$$

$$\rho_g \frac{\partial w}{\partial t} + \nabla \left[-\mu \left(\frac{\partial w}{\partial x} + \frac{\partial u}{\partial z} \right) - \mu \left(\frac{\partial w}{\partial y} + \frac{\partial v}{\partial z} \right) - 2\mu \frac{\partial w}{\partial z} \right] = - \left(\rho_g \left(u \frac{\partial w}{\partial x} + v \frac{\partial w}{\partial y} + w \frac{\partial w}{\partial z} \right) + \frac{\partial p}{\partial z} \right) \quad (3)$$

$$- \left(\rho_g \left(\frac{\partial u}{\partial x} + \frac{\partial v}{\partial y} + \frac{\partial w}{\partial z} \right) + \frac{\partial \rho_g}{\partial x} u + \frac{\partial \rho_g}{\partial y} v + \frac{\partial \rho_g}{\partial z} w \right) = 0 \quad (4)$$

For the adsorbent beds:

$$\rho_g \frac{\partial u}{\partial t} + \nabla \left[-2\mu \frac{\partial u}{\partial x} - \mu \left(\frac{\partial u}{\partial y} + \frac{\partial v}{\partial x} \right) - \mu \left(\frac{\partial u}{\partial z} + \frac{\partial w}{\partial x} \right) \right] = - \left(\frac{\mu}{k} u + \frac{\partial p}{\partial x} \right) \quad (5)$$

$$\rho_g \frac{\partial v}{\partial t} + \nabla \left[-\mu \left(\frac{\partial v}{\partial x} + \frac{\partial u}{\partial y} \right) - 2\mu \frac{\partial v}{\partial y} - \mu \left(\frac{\partial v}{\partial z} + \frac{\partial w}{\partial y} \right) \right] = - \left(\frac{\mu}{k} v + \frac{\partial p}{\partial y} \right) \quad (6)$$

$$\rho_g \frac{\partial w}{\partial t} + \nabla \left[-\mu \left(\frac{\partial w}{\partial x} + \frac{\partial u}{\partial z} \right) - \mu \left(\frac{\partial w}{\partial y} + \frac{\partial v}{\partial z} \right) - 2\mu \frac{\partial w}{\partial z} \right] = - \left(\frac{\mu}{k} w + \frac{\partial p}{\partial z} \right) \quad (7)$$

$$- \left(\rho_g \left(\frac{\partial u}{\partial x} + \frac{\partial v}{\partial y} + \frac{\partial w}{\partial z} \right) + \frac{\partial \rho_g}{\partial x} u + \frac{\partial \rho_g}{\partial y} v + \frac{\partial \rho_g}{\partial z} w \right) = 0 \quad (8)$$

The following notations are used in these equations: t - time, x, y, z - space coordinates, u, v and w - gas velocities in the x, y and z direction, respectively, p - pressure, ρ_g - inert gas density, k - adsorbent bed permeability and μ - dynamic viscosity of the carrier gas (inert).

2) Adsorbate balance for the gas phase:

For the central tubes:

$$\frac{\partial(\rho_g C)}{\partial t} + \nabla \left(-D_{mx}|_{i,o} \frac{\partial(\rho_g C)}{\partial x} - D_{my}|_{i,o} \frac{\partial(\rho_g C)}{\partial y} - D_{mz}|_{i,o} \frac{\partial(\rho_g C)}{\partial z} \right) = - \left(u \frac{\partial(\rho_g C)}{\partial x} + v \frac{\partial(\rho_g C)}{\partial y} + w \frac{\partial(\rho_g C)}{\partial z} \right) \quad (9)$$

For the adsorbent beds:

$$\begin{aligned} \frac{\partial(\rho_g C)}{\partial t} + \nabla \left(-D_{mx}|_b \frac{\partial(\rho_g C)}{\partial x} - D_{my}|_b \frac{\partial(\rho_g C)}{\partial y} - D_{mz}|_b \frac{\partial(\rho_g C)}{\partial z} \right) = k_m a (C^* - C) \\ - \left(u \frac{\partial(\rho_g C)}{\partial x} + v \frac{\partial(\rho_g C)}{\partial y} + w \frac{\partial(\rho_g C)}{\partial z} \right) \end{aligned} \quad (10)$$

For the space around the cartridges:

$$\begin{aligned} \frac{\partial(\rho_g C)}{\partial t} + \nabla \left(-D_{mx}|_{i,o} \frac{\partial(\rho_g C)}{\partial x} - D_{my}|_{i,o} \frac{\partial(\rho_g C)}{\partial y} - D_{mz}|_{i,o} \frac{\partial(\rho_g C)}{\partial z} \right) = \\ - \left(u \frac{\partial(\rho_g C)}{\partial x} + v \frac{\partial(\rho_g C)}{\partial y} + w \frac{\partial(\rho_g C)}{\partial z} \right) \end{aligned} \quad (11)$$

In these equations C is adsorbate concentration in the gas phase, defined as molar ratio, i.e., in moles of adsorbate per mole of inert gas and C^* the gas concentration in equilibrium with the solid phase. D_{mx} , D_{my} and D_{mz} are mass dispersion coefficients in the x , y and z direction, respectively, k_m mass transfer coefficient and a specific surface area of the adsorbent for mass and heat transfer. The subscript i,o corresponds to the inner (central) tubes and outer space and b to the adsorbent beds.

3) Heat balances for the gas phase:

For the central tubes and the space around the cartridges:

$$\begin{aligned} \frac{\partial}{\partial t} [\rho_g (c_{pg} + c_{pv} C) T_g] + \nabla \left(-D_{tx}^{hg}|_{i,o} \frac{\partial T_g}{\partial x} - D_{ty}^{hg}|_{i,o} \frac{\partial T_g}{\partial y} - D_{tz}^{hg}|_{i,o} \frac{\partial T_g}{\partial z} \right) = \\ - \rho_g c_{pg} \left(u \frac{\partial T_g}{\partial x} + v \frac{\partial T_g}{\partial y} + w \frac{\partial T_g}{\partial z} \right) - \rho_g c_{pv} \left(u \frac{\partial (CT_g)}{\partial x} + v \frac{\partial (CT_g)}{\partial y} + w \frac{\partial (CT_g)}{\partial z} \right) \end{aligned} \quad (12)$$

For the adsorbent beds:

$$\begin{aligned} \frac{\partial}{\partial t} [\rho_g (c_{pg} + c_{pv} C) T_g] + \nabla \left(-D_{tx}^{hg}|_b \frac{\partial T_g}{\partial x} - D_{ty}^{hg}|_b \frac{\partial T_g}{\partial y} - D_{tz}^{hg}|_b \frac{\partial T_g}{\partial z} \right) = \\ - \rho_g c_{pg} \left(u \frac{\partial T_g}{\partial x} + v \frac{\partial T_g}{\partial y} + w \frac{\partial T_g}{\partial z} \right) - \rho_g c_{pv} \left(u \frac{\partial (CT_g)}{\partial x} + v \frac{\partial (CT_g)}{\partial y} + w \frac{\partial (CT_g)}{\partial z} \right) + h_b a (T_s - T_g) \end{aligned} \quad (13)$$

In equations (12) and (13) T_g and T_s are the gas and solid temperature, respectively, D_{tx}^{hg} , D_{ty}^{hg} and D_{tz}^{hg} heat diffusivities of the gas phase, in the x , y and z direction, respectively, h_b heat transfer coefficient within the adsorbent bed and c_{pg} and c_{pv} molar heat capacities of the inert gas and adsorbate vapor, respectively.

4) Electric current balance for resistive heating:

$$-\nabla \left(\frac{1}{\rho} \frac{\partial U}{\partial x} + \frac{1}{\rho} \frac{\partial U}{\partial y} + \frac{1}{\rho} \frac{\partial U}{\partial z} \right) = 0 \quad (14)$$

In this equation U is electric potential and ρ electric resistivity of the adsorbent, which is temperature dependent. Linear temperature dependence of the electric resistivity, which was obtained experimentally in Ref. 5, was used in our models:

$$\rho = \rho_0(1 + b(T_s - T_R)) \quad (15)$$

5) Adsorbate balance for the solid phase within the adsorbent bed:

$$\rho_b \frac{\partial q}{\partial t} = k_m a(C - C^*) \quad (16)$$

where q is the adsorbate concentration in the solid phase and ρ_b the solid phase density. The Dubinin-Radushkevich equation, valid for adsorption of hydrocarbons on activated carbon [5] is used to describe the adsorption equilibrium. This equation can be put into the following form:

$$C^* = \frac{p^o}{p \exp \left[\frac{E}{R_g T_s} \sqrt{-\ln \left(\frac{M_A q}{\rho_A W_0} \right)} \right]} \quad (17)$$

where W_0 is the volume of the micropores per unit mass of adsorbent, E - the adsorbate energy of adsorption and M_A and ρ_A the adsorbate molar mass and density, respectively. p^o is the adsorbate saturation pressure. In our models, it was calculated using the Wagner equation:

$$\ln \left(\frac{p^o}{p_c} \right) = \left(\frac{VP_A x + VP_B x^{1.5} + VP_C x^3 + VP_D x^6}{1 - x} \right) \quad (18)$$

where $x = 1 - (T_s/T_c)$, p_c and T_c are the critical pressure and temperature and VP_A , VP_B , VP_C and VP_D are the Wagner constants.

6) Heat balance for the solid phase within the adsorbent bed:

$$\rho_b \frac{\partial}{\partial t} [(c_{ps} + c_{pl} q) T_s] - \nabla \left(D_{tx}^{hs} \frac{\partial T_s}{\partial x} + D_{ty}^{hs} \frac{\partial T_s}{\partial y} + D_{tz}^{hs} \frac{\partial T_s}{\partial z} \right) = \frac{\delta \dot{Q}_{el}}{dV} - h_b a (T_s - T_g) + \rho_b (\Delta H_{ads}) \frac{\partial q}{\partial t} \quad (19)$$

where $(\delta \dot{Q}_{el} / dV)$ is the electric power supply per unite volume of the adsorbent bed:

$$\frac{\delta \dot{Q}_{el}}{dV} = \frac{1}{\rho} \left(\left(\frac{\partial U}{\partial x} \right)^2 + \left(\frac{\partial U}{\partial y} \right)^2 + \left(\frac{\partial U}{\partial z} \right)^2 \right) \quad (20)$$

In the case of adsorption $(\delta \dot{Q}_{el} / dV) = 0$. In equation (19) (ΔH_{ads}) is the molar heat of adsorption, D_{tx}^{hs} , D_{ty}^{hs} and D_{tz}^{hs} the heat diffusivities of the solid phase in the x , y and z direction, respectively, and c_{ps} and c_{pl} the heat capacities of the solid and of the liquid adsorbate, respectively.

Boundary conditions:

The boundary conditions listed bellow correspond to the gas flow dirrections defined for the desorption step in Figure 4, and for one quarter of the four-cartridge adsorber.

- For the bottom surface of the central tube:

$$z = 0, \quad (x - x_0)^2 + (y - y_0)^2 < r_1^2: \quad u = 0, \quad v = 0, \quad w = 0, \\ -D_{tz}^{hg} \Big|_{i,o} \frac{\partial T_g}{\partial z} = 0, \quad -D_{mz} \Big|_{i,o} \frac{\partial C}{\partial z} = 0 \quad (21)$$

- For the top surface of the central tube:

$$z = H, \quad (x - x_0)^2 + (y - y_0)^2 < r_1^2: \quad u = 0, \quad v = 0, \quad w = \frac{G}{r_1^2 \pi \rho_g}, \quad T_g = T_{gin}, \quad C = C_{in} \quad (22)$$

- For the cylindrical surface of the central tube:

$$(x - x_0)^2 + (y - y_0)^2 = r_1^2, \quad z \in (0, H): \quad p|_{it} = p|_b, \quad u|_{it} = u|_b, \quad v|_{it} = v|_b, \quad w|_{it} = w|_b, \\ \left(-D_{tx}^{hg} \Big|_{i,o} \frac{\partial T_g}{\partial x} - D_{ty}^{hg} \Big|_{i,o} \frac{\partial T_g}{\partial y} + \rho_g (c_{pg} + c_{pv} C) T_g (u + v) \right)_{it} = \\ \left(-D_{tx}^{hg} \Big|_b \frac{\partial T_g}{\partial x} - D_{ty}^{hg} \Big|_b \frac{\partial T_g}{\partial y} + \rho_g (c_{pg} + c_{pv} C) T_g (u + v) \right)_b + h_{s1} (1 - \varepsilon_b) (T_s - T_g) \\ \left(-D_{mx} \Big|_{i,o} \frac{\partial C}{\partial x} - D_{my} \Big|_{i,o} \frac{\partial C}{\partial y} + (u + v) C \right)_{it} = \\ \left(-D_{mx} \Big|_b \frac{\partial C}{\partial x} - D_{my} \Big|_b \frac{\partial C}{\partial y} + (u + v) C \right)_b + k_{m1} (1 - \varepsilon_b) (C^* - C), \\ -D_{tx}^{hs} \frac{\partial T_s}{\partial x} - D_{ty}^{hs} \frac{\partial T_s}{\partial y} = h_{s1} (T_g - T_s), \quad J = 0 \quad (23)$$

- For the bottom surface of the adsorbent bed:

$$z = 0, \quad r_1^2 < (x - x_0)^2 + (y - y_0)^2 < r_2^2: \quad u = 0, \quad v = 0, \quad w = 0, \\ -D_{tz}^{hg} \Big|_b \frac{\partial T_g}{\partial z} = 0, \quad -D_{mz} \Big|_b \frac{\partial C}{\partial z} = 0, \quad U = 0, \quad -D_{tz}^{hs} \frac{\partial T_s}{\partial z} = 0 \quad (24)$$

- For the top surface of the adsorbent bed:

$$z = H, \quad r_1^2 < (x - x_0)^2 + (y - y_0)^2 < r_2^2: \quad u = 0, \quad v = 0, \quad w = 0, \\ -D_{tz}^{hg} \Big|_b \frac{\partial T_g}{\partial z} = 0, \quad -D_{mz} \Big|_b \frac{\partial C}{\partial z} = 0, \quad U = U_0, \quad -D_{tz}^{hs} \frac{\partial T_s}{\partial z} = 0 \quad (25)$$

- For the outer cylindrical surface of the adsorbent bed:

$$\begin{aligned}
(x - x_0)^2 + (y - y_0)^2 &= r_2^2, \quad z \in (0, H): \quad p|_b = p|_{ot}, \quad u|_b = u|_{ot}, \quad v|_b = v|_{ot}, \quad w|_b = w|_{ot}, \\
\left(-D_{tx}^{hg}|_b \frac{\partial T_g}{\partial x} - D_{ty}^{hg}|_b \frac{\partial T_g}{\partial y} + \rho_g (c_{pg} + c_{pv}C)T_g(u + v) \right)_b &= \\
\left(-D_{tx}^{hg}|_{i,o} \frac{\partial T_g}{\partial x} - D_{ty}^{hg}|_{i,o} \frac{\partial T_g}{\partial y} + \rho_g (c_{pg} + c_{pv}C)T_g(u + v) \right)_{ot} &+ h_{s2}(1 - \varepsilon_b)(T_s - T_g) \\
\left(-D_{mx}|_b \frac{\partial C}{\partial x} - D_{my}|_b \frac{\partial C}{\partial y} + (u + v)C \right)_b &= \\
\left(-D_{mx}|_{i,o} \frac{\partial C}{\partial x} - D_{my}|_{i,o} \frac{\partial C}{\partial y} + (u + v)C \right)_{ot} &+ k_{m2}(1 - \varepsilon_b)(C^* - C), \\
-D_{tx}^{hs} \frac{\partial T_s}{\partial x} - D_{ty}^{hs} \frac{\partial T_s}{\partial y} &= h_{s1}(T_g - T_s), \quad J = 0
\end{aligned} \tag{26}$$

- For the bottom surface of the space around the cartridges:

$$\begin{aligned}
z = 0, \quad (x - x_0)^2 + (x - y_0)^2 > r_2^2 \cap x^2 + y^2 < r_3^2 \cap x \in (0, r_3), y \in (0, r_3): \\
p|_{ot} = p_a, \quad -D_{tz}^{hg}|_{i,o} \frac{\partial T_g}{\partial z} = 0, \quad -D_{mz}|_{i,o} \frac{\partial C}{\partial z} = 0
\end{aligned} \tag{27}$$

- For the top surface of the space around the cartridges:

$$\begin{aligned}
z = H, \quad (x - x_0)^2 + (x - y_0)^2 > r_2^2 \cap x^2 + y^2 < r_3^2 \cap x \in (0, r_3), y \in (0, r_3): \\
u = 0, \quad v = 0, \quad w = 0, \quad -D_{tz}^{hg}|_{i,o} \frac{\partial T_g}{\partial z} = 0, \quad -D_{mz}|_{i,o} \frac{\partial C}{\partial z} = 0
\end{aligned} \tag{28}$$

- For the adsorber wall:

$$\begin{aligned}
x^2 + y^2 &= r_3^2, \quad z \in (0, H): \quad u = 0, \quad v = 0, \quad w = 0, \\
-D_{mx}|_{i,o} \frac{\partial C}{\partial x} - D_{my}|_{i,o} \frac{\partial C}{\partial y} &= 0, \quad -D_{tx}^{hg}|_{i,o} \frac{\partial T_g}{\partial x} - D_{ty}^{hg}|_{i,o} \frac{\partial T_g}{\partial y} = h_{wg}(T_g - T_a)
\end{aligned} \tag{29}$$

For the case of desorption accompanied with condensation, the boundary condition (29) has to be changed in the following way [11]:

$$\begin{aligned}
x^2 + y^2 &= r_3^2, \quad z \in (0, H): \quad u = 0, \quad v = 0, \quad w = 0, \\
C &= C_{sat}(T_w), \quad -D_{tx}^{hg}|_{i,o} \frac{\partial T_g}{\partial x} - D_{ty}^{hg}|_{i,o} \frac{\partial T_g}{\partial y} = - \left(D_{mx}|_{i,o} \frac{\partial C}{\partial x} + D_{my}|_{i,o} \frac{\partial C}{\partial y} \right) (\Delta H_{cond}) h_{wg}(T_g - T_a)
\end{aligned} \tag{29a}$$

In equations (21-29), r_1 , r_2 , r_3 , x_0 , y_0 and H are the adsorber dimensions which have been defined in Figure 7. G is the molar flow-rate of the gas inert, U_0 – the supply voltage (0 for adsorption) and ε_b – bed porosity. The subscript *in* denotes inlet, *it* the inner (central) tube, *ot* the space around the cartridges, *a* the ambient and *w* the adsorber wall. (ΔH_{cond}) is the molar

heat of condensation and $C_{sat}(T_w)$ the saturation concentration corresponding to the wall temperature T_w , which was calculated using the Wagner equation.

Although the heat and mass transfer coefficients, dispersion coefficients, etc., can be defined as functions of the dependent variables (velocities, concentrations, temperatures), in our current models constant values were used in order to reduce the convergence problems. The adsorption isotherm relation ($q=f(C, T_s)$) and the temperature dependence of the electric resistivity were included in the models (equations (17) and (15), respectively).

3. Solution of the models

The Models of the four-cartridge adsorber were developed and solved using a personal computer with AMD ATHLON™ 64 3000+ processor and with 4 GB RAM.

Our 3-D models are very complex, regarding both the geometry and the number of application modes used. Even with a “coarser” mesh, the number of degrees of freedom obtained is very high. This number was so high, that the models for the whole and for one half of the four-cartridge adsorber (Figs. 5 and 6) could not be solved at all, owing to the lack of computer memory. Because of that, we focused our efforts on solving the model of a quarter of the adsorber, assuming model symmetry in the x and y directions (Fig. 7).

3.1. The mesh parameters

In order to reduce the memory requirements, the following mesh parameters were used:

- Predefined mesh sizes: “coarser”
- Maximum element size scaling factor: 1.9
- Element growth rate: 1.7

With these mesh parameters the following mesh statistics was obtained:

- Number of mesh elements: 2237
- Minimum element quality: 0.2578.

The mesh with these characteristics is presented in Figure 8.

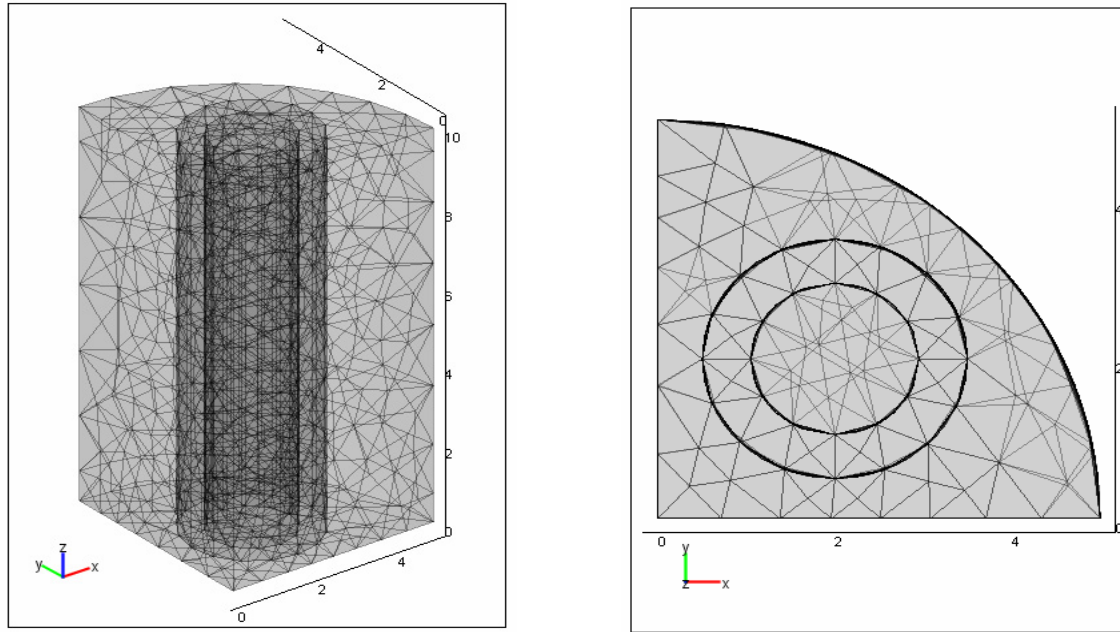


Figure 8. The mesh for one quarter of the four-cartridge adsorber: 3-D view (left) and 2-D x-y view (right)

3.2. The solver

The three models were solved using the **direct SPOOLS solver**, with default solver parameters. The following values of the absolute and relative tolerance and the time step were used:

- For Model_A3 (adsorption):
 - Absolute tolerance: 0.001
 - Relative tolerance: 0.01
 - Time step: 20 s
- For Model_D3 and Model_DC3 (desorption without and with condensation):
 - Absolute tolerance: 0.01
 - Relative tolerance: 0.1
 - Time step: 10 s

The solution of Model_D3 and Model_DC3 did not converge with lower tolerances and longer step time.

3.3. The solution procedure

Owing to their great complexity, the complete models couldn't be solved directly, i.e. the solution usually diverged. The main difficulty was to solve the nonisothermal Navier-Stokes equations simultaneously with the heat and mass balances, and especially to start the solution process with zero initial values of the velocities. In order to overcome these problems, a step-by-step solution procedure was developed:

Step 1. A simplified model, in which only *Non-Isothermal Flow* and *Brinkman Equations* application modes are active, is solved using the **stationary nonlinear solver**. In this way the steady state values of the velocities and pressures in the system are calculated, for the case of inert isothermal flow. In reality, the steady-state velocities and pressures are established very fast, during the first 3 seconds [11], so using a stationary solver is appropriate. In this step, the initial values of the solved variables (velocities in all three sub-domains and x , y and z direction and pressures in all three sub-domains) are taken from the *Initial value expression*. The values of the variables not solved in this step (temperatures and concentrations of both phases) are taken from *Initial value frame*. The solution of this stage is stored. The number of degrees of freedom for this calculation was 14599.

Step 2. The model is solved with the following application modes: *Convection and Conduction*, *Heat Transfer by Conduction*, *Convection and Diffusion*, *Diffusion* and *Conductive media DC* (this mode is not used in Model_A3). The application modes *Non-Isothermal Flow* and *Brinkman Equations* are inactive. In this step the heat and mass balances are solved, using the **time depended solver**, with fixed velocities and pressures, calculated previously in Step 1. The initial values of the variables solved in this step (temperatures and concentrations of both phases) were taken from the *Initial value frame* and the values of the variables not solved in this step (velocities and pressures) from the *Stored solution*. At the end, the solution at time 0 was stored. The number of degrees of freedom was 10508 for Model_A3 and 12425 for Model_D3 and Model_DC3.

Step 3. The integral model is solved using the **time depended solver**, with all application modes active. The initial values of all variables are taken from the *Stored solution at time 0*. For faster and more stable convergence this step is performed twice. The first run is interrupted and the solution at time 0 is stored, and a new run is performed with the initial values of all variables taken from that stored solution. If necessary, larger tolerance values can be used in the first run. The number of degrees of freedom in this step was 25107 for Model_A3 and 27024 for Model_D3.

Step 3 for Model_DC3 did not converge, even when the tolerance was further increased.

4. Simulation results

The simulations were performed for a laboratory scale adsorber, with the dimensions defined in Section 3.1.

The model parameters correspond to the following system:

- Adsorbate: methyl ethyl ketone (MEK)
- Adsorbent: American Kynol ACC-5092-20 (activated carbon fiber cloth - ACFC material)
- Carrier gas: nitrogen.

The bed permeability was calculated from the experimentally measured pressure drop of the ACFC bed given in Ref. 5. The numerical values of most of the other model parameters are also based on Ref. 5. A complete list of the model parameters used in these simulations is given in Table 1.

Table 1. The model parameters used for simulation

Parameter	Notation and value	
	Adsorption	Desorption
Adsorber dimensions	$r_1=0.95$ cm $r_2=1.5$ cm $r_3=5$ cm $x_0=2$ cm, $y_0=2$ cm $H=10$ cm	
Bed porosity	$\varepsilon_b=0.72$	
Specific surface area	$a=13.65$ cm ² /cm ³	
Ambient pressure	$p_{atm}=101.325$ kPa	
Inlet adsorbate concentration in the gas phase	$C_{in}=0.001$ mol/mol	$C_{in}=0$ mol/mol
Initial adsorbate concentration in the gas phase	$C_p=0$ mol/mol	$C_p=0.001$ mol/mol
Initial adsorbate concentration in the solid phase	$q_p=0$ mol/g	$q_p=0.00465$ mol/g
Saturation concentration	$C_{sat}=0.0147$ mol/mol	
Electric voltage supply U_0	$U_0=0$ V	$U_0=3$ V
Inert gas flow-rate	$G=0.03$ mol/s	$G=0.01$ mol/s
Inlet gas temperature	$T_{gin}=293.15$ K	
Ambient temperature	$T_a=293.15$ K	
Wall temperature	$T_w=293.15$ K	
Mass transfer coefficient in the bed	$k_m=0.0003$ mol/(cm ² s)	
Bed to central tube mass transfer coefficient	$k_{m1}=0.0003$ mol/(cm ² s)	
Bed to annular tube mass transfer coefficient	$k_{m2}=0.0003$ mol/(cm ² s)	
Heat transfer coefficient in the bed	$h_b=0.01$ W/(cm ² s)	
Bed to central tube heat transfer coefficient	$h_{s1}=0.02$ W/(cm ² s)	
Bed to central tube heat transfer coefficient	$h_{s2}=0.012$ W/(cm ² s)	
Heat transfer coefficient defining heat losses	$h_{wg}=5 \times 10^{-5}$ W/(cm ² K)	
Dispersion coefficient of the gas phase in the central and annular tubes	$D_{mx,y,z} _{it}=D_{mx,y,z} _{ot}=1$ mol/(cms)	
Heat diffusivity of the gas phase in the central and annular tubes	$D_{tx,y,z} _{it}^{hg}=D_{tx,y,z} _{ot}^{hg}=1$ W/(Kcm)	
Heat diffusivity of the solid phase	$D_t^{hs}=1.2 \times 10^{-3}$ W/(cmK)	
Molar heat of adsorption	$(-\Delta H_{ads})=66200$ J/mol	
Molar heat of condensation	$(-\Delta H_{cond})=31230$ J/mol	
Adsorbent bed density	$\rho_b=0.221$ g/cm ³	
Adsorbate density	$\rho_A=0.81$ g/cm ³	
Adsorbate molar mass	$M_A=72.107$ g/mol	
Inert gas molar mass (nitrogen)	$M_B=28.02$ g/mol	
Specific heat capacity of liquid adsorbate	$c_{pl}=157.9$ J/(molK)	
Specific heat capacity of the inert gas	$c_{pg}=29.13$ J/(molK)	
Heat capacity of the solid phase	$c_{ps}=0.71$ J/(gK)	
Specific heat capacity of gaseous adsorbate	$c_{pv}=100.9$ J/(molK)	
Dynamic viscosity of the inert gas	$\mu=1.78 \times 10^{-4}$ g/cm/s	
Inert gas thermal conductivity	$k=2.6 \times 10^{-4}$ W/cm/K	
Electric resistivity at referent temperature T_R	$\rho_0=0.202$ Ω cm	
Temperature coefficient of the bed electrical resistivity	$b=-2 \times 10^{-3}$ 1/K	
Critical temperature of the inert gas (nitrogen)	$T_{cB}=126.2$ K	
Critical pressure of the adsorbate	$p_{cA}=4.26 \times 10^{+6}$ Pa	
Critical pressure of the inert gas (nitrogen)	$p_{cB}=3.39 \times 10^{+6}$ Pa	
Wagner constants for the adsorbate	$VP_A=-7.71476$, $VP_B=1.71061$ $VP_C=-3.6877$, $VP_D=-0.75169$	
Total volume of micropores (D-R equation)	$W_0=0.748$ cm ³ /g	
Boltzmann's constant	$k_b=1.38048 \times 10^{-23}$ J/K	
Adsorption energy of the adsorbate (D-R eq.)	$E=14.43 \times 10^{+3}$ J/mol	

COMSOL Multiphysics offers a large number of options for graphical representation of the simulated results. The following results were chosen for presentation in this report:

- Velocity field in the adsorber;
- Gas and solid concentration fields in the adsorber;
- Gas and solid temperature fields in the adsorber.

For all three models, these results are shown in the form of 3-D slice plots, corresponding to 3 different times. The arrows, showing the direction of the gas flow are also shown in these 3-D plots. Also, 2-D plots are given, corresponding to two cross-sections of the adsorber, one corresponding to $x=y$, and the other to $z=H/2$ (planes A and B in Figure 9). Time profiles, in 1-D diagrams, corresponding to 5 points in the adsorber, also defined in Figure 9, are also given. These points lie at the intersection of planes A and B, two of them within the adsorbent bed, two in the space around the cartridges and one on the axes of the central tube.

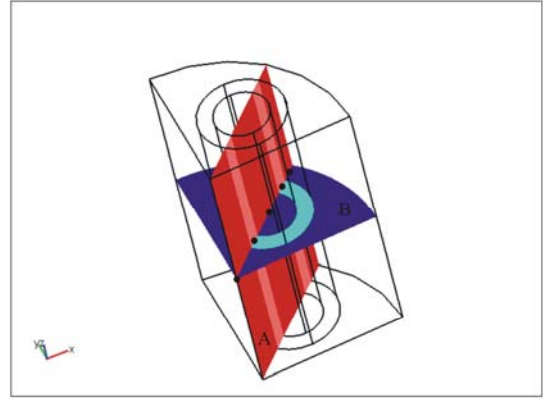


Figure 9. Definition of the planes and points used for 2-D and 1-D presentation

Two additional parameters, very important for the electrothermal desorption process, were calculated and presented: the amount of the condensed liquid and the used electric energy.

Calculation of the amount of the condensed liquid:

The flux of the condensed liquid is calculated from the gas concentration gradient at the adsorber wall:

$$J_{cond}(x, y, z, t) = -D_{mx}|_{i,o} \frac{\partial C}{\partial x} - D_{my}|_{i,o} \frac{\partial C}{\partial y} \Big|_{x^2+y^2=r_3^2} \quad (30)$$

It should be noticed that, owing to the 3-D geometry, this flux is a function of all three coordinates and time. The condensation rate is calculated by integration of the flux over the whole surface of the adsorber wall, and is a function of time only:

$$\dot{L}_{cond}(t) = 4 \int\limits_{\substack{x^2+y^2=r_3^2 \\ 0 \leq z \leq H}} J_{cond}(x, y, z, t) dx dy dz \quad (31)$$

The total amount of the condensed liquid is obtained by integrating the condensation rate over time:

$$L_{cond} = \int_0^{\tau_{cond}} \dot{L}_{cond}(t) dt \quad (32)$$

τ_{cond} is the total time of condensation.

Calculation of the used electric energy:

It is assumed that the used electric energy is equal to the Joule heat produced in the adsorbent material. This energy per unit volume and unit time is defined by equation (20).

The used electric power per cartridge is obtained by integration over the adsorbent bed volume:

$$\dot{Q}_{el}(t) = \iiint_{\substack{r_1^2 < (x-x_0)^2 + (y-y_0)^2 < r_2^2 \\ 0 \leq z \leq H}} \frac{1}{\rho_0(1+b(T_s - T_R))} \left(\left(\frac{\partial U}{\partial x} \right)^2 + \left(\frac{\partial U}{\partial y} \right)^2 + \left(\frac{\partial U}{\partial z} \right)^2 \right) dx dy dz \quad (33)$$

and the total used energy during desorption, by integration over time and multiplying by 4:

$$Q_{el} = \int_0^{\tau_{des}} \dot{Q}_{el}(t) dt \quad (34)$$

τ_{des} is the total desorption time.

4.1. Simulation of adsorption

The simulation of adsorption was performed using Model_A3. In order to reduce the memory and computer time demands, the Conductive Media DC application mode was not active in this model, i.e., only 6 application modes were used. Adsorption on an initially clean adsorbent bed was simulated. The model was solved for 7500 s, until complete saturation of the adsorbent bed was achieved. The results presented here were obtained for inlet concentration 0.001 mol/mol and inert gas flow rate 0.03 mol/s. All other parameters used in the model are listed in Table 1.

Velocity distribution

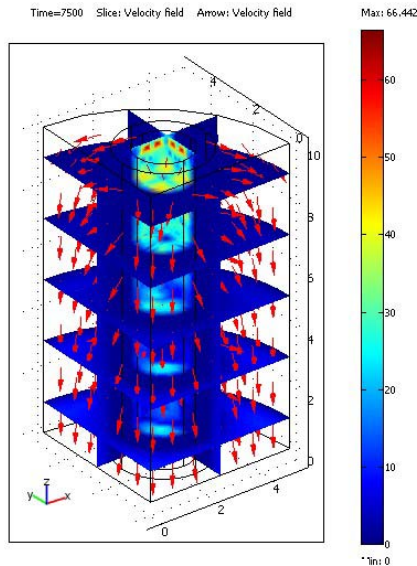


Figure 10. Velocity distribution at the end of adsorption, in one quarter of the 4-cartridge adsorber

The velocity during adsorption is practically not changing, so in Figure 10 we show the slice plot corresponding to only one time, $t=7500$ s. In Figure 11, the 2-D plots of the velocities, corresponding to planes A and B defined in Figure 5 are shown, and in Figure 12, a 1-D plot showing the time profiles of the velocities in 5 points at the intersection of planes A and B, also defined in Figure 5. The dashed lines in Figure 12 correspond to Step 2 of the solution procedure defined in Section 3.3. (solution of the model with fixed velocities). These lines were shown in order to check how big the discrepancies between the results obtained in Step 2 and Step 3, were.

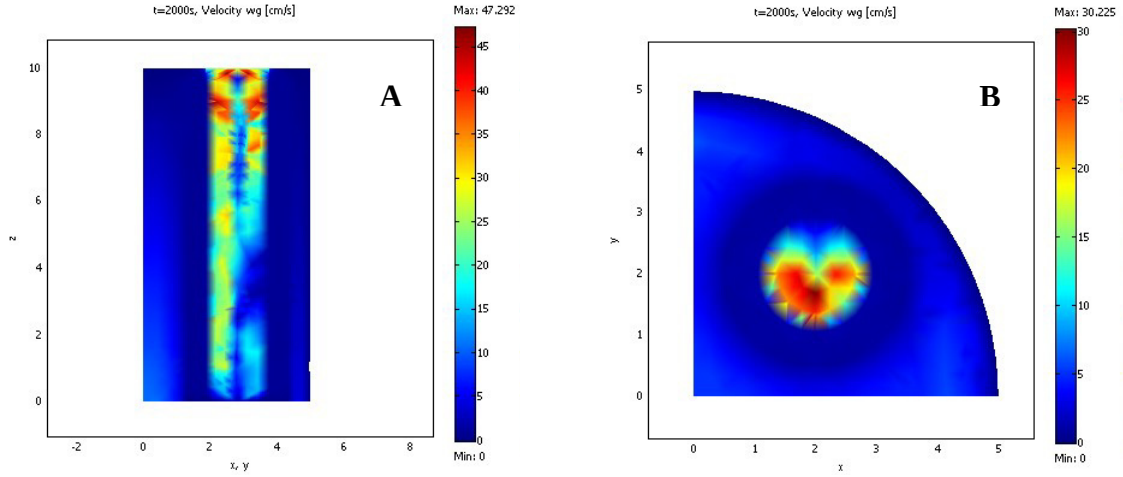


Figure 11. The velocity distributions during adsorption corresponding to cross-sections A and B (Fig. 9) and $t=2000$ s

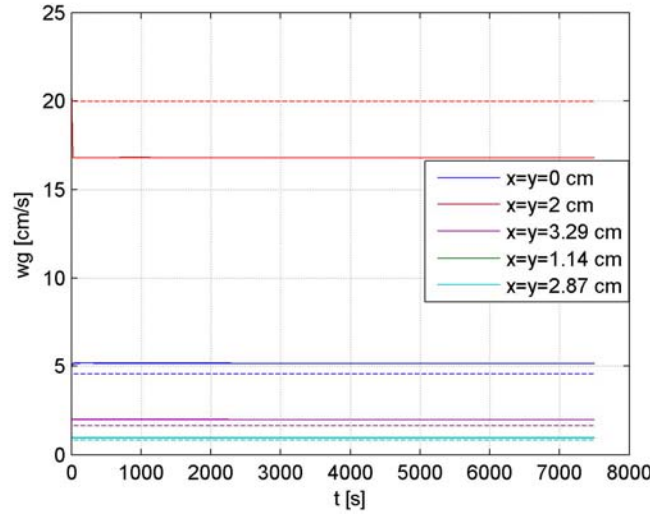


Figure 12. Time dependence of the velocity during adsorption in 5 points in the adsorber (Fig. 9)

Gas concentration distributions

Figure 13 shows the gas concentrations corresponding to 60, 2000 and 7500 s, in 3-D plots. Figure 14 shows the gas concentration distributions across the cross-sections A and B. These results correspond to $t=2000$ s. And finally, in Figure 15, we show the time profiles of the gas phase concentrations in the five points defined in Figure 9, together with the curves obtained with the fixed gas velocities, in Step 2. Figure 15 shows that the difference between the results obtained in Step 1 and Step 2 are rather small. The blue lines, corresponding to the center of the adsorber are not visible, as they overlap with the violet lines corresponding to the point $x=y=3.29$ cm. This is a result of high diffusion coefficients used in the simulations, resulting with practically uniform concentration in the space around the cartridges.

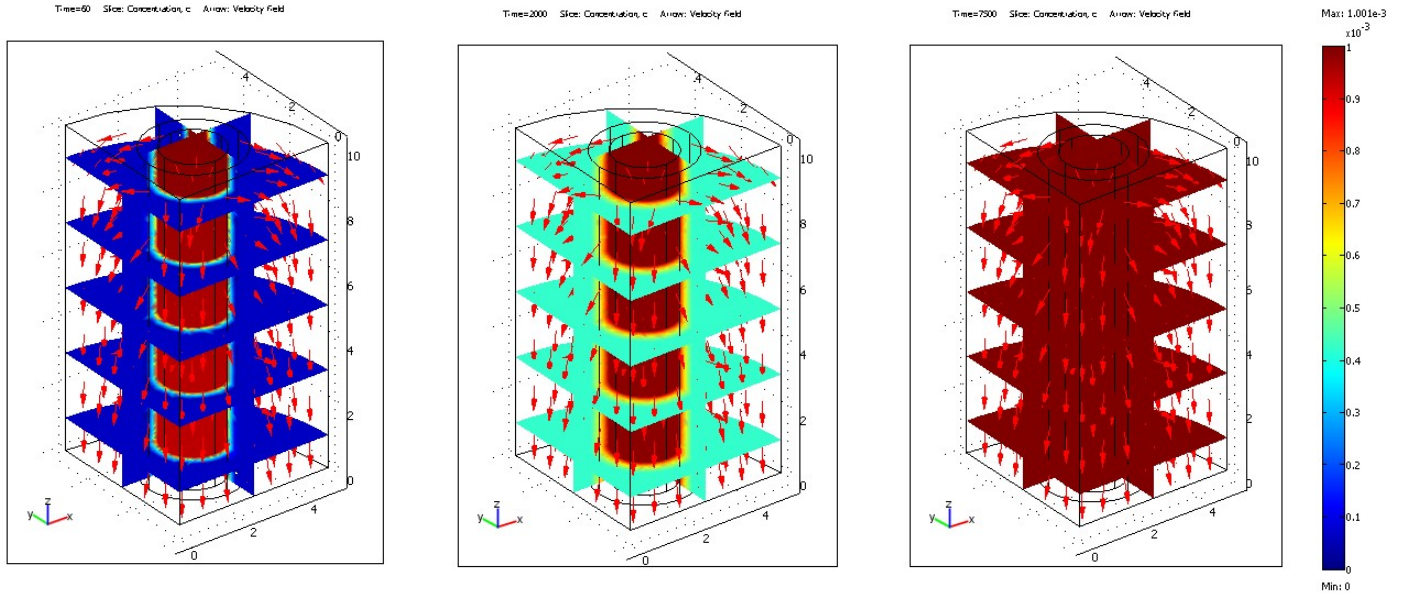


Figure 13. Concentration in the gas phase in one quarter of the 4-cartridge adsorber, during adsorption (from left to right $t=60$, 2000 and 7500 s)

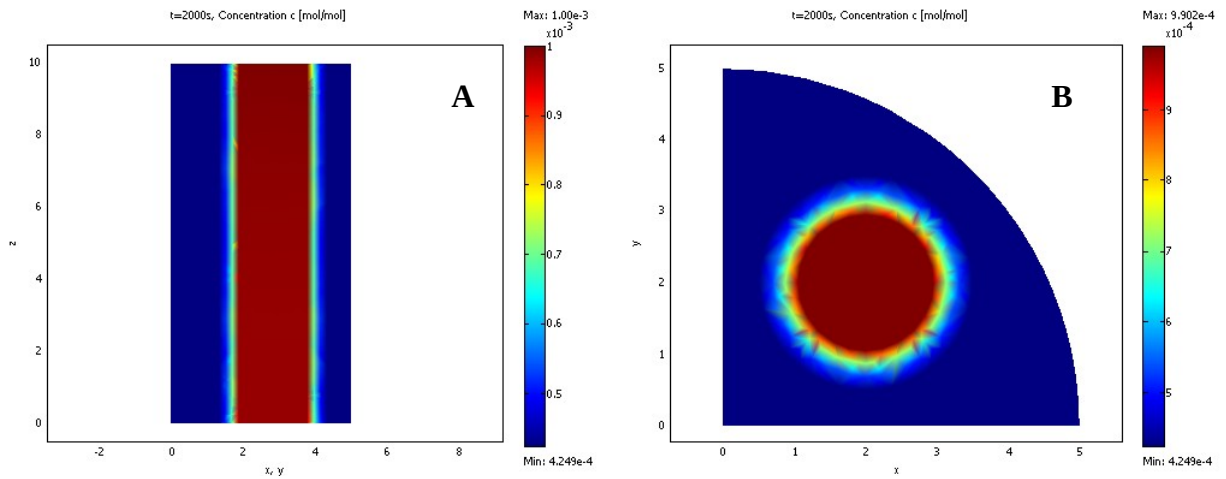


Figure 14. The gas concentrations during adsorption corresponding to cross-sections A and B (Fig. 9) and $t=2000$ s

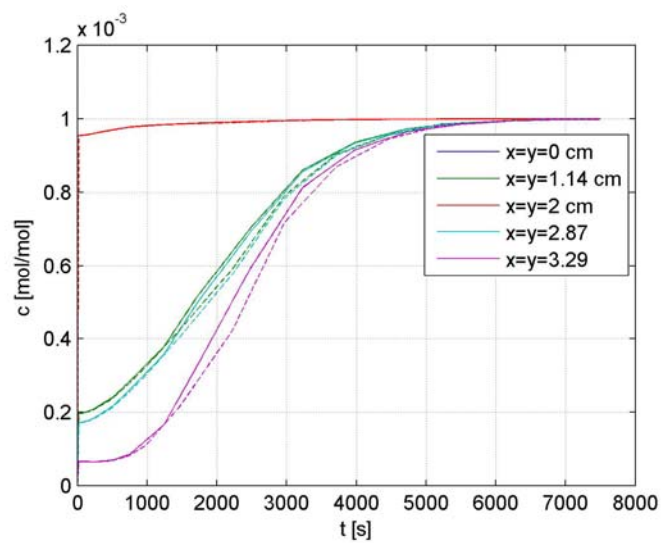


Figure 15. Time dependence of the gas concentration during adsorption in 5 points in the adsorber (Fig. 9)

Solid concentration distributions

The simulated concentrations in the solid phase corresponding to 60, 2000 and 7500 s, are first shown in 3-D plots, in Figure 16. The 2-D pictures of the solid concentrations at cross-sections A and B, for $t=2000$ s, are given in Figure 17 and their time profiles in Figure 18. Naturally, only the results corresponding to the two points inside the adsorbent bed are shown in this figure. As previously, the results of the approximate solution obtained in Step 2 are also shown in Figure 18, as dashed lines. The differences are again rather small.

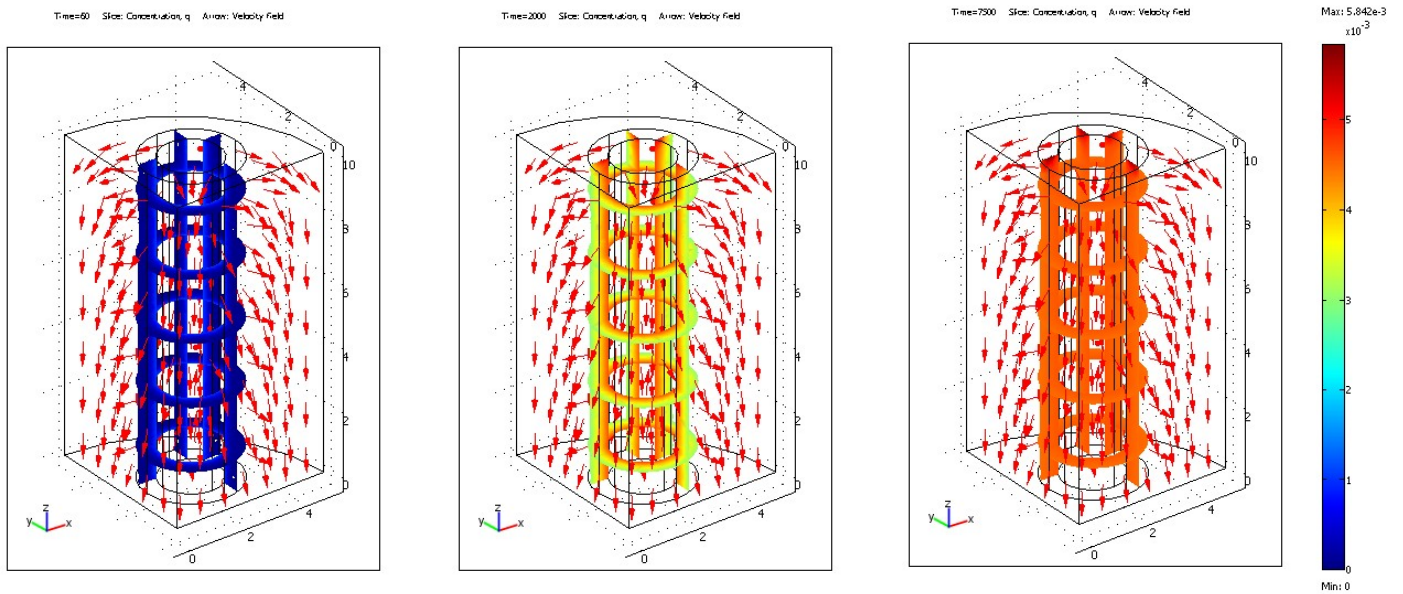


Figure 16. Concentrations in the sold phase in one quarter of the 4-cartridge adsorber, during adsorption (from left to right $t=60$, 2000 and 7500 s)

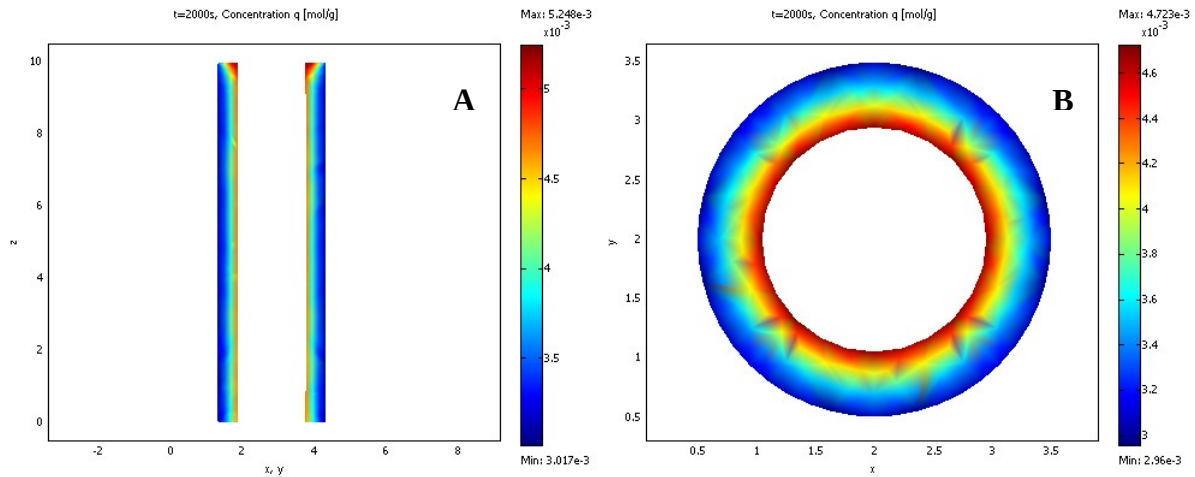


Figure 17. The solid concentrations during adsorption corresponding to cross-sections A and B (Fig. 9) and $t=2000$ s

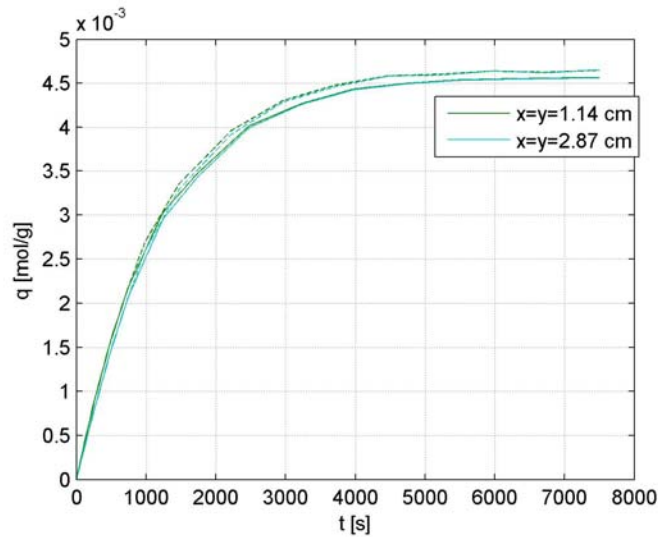


Figure 18. Time dependence of solid concentration during adsorption in 2 points in the adsorbent bed (Fig. 9)

Gas temperature distribution

The 3-D plots of the temperatures of the gas phase corresponding to 60, 2000 and 7500 s are given in Figure 19. Figure 20 shows the gas temperatures across cross-sections A and B, for $t=2000$ s, and Figure 21 the time profiles of the gas temperatures in the five points defined in Figure 5. The dashed lines again correspond to the solution obtained in Step 2, with fixed velocities. Although the differences between the solid and the dashed lines seem remarkable, actually they are less than 2 K. The lines corresponding to the center of the adsorber and to the point closest to the adsorber wall, again overlap (the blue lines are not visible).

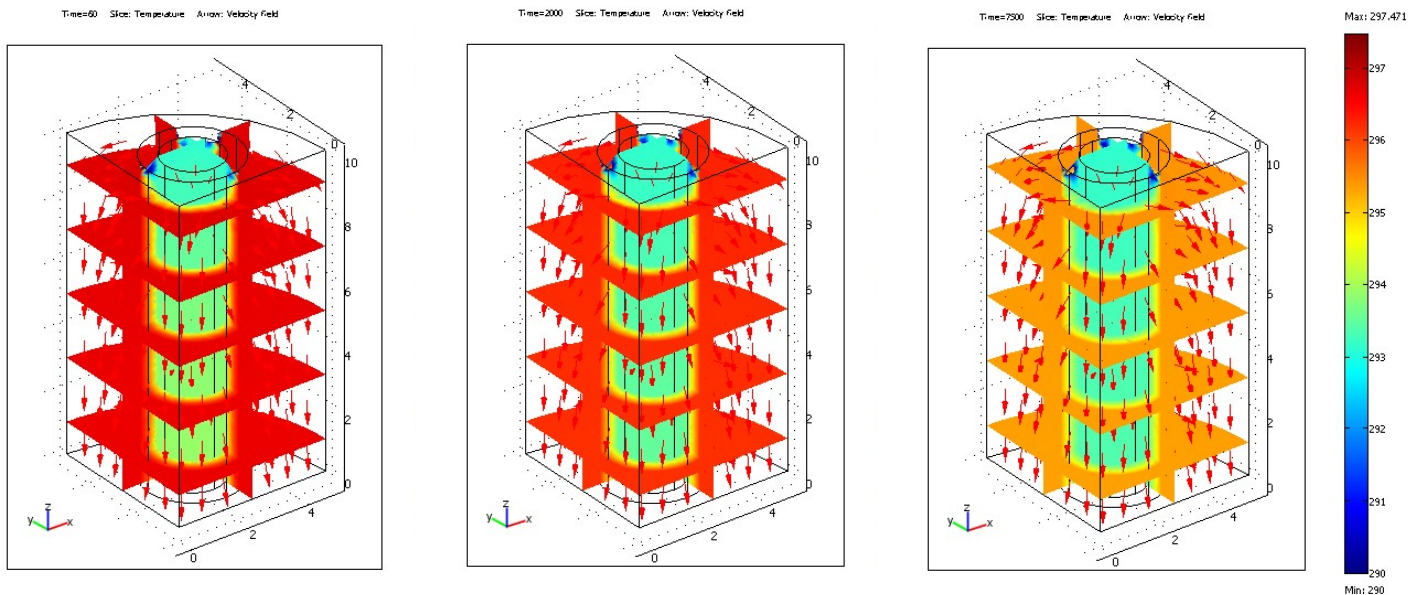


Figure 19. The temperatures of the gas phase in one quarter of the 4-cartridge adsorber, during adsorption (from left to right $t=60$, 2000 and 7500 s)

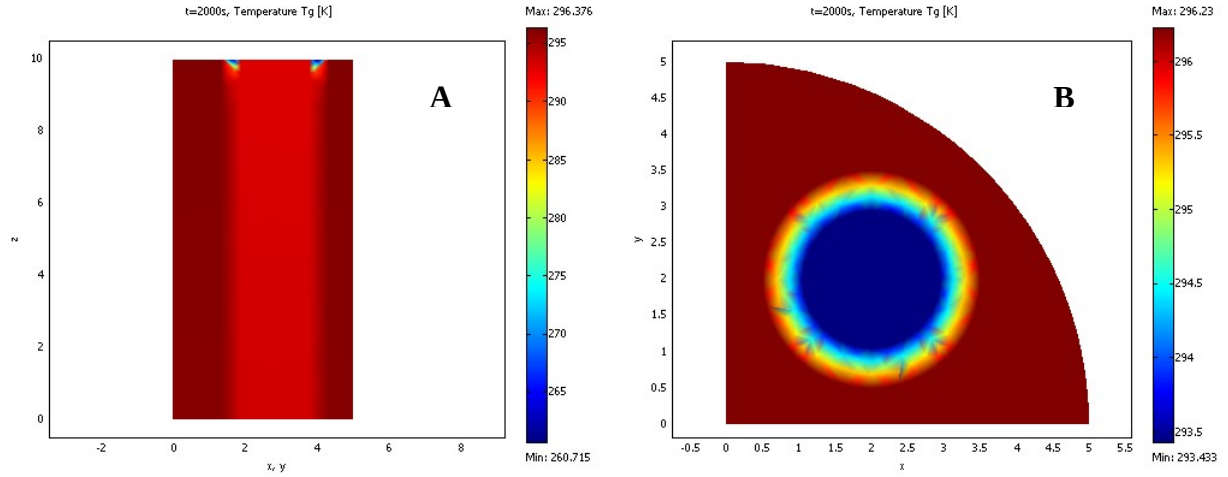


Figure 20. The gas temperatures during adsorption corresponding to cross-sections A and B (Fig. 9) and $t=2000$ s

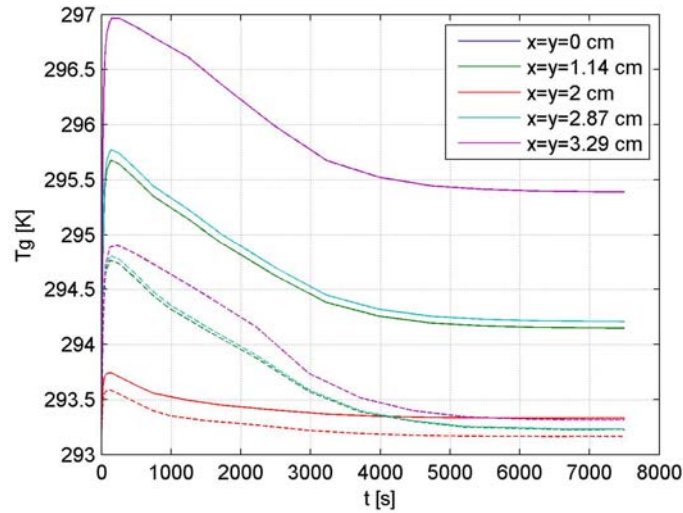


Figure 21. Time dependence of the gas temperature during adsorption in 5 points in the adsorber (Fig. 9)

Solid temperature distribution

The simulated temperatures of the solid phase are shown in Figure 22 (the 3-D plots), Figure 23 (the 2-D plots) and Figure 24 (the time profiles in the 1-D plot). The presentation is completely analogous to the one used for the solid concentrations, in Figures 16, 17 and 18. The discrepancies between the final solution and the solution obtained in Step 2 are of the same order of magnitude as for the gas temperature.

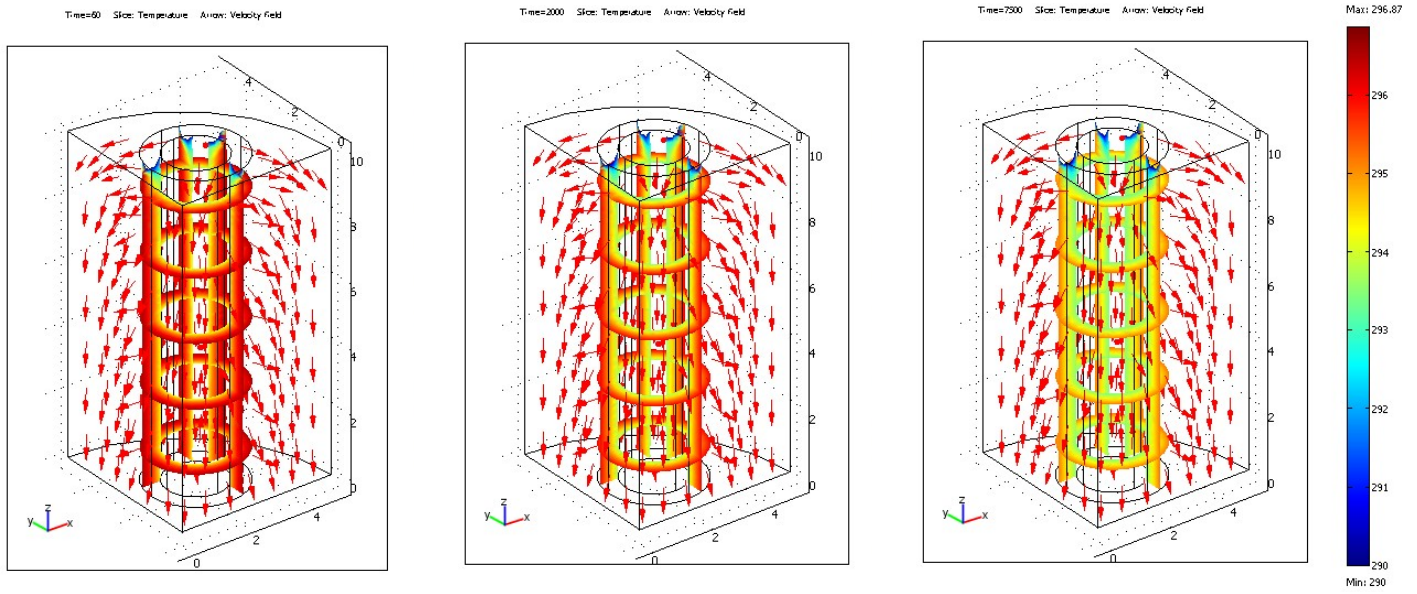


Figure 22. Temperatures of the solid phase in one quarter of the 4-cartridge adsorber, during adsorption (from left to right $t=60$, 2000 and 7500 s)

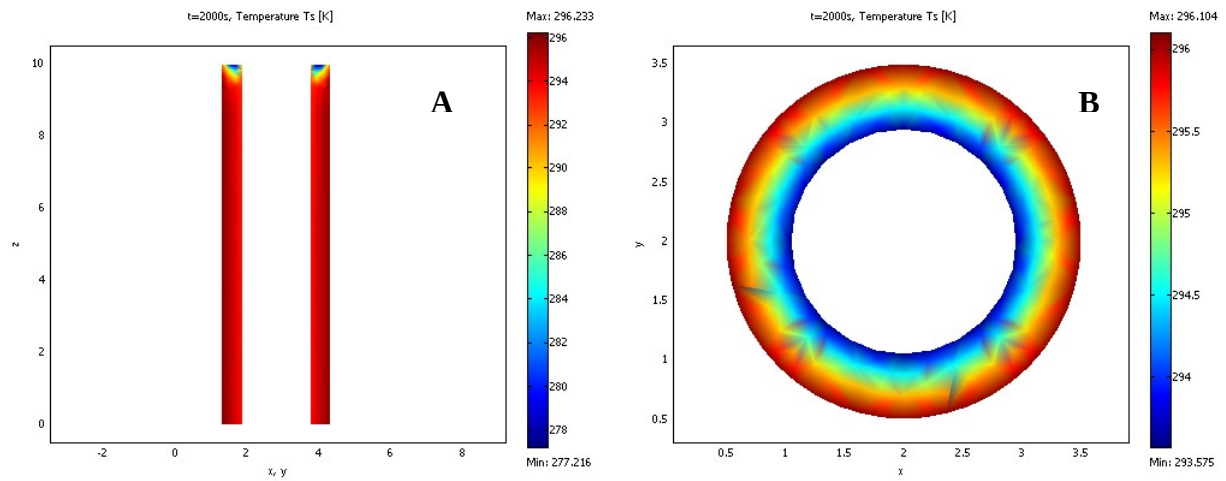


Figure 23. The solid temperature during adsorption corresponding to cross-sections A and B (Fig. 9) and $t=2000$ s

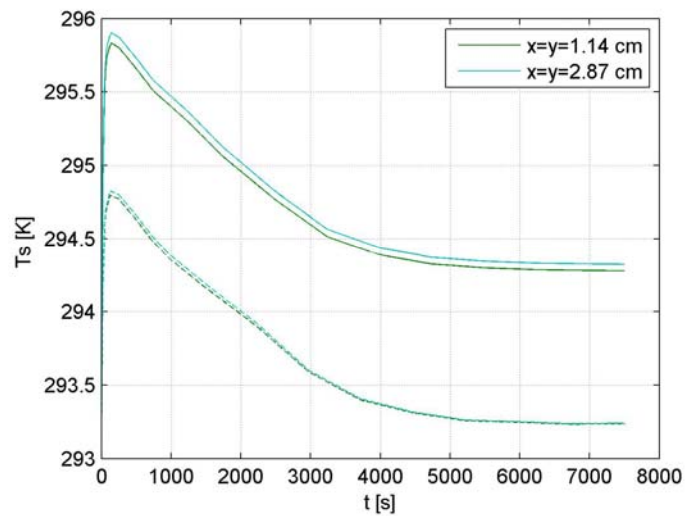


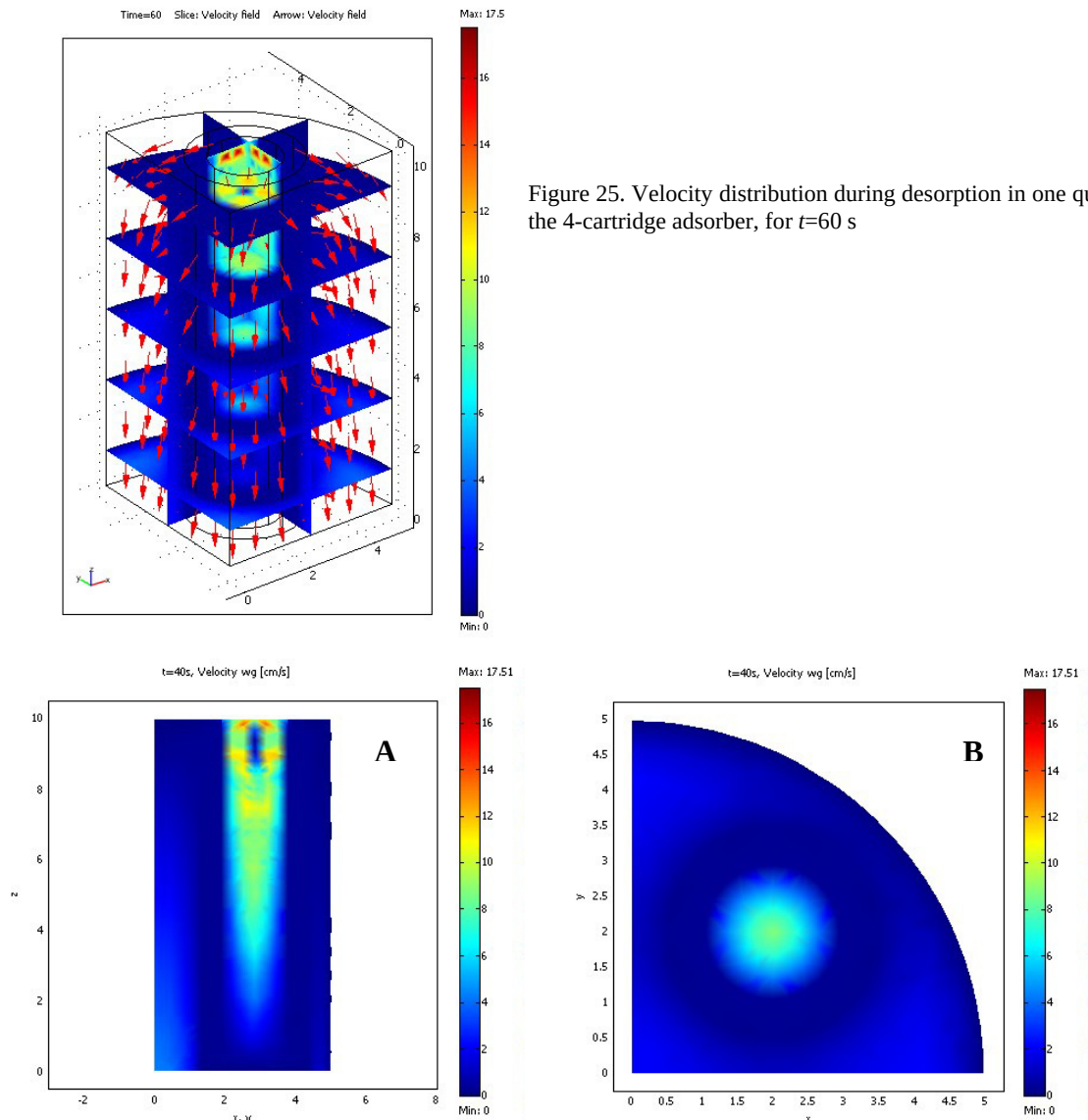
Figure 24. Time dependence of the solid temperature during ads. in 2 points in the adsorbent bed (Fig. 9)

4.2. Simulation of desorption without condensation

The desorption process before the start of condensation at the adsorber wall was simulated using Model_D3. All 7 application modes defined in Section 2.2. were active in this model. The simulation results presented in this report correspond to desorption from a previously saturated adsorption bed, in equilibrium with the gas phase of concentration 0.001 mol/mol. The inlet gas phase is pure inert, with flow rate 0.01 mol/s, and the voltage is 3 V. All other simulation parameters are defined in Table 1. For the chosen set of parameters, at $t=60$ s the gas concentration at the adsorber wall reaches 0.0147 mol/mol, which is the saturation concentration of MEK corresponding to the wall temperature of 293.15 K. Although the simulation was performed for longer times, only the results for times from 0 to 60 s correspond to this stage of the TSA cycle.

The simulation results for desorption without condensation are presented in an analogous way as for adsorption. The 3-D plots are shown for $t=10, 40$ and 60 s (the velocity only for 60 s), and the 2-D plots for $t=40$ s. The results are given in Figures 25 to 39.

Velocity distributions



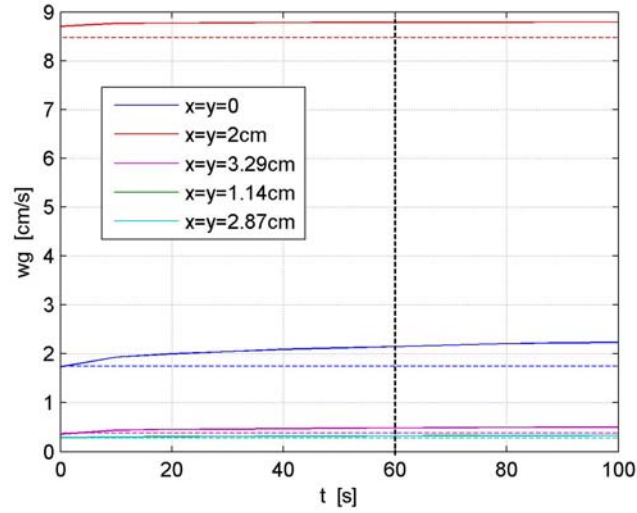


Figure 27. Time dependence of the velocity during adsorption in 5 points in the adsorber (Fig. 5)

Gas concentration distribution

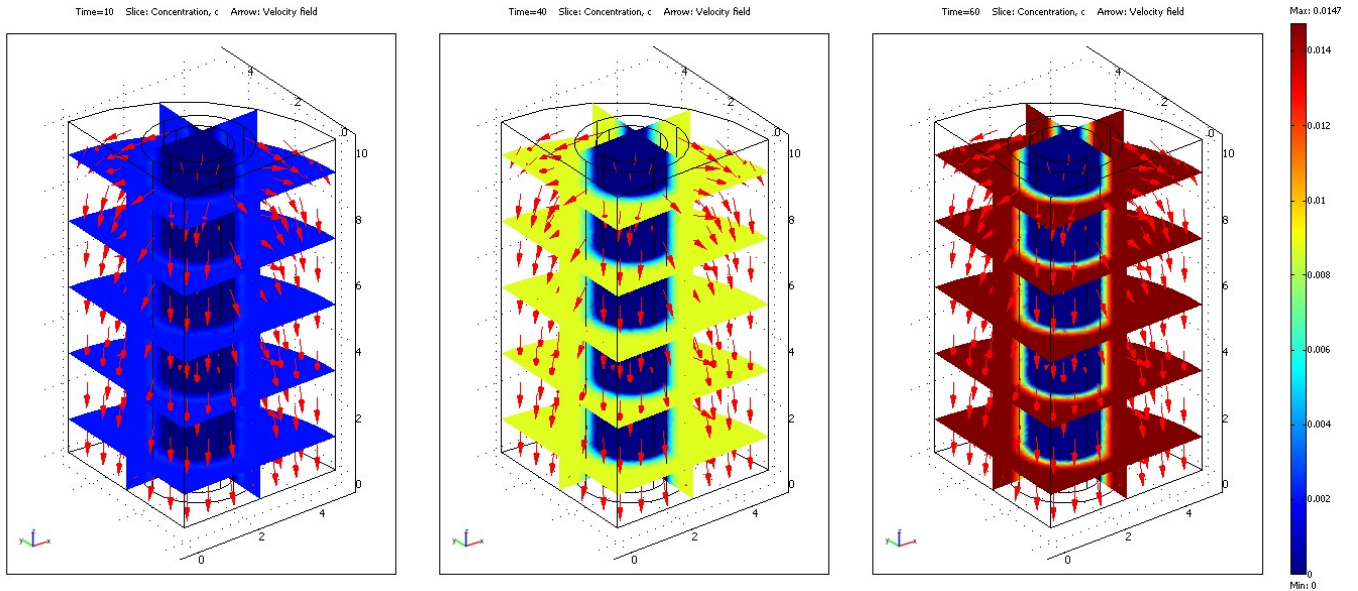


Figure 28. Concentration of the gas phase in one quarter of the 4-cartridge adsorber, during desorption (from left to right $t=10, 40$ and 60 s)

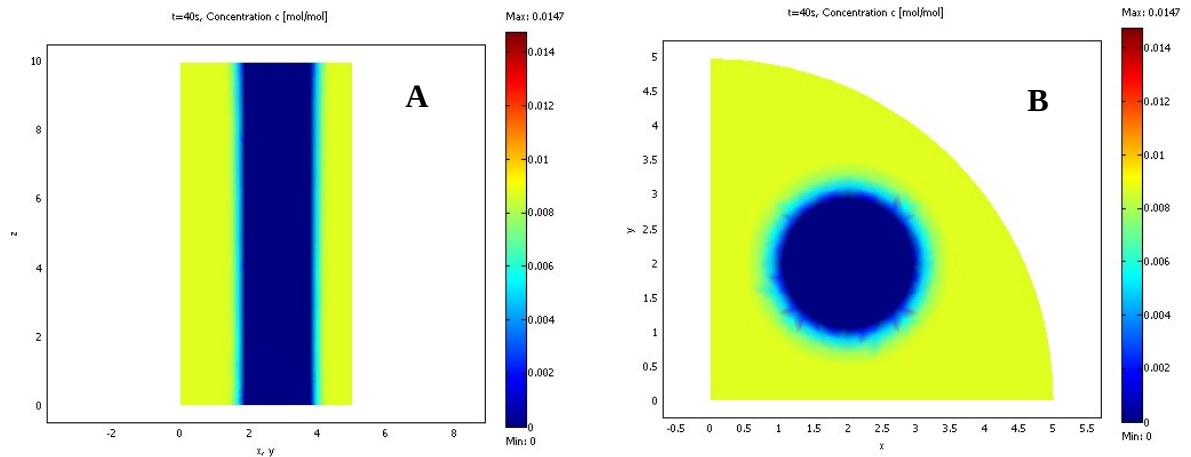


Figure 29. The gas concentration during desorption at cross-sections A and B (Fig. 9) and for $t=40$ s

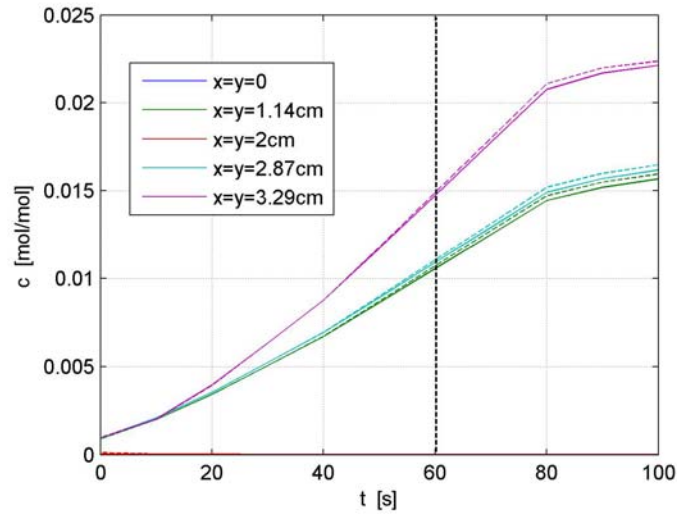


Figure 30. Time dependence of the gas concentration during desorption in 5 points in the adsorber (Fig. 9)

Solid concentration distribution

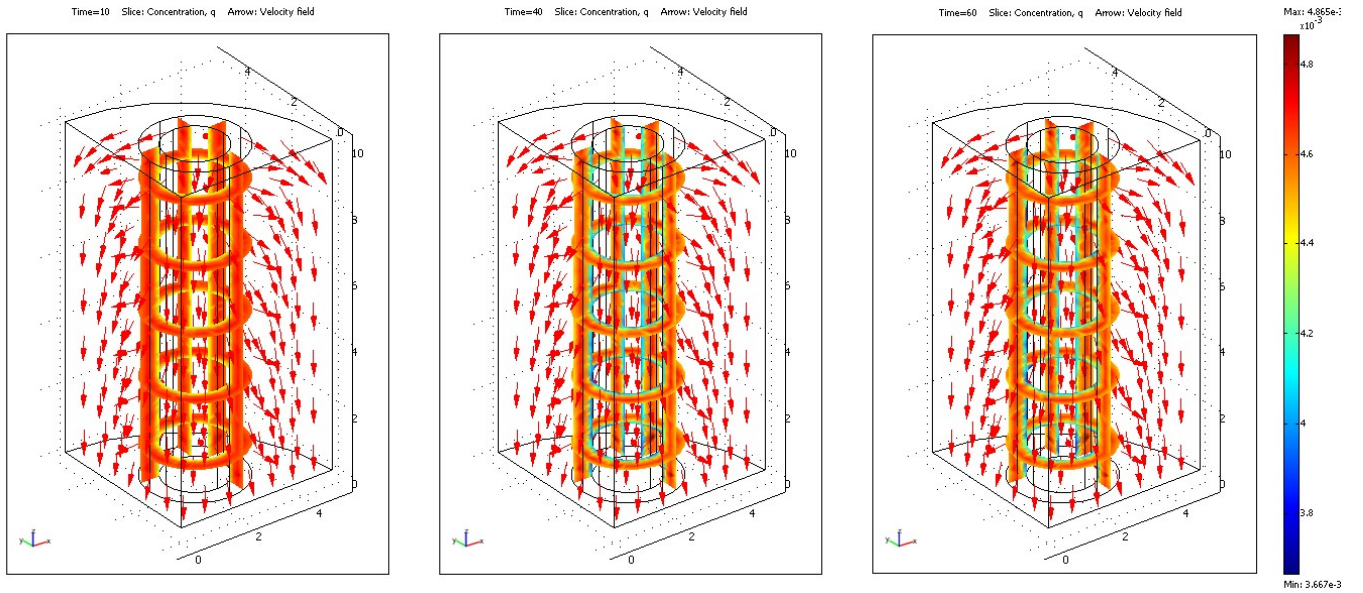


Figure 31. Concentration of the solid phase in one quarter of the 4-cartridge adsorber, during desorption (from left to right $t=10, 40$ and 60 s)

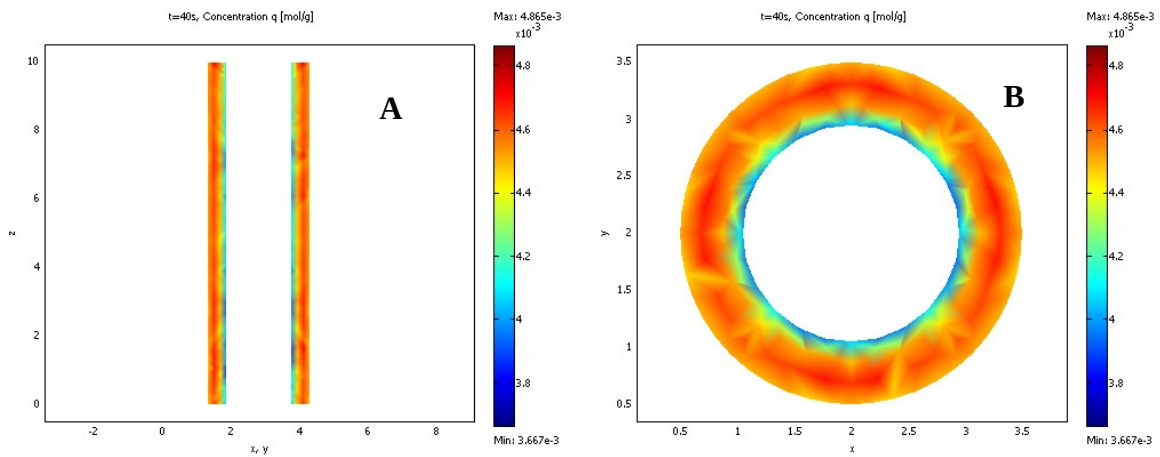


Figure 32. The solid concentration during desorption corresponding to cross-sections A and B (Fig. 9) and $t=60$ s

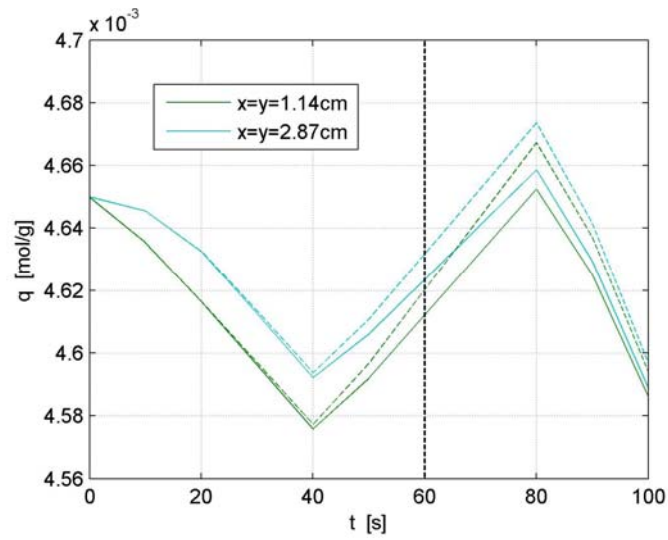


Figure 33. Time dependence of the solid concentration during desorption in 2 points in the adsorption bed (Fig. 9)

Gas temperature distribution

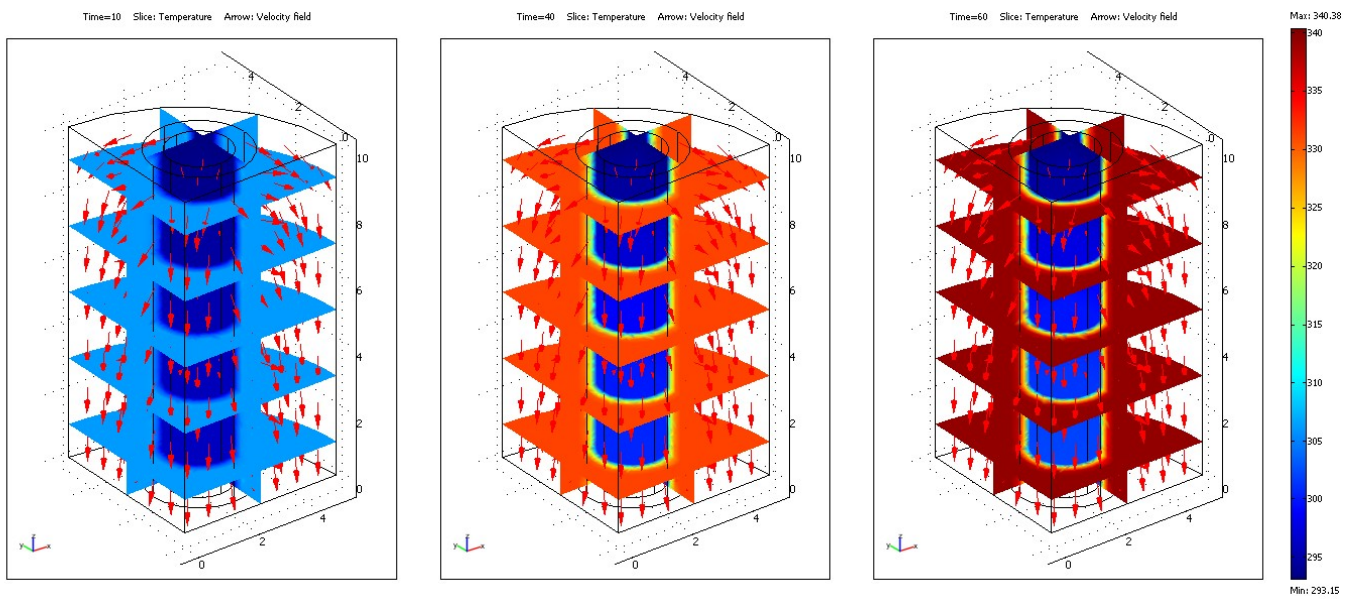


Figure 34. Temperature of the gas phase in one quarter of the 4-cartridge adsorber, during desorption (from left to right $t=10, 40$ and 60 s)

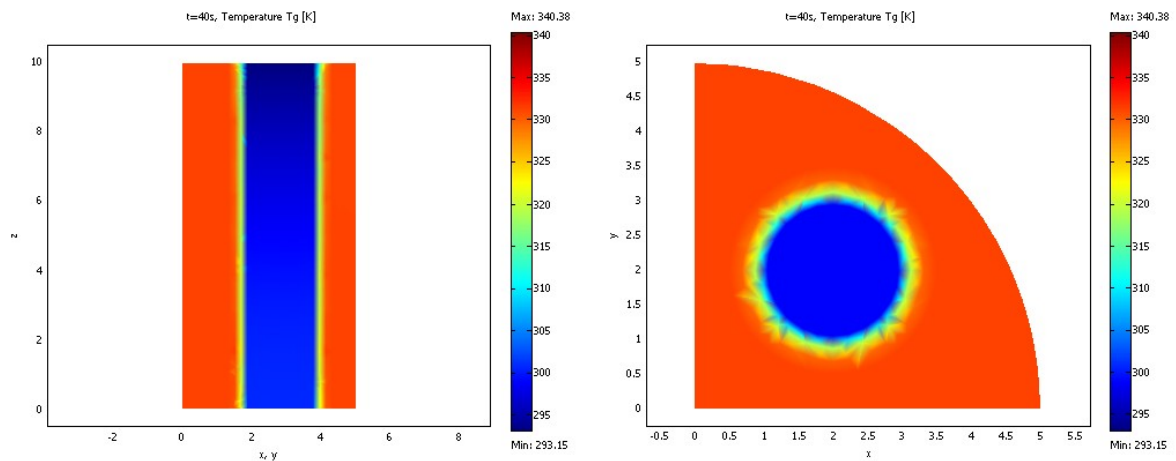


Figure 35. The gas temperature during desorption, corresponding to cross-sections A and B (Fig. 9) and $t=40$ s

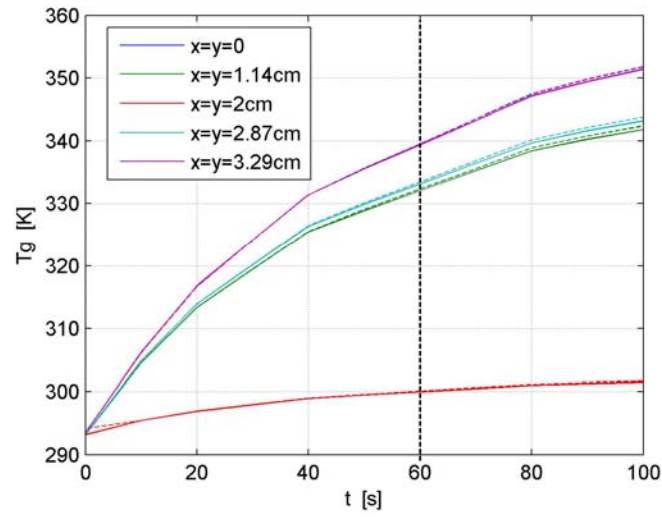


Figure 36. Time dependence of the gas temperature during desorption in 5 points in the adsorber (Fig. 9)

Solid temperature distribution

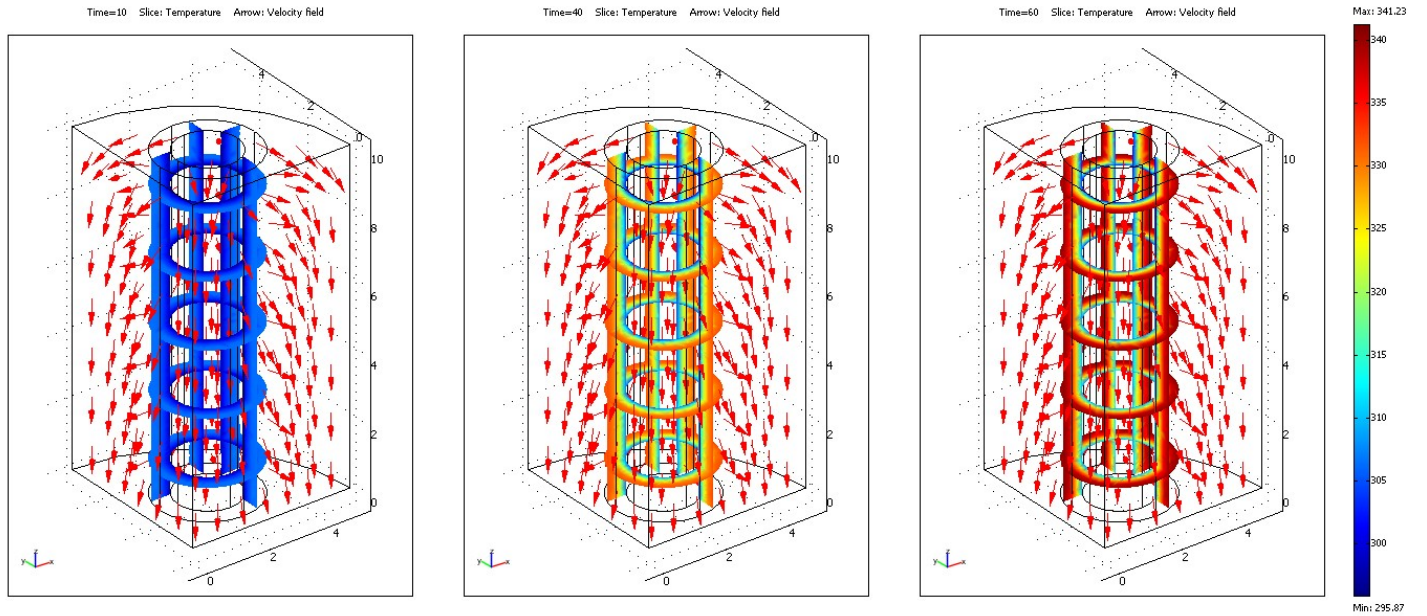


Figure 37. Temperature of the solid phase in one quarter of the 4-cartridge adsorber, during desorption (from left to right $t=10, 40$ and 60 s)

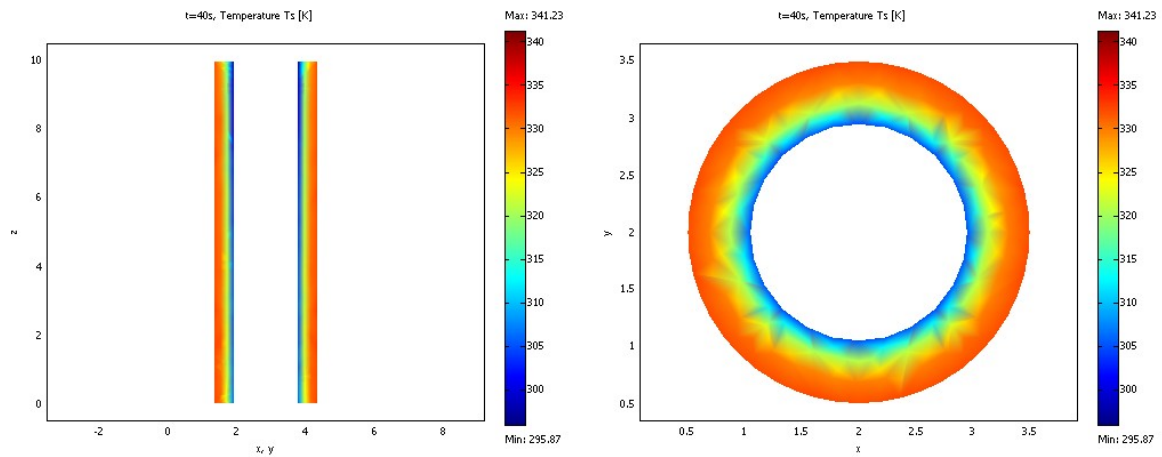


Figure 38. The solid temperature during desorption, corresponding to cross-sections A and B (Fig. 9) and $t=40$ s

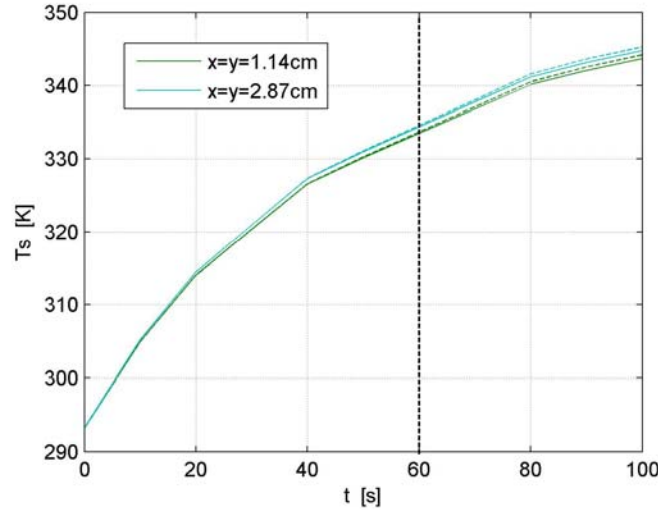


Figure 39. Time dependence of the solid temperature during desorption, in 2 points in the adsorbent bed (Fig. 9)

4.3. Simulation of desorption accompanied with condensation

The desorption process accompanied with condensation of the desorbed vapor at the adsorber wall was simulated using Model_DC3. All 7 application modes defined in Section 2.2. were again active in this model. The simulation of this stage was performed as a continuation of the previous one – desorption without condensation, i.e. the stored solution of Model_D3 for $t=60$ s was used for the initial values of the variables for Model_DC3. Model_D3 and Model_DC3 differ only in the boundary conditions for the mass and heat transfer equations at the adsorber wall (equations (29) and (29a)). Faster convergence was obtained if the change of these boundary conditions was performed in two steps: first only for the mass transfer equation and then for the heat transfer equation, as well. All model parameters, including the flow rate and voltage were the same as in the previous stage.

As mentioned in Section 3.3., the solution of the complete Model_DC3 did not converge, even with increased tolerances. Because of that, the simulation results presented here were obtained with fixed velocities (the final values of the velocities obtained with Model_D3 for $t=60$ s were used). The simulation results for adsorption and desorption without condensation show that the solution obtained with fixed velocities don't differ very much from the final solution of the integral models, so the approximate solution obtained for fixed velocities is acceptable.

The simulation results for desorption accompanied with condensation are presented in an analogous way as for adsorption and desorption without condensation. The only difference is in the time profiles, as the solution of the integral model with variable velocities is not available. Also, as the velocities are kept constant, the velocity distribution is the same as for desorption without condensation at $t=60$ s. so it is not shown here. The 3-D plots are shown for $t=60$, 140 and 330 s, and the 2-D plots for $t=140$ s. The results are given in Figures 40 to 51.

Gas concentration distribution

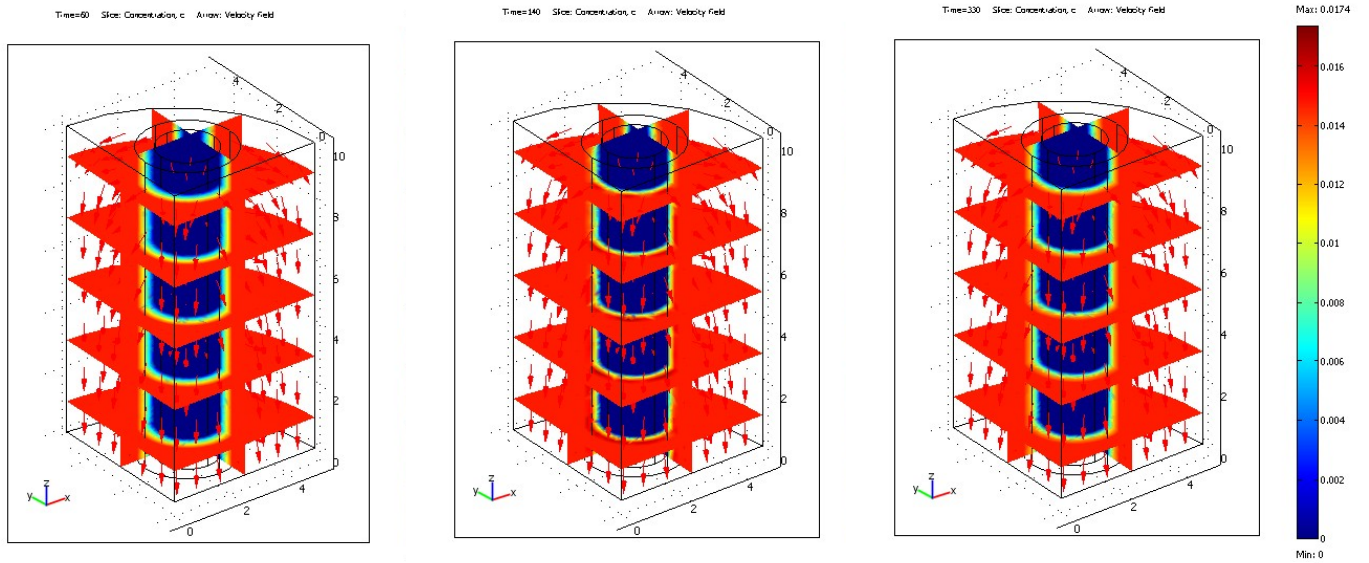


Figure 40. Temperature of the gas phase in one quarter of the 4-cartridge adsorber, during desorption with condensation (from left to right $t=60$, 140 and 330 s)

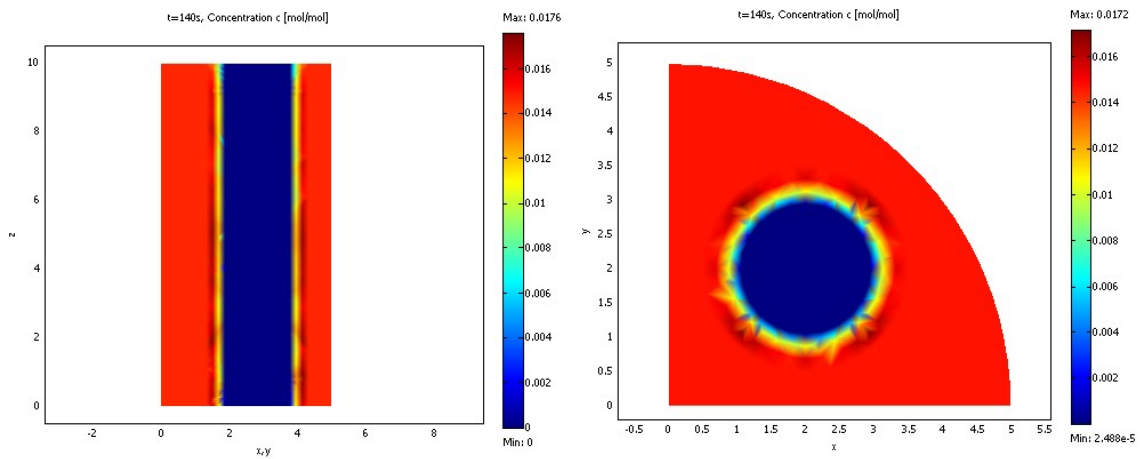


Figure 41. The gas concentration during desorption with condensation, corresponding to cross-sections A and B (Fig. 9) and $t=40$ s

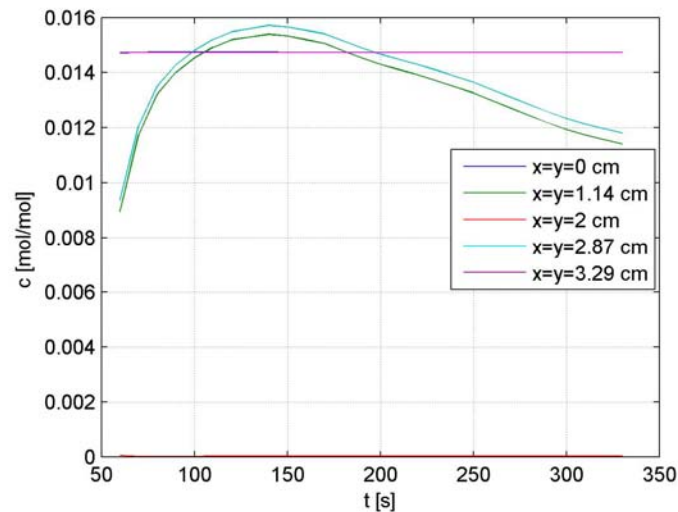


Figure 42. Time dependence of the gas concentration during desorption with condensation, in 5 points in the adsorber (Fig. 9)

Solid concentration distributions

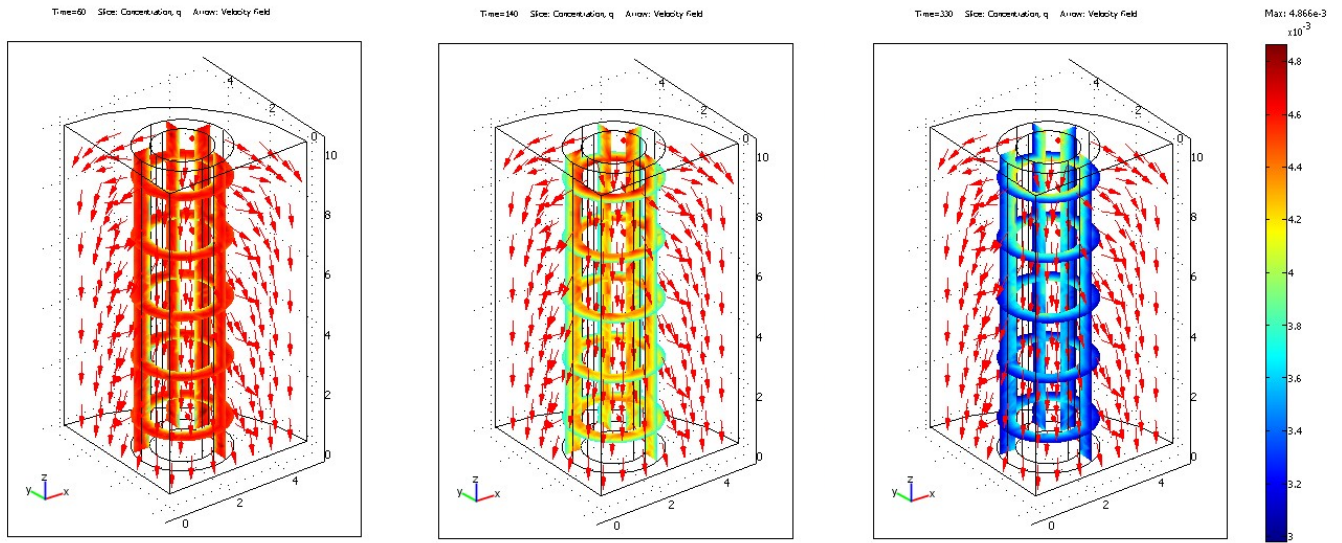


Figure 43. Concentration of the solid phase in one quarter of the 4-cartridge adsorber, during desorption with condensation (from left to right $t=60$, 140 and 330 s)

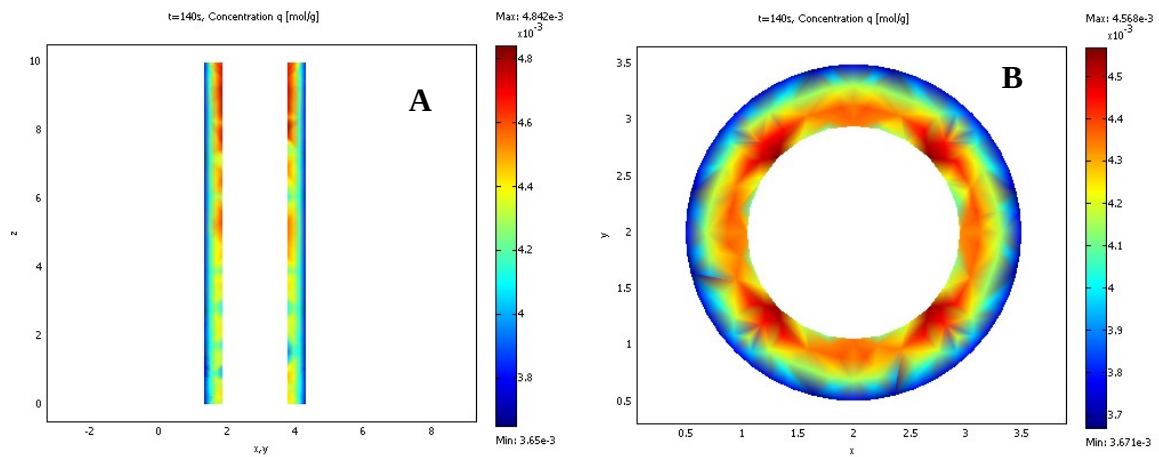


Figure 44. The solid concentration during desorption with condensation, corresponding to cross-sections A and B (Fig. 9) and $t=140$ s

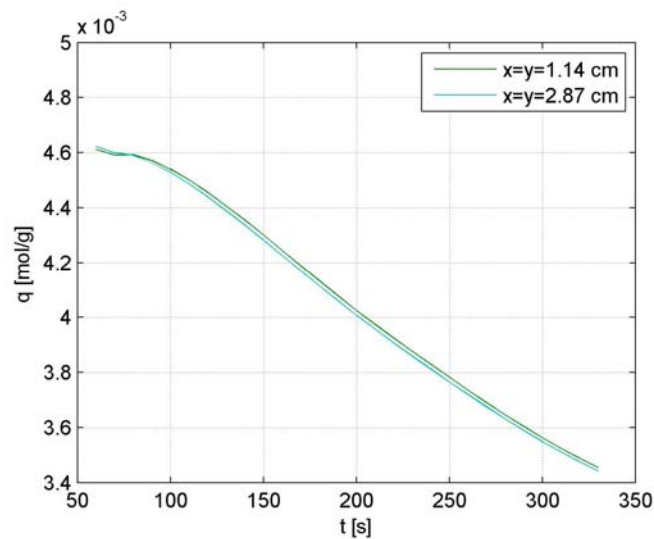


Figure 45. Time dependence of the solid concentration during desorption with condensation, in 2 points in the adsorbent bed

Gas temperature distributions

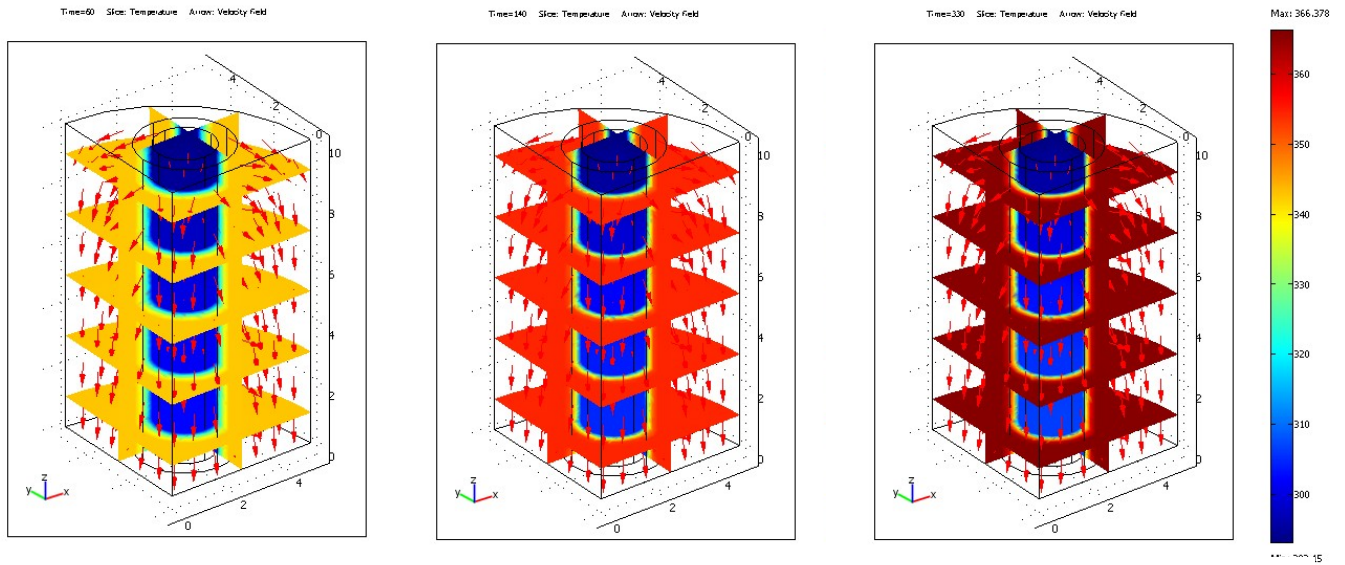


Figure 46. Temperature of the gas phase in one quarter of the 4-cartridge adsorber, during desorption with condensation (from left to right $t=60$, 140 and 330 s)

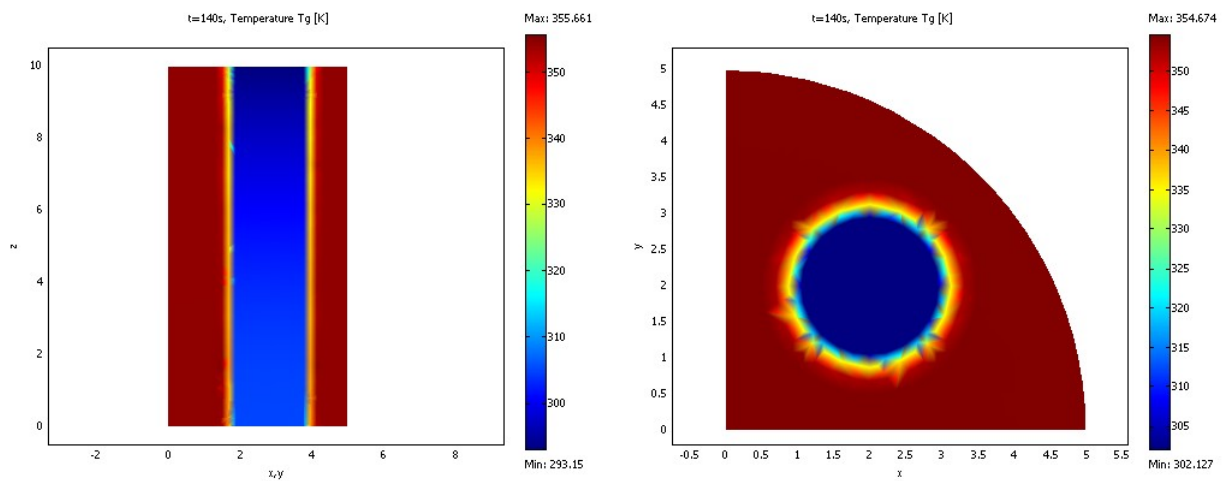


Figure 47. The gas temperature during desorption with condensation, corresponding to cross-sections A and B (Fig. 9) and $t=40$ s

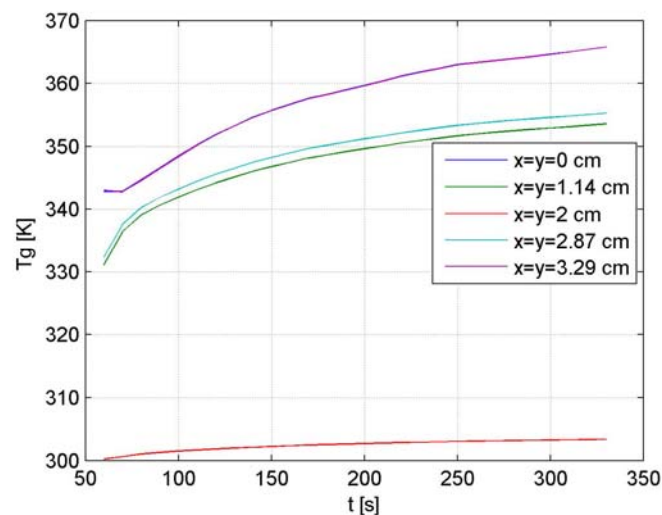


Figure 48. Time dependence of the gas temperature during desorption with condensation in 2 points in the adsorption bed

Solid temperature distributions

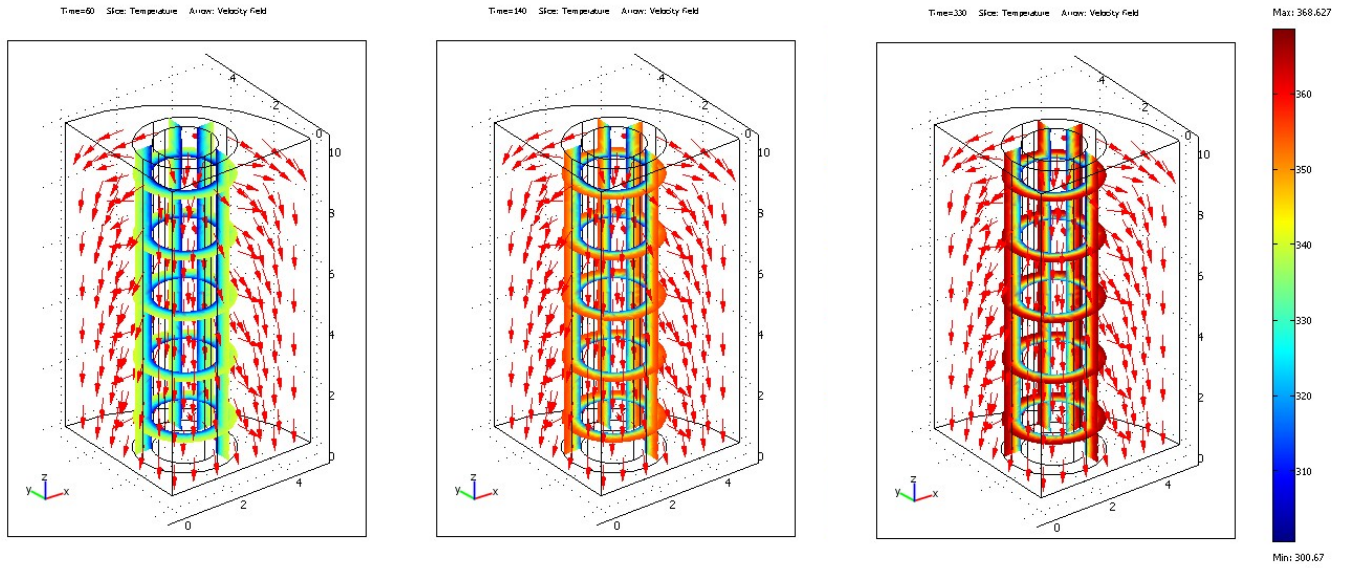


Figure 49. Temperature of the solid phase in one quarter of the 4-cartridge adsorber, during desorption with condensation (from left to right $t=60$, 140 and 330 s)

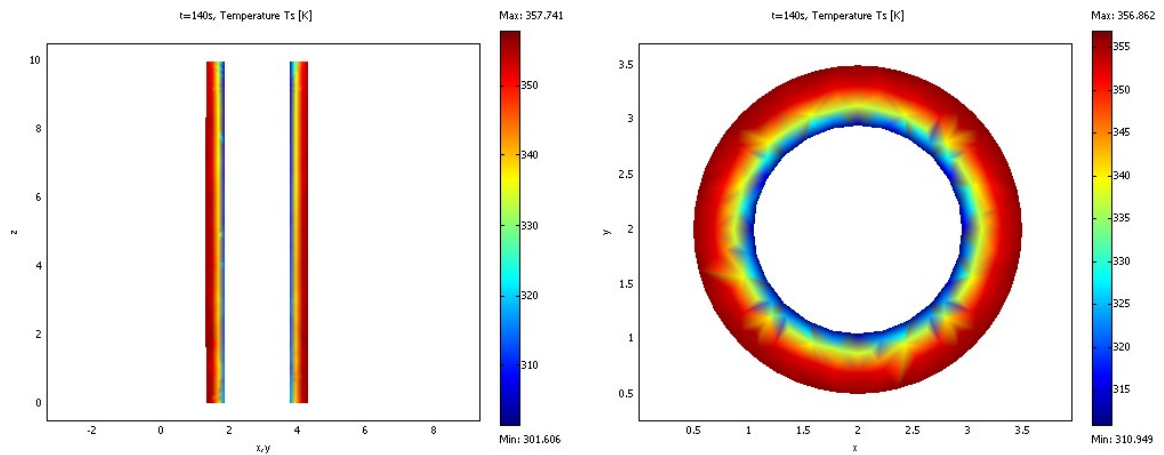


Figure 50. The solid temperature during desorption with condensation, corresponding to cross-sections A and B (Fig. 9) and $t=40$ s

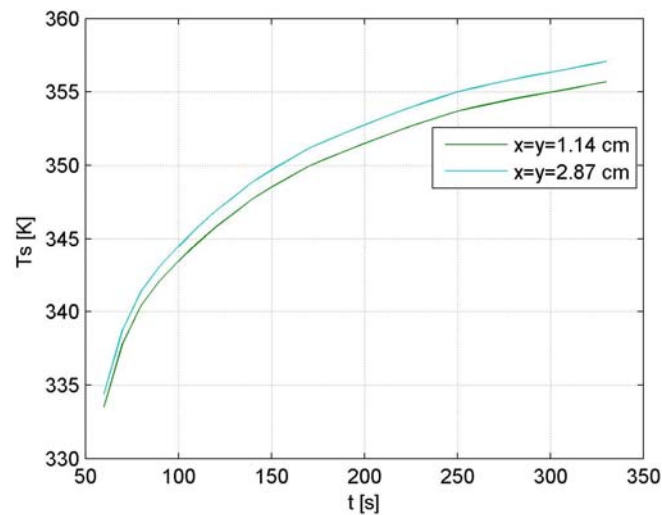


Figure 51. Time dependence of the solid temperature during desorption with condensation, in 2 points in the adsorption bed

The condensation rate has also been calculated, using equations (30) and (31). Its time dependence is shown in Figure 52.

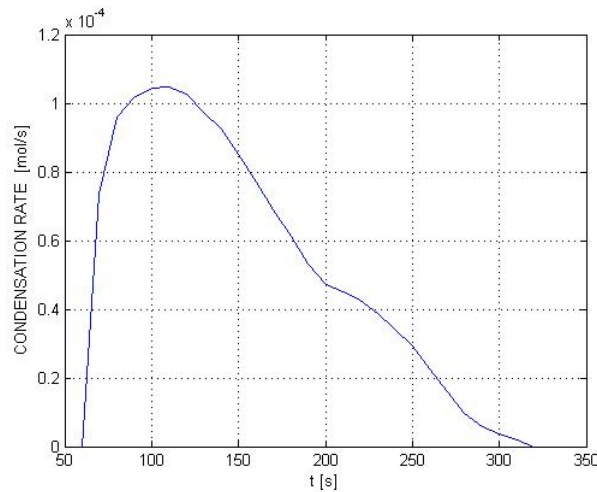


Figure 52. Condensation rate in the 4-cartridge adsorber

Figure 52 shows that the condensation starts at 60 s and ends at 330 s. Integration of the condensation rate over time gives the total amount of the condensed liquid:

$$L_{cond} = 1.415 \times 10^{-2} \text{ mol}$$

The used electric power per cartridge was also calculated, as defined by equation (33) and (34). It was calculated for both stages of desorption, before and after the start of condensation. It is shown in Figure 53.

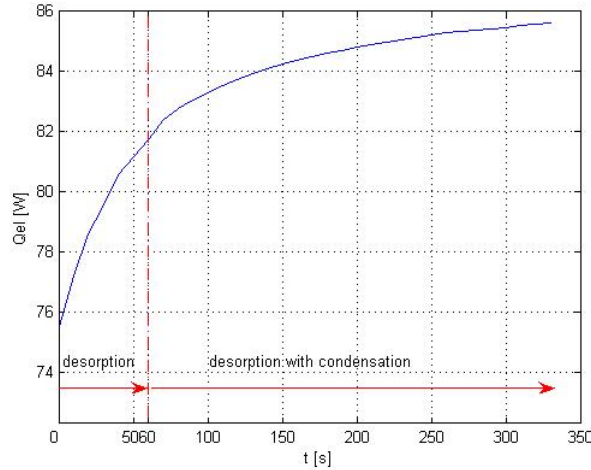


Figure 53. The used electric power during desorption, per cartridge

Integration of the electric power over time gives the spent electric energy during desorption. For the whole adsorber with four cartridges, this energy is:

$$E = 110.272 \text{ kJ}$$

5. Conclusions

The objective of this phase of the project: to model the TSA system with an adsorber with four parallel, cartridge-type adsorbers, raised a very difficult and demanding task, even when using an excellent modeling tool, such as COMSOL Multiphysics. It was necessary to build 3-D models of a very complex system, with complex geometry and a large number of interacting phenomena and processes. Even when the problem was reduced to modeling of only one quarter of the adsorber (assuming that all four cartridges in the adsorber were identical), seven interacting application modes needed to be incorporated into a 3-D model geometry consisting of 3 sub-domains. This resulted with a very large number of degrees of freedom, i.e., with a large number of coupled nonlinear differential equations that needed to be solved numerically.

The first problem associated with that was a very high demand regarding computer memory. A rather coarse mesh had to be used in order to enable solving of the models on a personal computer with 4 GB RAM.

The second problem was diverging of the solution. The complete models did not converge. For that reason, a step-by-step solution procedure had to be developed. For the model of desorption with condensation, even this procedure failed, and only an approximate solution of the model with fixed velocities could be obtained. We believe that this procedure could help other researchers in solving their difficult 3-D problems.

The third problem was in relatively long computing times. For example, when solving the model for adsorption, we needed ~200 s for Step 1, ~1150 s for Step 2 and ~4300 s for Step 3.

Nevertheless, in spite of all these problems, we succeeded in postulating and solving the models describing all three stages of the TSA process in the 4-cartridge adsorber.

We believe that the convergence problems are partly associated with the coarse mesh used, and that using a finer mesh would improve the convergence, on account of a larger computer memory and longer computer times. Of course, this would demand using of a much faster computer with more memory.

With this phase of the project we completed mathematical modeling of the adsorber configurations developed in the group of Prof. Rood. These models can be used for simulation, optimization of the process parameters and adsorber geometry, but also for comparison between different configurations. Nevertheless, for practical applications of these models, it would be necessary to have accurate physical and transport parameters for the particular system. Usually, such parameters have to be estimated from experimental data. Also, more accurate solution of the mathematical models, obtained with finer mesh and lower tolerances, would be needed, which is directly related to the computer hardware. This is especially important for the 3-D models developed in this phase of the project.

The set of mathematical models developed in the three phases of the project offer a good basis for optimization of the TSA processes with electrothermal desorption.

Nomenclature

a (m^2/m^3)	- Specific surface area
b (K^{-1})	- Temperature coefficient of the bed electrical resistivity
C (mol/mol)	- Adsorbate concentration in the gas phase
C^* (mol/mol)	- Adsorbate concentration in the gas phase in equilibrium with the solid phase
C_{sat} (mol/mol)	- Saturation concentration
c_{pg} ($\text{J}/\text{mol}/\text{K}$)	- Specific heat capacity of the inert gas
c_{pl} ($\text{J}/\text{mol}/\text{K}$)	- Specific heat capacity of liquid adsorbate
c_{ps} ($\text{J}/\text{g}/\text{K}$)	- Heat capacity of the solid phase
c_{pv} ($\text{J}/\text{mol}/\text{K}$)	- Specific heat capacity of gaseous adsorbate
D_m ($\text{mol}/\text{cm}/\text{s}$)	- Mass transfer dispersion coefficient
D_t^{hg} ($\text{W}/\text{K}/\text{cm}$)	- Heat diffusivity of the gas phase
D_t^{hs} ($\text{W}/\text{K}/\text{cm}$)	- Heat diffusivity of the solid phase
E (J/mol)	- Adsorption energy of the adsorbate (D-R eq.)
g (cm/s^2)	- Gravitation constant
G (mol/s)	- Flow rate of the inert gas
H (cm)	- Bed axial dimension (Fig. 7)
h_b ($\text{J}/\text{cm}^2/\text{K}$)	- Solid to gas heat transfer coefficient
h_{s1} ($\text{J}/\text{cm}^2/\text{K}$)	- Heat transfer coefficient from the solid phase to the gas phase in the central tube(s)
h_{s2} ($\text{J}/\text{cm}^2/\text{K}$)	- Heat transfer coefficient from the solid phase to the gas phase in the annular tube
h_{wg} ($\text{J}/\text{cm}^2/\text{K}$)	- Gas to ambient heat transfer coefficient (heat losses)
J (A/cm^2)	- Current density
J_{cond} ($\text{mol}/\text{cm}^2/\text{s}$)	- Condensation flux
k (cm^2)	- Adsorbent bed permeability
k_m ($\text{mol}/\text{cm}^2/\text{s}$)	- Mass transfer coefficient in the adsorbent bed
k_{m1} ($\text{mol}/\text{cm}^2/\text{s}$)	- Mass transfer coefficient between the adsorbent bed and the gas in the central tube
k_{m2} ($\text{mol}/\text{cm}^2/\text{s}$)	- Mass transfer coefficient between the adsorbent bed and the gas in the annular tube
$\dot{L}_{cond}$ (mol/s)	- Flow-rate of the condensed liquid
L_{cond} (mol)	- Total amount of the condensed liquid
p (Pa)	- Gas pressure
p_A (Pa)	- Adsorbate partial pressure
p_c (Pa)	- Critical pressure
p^0 (Pa)	- Adsorbate saturation pressure
$\dot{Q}_{el}$ (W)	- Rate of heat generation (equal to electric power)
Q_{el} (J)	- Heat generation (equal to electric energy)
q (mol/g)	- Adsorbate concentration in the solid phase
R_g ($\text{J}/\text{mol}/\text{K}$)	- Gas constant

r_1 (cm)	- radius of the central tube (Fig. 7)
r_2 (cm)	- outer radius of the adsorbent bed (Fig. 7)
r_1 (cm)	- radius of the adsorber vessel (Fig. 7)
T_a (K)	- Ambient temperature
T_g (K)	- Gas phase temperature
T_c (K)	- Critical temperature
T_R (K)	- Referent temperature
T_s (K)	- Solid phase temperature
T_w (K)	- Wall temperature
t (s)	- Time
U (V)	- Electric potential
U_0 (V)	- Supply voltage
u (cm/s)	- x-component of the gas velocity
VP_A, VP_B, VP_C, VP_D - Wagner constants	
v (cm/s)	- y-component of the gas velocity
W (cm ³ /g)	- Adsorbate volume per 1 g of the adsorbent mass
W_0 (cm ³ /g)	- Total volume of micropores (D-R equation)
w (cm/s)	- z-component of the gas velocity
x (cm)	- x - coordinate
x_0 (cm)	- position of the cartridge center in the x-direction
y (cm)	- y - coordinate
y_0 (cm)	- position of the cartridge center in the y-direction
z (cm)	- z - coordinate

Greek letters:

ΔH_{ads} (J/mol)	- Heat of adsorption
ΔH_{cond} (J/mol)	- Heat of condensation
ε_b	- Bed porosity
μ (Pa/s)	- Dynamic viscosity
ρ (Ω cm)	- Electric resistivity
ρ_0 (Ω cm)	- Electric resistivity at referent temperature T_R
ρ_g (mol/cm ³)	- Gas phase density
ρ_b (g/cm ³)	- Adsorbent bed density
ρ_A (g/cm ³)	- Adsorbate density

Subscripts:

<i>b</i>	- bed
<i>g</i>	- gas
<i>in</i>	- inlet
<i>i,o</i>	- inner and outer tube
<i>out</i>	- outlet
<i>p</i>	- previous (initial)
<i>s</i>	- solid phase
<i>it</i>	- inner (central) tube
<i>ot</i>	- outer (annular) tube
<i>x</i>	- x-direction
<i>y</i>	- y-direction
<i>z</i>	- z-direction

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